Spoon Feeding Empirical Formulae

Simplified Knowledge Management Classes Bangalore

My name is Subhashish Chattopadhyay. I have been teaching for IIT-JEE, Various International Exams (such as IMO [International Mathematics Olympiad], IPhO [International Physics Olympiad], IChO [International Chemistry Olympiad]), IGCSE (IB), CBSE, I.Sc, Indian State Board exams such as WB-Board, Karnataka PU-II etc since 1989. As I write this book in 2016, it is my 25 th year of teaching. I was a Visiting Professor to BARC Mankhurd, Chembur, Mumbai, Homi Bhabha Centre for Science Education (HBCSE) Physics Olympics camp BARC Campus.

I am Life Member of ...

- IAPT (Indian Association of Physics Teachers)

- IPA (Indian Physics Association)

- AMTI (Association of Mathematics Teachers of India)

- National Human Rights Association

- Men's Rights Movement (India and International)

- MGTOW Movement (India and International)

And also of

IACT (Indian Association of Chemistry Teachers)

The selection for National Camp (for Official Science Olympiads - Physics, Chemistry, Biology, Astronomy) happens in the following steps ....

1) **NSEP** (National Standard Exam in Physics) and **NSEC** (National Standard Exam in Chemistry) held around 24 rth November. Approx 35,000 students appear for these exams every year. The exam fees is Rs 100 each. Since 1998 the IIT JEE toppers have been topping these exams and they get to know their rank / performance ahead of others.

2) **INPhO** (Indian National Physics Olympiad) and **INChO** (Indian National Chemistry Olympiad). Around 300 students in each subject are allowed to take these exams. Students coming from outside cities are paid fair from the Govt of India.

3) The Top 35 students of each subject are invited at HBCSE (Homi Bhabha Center for Science Education) Mankhurd, near Chembur, BARC, Mumbai. After a 2-3 weeks camp the top 5 are selected to represent India. The flight tickets and many other expenses are taken care by Govt of India.

Since last 50 years there has been no dearth of “Good Books“. Those who are interested in studies have been always doing well. This e-Book does not intend to replace any standard text book. These topics are very old and already standardized.

There are 3 kinds of Text Books

- The thin Books - Good students who want more details are not happy with these. Average students who need more examples are not happy with these. Most students who want to “Cram” quickly and pass somehow find the thin books “good” as they have to read less !!

- The Thick Books - Most students do not like these, as they want to read as less as possible. Average students are “busy” with many other things and have no time to read all these.

- The Average sized Books - Good students do not get all details in any one book. Most bad students do not want to read books of “this much thickness“ also !!

We know there can be no shoe that’s fits in all.

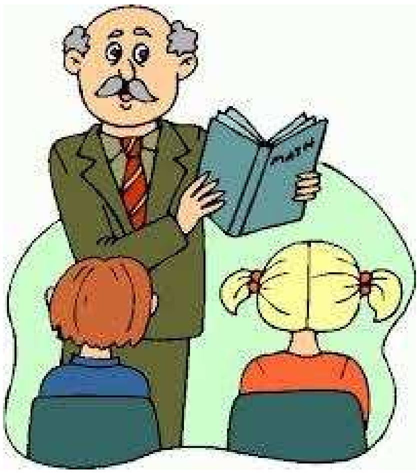
Printed books are not e-Books! Can’t be downloaded and kept in hard-disc for reading “later” ....

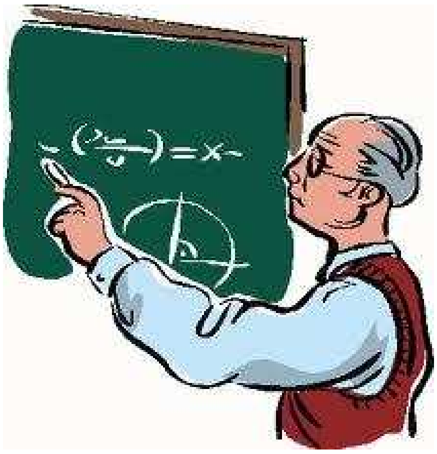
So if you read this book later, you will get all kinds of examples in a single place. This becomes a very good “Reference Material”. I sincerely wish that all find this “very useful”.

Students who do not practice lots of problems, do not do well. The rules of “doing well” had never changed .... Will never change !

In this e-Book I am trying to solve this problem. Those students who practice can learn.

No one can help those who are not studying, or practicing.





CBSE Standard 12 Chem Survival Guide - Empirical Formulae by Prof. Subhashish Chattopadhyay SKMClasses Bangalore Useful for IIT-JEE, I.Sc. PU-II, Boards, IGCSE IB AP-Chemistry and other exams

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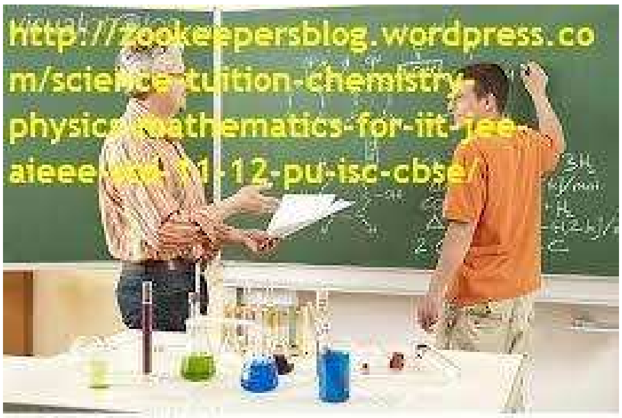


A very polite request :

I wish these e-Books are read only by Boys and Men. Girls and Women, better read something else; learn from somewhere else.

Preface

We all know that in the species “Homo Sapiens “, males are bigger than females. The reasons are explained in standard 10, or 11 (high school) Biology texts. This shapes or size, influences all of our culture. Before we recall / understand the reasons once again, let us see some random examples of the influence

Random - 1

If there is a Road rage, then who all fight ? (generally ?). Imagine two cars driven by adult drivers. Each car has a woman of similar age as that of the Man. The cars “ touch “ or “ some issue happens”. Who all comes out and fights ? Who all are most probable to drive the cars ?

(Men are eager to fight, eager to rule, eager for war. Men want to drive. Men want to win)

Random - 2

Heavy metal music artists are all Men. Metallica, Black Sabbath, Motley Crue, Megadeth, Motorhead, AC/DC, Deep Purple, Slayer, Guns & Roses, Led Zeppelin, Aerosmith ….. the list can be in thousands. All these are grown-up Boys, known as Men.

(Men strive for perfection. Men are eager to excel. Men work hard. Men want to win.)

Random - 3

Apart from Marie Curie, only one more woman got Nobel Prize in Physics. (Maria Goeppert Mayer - 1963). So, … almost all are men.

(Men want to excel. Men strive for perfection. Men want to win. Men work hard. Men do better than women.)

Random - 4

The best Tabla Players are all Men.

(Men want to excel. Men strive for perfection. Men want to win. Men work hard. Men do better than women.)

Random - 5

History is all about, which Kings ruled. Kings, their men, and Soldiers went for wars. History is all about wars, fights, and killings by men.

Boys start fighting from school days. Girls do not fight like this

(Men are eager to fight, eager to rule, eager for war. Men want to drive. Men want to win.)

Random - 6

The highest award in Mathematics, the " Fields Medal " is around since decades. Till date only one woman could get that. (Maryam Mirzakhani - 2014). So, … almost all are men.

(Men want to excel. Men strive for perfection. Men want to win. Men work hard. Men do better than women.)

Random - 7

Actor is a gender neutral word. Could the movie like " Top Gun " be made with Female actors ? The best pilots, astronauts, Fighters are all Men.

Random - 8

In my childhood had seen a movie named " The Tower in Inferno ". In the movie when the tall tower is in fire, women were being saved first, as only one lift was working....

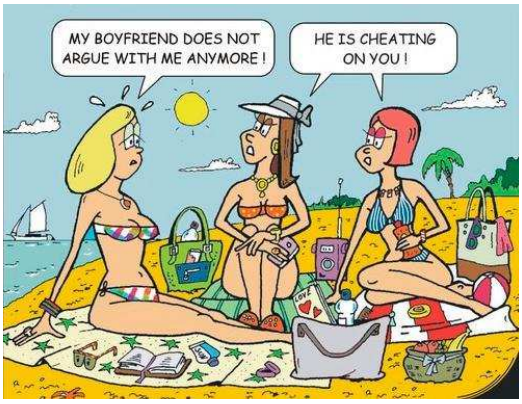
Many decades later another movie is made. A box office hit. " The Titanic ". In this also .... As the ship is sinking women are being saved. **Men are disposable**. Men may get their turn later...

Movies are not training programs. Movies do not teach people what to do, or not to do. Movies only reflect the prevalent culture. Men are disposable, is the culture in the society. Knowingly, unknowingly, the culture is depicted in Movies, Theaters, Stories, Poems, Rituals, etc. I or you can’t write a story, or make a movie in which after a minor car accident the Male passengers keep seating in the back seat, while the both the women drivers come out of the car and start fighting very bitterly on the road. There has been no story in this world, or no movie made, where after an accident or calamity, Men are being helped for safety first, and women are told to wait.

Random - 9

Artists generally follow the prevalent culture of the Society. In paintings, sculptures, stories, poems, movies, cartoon, Caricatures, knowingly / unknowingly, “ the prevalent Reality “ is depicted. The opposite will not go well with people. If deliberately “ the opposite “ is shown then it may only become a special art, considered as a special mockery.

पत्नी (सल्टू से) : मुझे नई साड़ी ला दो प्लीज।
सल्टू : पर तुम्हारी दो-दो अलमारियां साड़ियों से ही तो भरी है।
पत्नी – वह सारी तो पूरे मोहल्ले वालों ने देख रखी है।
सल्टू – तो साड़ी लेने के बजाए मोहल्ला बदल लेते हैं।



Random - 10

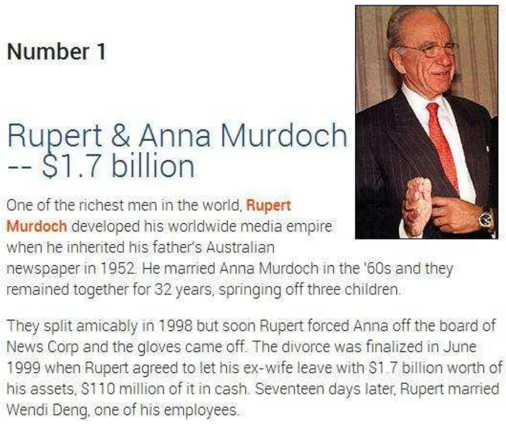
Men go to “girl / woman’s house” to marry / win, and bring her to his home. That is a sort of winning her. When a boy gets a “ Girl-Friend “, generally he and his friends consider that as an achievement. The boy who “ got / won “ a girl-friend feels proud. His male friends feel, jealous, competitive and envious. Millions of stories have been written on these themes. Lakhs of movies show this. Boys / Men go for “ bike race “, or say “ Car Race “, where the winner “ gets “ the most beautiful girl of the college.

(Men want to excel. Men are eager to fight, eager to rule, eager for war. Men want to drive. Men want to win.)

Prithviraj Chauhan ‘ went ` to “ pickup “ or “ abduct “ or “ win “ or “ bring “ his love. There was a Hindi movie (hit) song … “ Pasand ho jaye, to ghar se utha laye “. It is not other way round. Girls do not go to Boy’s house or man’s house to marry. Nor the girls go in a gang to “ pick-up “ the boy / man and bring him to their home / place / den.

Random - 11

Rich people; often are very hard working. Successful business men, establish their business (empire), amass lot of wealth, with lot of difficulty. Lots of sacrifice, lots of hard work, gets into this. Rich people's wives had no contribution in this wealth creation. Women are smart, and successful upto the extent to choose the right/rich man to marry. So generally what happens in case of Divorces ? Search the net on " most costly divorces " and you will know. The women;(who had no contribution at all, in setting up the business / empire), often gets in Billions, or several Millions in divorce settlements.



Number 1

Rupert & Anna Murdoch -- $1.7 billion

One of the richest men in the world, Rupert Murdoch developed his worldwide media empire when he inherited his father's Australian newspaper in 1952. He married Anna Murdoch in the '60s and they remained together for 32 years, springing off three children.

They split amicably in 1998 but soon Rupert forced Anna off the board of News Corp and the gloves came off. The divorce was finalized in June 1999 when Rupert agreed to let his ex-wife leave with $1.7 billion worth of his assets, $110 million of it in cash. Seventeen days later, Rupert married Wendi Deng, one of his employees.

Ted Danson & Casey Coates -- $30 million

Ted Danson's claim to fame is undoubtedly his decade-long stint as Sam Malone on NBC's celebrated sitcom Cheers . While he did other TV shows and movies, he will always be known as the bartender of that place where everybody knows your name. He met his future first bride Casey, a designer, in 1976 while doing Erhard Seminars Training.

Ten years his senior, she suffered a paralyzing stroke while giving birth to their first child in 1979. In order to nurse her back to health, Danson took a break from acting for six months. But after two children and 15 years of marriage, the infatuation fell to pieces. Danson had started seeing Whoopi Goldberg while filming the comedy, Made in America and this precipitated the 1992 divorce. Casey got $30 million for her trouble.

See https://zookeepersblog.wordpress.com/misandry-and-men-issues-a-short-summary-at-single-place/

See http://skmclasses.kinja.com/save-the-male-1761788732

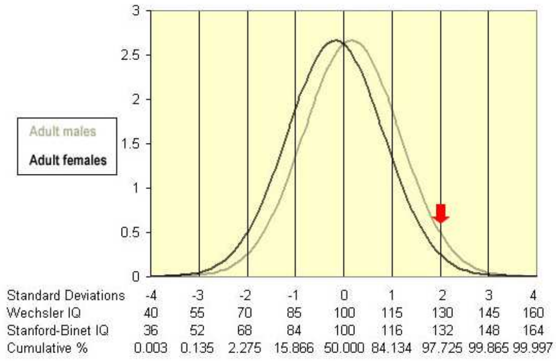
It was Boys and Men, who brought the girls / women home. The Laws are biased, completely favoring women. The men are paying for their own mistakes.

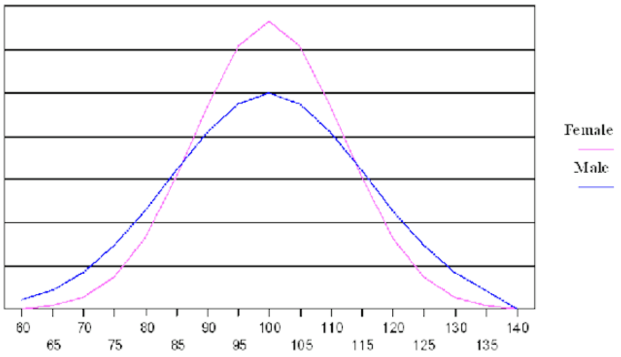
See https://zookeepersblog.wordpress.com/biased-laws/

(Man brings the Woman home. When she leaves, takes away her share of big fortune!)

Random - 12

A standardized test of Intelligence will never be possible. It never happened before, nor ever will happen in future; where the IQ test results will be acceptable by all. In the net there are thousands of charts which show that the intelligence scores of girls / women are lesser. Debates of Trillion words, does not improve performance of Girls.

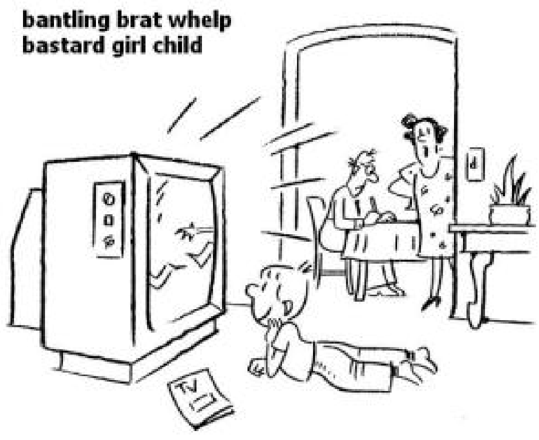




I am not wasting a single second debating or discussing with anyone, on this. I am simply accepting ALL the results. IQ is only one of the variables which is required for success in life. Thousands of books have been written on “ Networking Skills “, EQ (Emotional Quotient), Drive, Dedication, Focus, “ Tenacity towards the end goal “ … etc. In each criteria, and in all together, women (in general) do far worse than men. Bangalore is known as “ ….. capital of India “. [Fill in the blanks]. The blanks are generally filled as “ Software Capital “, “ IT Capital “, “ Startup Capital “, etc. I am member in several startup eco-systems / groups. I have attended hundreds of meetings, regarding “ technology startups “, or “ idea startups “. These meetings have very few women. Starting up new companies are all “ Men’s Game “ / “ Men’s business “. Only in Divorce settlements women will take their goodies, due to Biased laws. There is no dedication, towards wealth creation, by women.

Random - 13

Many men, as fathers, very unfortunately treat their daughters as “ Princess “. Every “ non-performing “ woman / wife was “ princess daughter “ of some loving father. Pampering the girls, in name of “ equal opportunity “, or “ women empowerment “, have led to nothing.



"Please turn it down - Daddy is trying to do your homework."

See http://skmclasses.kinja.com/progressively-daughters-become-monsters-1764484338

See http://skmclasses.kinja.com/vivacious-vixens-1764483974

There can be thousands of more such random examples, where “ Bigger Shape / size “ of males have influenced our culture, our Society. Let us recall the reasons, that we already learned in standard 10 - 11, Biology text Books. In humans, women have a long gestation period, and also spends many years (almost a decade) to grow, nourish, and stabilize the child. (Million years of habit) Due to survival instinct Males want to inseminate. Boys and Men fight for the “ facility (of womb + care) “ the girl / woman may provide. Bigger size for males, has a winning advantage. Whoever wins, gets the “ woman / facility “. The male who is of “ Bigger Size “, has an advantage to win…. Leading to Natural selection over millions of years. In general “ Bigger Males “; the “ fighting instinct “ in men; have led to wars, and solving tough problems (Mathematics, Physics, Technology, startups of new businesses, Wealth creation, Unreasonable attempts to make things [such as planes], Hard work ….)

So let us see the IIT-JEE results of girls. Statistics of several years show that there are around 17, (or less than 20) girls in top 1000 ranks, at all India level. Some people will yet not understand the performance, till it is said that … year after year we have around 980 boys in top 1000 ranks. Generally we see only 4 to 5 girls in top 500. In last 50 years not once any girl topped in IIT-JEE advanced. Forget about Single digit ranks, double digit ranks by girls have been extremely rare. It is all about “ good boys “, “ hard working “, “ focused “, “**Bel-esprit** “ boys.

In 2015, Only 2.6% of total candidates who qualified are girls (upto around 12,000 rank). while 20% of the Boys, amongst all candidates qualified. The Total number of students who appeared for the exam were around 1.4 million for IIT-JEE main. Subsequently 1.2 lakh (around 120 thousands) appeared for IIT-JEE advanced.

IIT-JEE results and analysis, of many years is given at https://zookeepersblog.wordpress.com/iit-jee-iseet-main-and-advanced-results/

In Bangalore it is rare to see a girl with rank better than 1000 in IIT-JEE advanced. We hardly see 6-7 boys with rank better than 1000. Hardly 2-3 boys get a rank better than 500.

See http://skmclasses.weebly.com/everybody-knows-so-you-should-also-know.html

Professor Subhashish Chattopadhyay

-

HURT FEMINISM BY DOING NOTHING

- ✗ DON'T HELP WOMEN
- ✗ DON'T FIX THINGS FOR WOMEN
- ✗ DON'T SUPPORT WOMEN'S ISSUES
- ✗ DON'T COME TO WOMEN'S DEFENSE[1]
- ✗ DON'T SPEAK FOR WOMEN
- ✗ DON'T VALUE WOMEN'S FEELINGS
- ✗ DON'T PORTRAY WOMEN AS VICTIMS
- ✗ DON'T PROTECT WOMEN[2]

✓ WITHOUT WHITE KNIGHTS FEMINISM WOULD END TODAY

[1] Don't even nawalt ("Not All Women Are Like That")

[2] for example from criticism or insults

How Society prioritize Men

Rich women	They can get away with murder.
Women	They get all the rights with no responsibility and Shelters for Homeless women.
Rich Men	They get tax bail outs and short prison sentence.
Girls	They get educational benefits but no violence against kids Act.
Boys	They have some support but don't have any education that fits boys.
Animals	They have animal rights and PETA.
Prisoners	They get conjugal visits and 3 squares and a roof.
Men	Paid slaves.
Poor Men	Nothing.

Who pays the most Taxes?
This is why MGTOW exist.

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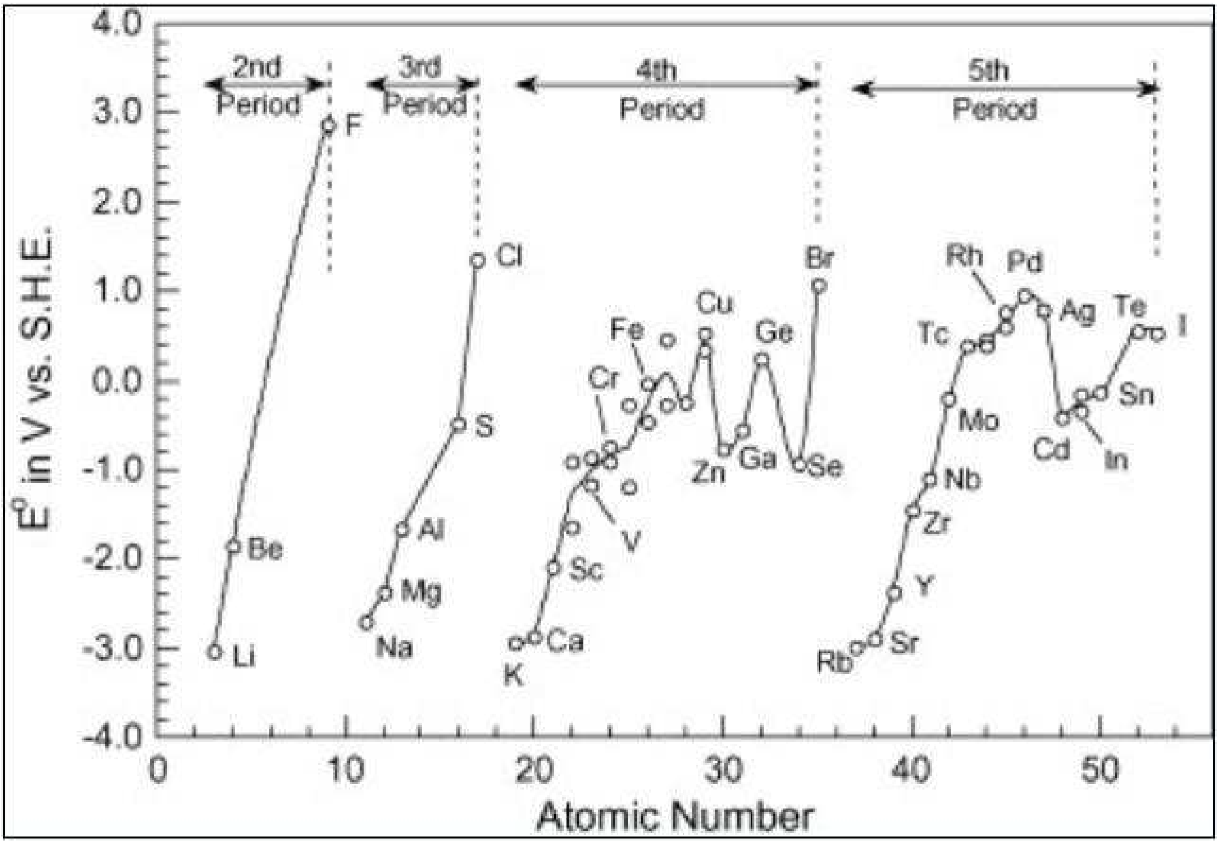
Spoon Feeding Series - Empirical Formulae in Chemistry

Most Important graph or Image of Chemistry



It is very strange, that so many Chemistry books exists in this world, which do not even have a single Periodic table in them. Almost all books do not have this graph either. This is the graph of electrode potential of various elements with respect to (standard) Hydrogen electrode.

All chemical reactions happen because of this electrode potential difference. When a Solid is added to a solution of another compound such as a Gold Spoon in $CuSO_4$ Solution, of a Copper container having $AgNO_3$ Solution, then the displacement reaction happens or not is decided by these potential values. Battery and Electrochemistry chapter uses this knowledge. Oxidation-Reduction reactions, titrations etc all need this understanding.

-

The e.m.f of the standard cell Zn | Zn^{2+} || Ag^{+} | Ag E^{o}_{cell} = 1.562 Volt

If the standard electrode potential of Zinc is - 0.762 Volt, then that of Silver is ?

Solution : $E^{o}_{cell} = E_{Right} - E_{Left} = E_{Ag+ \mid Ag} - E_{Zn2+ \mid Zn} = (E_{Reduction\ (\ Right\ )} - E_{Reduction\ (\ Left\ )})$

So 1.562 = $E(Ag^+ \mid Ag) - (-0.762)$ => $E_{Ag+ \mid Ag}$ = 0.8 Volt

-

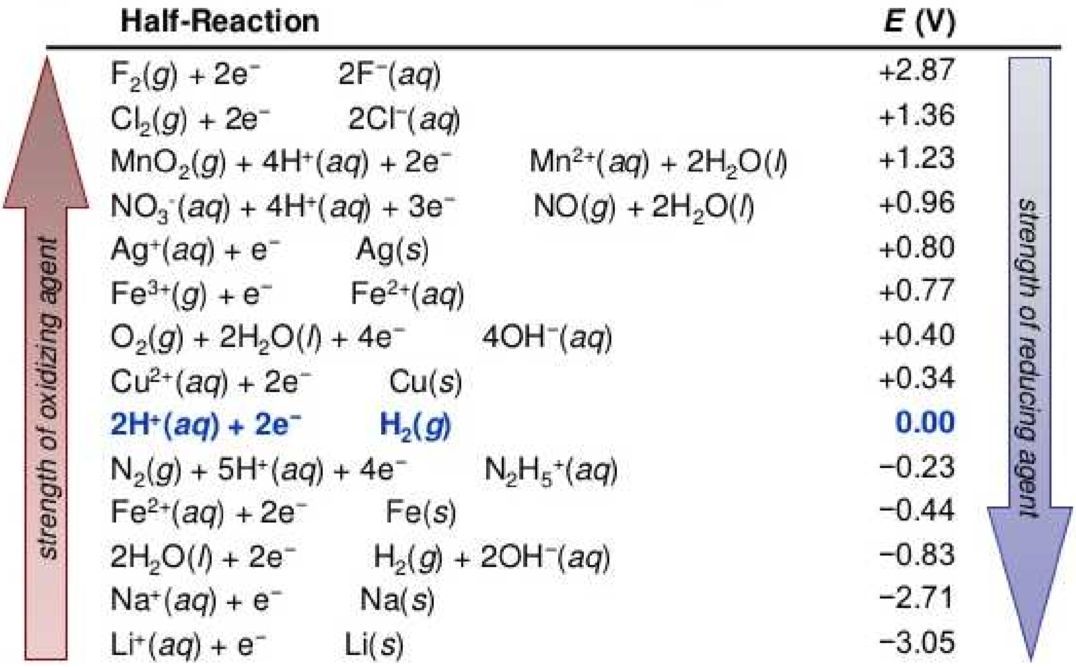
What will happen if a solution of copper sulphate put into an iron pot ?

Note : Electrode potential of Cu is more positive (so higher) than Fe (Iron).

A solution of copper sulphate put into an iron pot will damage. This is because the standard electrode potential for Fe2+/Fe is more negative than that for Cu2+/Cu. The iron oxidises first, while copper ions are reduced. Iron will corrode (rust). Or in more simple words it is easier to replace Cu ions by Fe ions from the pot. Any element lower in the series tends to replace the element which is higher up the series.

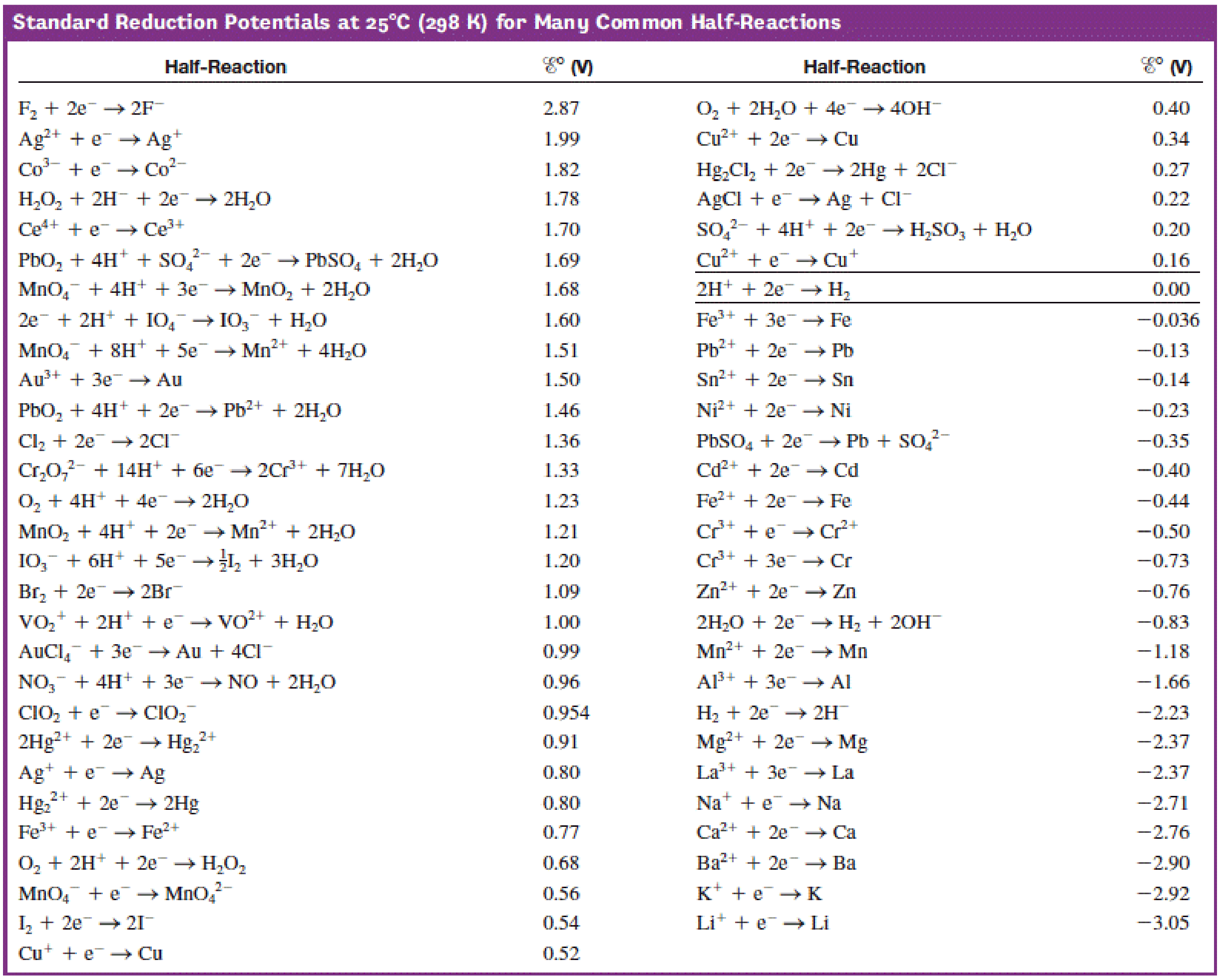
This is called a redox reaction, since the iron is being oxidised, and copper ions reduced.

Note : A solution of Iron Salt in a Copper Vessel would have just done fine. In this case copper is in solid form, as vessel. So will not replace the Iron ions.



Half-Reaction		E (V)
$F_2(g) + 2e^-$	$2F^-(aq)$	+2.87
$Cl_2(g) + 2e^-$	$2Cl^-(aq)$	+1.36
$MnO_2(g) + 4H^+(aq) + 2e^-$	$Mn^{2+}(aq) + 2H_2O(l)$	+1.23
$NO_3^-(aq) + 4H^+(aq) + 3e^-$	$NO(g) + 2H_2O(l)$	+0.96
$Ag^+(aq) + e^-$	$Ag(s)$	+0.80
$Fe^{3+}(g) + e^-$	$Fe^{2+}(aq)$	+0.77
$O_2(g) + 2H_2O(l) + 4e^-$	$4OH^-(aq)$	+0.40
$Cu^{2+}(aq) + 2e^-$	$Cu(s)$	+0.34
$2H^+(aq) + 2e^-$	**$H_2(g)$**	**0.00**
$N_2(g) + 5H^+(aq) + 4e^-$	$N_2H_5^+(aq)$	−0.23
$Fe^{2+}(aq) + 2e^-$	$Fe(s)$	−0.44
$2H_2O(l) + 2e^-$	$H_2(g) + 2OH^-(aq)$	−0.83
$Na^+(aq) + e^-$	$Na(s)$	−2.71
$Li^+(aq) + e^-$	$Li(s)$	−3.05

A more detailed table



Standard Reduction Potentials at 25°C (298 K) for Many Common Half-Reactions

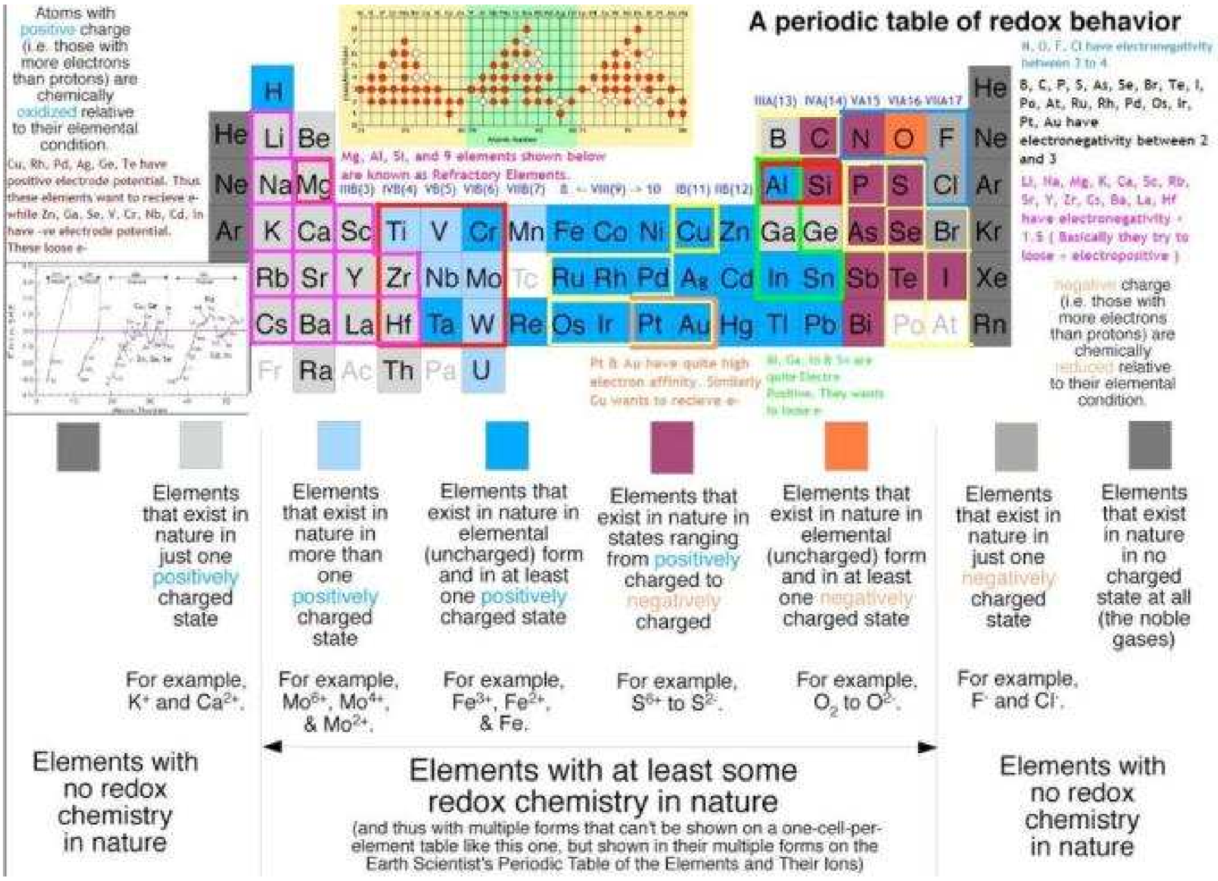
Half-Reaction	$\mathscr{E}^\circ$ (V)	Half-Reaction	$\mathscr{E}^\circ$ (V)
$F_2 + 2e^- \rightarrow 2F^-$	2.87	$O_2 + 2H_2O + 4e^- \rightarrow 4OH^-$	0.40
$Ag^{2+} + e^- \rightarrow Ag^+$	1.99	$Cu^{2+} + 2e^- \rightarrow Cu$	0.34
$Co^{3-} + e^- \rightarrow Co^{2-}$	1.82	$Hg_2Cl_2 + 2e^- \rightarrow 2Hg + 2Cl^-$	0.27
$H_2O_2 + 2H^- + 2e^- \rightarrow 2H_2O$	1.78	$AgCl + e^- \rightarrow Ag + Cl^-$	0.22
$Ce^{4+} + e^- \rightarrow Ce^{3+}$	1.70	$SO_4^{2-} + 4H^+ + 2e^- \rightarrow H_2SO_3 + H_2O$	0.20
$PbO_2 + 4H^+ + SO_4^{2-} + 2e^- \rightarrow PbSO_4 + 2H_2O$	1.69	$Cu^{2+} + e^- \rightarrow Cu^+$	0.16
$MnO_4^- + 4H^+ + 3e^- \rightarrow MnO_2 + 2H_2O$	1.68	$2H^+ + 2e^- \rightarrow H_2$	0.00
$2e^- + 2H^+ + IO_4^- \rightarrow IO_3^- + H_2O$	1.60	$Fe^{3+} + 3e^- \rightarrow Fe$	−0.036
$MnO_4^- + 8H^+ + 5e^- \rightarrow Mn^{2+} + 4H_2O$	1.51	$Pb^{2+} + 2e^- \rightarrow Pb$	−0.13
$Au^{3+} + 3e^- \rightarrow Au$	1.50	$Sn^{2+} + 2e^- \rightarrow Sn$	−0.14
$PbO_2 + 4H^+ + 2e^- \rightarrow Pb^{2+} + 2H_2O$	1.46	$Ni^{2+} + 2e^- \rightarrow Ni$	−0.23
$Cl_2 + 2e^- \rightarrow 2Cl^-$	1.36	$PbSO_4 + 2e^- \rightarrow Pb + SO_4^{2-}$	−0.35
$Cr_2O_7^{2-} + 14H^+ + 6e^- \rightarrow 2Cr^{3+} + 7H_2O$	1.33	$Cd^{2+} + 2e^- \rightarrow Cd$	−0.40
$O_2 + 4H^+ + 4e^- \rightarrow 2H_2O$	1.23	$Fe^{2+} + 2e^- \rightarrow Fe$	−0.44
$MnO_2 + 4H^+ + 2e^- \rightarrow Mn^{2+} + 2H_2O$	1.21	$Cr^{3+} + e^- \rightarrow Cr^{2+}$	−0.50
$IO_3^- + 6H^+ + 5e^- \rightarrow \frac{1}{2}I_2 + 3H_2O$	1.20	$Cr^{3+} + 3e^- \rightarrow Cr$	−0.73
$Br_2 + 2e^- \rightarrow 2Br^-$	1.09	$Zn^{2+} + 2e^- \rightarrow Zn$	−0.76
$VO_2^+ + 2H^+ + e^- \rightarrow VO^{2+} + H_2O$	1.00	$2H_2O + 2e^- \rightarrow H_2 + 2OH^-$	−0.83
$AuCl_4^- + 3e^- \rightarrow Au + 4Cl^-$	0.99	$Mn^{2+} + 2e^- \rightarrow Mn$	−1.18
$NO_3^- + 4H^+ + 3e^- \rightarrow NO + 2H_2O$	0.96	$Al^{3+} + 3e^- \rightarrow Al$	−1.66
$ClO_2 + e^- \rightarrow ClO_2^-$	0.954	$H_2 + 2e^- \rightarrow 2H^-$	−2.23
$2Hg^{2+} + 2e^- \rightarrow Hg_2^{2+}$	0.91	$Mg^{2+} + 2e^- \rightarrow Mg$	−2.37
$Ag^+ + e^- \rightarrow Ag$	0.80	$La^{3+} + 3e^- \rightarrow La$	−2.37
$Hg_2^{2+} + 2e^- \rightarrow 2Hg$	0.80	$Na^+ + e^- \rightarrow Na$	−2.71
$Fe^{3+} + e^- \rightarrow Fe^{2+}$	0.77	$Ca^{2+} + 2e^- \rightarrow Ca$	−2.76
$O_2 + 2H^+ + 2e^- \rightarrow H_2O_2$	0.68	$Ba^{2+} + 2e^- \rightarrow Ba$	−2.90
$MnO_4^- + e^- \rightarrow MnO_4^{2-}$	0.56	$K^+ + e^- \rightarrow K$	−2.92
$I_2 + 2e^- \rightarrow 2I^-$	0.54	$Li^+ + e^- \rightarrow Li$	−3.05
$Cu^+ + e^- \rightarrow Cu$	0.52		

Spoon Feeding

What will happen if a solution of Magnesium sulphate put into an Copper pot ?

Note : Magnesium looses electron very easily, metallic character, or basic character being high. Electrode potential is - 2.37 While we know Copper has +ve electrode potential. (This is an important exception and the data is better remembered as 0.5)

So Copper is solid in Copper pot (or even if a copper spoon was inserted, the same logic). Copper will not replace Magnesium ions. But if a Magnesium spoon was put into Copper-Sulphate solution then the Magnesium ions / atoms will start replacing the Cu ions and Copper will start getting precipitated. The Magnesium spoon, or container will start getting corroded or damaged.



A periodic table of redox behavior

Atoms with positive charge (i.e. those with more electrons than protons) are chemically oxidized relative to their elemental condition.

Cu, Rh, Pd, Ag, Ge, Te have positive electrode potential. Thus these elements want to recieve e- while Zn, Ga, Se, V, Cr, Nb, Cd, In have -ve electrode potential. These loose e-

Mg, Al, Si, and 9 elements shown below are known as Refractory Elements.

H, O, F, Cl have electronegativity between 3 to 4

B, C, P, S, As, Se, Br, Te, I, Po, At, Ru, Rh, Pd, Os, Ir, Pt, Au have electronegativity between 2 and 3

Li, Na, Mg, K, Ca, Sc, Rb, Sr, Y, Zr, Cs, Ba, La, Hf have electronegativity < 1.5 (Basically they try to loose = electropositive)

negative charge (i.e. those with more electrons than protons) are chemically reduced relative to their elemental condition.

Pt & Au have quite high electron affinity. Similarly Cu wants to recieve e-

Al, Ga, In & Sn are quite Electro Positive. They wants to loose e-

Elements that exist in nature in just one positively charged state. For example, K^+ and Ca^{2+}.

Elements that exist in nature in more than one positively charged state. For example, Mo^{6+}, Mo^{4+}, & Mo^{2+}.

Elements that exist in nature in elemental (uncharged) form and in at least one positively charged state. For example, Fe^{3+}, Fe^{2+}, & Fe.

Elements that exist in nature in states ranging from positively charged to negatively charged. For example, S^{6+} to S^{2-}.

Elements that exist in nature in elemental (uncharged) form and in at least one negatively charged state. For example, O_2 to O^{2-}.

Elements that exist in nature in just one negatively charged state. For example, F^- and Cl^-.

Elements that exist in nature in no charged state at all (the noble gases)

Elements with no redox chemistry in nature

Elements with at least some redox chemistry in nature

(and thus with multiple forms that can't be shown on a one-cell-per-element table like this one, but shown in their multiple forms on the Earth Scientist's Periodic Table of the Elements and Their Ions)

Elements with no redox chemistry in nature

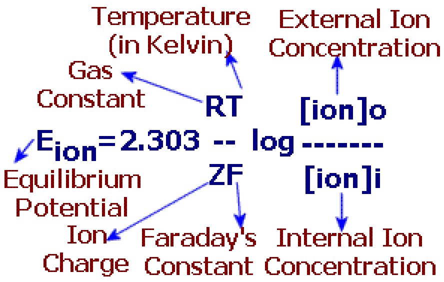
Nernst Equation also uses Electrode potential

Walther Hermann Nernst (1864 - 1941) German Physicist & Chemist received the Nobel prize in 1920 Chemistry.

$$At\ 25^0C;\ E_{cell} = E^0_{cell} - \frac{0.059}{n}\log\frac{M}{[M^{n+}]}\ or\ \frac{1}{[M^{n+}]}$$

$$E_{cell} = E^0_{cell} - \frac{0.059}{n}\log[\frac{Products}{Reactants}]or Q.$$

$$E_{cell} = E^0_{cathode} - E^0_{anode} - \frac{0.059}{n}\log\frac{[P]}{[R]}$$



Consider the following set of half-reactions

Reduction $Fe^{2+} + 2e^- \rightleftharpoons Fe$ $E^\circ_{red} = -0.44$ V

Oxidation $Zn \rightleftharpoons Zn^{2+} + 2e^-$ $E^\circ_{ox} = +0.76$ V

Overall $Zn + Fe^{2+} \rightleftharpoons Zn^{2+} + Fe$ $E^\circ_{cell} = 0.32$ V

This is valid only if all species have concentrations of 1 *M*.

Le Châtelier's principle gives us some ideas. Look at the overall reaction. Consider the reaction to be at equilibrium (E_{cell} = 0 V). If we add more Fe^{2+} to the reaction mixture, The equilibrium will shift to the right to use up the excess Fe^{2+}. This will, of course cause a positive cell potential to be measurable.

Conversely, if we removed Fe^{2+} from the reaction mixture, the equilibrium would shift to the left to try to replace some of the missing Fe^{2+}. This would result in a measurable negative potential (reaction goes left).

If the reaction were not at equilibrium then these preceding two changes would me additive to the measured voltage (say standard voltage). So, if we add Fe^{2+} to a mixture at Standard Conditions, then the voltage would be more positive and if we removed some Fe^{2+} from a mixture at Standard Conditions the measured voltage would be lower (less positive).

The reverse results would be found for addition or removal of Zn^{2+} (a product). More Zn^{2+} would give a lower positive potential and *vice versa*.

Recall the relation: $\Delta G = \Delta G^\circ + RT \ln Q$ (for non-standard conditions)

Substitute the new definition of ΔG and we get

$-nFE = -nFE^\circ + RT \ln Q$

$$E = E^\circ - \frac{RT}{nF} \ln Q$$

(This is called the *Nernst Equation*).

Thus, we can calculate the cell potential of any cell, given only the initial conditions (to calculate Q) and the reduction potentials (to calculate E°)

We can calculate the voltage for the situation where the concentrations are not standard. For example, consider $[Fe^{2+}] = 0.1$ *M* and $[Zn^{2+}] = 1.9$ *M*

For this reaction, we also know that there are two electrons transferred for each equation (n = 2).

$$E = E^{\circ} - \frac{RT}{nF} \times \ln \frac{[Zn^{2+}]}{[Fe^{2+}]}$$

$$= 0.32V - \frac{8.3145\frac{J}{mol\,K} \times 298.15K}{2 \times 96485\frac{C}{mol}} \times \ln \frac{1.9}{0.10}$$

$$= 0.28V$$

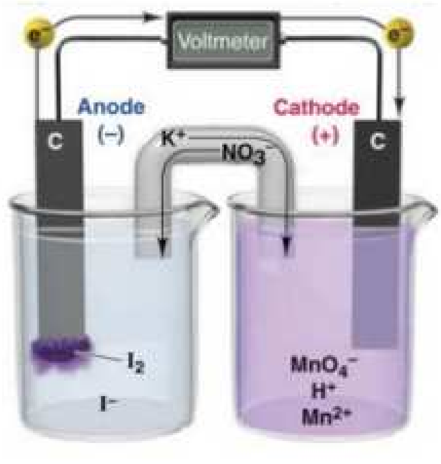
Although the reaction is 90% complete (from Standard Conditions initial) the cell potential has only dropped by a small amount (0.04 V). This is normal. For example, the batteries in your flashlight will put out almost full voltage until the last traces of chemical are almost used up at which point the voltage will drop off rather sharply. This is an especially useful characteristic of cells powering electronic equipment.

-

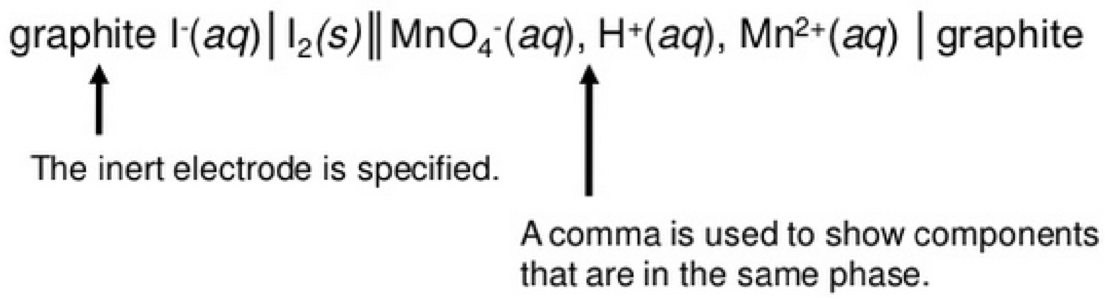
A voltaic cell using inactive electrodes.

Oxidation half-reaction
$2I^-(aq) \rightarrow I_2(s) + 2e^-$

Reduction half-reaction
$MnO_4^-(aq) + 8H^+(aq) + 5e^- \rightarrow Mn^{2+}(aq) + 4H_2O(l)$



Overall (cell) reaction
$2MnO_4^-(aq) + 16H^+(aq) + 10I^-(aq) \rightarrow 2Mn^{2+}(aq) + 5I_2(s) + 8H_2O(l)$



graphite $I^-(aq) | I_2(s) \| MnO_4^-(aq), H^+(aq), Mn^{2+}(aq) |$ graphite

The inert electrode is specified.

A comma is used to show components that are in the same phase.

A voltaic cell is made of aqueous Bromine and Zinc. Given -

$Br_2(aq) + Zn(s) \rightarrow Zn^{2+}(aq) + 2Br^{-}(aq) \qquad E_{cell} = 1.83$ V.

Calculate $E_{bromine}$, given that $E_{zinc} = -0.76$ V

Solution :

$Br_2(aq) + 2e^- \rightarrow 2Br^{-}(aq)$ [reduction; cathode]

$Zn(s) \rightarrow Zn^{2+}(aq) + 2e^-$ [oxidation; anode] $E°_{zinc} = -0.76$ V

$E_{cell} = E_{cathode} - E_{anode}$

$1.83 = E_{bromine} - (-0.76)$

$1.83 - 0.76 = E_{bromine}$ **= 1.07 Volt**

-

What is the equilibrium constant for the reaction of copper metal with bromine to form copper(II) ions and bromide ions in aqueous solution at 25°C?

Cathode	$Br_2 + 2e^- \rightleftharpoons 2\ Br^-$	$E°_{red} = 1.09$ V
Anode	$Cu \rightleftharpoons Cu^{2+} + 2e^-$	$E°_{ox} = -E°_{red} = -0.34$ V
Cell	$Cu + Br_2 \rightleftharpoons Cu^{2+} + 2\ Br^-$	$E°_{cell} = 0.75$

$$\ln K = \frac{nFE°}{RT} = \frac{2\times 96458\frac{C}{mol}\times 0.75\ V}{8.31451\frac{J}{mol\ K}\times 298.15 K} = 58.38_{208}$$

$$K = e^{58.38_{208}} = 2.2\times 10^{25}$$

-

Using half reactions write the spontaneous Redox reactions

$Sn^{2+}(aq) + 2e^- \rightarrow Sn(s) \quad E°_{tin} = -0.14$ V

$Ag^{+}(aq) + e^- \rightarrow Ag(s) \quad E°_{silver} = 0.80$ V

We have to multiply the Silver reaction with 2 to balance and match electrons

Note : This does not change the voltage obtained

$\mathbf{2}Ag^{+}(aq) + \mathbf{2}e^- \rightarrow \mathbf{2}Ag(s) \qquad E°_{silver} = 0.80$ V

Now we need to add the half reactions to eliminate the electrons, but subtract to get the Voltage correctly. **Recall** $E^{o}_{cell} = E_{Right} - E_{Left} = E_{Cathode} - E_{Anode} = (E_{Reduction\ (\ Right\ )} - E_{Reduction\ (\ Left\ )})$

$$\begin{array}{rl} Sn(s) \rightarrow Sn^{2+}(aq) + 2e^{-} & E^{\circ}_{tin} = -0.14\ V \\ 2Ag^{+}(aq) + 2e^{-} \rightarrow 2Ag(s) & E^{\circ}_{silver} = 0.80\ V \\ \hline Sn(s) + 2Ag^{+}(aq) \rightarrow 2Ag(s) + Sn^{2+}(aq) & E^{\circ}_{cell} = 0.94\ V \end{array}$$

$$E^{\circ}_{cell} = E^{\circ}_{silver} - E^{\circ}_{tin} = 0.80 - (-0.14) = 0.94\ V$$

Cathode provides electrons. Ag ions are receiving electrons at Cathode. Anode is positive where electrons are lost. See Tin is more metallic (than Silver) and is looses electrons in Anode.

It is easier to visualize the Voltage as gap in electrode potential. Silver is higher up than Hydrogen in the series. Thus has positive electrode potential. While Tin is more metallic, below Hydrogen in the series. The gap is of 0.94 Volts.

-

Zinc and Magnesium Oxide Battery

Anode (oxidation): $Zn(s) + 2OH^{-}(aq) \rightarrow ZnO(s) + H_2O(l) + 2e^{-}$

Cathode (reduction): $MnO_2(s) + 2H_2O(l) + 2e^{-} \rightarrow Mn(OH)_2(s) + 2OH^{-}(aq)$

Overall (cell) reaction:

$$Zn(s) + MnO_2(s) + H_2O(l) \rightarrow ZnO(s) + Mn(OH)_2(s) \quad E_{cell} = 1.5\ V$$

-

Zinc and Magnesium Oxide Battery

Anode (oxidation): $Zn(s) + 2OH^{-}(aq) \rightarrow ZnO(s) + H_2O(l) + 2e^{-}$

Cathode (reduction): $Ag_2O(s) + H_2O(l) + 2e^{-} \rightarrow 2Ag(s) + 2OH^{-}(aq)$

Overall (cell) reaction: $Zn(s) + Ag_2O(s) \rightarrow ZnO(s) + 2Ag(s)$

$E_{cell} = 1.6\ V$

The mercury battery uses HgO as the oxidizing agent instead of Ag_2O and has cell potential of 1.3 V.

-

The Lithium Batteries use Silver Vanadium Pentoxide as cathode

Anode (oxidation):

$$3.5Li(s) \rightarrow 3.5Li^{+} + 3.5e^{-}$$

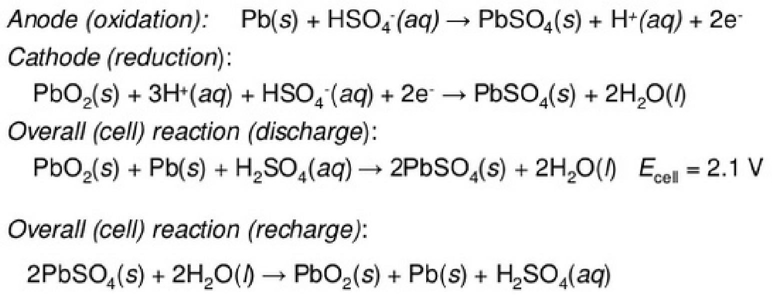
Cathode (reduction):

$$AgV_2O_{5.5} + 3.5Li^{-} + 3.5e^{-} \rightarrow Li_{3.5}V_2O_{5.5}$$

Overall (cell) reaction:

$$AgV_2O_{5.5} + 3.5Li(s) \rightarrow Li_{3.5}V_2O_{5.5}$$

The most commonly used Battery is Lead-Acid Battery



Anode (oxidation): $Pb(s) + HSO_4^-(aq) \rightarrow PbSO_4(s) + H^+(aq) + 2e^-$

Cathode (reduction):

$PbO_2(s) + 3H^+(aq) + HSO_4^-(aq) + 2e^- \rightarrow PbSO_4(s) + 2H_2O(l)$

Overall (cell) reaction (discharge):

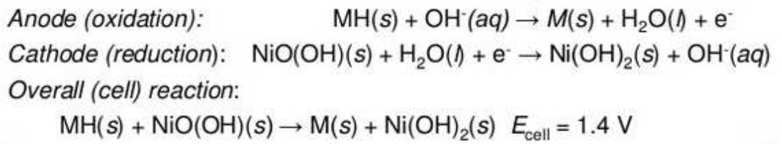
$PbO_2(s) + Pb(s) + H_2SO_4(aq) \rightarrow 2PbSO_4(s) + 2H_2O(l) \quad E_{cell} = 2.1\ V$

Overall (cell) reaction (recharge):

$2PbSO_4(s) + 2H_2O(l) \rightarrow PbO_2(s) + Pb(s) + H_2SO_4(aq)$

-

Nickel Metal Hydride Battery



Anode (oxidation): $MH(s) + OH^-(aq) \rightarrow M(s) + H_2O(l) + e^-$

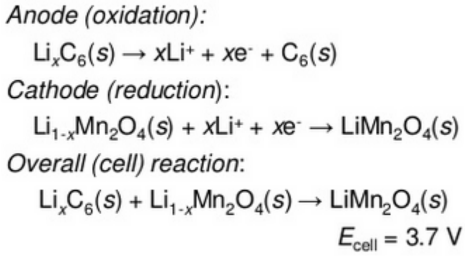
Cathode (reduction): $NiO(OH)(s) + H_2O(l) + e^- \rightarrow Ni(OH)_2(s) + OH^-(aq)$

Overall (cell) reaction:

$MH(s) + NiO(OH)(s) \rightarrow M(s) + Ni(OH)_2(s) \quad E_{cell} = 1.4\ V$

-

Laptops, Cell-phones most commonly use Lithium ion Batteries



Anode (oxidation):

$Li_xC_6(s) \rightarrow xLi^+ + xe^- + C_6(s)$

Cathode (reduction):

$Li_{1-x}Mn_2O_4(s) + xLi^+ + xe^- \rightarrow LiMn_2O_4(s)$

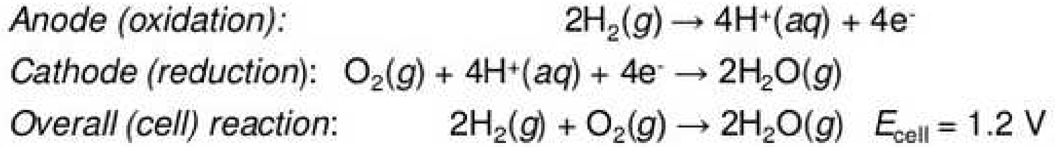
Overall (cell) reaction:

$Li_xC_6(s) + Li_{1-x}Mn_2O_4(s) \rightarrow LiMn_2O_4(s)$

$E_{cell} = 3.7\ V$

-

Hydrogen Fuel Cell



Anode (oxidation): $2H_2(g) \rightarrow 4H^+(aq) + 4e^-$

Cathode (reduction): $O_2(g) + 4H^+(aq) + 4e^- \rightarrow 2H_2O(g)$

Overall (cell) reaction: $2H_2(g) + O_2(g) \rightarrow 2H_2O(g) \quad E_{cell} = 1.2\ V$

-

What is the value of the solubility product constant for AgCl?

We could calculate this by breaking the overall K_{sp} reaction into a series of redox reactions as follows:

Anode	$Ag(s) \rightarrow Ag^+ + e^-$	$E°_{ox}$ = -0.80 V
Cathode	$AgCl + e^- \rightarrow Ag(s) + Cl^-$	$E°_{red}$ = 0.22 V
Overall **(° K_{sp} reaction)**	$AgCl \rightleftarrows Ag^+ + Cl^-$	$E°_{cell}$ = -0.58 V (not very spontaneous)

$$\ln K = \frac{nFE°}{RT} = \frac{1 \times 96485\frac{C}{mol}(-0.58V)}{8.3145\frac{J}{mol\,K} \times 298.15K} = 22.5_{744}$$

$K = 6.3 \times 10^{-9}$

-

Ion Concentrations can also be calculated...

Consider the Platinum-Hydrogen electrode coupled with a copper/copper(II) electrode.

The overall reaction is $Cu^{2+} + H_2 \rightarrow Cu + 2H^+$.

The Nernst equation is

$$E = E° - \frac{8.31451\frac{J}{mol\,K} \times 298.15\,K}{2\,mol \times 96485\frac{C}{mol}} \times \ln\left(\frac{a(H^+)^2}{a(Cu^{2+}) \times a(H_2)}\right)$$

$$E = E° - \frac{.0257\,V}{2} \times \ln\left(\frac{a(H^+)^2}{a(Cu^{2+}) \times a(H_2)}\right) \qquad \text{since}$$

$$\frac{8.31451\frac{J}{mol\,K} \times 298.15\,K}{96485\frac{C}{mol}} = .0256_{928}\,V$$

If $[Cu^{2+}]$ = 1 *M* and $P(H_2)$ = 1 bar (both their activities are 1) then

$$E = E° - \frac{0.0257\,V}{2} \ln[H^+]^2$$

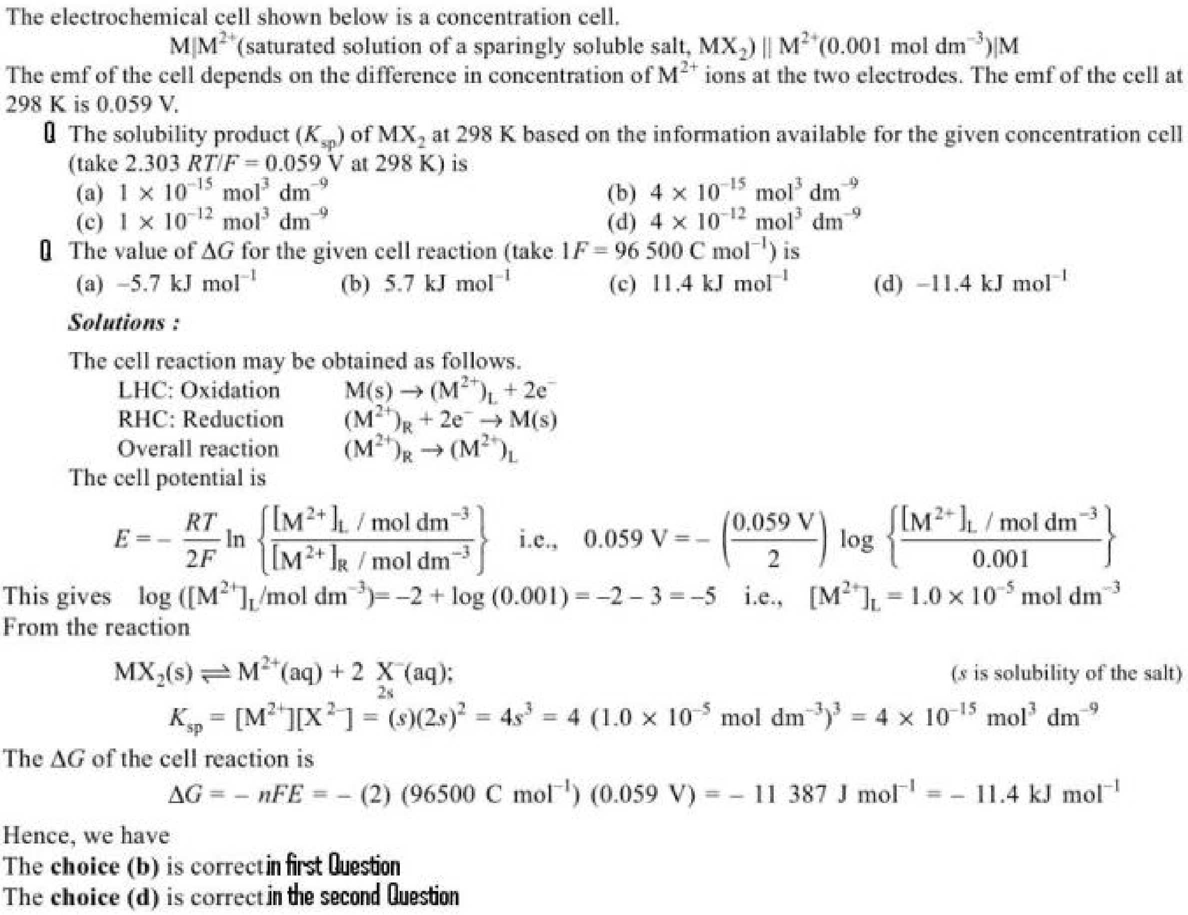
$E = E° - 0.0257\,V \ln[H^+]$ (Note: $\ln Y = 2.303 \times \log Y$ and .-257×2.303=.05917)

$E = E° + 0.05917\,V\,(-\log[H^+])$

$E = E^\circ + 0.05917\ \text{V} \times \text{pH}$.

We see here that the cell potential is a function of pH. The probes in pH meters are set up this way. A complete electrochemical cell is contained within the probe casing. All chemicals are at standard conditions and a porous glass membrane allows only H^+ ions to pass through.

Electrochemical Cell



The electrochemical cell shown below is a concentration cell.

$$M|M^{2+}(\text{saturated solution of a sparingly soluble salt, } MX_2)\ ||\ M^{2+}(0.001\ \text{mol dm}^{-3})|M$$

The emf of the cell depends on the difference in concentration of M^{2+} ions at the two electrodes. The emf of the cell at 298 K is 0.059 V.

Q The solubility product (K_{sp}) of MX_2 at 298 K based on the information available for the given concentration cell (take $2.303\ RT/F = 0.059$ V at 298 K) is

(a) $1 \times 10^{-15}\ \text{mol}^3\ \text{dm}^{-9}$ (b) $4 \times 10^{-15}\ \text{mol}^3\ \text{dm}^{-9}$
(c) $1 \times 10^{-12}\ \text{mol}^3\ \text{dm}^{-9}$ (d) $4 \times 10^{-12}\ \text{mol}^3\ \text{dm}^{-9}$

Q The value of ΔG for the given cell reaction (take $1F = 96\,500\ \text{C mol}^{-1}$) is

(a) $-5.7\ \text{kJ mol}^{-1}$ (b) $5.7\ \text{kJ mol}^{-1}$ (c) $11.4\ \text{kJ mol}^{-1}$ (d) $-11.4\ \text{kJ mol}^{-1}$

Solutions :

The cell reaction may be obtained as follows.

LHC: Oxidation $M(s) \rightarrow (M^{2+})_L + 2e^-$
RHC: Reduction $(M^{2+})_R + 2e^- \rightarrow M(s)$
Overall reaction $(M^{2+})_R \rightarrow (M^{2+})_L$

The cell potential is

$$E = -\frac{RT}{2F}\ln\left\{\frac{[M^{2+}]_L / \text{mol dm}^{-3}}{[M^{2+}]_R / \text{mol dm}^{-3}}\right\} \quad \text{i.e.,} \quad 0.059\ \text{V} = -\left(\frac{0.059\ \text{V}}{2}\right)\log\left\{\frac{[M^{2+}]_L / \text{mol dm}^{-3}}{0.001}\right\}$$

This gives $\log([M^{2+}]_L/\text{mol dm}^{-3}) = -2 + \log(0.001) = -2 - 3 = -5$ i.e., $[M^{2+}]_L = 1.0 \times 10^{-5}\ \text{mol dm}^{-3}$

From the reaction

$$MX_2(s) \rightleftharpoons M^{2+}(aq) + 2\ \underset{2s}{X^-}(aq);$$

(*s* is solubility of the salt)

$$K_{sp} = [M^{2+}][X^{2-}] = (s)(2s)^2 = 4s^3 = 4\ (1.0 \times 10^{-5}\ \text{mol dm}^{-3})^3 = 4 \times 10^{-15}\ \text{mol}^3\ \text{dm}^{-9}$$

The ΔG of the cell reaction is

$$\Delta G = -nFE = -(2)\ (96500\ \text{C mol}^{-1})\ (0.059\ \text{V}) = -11\,387\ \text{J mol}^{-1} = -11.4\ \text{kJ mol}^{-1}$$

Hence, we have

The **choice (b)** is correctin first Question
The **choice (d)** is correctin the second Question

-

Empirical Formulae in Chemistry

In chemistry, the empirical formula of a chemical compound is the simplest positive integer ratio of atoms present in a compound. A simple example of this concept is that the empirical formula of hydrogen peroxide, or H2O2, would simply be HO. Glucose (C6H12O6), ribose (C5H10O5), acetic acid (C2H4O2), and formaldehyde (CH2O) all have different molecular formulas but the same empirical formula: CH2O. This is the actual molecular formula for formaldehyde, but acetic acid has double the number of atoms, ribose has five times the number of atoms, and glucose has six times the number of atoms.

In this article we are not discussing the above " Empirical formulas ". There is another kind of " Empirical formulas " where the constants are determined by experiments, rather than derived. The dimesions on the left side or in the right sides may or may not match, case to case basis.

1) Slater's rule

In quantum chemistry, Slater's rules provide numerical values for the effective nuclear charge concept. In a many-electron atom, each electron is said to experience less than the actual nuclear charge owing to shielding or screening by the other electrons. For each electron in an atom, Slater's rules provide a value for the screening constant, denoted by s, S, or s, which relates the effective and actual nuclear charges as $Z_{effective} = Z - s$ (sigma)

The rules were devised semi-empirically by John C. Slater and published in 1930.

Revised values of screening constants based on computations of atomic structure by the Hartree-Fock method were obtained by Enrico Clementi et al in the 1960s

Steps to follow -

1.1) Write the electron configuration for the atom using the following design;
(1s)(2s,2p)(3s,3p) (3d) (4s,4p) (4d) (4f) (5s,5p)

1.2) Any electrons to the right of the electron of interest contributes no shielding. (Approximately correct statement.)
1.3) All other electrons in the same group as the electron of interest shield to an extent of 0.35 nuclear charge units
1.4) If the electron of interest is an s or p electron: All electrons with one less value of the principal quantum number shield to an extent of 0.85 units of nuclear charge. All electrons with two less values of the principal quantum number shield to an extent of 1.00 units.
1.5) If the electron of interest is an d or f electron: All electrons to the left shield to an extent of 1.00 units of nuclear charge.
1.6) Sum the shielding amounts from steps 2 through 5 and subtract from the nuclear charge value to obtain the effective nuclear charge.

Examples:
Calculate Z* for a valence electron in fluorine.
(1s2)(2s2,2p5)
Rule 2 does not apply; $0.35 \cdot 6 + 0.85 \cdot 2 = 3.8$
$Z^* = 9 - 3.8 = 5.2$ for a valence electron.

-

Calculate Z* for a 6s electron in Platinum.
(1s2)(2s2,2p6)(3s2,3p6) (3d10) (4s2,4p6) (4d10) (4f14) (5s2,5p6) (5d8) (6s2)
Rule 2 does not apply; 0.35 · 1 + 0.85 · 16 + 60 · 1.00 = 73.95
Z* = 78 – 73.95 = 4.15 for a valence electron.

Shielding

The first ionization energy for hydrogen is 1310 kJ·mol-1 while the first ionization energy for lithium is 520 kJ·mol-1. The IE for lithium is lower for two reasons-

1.7) The average distance from the nucleus for a 2s electron is greater than a 1s electron;

1.8) The 2s1 electron in lithium is repelled by the inner core electrons, so the valence electron is easily removed.

The inner core electrons shield the valence electron from the nucleus so the outer most electron only experiences an effective nuclear charge. In the case of the lithium the bulk of the 1s electron density lies between the nucleus and the 2s1 electron. So the valence electron `sees' the sum of the charges or approximately +1. In reality the charge the valence electron experiences is greater than 1 because the radial distribution show there is some probability of finding the 2s electron close to the nucleus.

Allred and Rochow’s empirical formula can closely predict the values of electronegativity from $Z_{effective}$

Allred and Rochow were two chemists who came up with the Allred-Rochow Electronegativity values by taking the electrostatic force exerted by effective nuclear charge, Z_{eff}, on the valence electron. To do so, they came up with an equation:

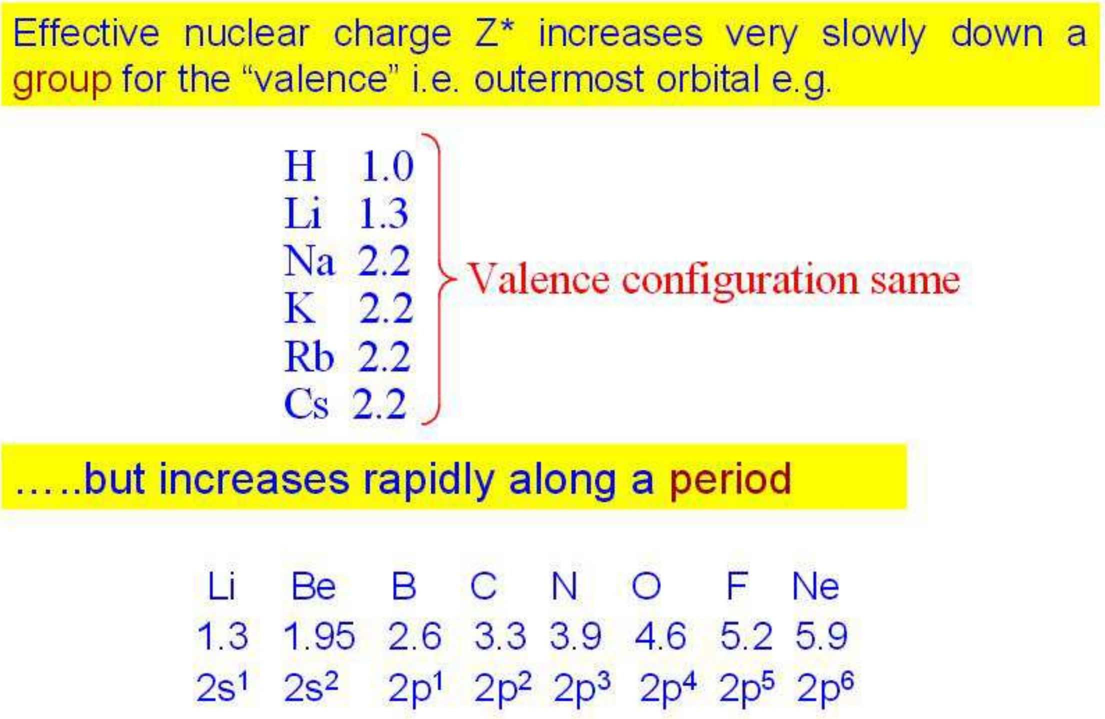
$$\chi^{AR} = \left(\frac{3590 \times Z_{eff}}{r_{cov}^{2}}\right) + 0.744 \quad \mathbf{(1)}$$

At that time, the values for the covalent radius, r_{cov} were inaccurate. Allred and Rochow added certain perimeters so that it would more closely correspond to Pauling's electronegativity scale. In the equation radius is in Angstroms.

Allred-Rochow Electronegativity Values

H **2.20**																
Li **0.97**	Be **1.47**											B **2.01**	C **2.50**	N **3.07**	O **3.50**	F **4.10**
Na **1.01**	Mg **1.23**											Al **1.47**	Si **1.74**	P **2.06**	S **2.44**	Cl **2.83**
K **0.91**	Ca **1.04**	Sc **1.20**	Ti **1.32**	V **1.45**	Cr **1.56**	Mn **1.60**	Fe **1.64**	Co **1.70**	Ni **1.75**	Cu **1.75**	Zn **1.66**	Ga **1.82**	Ge **2.02**	As **2.20**	Se **2.48**	Br **2.74**
Rb **0.89**	Sr **0.99**	Y **1.11**	Zr **1.22**	Nb **1.23**	Mo **1.30**	Te **1.36**	Ru **1.42**	Rh **1.45**	Pd **1.35**	Ag **1.42**	Cd **1.46**	In **1.49**	Sn **1.72**	Sb **1.82**	Te **2.01**	I **2.21**
Cs **0.86**	Ba **0.97**	La **1.08**	Hf **1.23**	Ta **1.33**	W **1.40**	Re **1.46**	Os **1.52**	Ir **1.55**	Pt **1.44**	Au **1.42**	Hg **1.44**	Tl **1.44**	Pb **1.55**	Bi **1.67**	Po **1.76**	At **1.90**

In this table, the electronegativities increases from left to right just like Pauling's scale because the Z is increasing. As we go down the group, it decreases because of the larger atomic size that increases the distance between the electrons and nucleus.



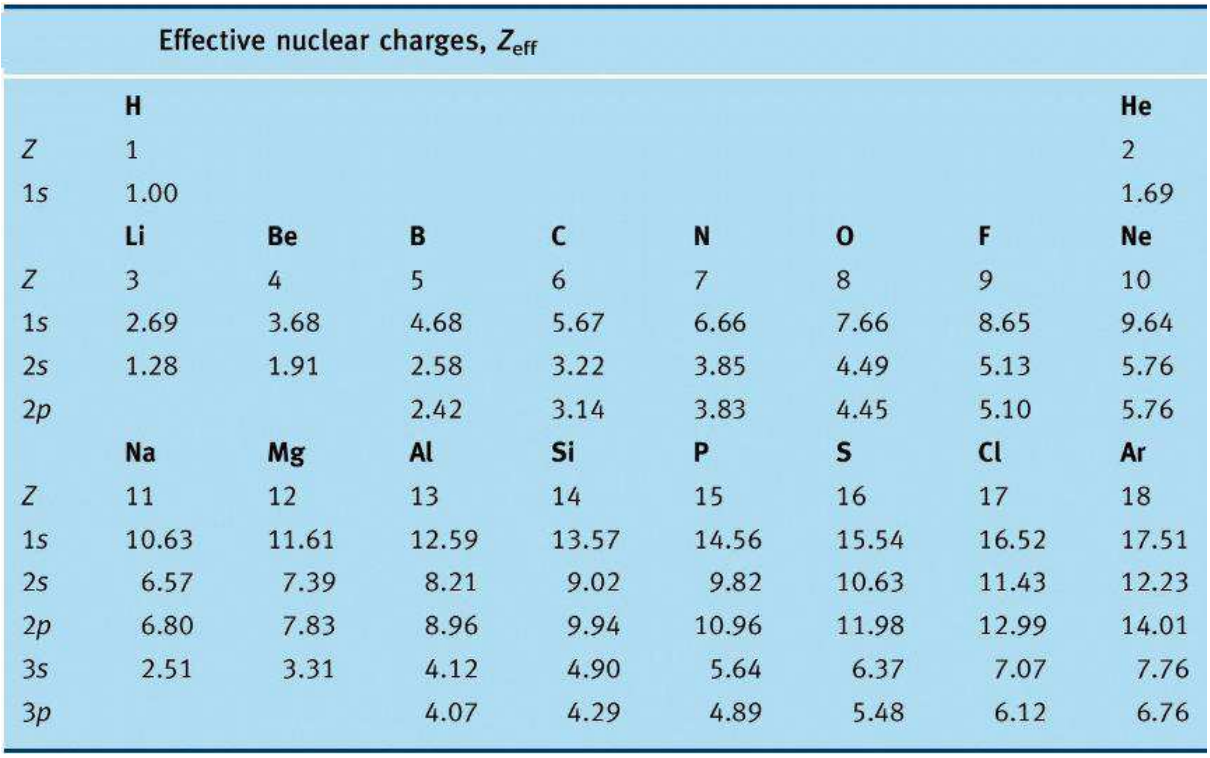
Effective nuclear charge Z* increases very slowly down a group for the “valence” i.e. outermost orbital e.g.

H	1.0
Li	1.3
Na	2.2
K	2.2
Rb	2.2
Cs	2.2

Valence configuration same

…..but increases rapidly along a period

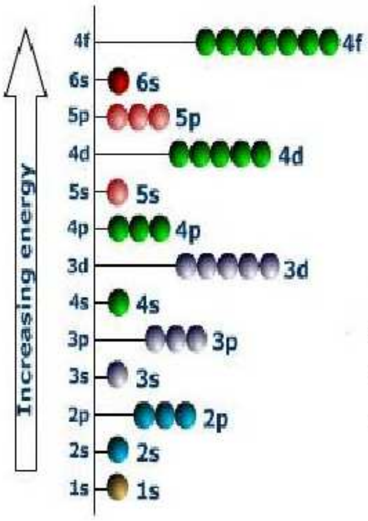
Li	Be	B	C	N	O	F	Ne
1.3	1.95	2.6	3.3	3.9	4.6	5.2	5.9
$2s^1$	$2s^2$	$2p^1$	$2p^2$	$2p^3$	$2p^4$	$2p^5$	$2p^6$

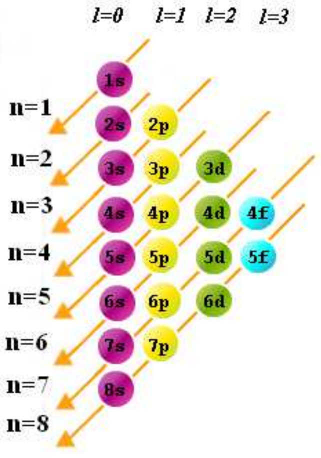


Effective nuclear charges, Z_{eff}

	H							**He**
Z	1							2
1*s*	1.00							1.69
	Li	**Be**	**B**	**C**	**N**	**O**	**F**	**Ne**
Z	3	4	5	6	7	8	9	10
1*s*	2.69	3.68	4.68	5.67	6.66	7.66	8.65	9.64
2*s*	1.28	1.91	2.58	3.22	3.85	4.49	5.13	5.76
2*p*			2.42	3.14	3.83	4.45	5.10	5.76
	Na	**Mg**	**Al**	**Si**	**P**	**S**	**Cl**	**Ar**
Z	11	12	13	14	15	16	17	18
1*s*	10.63	11.61	12.59	13.57	14.56	15.54	16.52	17.51
2*s*	6.57	7.39	8.21	9.02	9.82	10.63	11.43	12.23
2*p*	6.80	7.83	8.96	9.94	10.96	11.98	12.99	14.01
3*s*	2.51	3.31	4.12	4.90	5.64	6.37	7.07	7.76
3*p*			4.07	4.29	4.89	5.48	6.12	6.76

So using these equations we can calculate **Z**$_{effective}$ from electron configurations. We know electron configuration from Afbau principle. Next we can know the electronegativity values of the elements. The reactions are decided by these values. Also there are equations to calculate Ionizations potential values from these.





2) Moseley's Law

This law relates to the frequency of the spectral lines of the characteristic X-radiation of a chemical element to its atomic number. This law was experimentally established by Henry G. J. Moseley in 1913. According to Moseley's law, the square root of the frequency v of a spectral line of the characteristic radiation of an element is a linear function of its atomic number Z:

$$\sqrt{f} = k_1 \cdot (Z - k_2)$$

Or another form

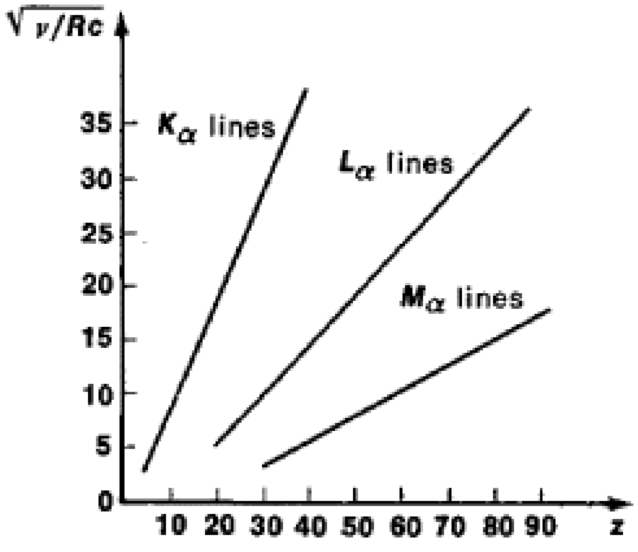
$$\sqrt{\frac{\nu}{R}} = \frac{Z - S_n}{n}$$

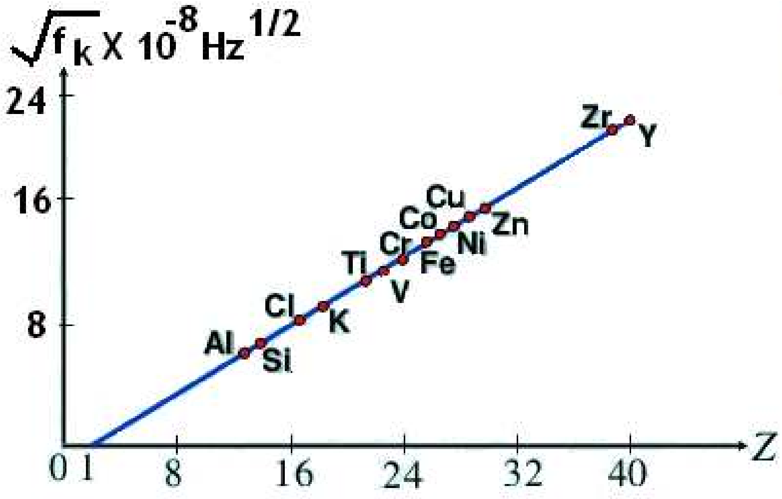
where R is the Rydberg constant, S_n is the screening constant, and n is the principal quantum number. On a Moseley plot (see Figure 1), the dependence of $\sqrt{\nu}$ on Z is a series of lines (such as the K_α lines, L_α lines, and M_α lines, which correspond to the values n = 1,2, and 3).

Moseley's law was incontrovertible proof of the correctness of the arrangement of the elements in D. I. Mendeleev's periodic system of the elements and the law helped to clarify the physical significance of *Z*.

According to Moseley's law, the characteristic X-ray spectra do not display the periodic regularities that are inherent in optical spectra. This indicates that the inner electron shells of the atoms of all elements, which are manifested in the characteristic X-ray spectra, have an analogous structure.

Subsequent experiments revealed some deviations from a linear





3) Trouton's law

Trouton’s rule states that the entropy of vaporization is almost the same value, about 85-88 J K-1 mol-1, for various kinds of liquids at their boiling points.The entropy of vaporization is defined as the ratio between the enthalpy of vaporization and the boiling temperature. It is named after Frederick Thomas Trouton.

It also can be expressed as Delta S = 10.5 R

Or better approximation will be S = 4.5 R + R Ln T where T is Temperature in Kelvin. if T = 400 K, we find the original formulation for Trouton's rule.

Trouton’s rule is valid for many liquids; for instance, the entropy of vaporization of toluene is 87.30 J K-1 mol-1, that of benzene is 89.45 J K-1 mol-1, and that of chloroform is 87.92 J K-1 mol-1. Because of its convenience, the rule is used to estimate the enthalpy of vaporization of liquids whose boiling points are known.

The rule, however, has some exceptions. For example, the entropies of vaporization of water, ethanol, and formic acid are far from the predicted values. The entropy of vaporization of XeF6 at its boiling point has the extraordinarily high value of 136.9 J K-1 mol-1. The characteristic of those liquids to which Trouton’s rule cannot be applied is their special interaction between molecules such as hydrogen bonding.

The entropy of vaporization of water and ethanol shows positive deviance from the rule; this is because the hydrogen bonding in the liquid phase lessens the entropy of the phase. In contrast, the entropy of vaporization of formic acid has negative deviance. This fact indicates the existence of an orderly structure in the gas phase; it is known that formic acid forms a dimer structure even in the gas phase. Negative deviance can also occur as a result of a small gas phase entropy owing to a low population of excited rotational states in the gas phase, particularly in small molecules such as methane - a small moment of inertia I giving rise to a large rotational constant B, with correspondingly widely separated rotational energy levels and, via Maxwell-Boltzmann distribution, a small population of excited rotational states and hence a low rotational entropy.

4) Einstein-Debey equation (Dulong & Petit)

Dulong and Petit gave an Emperical Law for molar specific heat of Solids. The Dulong–Petit law, a thermodynamic rule proposed in 1819 by French physicists Pierre Louis Dulong and Alexis Thérèse Petit, states the classical expression for the molar specific heat capacity of a crystal. Experimentally the two scientists had found that the heat capacity per weight (the mass-specific heat capacity) for a number of substances became close to a constant value, after it had been multiplied by number-ratio representing the presumed relative atomic weight of the substance. These atomic weights had shortly before been suggested by Dalton.

In modern terms, Dulong and Petit found that the heat capacity of a mole of many solid substances is about 3R, where R is the modern constant called the universal gas constant. Dulong and Petit were unaware of the relationship with R, since this constant had not yet been defined from the later kinetic theory of gases. The value of 3R is about 25 joules per kelvin (Close to 6 Calories per Kelvin), and Dulong and Petit essentially found that this was the heat capacity of crystals, per mole of atoms they contained.

The modern theory of the heat capacity of solids states that it is due to lattice vibrations in the solid, and was first derived in crude form from this assumption by Albert Einstein, in 1907. The Einstein solid model thus gave for the first time a reason why the Dulong–Petit law should be stated in terms of the classical heat capacities for gases.

Einstein's oscillator treatment of specific heat gave qualitative agreement with experiment and gave the correct high temperature limit (the Law of Dulong and Petit). The quantitative fit to experiment was improved by Debye's recognition that there was a maximum number of modes of vibration in a solid. He pictured the vibrations as standing wave modes in the crystal, similar to the electromagnetic modes in a cavity which successfully explained blackbody radiation. The density of states for these modes, which are called "phonons", is of the same form as the photon density of states in a cavity.

In thermodynamics and solid state physics, the Debye model is a method developed by Peter Debye in 1912 for estimating the phonon contribution to the specific heat (heat capacity) in a solid.[1] It treats the vibrations of the atomic lattice (heat) as phonons in a box, in contrast to the Einstein model, which treats the solid as many individual, non-interacting quantum harmonic oscillators. The Debye model correctly predicts the low temperature dependence of the heat capacity, which is proportional to T3 (T Cube)

5) Reynolds number

In fluid mechanics, the Reynolds number (Re) is a dimensionless quantity that is used to help predict similar flow patterns in different fluid flow situations. The concept was introduced by George Gabriel Stokes in 1851, but the Reynolds number is named after Osborne Reynolds (1842–1912), who popularized its use in 1883

The Reynolds number is defined as the ratio of inertial forces to viscous forces and consequently quantifies the relative importance of these two types of forces for given flow conditions

Reynolds numbers frequently arise when performing scaling of fluid dynamics problems, and as such can be used to determine dynamic similitude between two different cases of fluid flow. They are also used to characterize different flow regimes within a similar fluid, such as laminar or turbulent flow:

laminar flow occurs at low Reynolds numbers, where viscous forces are dominant, and is characterized by smooth, constant fluid motion;

turbulent flow occurs at high Reynolds numbers and is dominated by inertial forces, which tend to produce chaotic eddies, vortices and other flow instabilities.

In practice, matching the Reynolds number is not on its own sufficient to guarantee similitude. Fluid flow is generally chaotic, and very small changes to shape and surface roughness can result in very different flows. Nevertheless, Reynolds numbers are a very important guide and are widely used.

R = Inertial Forces / Viscous Forces = Rho v L by Mu ($\rho vL/\mu$)

Inertial Force = Rho (V square) (L square)

Viscous Force = Mu V L

where

Rho = is the density of the fluid (kg/m^3).

v = is the kinematic viscosity mu/rho (m^2/s)

L = is a characteristic linear dimension, (travelled length of the fluid; hydraulic diameter when dealing with river systems) (m)

Mu(μ) = is the dynamic viscosity of the fluid (Pa·s or $N \cdot s/m^2$ or kg/(m·s))

6) Raoult's law

Raoult's law states that the vapor pressure of a solvent above a solution is equal to the vapor pressure of the pure solvent at the same temperature scaled by the mole fraction of the solvent present

In the 1880s (1882), French chemist François-Marie Raoult discovered that when a substance is dissolved in a solution, the vapor pressure of the solution will generally decrease. This observation depends on two variables:

The mole fraction of the amount of dissolved solute present and the original vapor pressure (pure solvent).

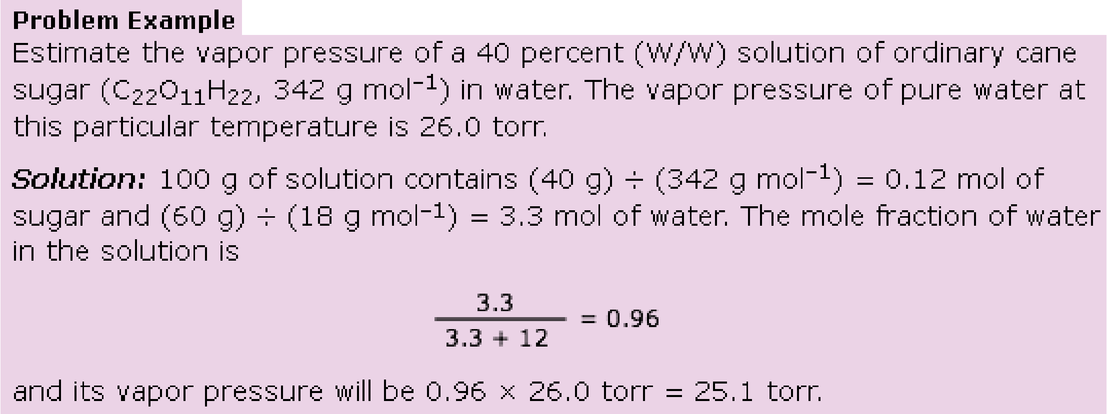
At any given temperature for a particular solid or liquid, there is a pressure at which the vapor formed above the substance is in dynamic equilibrium with its liquid or solid form. This is the vapor pressure of the substance at that temperature. At equilibrium, the rate at which the solid or liquid evaporates is equal to the rate that the gas is condensing back to its original form. All solids and liquids have a vapor pressure, and this pressure is constant regardless of how much of the substance is present.

If you look review the concepts of colligative properties, you will find that adding a solute lowers vapor pressure because the additional solute particles will fill the gaps between the solvent particles and take up space. This means less of the solvent will be on the surface and less will be able to break free to enter the gas phase, resulting in a lower vapor pressure.

In mixtures showing a positive deviation from Raoult's Law, the vapour pressure of the mixture is always higher than you would expect from an ideal mixture.

Also you can have mixtures with vapour pressures which are less than would be expected by Raoult's Law. In some cases, the deviations are small, but in others they are much greater giving a minimum value for vapour pressure lower than that of either pure component.

Let Take a binary solution with have components A and B. If the force of attraction between molecular of and A and B in the solution are stronger than that of between A — A and B — B, then the tendency of escaping of molecules A—B from the solution becomes less than that of pure liquids. The total pressure of the solution will be lower than the corresponding vapour pressure of ideal solution of the same component A and B. This type of solution shows negative deviation from Raoult's law. The negative deviation of solution have been shown in figure. Some energy released when we mix both solutions. This reaction is exothermic reaction. For Exothermic reaction delta H is always negative.



Problem Example

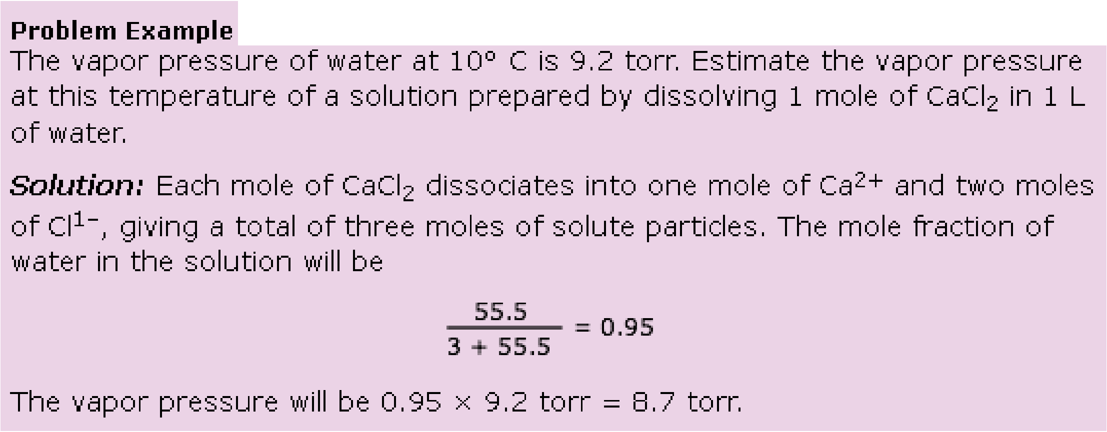
Estimate the vapor pressure of a 40 percent (W/W) solution of ordinary cane sugar ($C_{22}O_{11}H_{22}$, 342 g mol^{-1}) in water. The vapor pressure of pure water at this particular temperature is 26.0 torr.

Solution: 100 g of solution contains (40 g) ÷ (342 g mol^{-1}) = 0.12 mol of sugar and (60 g) ÷ (18 g mol^{-1}) = 3.3 mol of water. The mole fraction of water in the solution is

$$\frac{3.3}{3.3 + 12} = 0.96$$

and its vapor pressure will be 0.96 × 26.0 torr = 25.1 torr.

Since the sum of all mole fractions in a mixture must be unity, it follows that the more moles of solute, the smaller will be the mole fraction of the solvent. Also, if the solute is a salt that dissociates into ions, then the proportion of solvent molecules will be even smaller.



Problem Example

The vapor pressure of water at 10° C is 9.2 torr. Estimate the vapor pressure at this temperature of a solution prepared by dissolving 1 mole of $CaCl_2$ in 1 L of water.

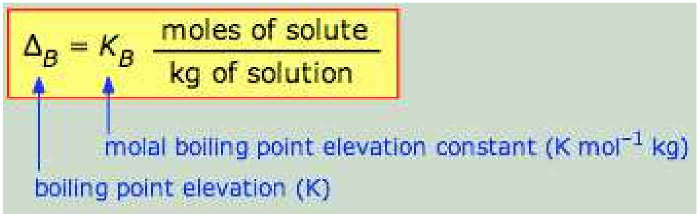
Solution: Each mole of $CaCl_2$ dissociates into one mole of Ca^{2+} and two moles of Cl^{1-}, giving a total of three moles of solute particles. The mole fraction of water in the solution will be

$$\frac{55.5}{3 + 55.5} = 0.95$$

The vapor pressure will be 0.95 × 9.2 torr = 8.7 torr.

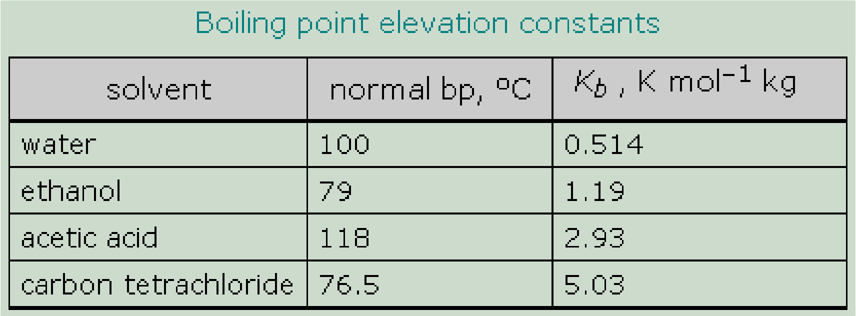
If addition of a nonvolatile solute lowers the vapor pressure of the solution, then it follows that the temperature must be raised to restore the vapor pressure to the value corresponding to the pure solvent. In particular, the temperature at which the vapor pressure is 1 atm will be higher than the normal boiling point by an amount known as the *boiling point elevation*.

The exact relation between the boiling point of the solution and the mole fraction of the solvent is rather complicated, but for dilute solutions the elevation of the boiling point is directly proportional to the *molal* concentration of the solute:



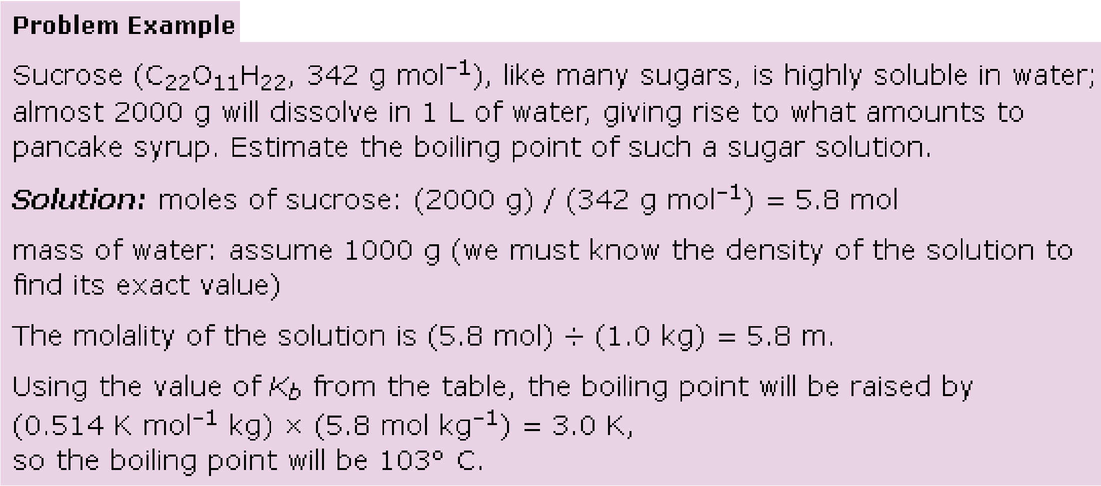
$$\Delta_B = K_B \frac{\text{moles of solute}}{\text{kg of solution}}$$

The proportionality constant K_B is a property of the ***solvent*** because this is the only component that contributes to the vapor pressure in the model we are considering in this section.



Boiling point elevation constants

solvent	normal bp, °C	K_b , K mol^{-1} kg
water	100	0.514
ethanol	79	1.19
acetic acid	118	2.93
carbon tetrachloride	76.5	5.03



Problem Example

Sucrose ($C_{22}O_{11}H_{22}$, 342 g mol^{-1}), like many sugars, is highly soluble in water; almost 2000 g will dissolve in 1 L of water, giving rise to what amounts to pancake syrup. Estimate the boiling point of such a sugar solution.

Solution: moles of sucrose: (2000 g) / (342 g mol^{-1}) = 5.8 mol

mass of water: assume 1000 g (we must know the density of the solution to find its exact value)

The molality of the solution is (5.8 mol) ÷ (1.0 kg) = 5.8 m.

Using the value of K_b from the table, the boiling point will be raised by
(0.514 K mol^{-1} kg) × (5.8 mol kg^{-1}) = 3.0 K,
so the boiling point will be 103° C.

The freezing point of a substance is the temperature at which the solid and liquid forms can coexist indefinitely — that is, they are in equilibrium. Under these conditions molecules pass between the two phases at equal rates because their escaping tendencies from the two phases are identical.

As with boiling point elevation, in dilute solutions there is a simple linear relation between the freezing point depression and the molality of the solute:

$$\Delta T_F = K_F \times \frac{\text{moles of solute}}{\text{kg of solvent}}$$

Freezing point depression constants

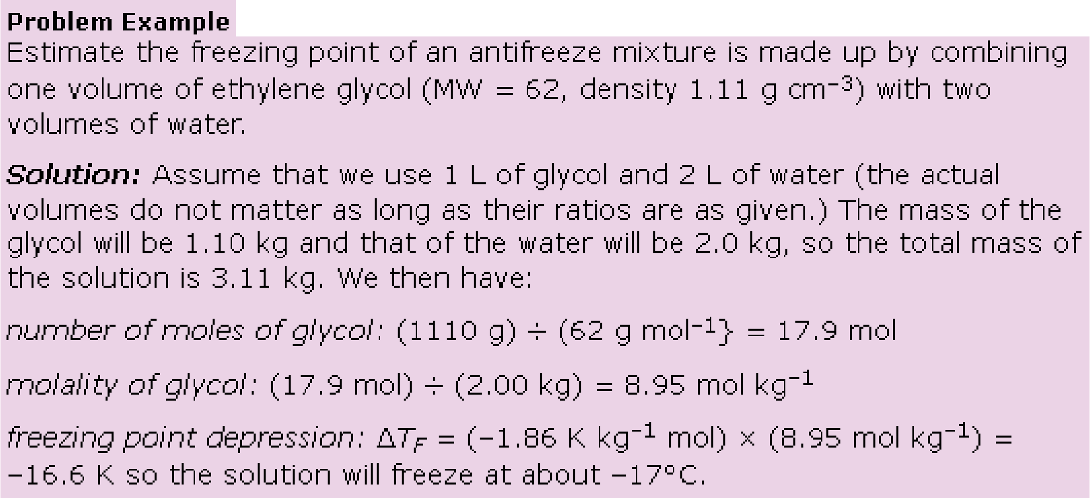
solvent	**normal fp , °C**	**K_f , K mol^{-1} kg**
water	0.0	-1.86
acetic acid	16.7	-3.90
benzene	5.5	-5.10
camphor	180	-40.0
cyclohexane	6.5	-20.2
phenol	40	-7.3

The use of salt to de-ice roads is a common application of this principle. The solution formed when some of the salt dissolves in the moist ice reduces the freezing point of the ice. If the freezing point falls below the ambient temperature, the ice melts. In very cold weather, the ambient temperature may be below that of the salt solution, and the salt will have no effect.

The effectiveness of a de-icing salt depends on the number of particles it releases on dissociation and on its solubility in water:

name	**formula**	**lowest practical T, °C**
ammonium sulfate	$(NH_4)_2SO_4$	-7
calcium chloride	$CaCl_2$	-29
potassium chloride	KCl	-15
sodium chloride	NaCl	-9
urea	$(NH_2)_2CO$	-7

Automotive radiator antifreezes are mostly based on ethylene glycol, $(CH_2OH)_2$. Owing to the strong hydrogen-bonding properties of this double alcohol, this substance is miscible with water in all proportions, and contributes only a very small vapor pressure of its own. Besides lowering the freezing point, antifreeze also raises the boiling point, increasing the operating range of the cooling system. The pure glycol freezes at -12.9°C and boils at 197°C, allowing water-glycol mixtures to be tailored to a wide range of conditions.



Problem Example

Estimate the freezing point of an antifreeze mixture is made up by combining one volume of ethylene glycol (MW = 62, density 1.11 g cm^{-3}) with two volumes of water.

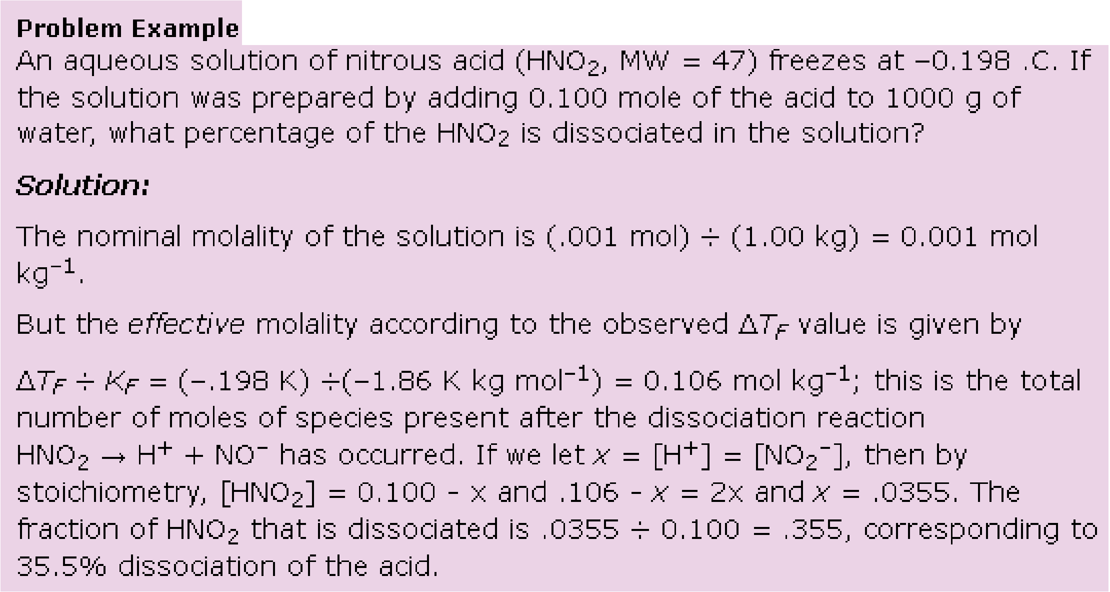
Solution: Assume that we use 1 L of glycol and 2 L of water (the actual volumes do not matter as long as their ratios are as given.) The mass of the glycol will be 1.10 kg and that of the water will be 2.0 kg, so the total mass of the solution is 3.11 kg. We then have:

number of moles of glycol: (1110 g) ÷ (62 g mol^{-1}} = 17.9 mol

molality of glycol: (17.9 mol) ÷ (2.00 kg) = 8.95 mol kg^{-1}

freezing point depression: ΔT_F = (−1.86 K kg^{-1} mol) × (8.95 mol kg^{-1}) = −16.6 K so the solution will freeze at about −17°C.

Any ionic species formed by dissociation will also contribute to the freezing point depression. This can serve as a useful means of determining the fraction of a solute that is dissociated.



Problem Example

An aqueous solution of nitrous acid (HNO_2, MW = 47) freezes at -0.198 .C. If the solution was prepared by adding 0.100 mole of the acid to 1000 g of water, what percentage of the HNO_2 is dissociated in the solution?

Solution:

The nominal molality of the solution is (.001 mol) ÷ (1.00 kg) = 0.001 mol kg^{-1}.

But the *effective* molality according to the observed ΔT_F value is given by

$\Delta T_F \div K_F = (-.198 \text{ K}) \div (-1.86 \text{ K kg mol}^{-1}) = 0.106 \text{ mol kg}^{-1}$; this is the total number of moles of species present after the dissociation reaction $HNO_2 \rightarrow H^+ + NO^-$ has occurred. If we let $x = [H^+] = [NO_2^-]$, then by stoichiometry, $[HNO_2] = 0.100 - x$ and $.106 - x = 2x$ and $x = .0355$. The fraction of HNO_2 that is dissociated is .0355 ÷ 0.100 = .355, corresponding to 35.5% dissociation of the acid.

7) Variation of viscosity with temperature

With an increase in temperature, there is typically an increase in the molecular interchange as molecules move faster in higher temperatures.

The gas viscosity will increase with temperature. According to the kinetic theory of gases, viscosity should be proportional to the square root of the absolute temperature, in practice, it increases more rapidly.

In a liquid there will be molecular interchange similar to those developed in a gas, but there are additional substantial attractive, cohesive forces between the molecules of a liquid (which are much closer together than those of a gas). Both cohesion and molecular interchange contribute to liquid viscosity.

The impact of increasing the temperature of a liquid is to reduce the cohesive forces while simultaneously increasing the rate of molecular interchange.

The former effect causes a decrease in the shear stress while the latter causes it to increase. The result is that liquids show a reduction in viscosity with increasing temperature. With high temperatures, viscosity increases in gases and decreases in liquids, the drag force will do the same.

The impact of increasing temperature will be to slow down the sphere in gases and to accelerate it in liquids. When you consider a liquid at room temperature, the molecules are tightly bound together by attractive inter-molecular forces (e.g. Van der Waal forces).

It is these attractive forces that are responsible for the viscosity since it is difficult for individual molecules to move because they are tightly bound to their neighbors.

The increase in temperature causes the kinetic or thermal energy to increase and the molecules become more mobile.

The attractive binding energy is reduced and therefore the viscosity is reduced. If you continue to heat the liquid the kinetic energy will exceed the binding energy and molecules will escape from the liquid and it can become a vapor.

So the temperature dependence of liquid viscosity is the phenomenon by which liquid viscosity tends to decrease (or, alternatively, its fluidity tends to increase) as its temperature increases. This can be observed, for example, by watching how cooking oil appears to move more fluidly upon a frying pan after being heated by a stove.

Exponential Model

An exponential model for the temperature-dependence of shear viscosity (μ) was first proposed by Reynolds in 1886.

$$\mu(T) = \mu_0 \exp(-bT)$$

where T is temperature and $\mu_0$and $b$are coefficients. This is an empirical model that usually works for a limited range of temperatures.

Arrhenius model

The model is based on the assumption that the fluid flow obeys the Arrhenius equation for molecular kinetics:

$$\mu(T) = \mu_0 \exp\left(\frac{E}{RT}\right)$$

where T is temperature, $\mu_0$is a coefficient, E is the activation energy and R is the universal gas constant. A first-order fluid is another name for a power-law fluid with exponential dependence of viscosity on temperature.

Williams-Landel-Ferry model

The **Williams-Landel-Ferry** model, or **WLF** for short, is usually used for polymer melts or other fluids that have a glass transition temperature.

The model is:

$$\mu(T) = \mu_0 \exp\left(\frac{-C_1(T-T_r)}{C_2+T-T_r}\right)$$

where T-temperature, C_1, C_2, $T_r$and $\mu_0$are empiric parameters (only three of them are independent from each other).

If one selects the parameter T_r based on the glass transition temperature, then the parameters C_1, C_2 become very similar for the wide class of polymers. Typically, if T_r is set to match the glass transition temperature T_g, we get

$$C_1 \approx 17.44$$

and

$$C_2 \approx 51.6\ \text{K}.$$

Van Krevelen recommends to choose

$T_r = T_g + 43$ K, then

$$C_1 \approx 8.86$$

and

$$C_2 \approx 101.6\ \text{K}.$$

Using such *universal parameters* allows one to guess the temperature dependence of a polymer by knowing the viscosity at a single temperature.

In reality the *universal parameters* are not that universal, and it is much better to fit the **WLF** parameters from the experimental data.

Masuko and Magill model

The model is usually used for **polymer melts** or other fluids that have a **glass transition temperature** as well as the **WLF model**. Ordinarily, The WLF model is limited to the temperature interval between Tg and Tg+ 100 K, But this model can be applied to more wide temperature range.

The model is:

$$\log(\eta/\eta_g) = A\left[\exp\left\{\frac{B(T_g - T)}{T}\right\} - 1\right]$$

The **A** and **B** are empirical parameters that does not depend on the materials. The average values are:

A = 14.25 to 16.24,

and

B = 5.34 to 7.60.

Viscosity of water

Viscosity of water equation accurate to within 2.5% from 0 °C to 370 °C:

$$\mu(T) = 2.414 \times 10^{-5} \times 10^{247.8/(T-140)}$$

where *T* has units of Kelvin, and *μ* has units of Kg/m.s.

Models for kinematic viscosity

The effect of temperature on the kinematic viscosity (*ν*) has also been described by a number of empirical equations.

Walther formula

The Walther formula is typically written in the form

$$\log_{10}[\log_{10}(\nu + \lambda)] = A - B\ \log_{10}(T)$$

where *λ* is a shift constant, and *A*, *B* are empirical parameters.

Wright model

The Wright model has the form

$$\log_{10}[\log_{10}[\nu + \lambda + f(\nu)]] = A - B\ \log_{10}(T)$$

where an addition function *f(ν)*, often a polynomial fit to experimental data, has been added to the Walther formula.

Seeton model

The Seeton model is based on curve fitting the viscosity dependence of many liquids (refrigerants, hydrocarbons and lubricants) versus temperature and applies over a large temperature and viscosity range:

$$\ln\left(\ln\left(\nu + 0.7 + e^{-\nu} K_0\left(\nu + 1.244067\right)\right)\right) = A - B * \ln\left(T\right)$$

where *T* is absolute temperature in kelvins, ν is the kinematic viscosity in centistokes, K_0 is the zero order modified Bessel function of the second kind, and *A* and *B* are liquid specific values. This form should not be applied to ammonia or water viscosity over a large temperature range.

For liquid metal viscosity as a function of temperature, Seeton proposed:

$$\ln\left(\ln\left(\nu + 0.7 + e^{-\nu} K_0\left(\nu + 1.244067\right)\right)\right) = A - \frac{B}{T}$$

8) Variation of surface tension with temperature

Surface tension is dependent on temperature. For that reason, when a value is given for the surface tension of an interface, temperature must be explicitly stated. The general trend is that surface tension decreases with the increase of temperature, reaching a value of 0 at the critical temperature. For further details see the Eötvös rule below. There are only empirical equations to relate surface tension and temperature:

- Eötvös:

$$\gamma V^{2/3} = k(T_C - T)$$

Here *V* is the molar volume of a substance, T_C is the critical temperature and *k* is a constant valid for almost all substances. A typical value is $k = 2.1 \times 10^{-7}$ [J K^{-1} $mol^{-2/3}$]. For water one can further use V = 18 ml/mol and T_C = 647 K (374 °C).

A variant on Eötvös is described by Ramay and Shields:

$$\gamma V^{2/3} = k\left(T_C - T - 6\right)$$

where the temperature offset of 6 kelvins provides the formula with a better fit to reality at lower temperatures.

- Guggenheim-Katayama:

$$\gamma = \gamma^o \left(1 - \frac{T}{T_C}\right)^n$$

$\gamma^o$is a constant for each liquid and n is an empirical factor, whose value is 11/9 for organic liquids. This equation was also proposed by van der Waals, who further proposed that γ^0could be given by the expression, $K_2 T_c^{\frac{1}{3}} P_c^{\frac{2}{3}}$, where $K_2$is a universal constant for all liquids, and $P_c$is the critical pressure of the liquid (although later experiments found $K_2$to vary to some degree from one liquid to another).

Both Guggenheim-Katayama and Eötvös take into account the fact that surface tension reaches 0 at the critical temperature, whereas Ramay and Shields fails to match reality at this endpoint.

9) Debye-Huckel-Onsager theory

A theory providing quantitative results for the conductivity of ions in dilute solutions of strong electrolytes, which enables the Kohlrausch equation to be derived. This theory can be stated as: K = $A + B\Lambda^0_m$, where Λ^0_m is the limiting molar conductivity.$A = z^2eF^2/3\pi\eta(2/\varepsilon RT)^{½}$,$B = \mathbf{q}\mathbf{z}^3eF/24\pi\varepsilon RT(2/\varepsilon RT)^{½}$,where *z* is the charge of an ion, *e* is the charge of an electron, *F* is Faraday's constant, η is the viscosity of the liquid, *R* is the gas constant, *T* is the thermodynamic temperature, and *q* = 0.586 in the case of a 1,1 electrolyte. The Debye-Hückel-Onsager theory uses the same assumptions and approximations as the Debye-Hückel theory and is also limited to very dilute solutions (usually less than 10^{-3} M) for which there is good agreement between theory and experiment. The modifications were made by the Norwegian-born US chemist Lars Onsager (1903–76)

$A = z^2eF^2/3\pi\eta(2/\varepsilon RT)^{½}$,$B = \mathbf{q}\mathbf{z}^3eF/24\pi\varepsilon RT(2/\varepsilon RT)^{½}$

10) Liquid drop model of Nucleus

In nuclear physics, description of atomic nuclei formulated (1936) by Niels Bohr and used (1939) by him and John A. Wheeler to explain nuclear fission. According to the model, the nucleons (neutrons and protons) behave like the molecules in a drop of liquid. If given sufficient extra energy (as by the absorption of a neutron), the spherical nucleus may be distorted into a dumbbell shape and then split at the neck into two nearly equal fragments, releasing energy. Although inadequate to explain all nuclear phenomena, the theory underlying the model provides excellent estimates of average properties of nuclei.

The semi-empirical mass formula (SEMF) (sometimes also called Weizsäcker's formula, or the Bethe-Weizsäcker formula, or the Bethe-Weizsäcker mass formula to distinguish it from the Bethe-Weizsäcker process) is used to approximate the mass and various other properties of an atomic nucleus from its number of protons and neutrons. As the name suggests, it is based partly on theory and partly on empirical measurements. The theory is based on the liquid drop model proposed by George Gamow, which can account for most of the terms in the formula and gives rough estimates for the values of the coefficients. It was first formulated in 1935 by German physicist Carl Friedrich von Weizsäcker, and although refinements have been made to the coefficients over the years, the structure of the formula remains the same today.

The SEMF gives a good approximation for atomic masses and several other effects, but does not explain the appearance of magic numbers of protons and neutrons, and the extra binding-energy and measure of stability that are associated with these numbers of nucleons.

Nuclear Shell Model

Maria Goeppert Mayer, who made important discoveries about nuclear structure, is one of only two women to have won the Nobel Prize in physics.

In August 1948, Goeppert Mayer published her first paper detailing the evidence for the nuclear shell model, which accounts for many properties of atomic nuclei.

During her time at Chicago and Argonne in the late 1940s, Goeppert Mayer developed a mathematical model for the structure of nuclear shells, which she published in 1950. Her model explained why certain numbers of nucleons in an atomic nucleus result in particularly stable configurations. These numbers are what Eugene Wigner called magic numbers: 2, 8, 20, 28, 50, 82, and 126. Enrico Fermi provided a critical insight by asking her: "Is there any indication of spin orbit coupling?" She realised that this was indeed the case, and postulated that the nucleus is a series of closed shells and pairs of neutrons and protons tend to couple together. She described the idea as follows:

Think of a room full of waltzers.

Suppose they go round the room in circles, each circle enclosed within another. Then imagine that in each circle, you can fit twice as many dancers by having one pair go clockwise and another pair go counterclockwise. Then add one more variation; all the dancers are spinning twirling round and round like tops as they circle the room, each pair both twirling and circling. But only some of those that go counterclockwise are twirling counterclockwise. The others are twirling clockwise while circling counterclockwise. The same is true of those that are dancing around clockwise: some twirl clockwise, others twirl counterclockwise.

Three German scientists, Otto Haxel, J. Hans D. Jensen, and Hans Suess, were also working on solving the same problem, and arrived at the same conclusion independently. Their results were announced in the issue of the Physical Review before Goeppert Mayer 's announcement in June 1949. Afterwards, she collaborated with them. Hans Jensen co-authored a book with Goeppert Mayer in 1950 titled Elementary Theory of Nuclear Shell Structure. In 1963, Goeppert Mayer, Jensen, and Wigner shared the Nobel Prize for Physics "for their discoveries concerning nuclear shell structure." She was the second and most recent female Nobel laureate in physics, after Marie Curie.

The shell model of the nucleus presumes that a given nucleon moves in an effective attractive potential formed by all the other nucleons. If that is true, then the potential is probably roughly proportional to the nuclear density and therefore could be expressed in the form

$$V = \frac{-V_0}{1 + \exp\left(\frac{r-R}{a}\right)}$$

The parameters in this model of the potential have been evaluated to be approximately evaluated:

$$\begin{aligned} &V_0 \approx 57\,MeV + corrections \\ &R \approx 1.25A^{1/3}\,fermi \\ &a \approx 0.65\,fermi \end{aligned}$$

Note that the radius above is larger than that given by the nuclear radius formula since it is related to the nuclear force which extends beyond the radius. Two other corrections are typically applied to more nearly fit observations. The first is called the symmetry energy, arising when there is an unequal number of protons and neutrons. Empirically, it is evaluated as

$$\Delta V_s = \pm 27\,MeV\left[\frac{N-Z}{A}\right] \quad \begin{array}{l} -\ neutrons \\ +\ protons \end{array}$$

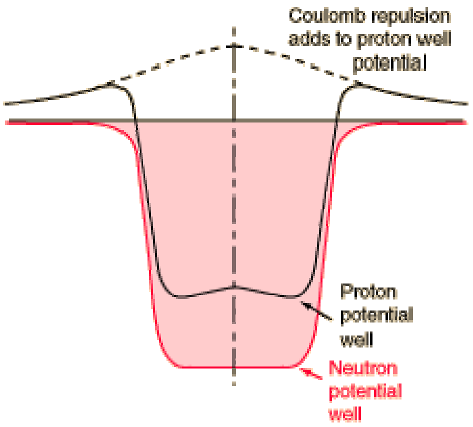
The other correction for protons is the electrostatic repulsion energy, which takes the form

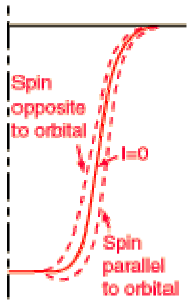
$$V(r) = \frac{Zke^2}{R_c}\left(1 + \frac{1}{2}\left[1 - \left(\frac{r}{R_c}\right)^2\right]\right) \qquad r < R_c$$

$$V(r) = \frac{Zke^2}{r} \qquad r > R_c$$

R_c = *charge radius, distinct from* R, *the model radius for the nuclear potential.*

The approximate potentials for neutrons and protons take the general form shown at left below.





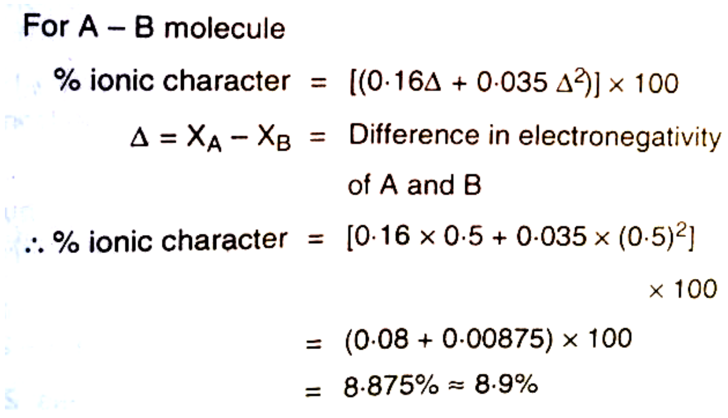
The additional correction which must be made is the spin-orbit interaction. It's general effect on the potential well is shown in the sketch above right. If the spin is opposite to the orbital angular momentum, the effective potential well is narrower, giving higher energy in the same manner as the square well potential. This can be seen in the spin-orbit splittings in the shell model level diagram.

11) Ionic character percentage of a diatomic molecule can be calculated by an Empirical formula if electronegativity values are given

The electronegativity difference between two elements A and B is 0.5 The magnitude of the percentage ionic character of a single A-B bond would be

a) 4.45 b) 8.9 c) 17.8 d) 26.7

The solution to this uses an empirical formula.



For A – B molecule

$$\% \text{ ionic character} = [(0{\cdot}16\Delta + 0{\cdot}035\,\Delta^2)] \times 100$$

$$\Delta = X_A - X_B = \text{Difference in electronegativity of A and B}$$

$$\therefore \% \text{ ionic character} = [0{\cdot}16 \times 0{\cdot}5 + 0{\cdot}035 \times (0{\cdot}5)^2] \times 100$$

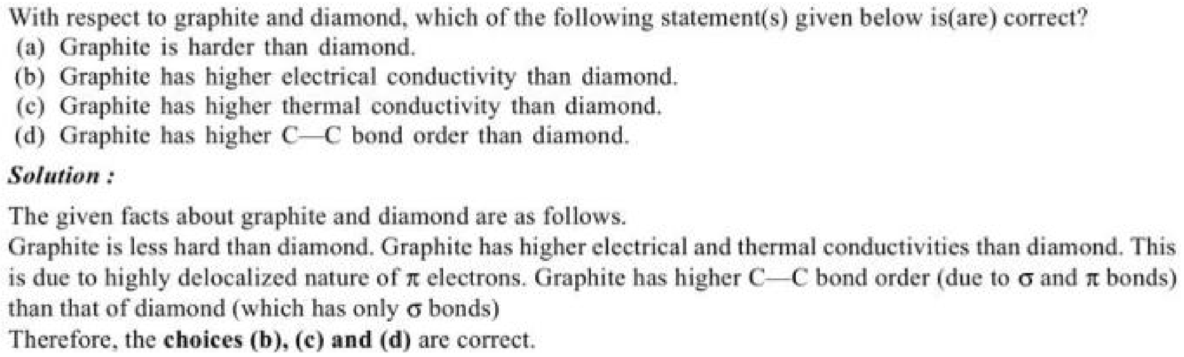
$$= (0{\cdot}08 + 0{\cdot}00875) \times 100$$

$$= 8{\cdot}875\% \approx 8{\cdot}9\%$$

-

Topic wise grouping of information kind of Dictionary of General Chemistry

Graphite and Diamond comparison



With respect to graphite and diamond, which of the following statement(s) given below is(are) correct?

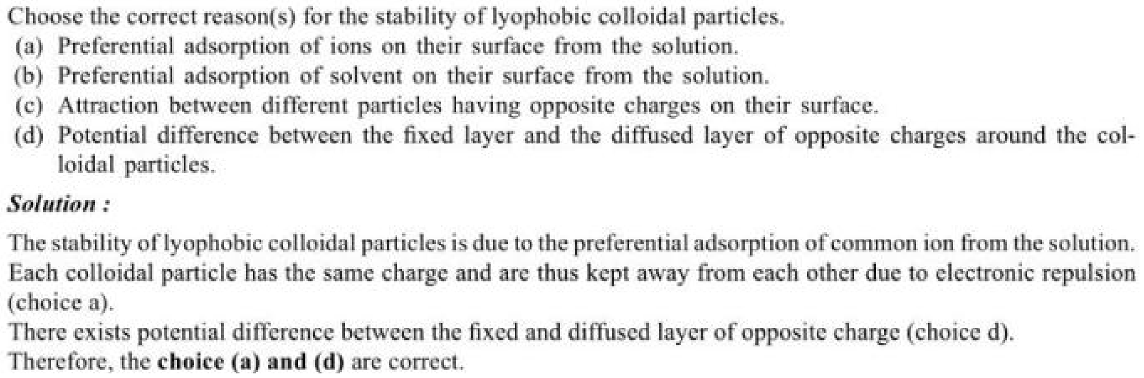
(a) Graphite is harder than diamond.
(b) Graphite has higher electrical conductivity than diamond.
(c) Graphite has higher thermal conductivity than diamond.
(d) Graphite has higher C—C bond order than diamond.

Solution :

The given facts about graphite and diamond are as follows.
Graphite is less hard than diamond. Graphite has higher electrical and thermal conductivities than diamond. This is due to highly delocalized nature of π electrons. Graphite has higher C—C bond order (due to σ and π bonds) than that of diamond (which has only σ bonds)
Therefore, the **choices (b), (c) and (d)** are correct.

-

Stability of lyophobic colloidal particles



Choose the correct reason(s) for the stability of lyophobic colloidal particles.

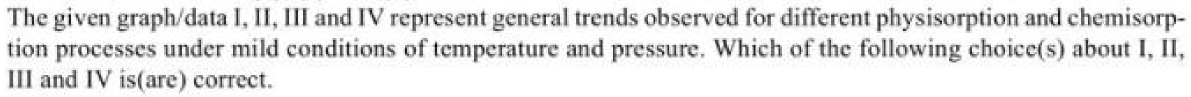
(a) Preferential adsorption of ions on their surface from the solution.
(b) Preferential adsorption of solvent on their surface from the solution.
(c) Attraction between different particles having opposite charges on their surface.
(d) Potential difference between the fixed layer and the diffused layer of opposite charges around the colloidal particles.

Solution :

The stability of lyophobic colloidal particles is due to the preferential adsorption of common ion from the solution. Each colloidal particle has the same charge and are thus kept away from each other due to electronic repulsion (choice a).
There exists potential difference between the fixed and diffused layer of opposite charge (choice d).
Therefore, the **choice (a) and (d)** are correct.

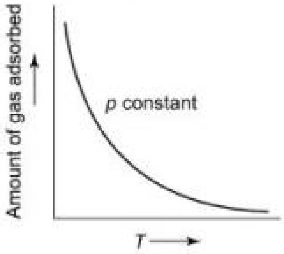
-

Physiorption and Chemisorption comparison

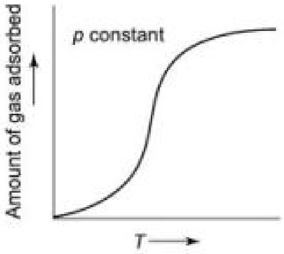


The given graph/data I, II, III and IV represent general trends observed for different physisorption and chemisorption processes under mild conditions of temperature and pressure. Which of the following choice(s) about I, II, III and IV is(are) correct.

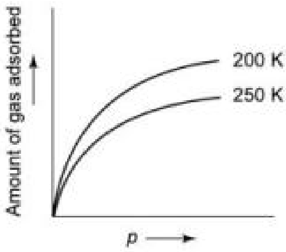
(I)



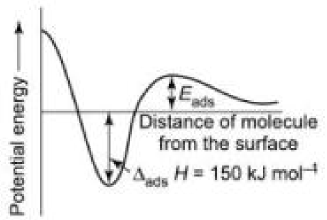
(II)



(III)



(IV)



(a) I is physisorption and II is chemisorption
(b) I is physisorption and III is chemisorption
(c) IV is chemisorption and II is chemisorption
(d) IV is chemisorption and III is chemisorption

Solution:

We have the following facts.

- Adsorption decreases with increase in temperature at constant pressure in case of physisorption
- Adsorption increases initially with increase in temperature at constant pressure in case of chemisorption as more molecules satisfy the energy of activation.
- The enthalpy change during chemisorptions is high due to the bond formation between adsorbent and adsorbate.

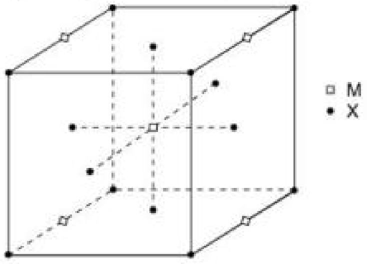
With these guidelines, we will have

Graph I represents physisorption
Graph II represents chemisorption
Graph III represents physisorption
Graph IV represents chemisorptions

Therefore, The **choices (a) and (c)** are correct.

-

A compound M_pX_q has cubic close packing (CCP) arrangement of X. Its unit cell structure is shown below.



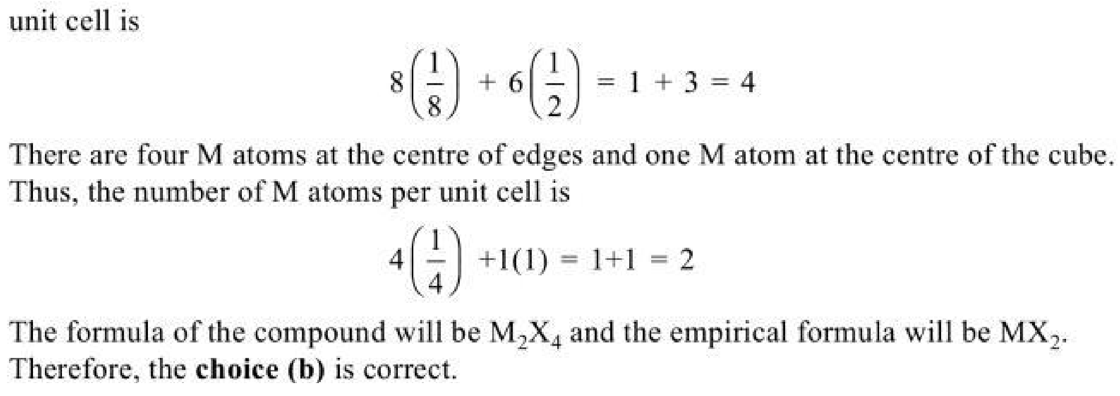
The empirical formula of the compound is

(a) MX
(b) MX_2
(c) M_2X
(d) M_5X_{14}

Solution :

The atom at the corner of the cube is shared amongst 8 unit cells; the atom at the centre of a face is shared between 2 unit cells, the atom at the centre of edges is shared amongst 4 unit cells and that at the centre of the cube belongs wholly to the unit cell.

There are eight X atoms at the corners and six X atoms at the centre of faces. Thus, the number of X atoms per



unit cell is

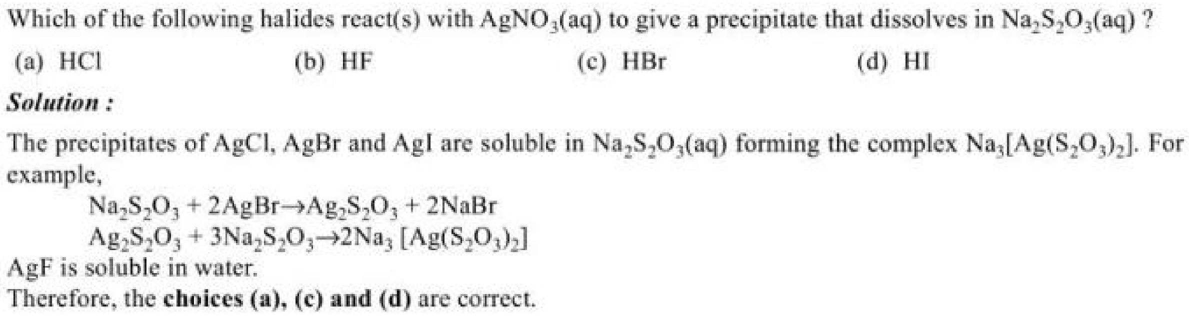
$$8\left(\frac{1}{8}\right) + 6\left(\frac{1}{2}\right) = 1 + 3 = 4$$

There are four M atoms at the centre of edges and one M atom at the centre of the cube. Thus, the number of M atoms per unit cell is

$$4\left(\frac{1}{4}\right) + 1(1) = 1+1 = 2$$

The formula of the compound will be M_2X_4 and the empirical formula will be MX_2. Therefore, the **choice (b)** is correct.

HCl HBr and HI react with AgNO3 to give a precipitate that dissolves in Na2S2O3



Which of the following halides react(s) with $AgNO_3(aq)$ to give a precipitate that dissolves in $Na_2S_2O_3(aq)$?

(a) HCl (b) HF (c) HBr (d) HI

Solution :

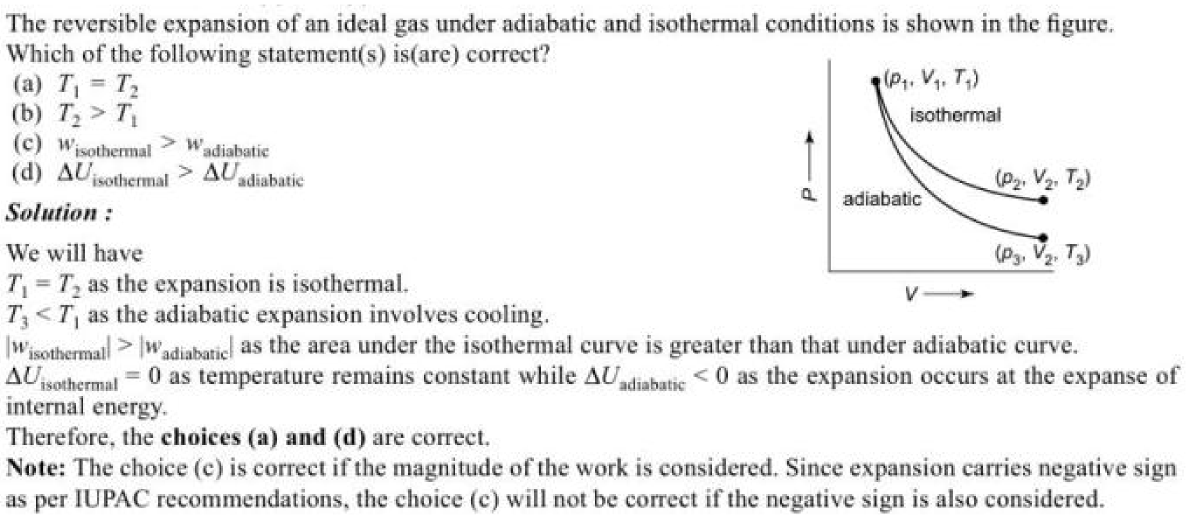
The precipitates of AgCl, AgBr and AgI are soluble in $Na_2S_2O_3(aq)$ forming the complex $Na_3[Ag(S_2O_3)_2]$. For example,

$$Na_2S_2O_3 + 2AgBr \rightarrow Ag_2S_2O_3 + 2NaBr$$
$$Ag_2S_2O_3 + 3Na_2S_2O_3 \rightarrow 2Na_3[Ag(S_2O_3)_2]$$

AgF is soluble in water.

Therefore, the **choices (a), (c) and (d)** are correct.

Reversible expansion of an ideal gas



The reversible expansion of an ideal gas under adiabatic and isothermal conditions is shown in the figure. Which of the following statement(s) is(are) correct?

(a) $T_1 = T_2$

(b) $T_2 > T_1$

(c) $w_{\text{isothermal}} > w_{\text{adiabatic}}$

(d) $\Delta U_{\text{isothermal}} > \Delta U_{\text{adiabatic}}$

Solution :

We will have

$T_1 = T_2$ as the expansion is isothermal.

$T_3 < T_1$ as the adiabatic expansion involves cooling.

$|w_{\text{isothermal}}| > |w_{\text{adiabatic}}|$ as the area under the isothermal curve is greater than that under adiabatic curve.

$\Delta U_{\text{isothermal}} = 0$ as temperature remains constant while $\Delta U_{\text{adiabatic}} < 0$ as the expansion occurs at the expanse of internal energy.

Therefore, the **choices (a) and (d)** are correct.

Note: The choice (c) is correct if the magnitude of the work is considered. Since expansion carries negative sign as per IUPAC recommendations, the choice (c) will not be correct if the negative sign is also considered.

Acid Ionization Constant	$K_a = \frac{[H^+][A^-]}{[HA]}$
Adibiatic change	$PV = k$
Base Ionization Constant	$K_b = \frac{[OH^-][HB^+]}{[B]}$
Bohr Radius	$a_0 = \frac{\hbar^2}{m_e k e^2}$
Boiling Point Elevation	$\Delta T_b = i K_b \times molality$
Buffer Design Equation	$pH \approx pK_a - \log \frac{[HA]_0}{[A^-]_0}$
Cell Voltage	$E_{cell} = E^\circ_{cell} - \frac{RT}{n\mathfrak{F}} \ln Q = E^\circ_{cell} - \frac{0.0592}{n} \log Q$
Charles' Law	$\frac{V}{t} = k$

Atomic Structure

DESCRIPTION	EQUATION
Bohr Radius	$a_0 = \frac{\hbar^2}{m_e k e^2}$
De Broglie Wavelength	$\lambda = \frac{h}{mv}$
Linear Momentum	$p = mv$
Planck's Quantized (Quantum) Energy Equation	$\Delta E = h\nu$
Radii of stable orbits in the Bohr model	$r = n^2 \frac{\hbar^2}{m_e k Z e^2} = n^2 \frac{a_0}{Z}$
Relationship between Energy and Principal Quantum Number	$E_n = -R_H \left(\frac{1}{n^2}\right) = \frac{-2.178 \times 10^{-18}}{n^2} joule$

Rydberg Equation	$\Delta E = R_H \left(\frac{1}{n_i^2} - \frac{1}{n_f^2} \right)$
Speed of Light to Wavelength and Frequency Relationship	$c = \lambda \nu$

Electrochemistry

DESCRIPTION	EQUATION
Cell Voltage	$E_{cell} = E^{\circ}_{cell} - \frac{RT}{n\Im} \ln Q = E^{\circ}_{cell} - \frac{0.0592}{n} \log Q$
Electric Current	$I = \frac{q}{t}$
Reaction Quotient	$Q = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ $where \quad aA + bB \rightarrow cC + dD$
Relationship between Equilibrium Constant and Cell Voltage	$\log K = \frac{nE^{\circ}}{0.0592}$

Equilibrium

DESCRIPTION	EQUATION
Acid Ionization Constant	$K_a = \frac{[H^+][A^-]}{[HA]}$
Base Ionization Constant	$K_b = \frac{[OH^-][HB^+]}{[B]}$
Buffer Design Equation	$pH \approx pK_a - \log \frac{[HA]_0}{[A^-]_0}$
Gas Pressure and Concentration Relationship	$K_p = K_c (RT)^{\Delta n}$
Ion Product Constant for Water	$K_w = [OH^-][H^+] = K_a \times K_b$ $= 1.0 \times 10^{-14} \quad at \quad 25^{\circ}C$
pH and pOH Relationship	$14 = pH + pOH$
pH Defined	$pH = -\log [H^+]$
pK_a Definition	$pK_a = -\log K_a$

pK_b Definition	$pK_b = -\log K_b$
pOH and Base Ionization Equilibrium Constant Relationship	$pOH = pK_b + \log \frac{[HB^+]}{[B]}$

Gases, Liquids, and Solutions

DESCRIPTION	EQUATION
Adibiatic change	$PV = k$
Boiling Point Elevation	$\Delta T_b = iK_b \times molality$
Charles' Law	$\frac{V}{t} = k$
Combined Gas Law	$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$
Density of a Material	$D = \frac{m}{V}$
Freezing Point Depression	$\Delta T_f = iK_f \times molality$
Graham's Law of Effusion	$\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$
Ideal gas equation	$PV = nRT$
Kinetic Energy per Mole	$\frac{KE}{mole} = \frac{3}{2}RTn$
Kinetic Energy per Mole	$\frac{KE}{mole} = \frac{3}{2}RTn$

Thermochemistry

DESCRIPTION	EQUATION
Acid Ionization Constant	$K_a = \frac{[H^+][A^-]}{[HA]}$
Adibiatic change	$PV = k$
Base Ionization Constant	$K_b = \frac{[OH^-][HB^+]}{[B]}$

Bohr Radius $$a_0 = \frac{\hbar^2}{m_e k e^2}$$

Boiling Point Elevation $$\Delta T_b = i K_b \times molality$$

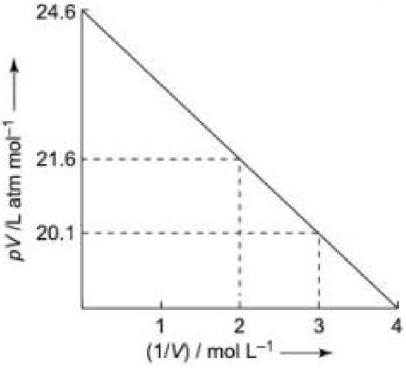
Buffer Design Equation $$pH \approx pK_a - \log \frac{[HA]_0}{[A^-]_0}$$

Cell Voltage $$E_{cell} = E^\circ_{cell} - \frac{RT}{n\Im} \ln Q = E^\circ_{cell} - \frac{0.0592}{n} \log Q$$

Charles' Law $$\frac{V}{t} = k$$

Vander waal's force

For one mole of a van der Waals gas when $b = 0$ and $T = 300$ K, the plot of pV versus $1/V$ is shown below.



The value of the van der Waals constant a is

(a) 1.0 L^2 atm mol^{-2} (b) 4.5 L^2 atm mol^{-2} (c) 1.5 L^2 atm mol^{-2} (d) 3.0 L^2 atm mol^{-2}

Solution :

For one mole of a gas, the van der Waals equation is $\left(p + \frac{a}{V^2}\right)(V - b) = RT$

When $b = 0$, we have $\left(p + \frac{a}{V^2}\right)V = RT$ i.e. $pV = RT - \frac{a}{V}$

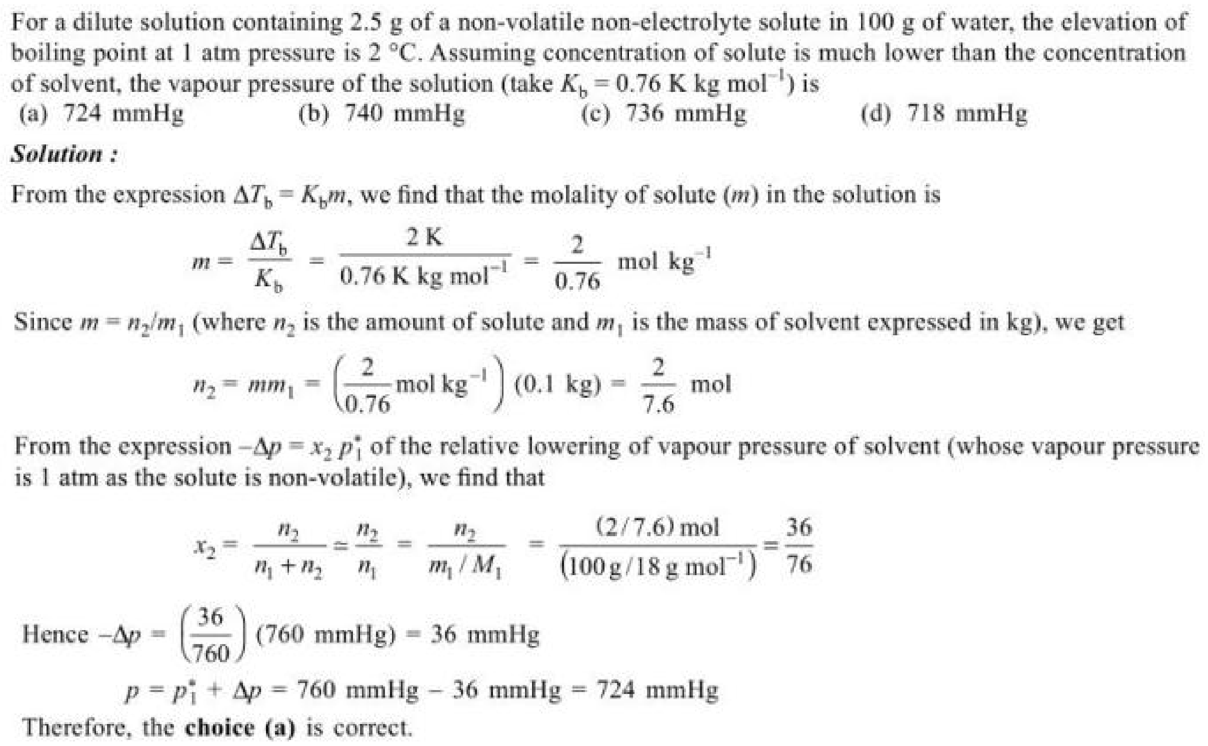
The plot of pV versus $1/V$ will be a straight line with slope equal to $-a$. Hence

Slope of the given straight line is $\frac{(20.1 - 21.6)\,\mathrm{L\,atm\,mol^{-1}}}{(3.0 - 2.0)\,\mathrm{mol\,L^{-1}}} = -1.5\ \mathrm{L^2\ atm\ mol^{-2}}$

Equating this to $-a$, we get $a = 1.5\ \mathrm{L^2\ atm\ mol^{-2}}$

Therefore, the **choice (c)** is correct.

dilute solution containing 2.5 gm of nonvolatile



For a dilute solution containing 2.5 g of a non-volatile non-electrolyte solute in 100 g of water, the elevation of boiling point at 1 atm pressure is 2 °C. Assuming concentration of solute is much lower than the concentration of solvent, the vapour pressure of the solution (take $K_b = 0.76$ K kg mol^{-1}) is

(a) 724 mmHg (b) 740 mmHg (c) 736 mmHg (d) 718 mmHg

Solution :

From the expression $\Delta T_b = K_b m$, we find that the molality of solute (m) in the solution is

$$m = \frac{\Delta T_b}{K_b} = \frac{2\ \text{K}}{0.76\ \text{K kg mol}^{-1}} = \frac{2}{0.76}\ \text{mol kg}^{-1}$$

Since $m = n_2/m_1$ (where n_2 is the amount of solute and m_1 is the mass of solvent expressed in kg), we get

$$n_2 = mm_1 = \left(\frac{2}{0.76}\text{mol kg}^{-1}\right)(0.1\ \text{kg}) = \frac{2}{7.6}\ \text{mol}$$

From the expression $-\Delta p = x_2\, p_1^*$ of the relative lowering of vapour pressure of solvent (whose vapour pressure is 1 atm as the solute is non-volatile), we find that

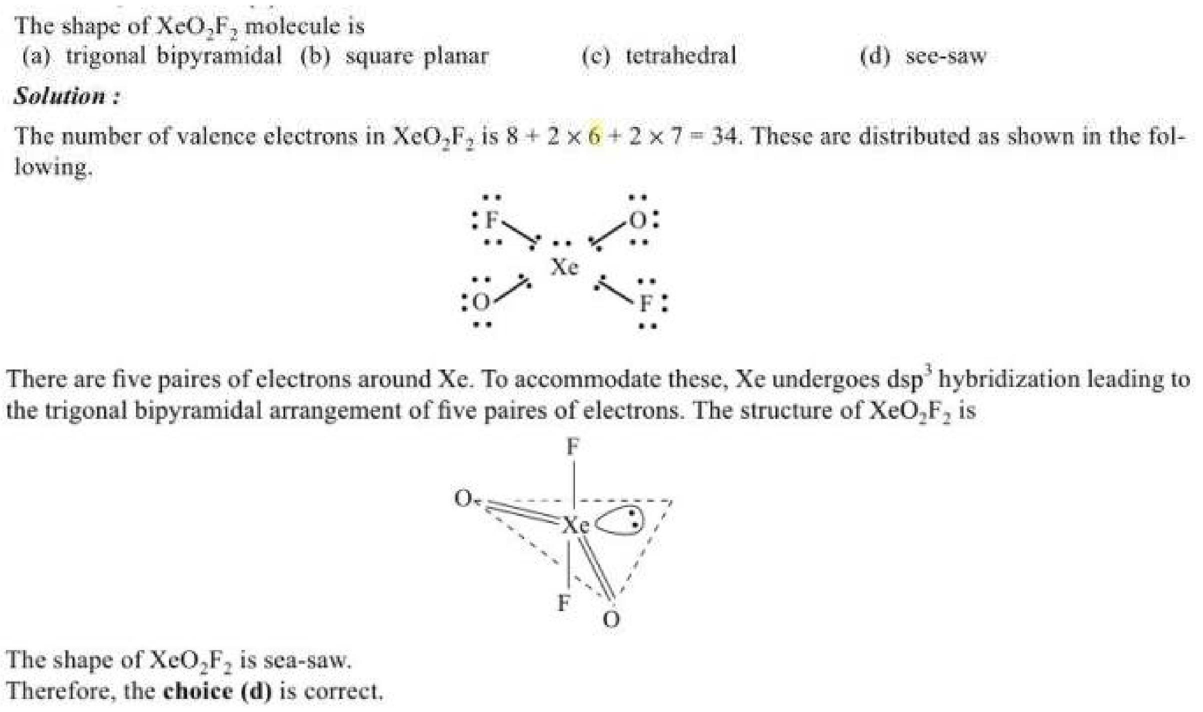
$$x_2 = \frac{n_2}{n_1 + n_2} \simeq \frac{n_2}{n_1} = \frac{n_2}{m_1/M_1} = \frac{(2/7.6)\ \text{mol}}{(100\ \text{g}/18\ \text{g mol}^{-1})} = \frac{36}{76}$$

$$\text{Hence } -\Delta p = \left(\frac{36}{760}\right)(760\ \text{mmHg}) = 36\ \text{mmHg}$$

$$p = p_1^* + \Delta p = 760\ \text{mmHg} - 36\ \text{mmHg} = 724\ \text{mmHg}$$

Therefore, the **choice (a)** is correct.

shape of XeO2F2 molecule is seesaw



The shape of XeO_2F_2 molecule is

(a) trigonal bipyramidal (b) square planar (c) tetrahedral (d) see-saw

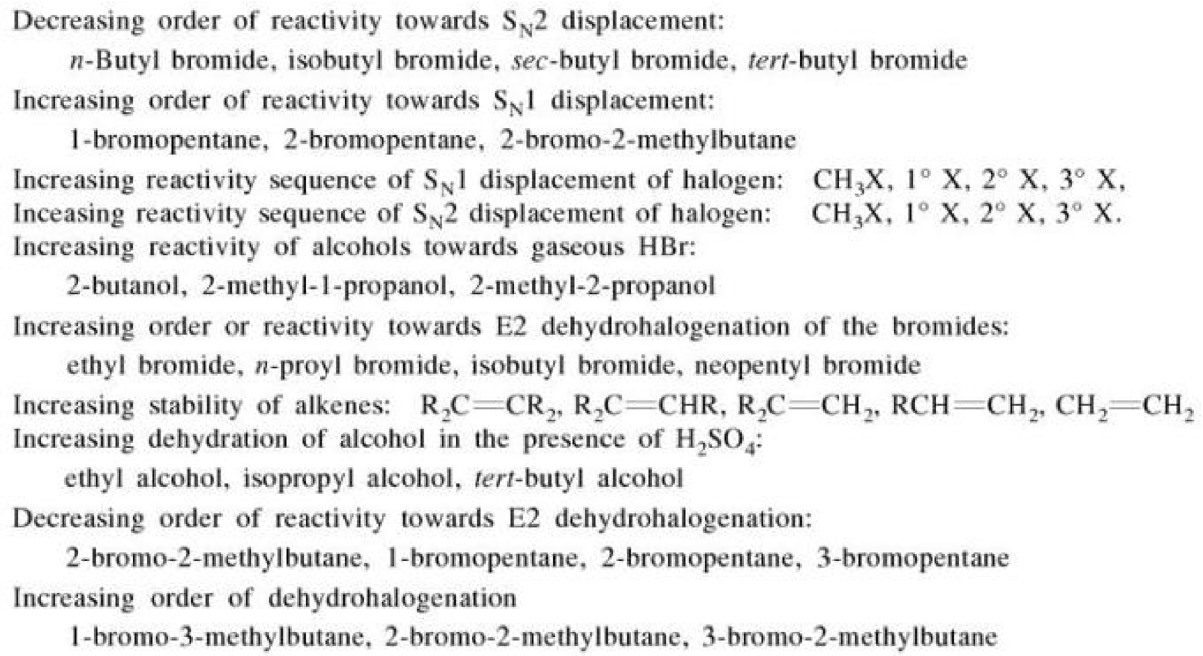
Solution :

The number of valence electrons in XeO_2F_2 is $8 + 2 \times 6 + 2 \times 7 = 34$. These are distributed as shown in the following.

There are five paires of electrons around Xe. To accommodate these, Xe undergoes dsp^3 hybridization leading to the trigonal bipyramidal arrangement of five paires of electrons. The structure of XeO_2F_2 is

The shape of XeO_2F_2 is sea-saw.
Therefore, the **choice (d)** is correct.

Various increasing and Decreasing properties



Decreasing order of reactivity towards S_N2 displacement:

n-Butyl bromide, isobutyl bromide, *sec*-butyl bromide, *tert*-butyl bromide

Increasing order of reactivity towards S_N1 displacement:

1-bromopentane, 2-bromopentane, 2-bromo-2-methylbutane

Increasing reactivity sequence of S_N1 displacement of halogen: CH_3X, 1° X, 2° X, 3° X,
Inceasing reactivity sequence of S_N2 displacement of halogen: CH_3X, 1° X, 2° X, 3° X.

Increasing reactivity of alcohols towards gaseous HBr:

2-butanol, 2-methyl-1-propanol, 2-methyl-2-propanol

Increasing order or reactivity towards E2 dehydrohalogenation of the bromides:

ethyl bromide, *n*-proyl bromide, isobutyl bromide, neopentyl bromide

Increasing stability of alkenes: $R_2C{=}CR_2$, $R_2C{=}CHR$, $R_2C{=}CH_2$, $RCH{=}CH_2$, $CH_2{=}CH_2$

Increasing dehydration of alcohol in the presence of H_2SO_4:

ethyl alcohol, isopropyl alcohol, *tert*-butyl alcohol

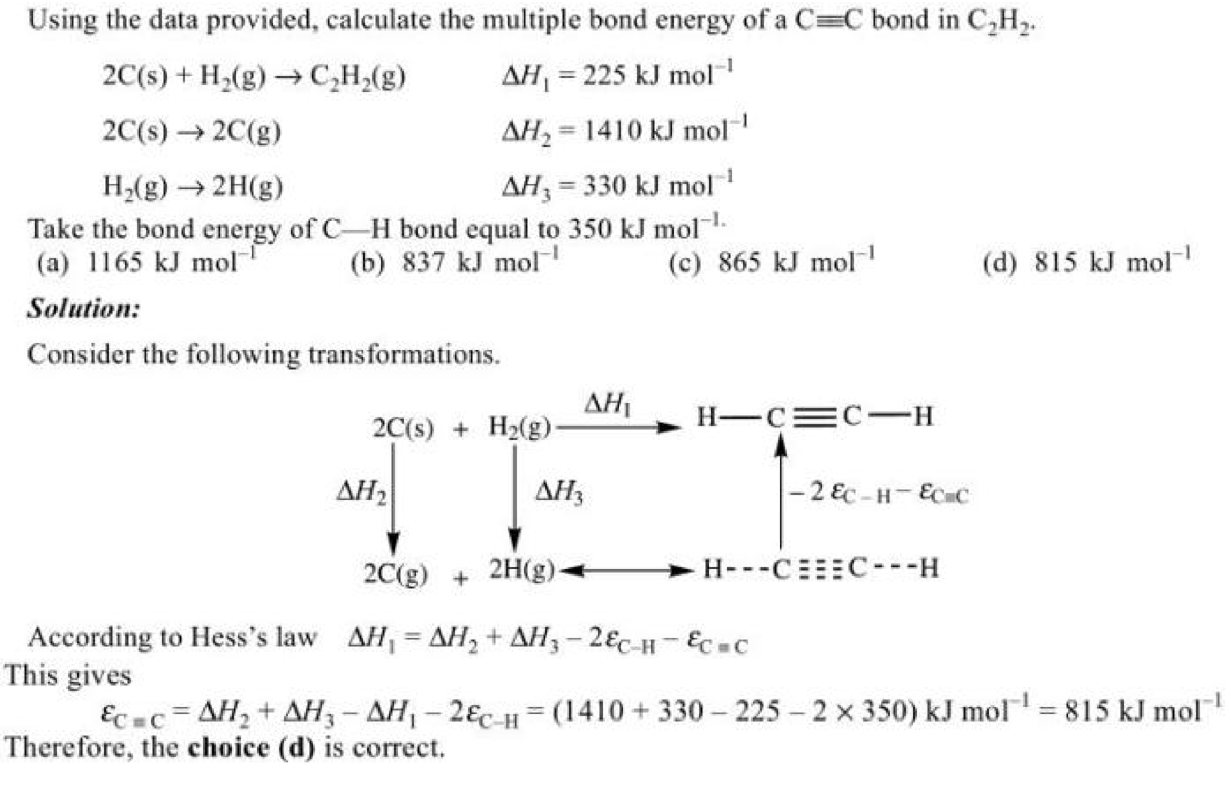
Decreasing order of reactivity towards E2 dehydrohalogenation:

2-bromo-2-methylbutane, 1-bromopentane, 2-bromopentane, 3-bromopentane

Increasing order of dehydrohalogenation

1-bromo-3-methylbutane, 2-bromo-2-methylbutane, 3-bromo-2-methylbutane

calculate multiple bond energy of a C triple bond



Using the data provided, calculate the multiple bond energy of a $C{\equiv}C$ bond in C_2H_2.

$2C(s) + H_2(g) \rightarrow C_2H_2(g) \qquad \Delta H_1 = 225 \text{ kJ mol}^{-1}$

$2C(s) \rightarrow 2C(g) \qquad \Delta H_2 = 1410 \text{ kJ mol}^{-1}$

$H_2(g) \rightarrow 2H(g) \qquad \Delta H_3 = 330 \text{ kJ mol}^{-1}$

Take the bond energy of C—H bond equal to 350 kJ mol^{-1}.

(a) 1165 kJ mol^{-1} (b) 837 kJ mol^{-1} (c) 865 kJ mol^{-1} (d) 815 kJ mol^{-1}

Solution:

Consider the following transformations.

$$\begin{tikzcd}
2C(s) \arrow[d, "\Delta H_2"'] & + & H_2(g) \arrow[r, "\Delta H_1"] \arrow[d, "\Delta H_3"] & H-C\equiv C-H \\
2C(g) & + & 2H(g) \arrow[r, leftrightarrow] & H\cdots C \,\vdots\vdots\vdots\, C\cdots H \arrow[u, "-2\varepsilon_{C-H}-\varepsilon_{C\equiv C}"']
\end{tikzcd}$$

According to Hess's law $\Delta H_1 = \Delta H_2 + \Delta H_3 - 2\varepsilon_{C-H} - \varepsilon_{C\equiv C}$

This gives

$$\varepsilon_{C\equiv C} = \Delta H_2 + \Delta H_3 - \Delta H_1 - 2\varepsilon_{C-H} = (1410 + 330 - 225 - 2 \times 350) \text{ kJ mol}^{-1} = 815 \text{ kJ mol}^{-1}$$

Therefore, the **choice (d)** is correct.

Some facts student must know

Largest Anion - At^-

Metals having highest b.pt and m.pt. = mercury and tungsten

Most reactive solid element - Li

Most reactive liquid element - Cs

total no. of gaseous element in periodic table - 11 (H , He , N , O , F , Ne , Cl , Ar , Kr , Xe , Ra)

TOTAL NO. OF LIQUID elements - 6 (Ga , Br , Cs , Hg , Fr , EKa)

Liquid radioactive element - Francium

N.metal wid highest M.Pt - Carbon

Metal wid highest valency = Plutonium

highest tensile strength - Boron

Most ionic compound = CsF

Strongest base = Cs (OH)

Strongest basic oxide = Cs_2O

Most conducting metal = Ag

-

Topic wise grouping of information kind of Dictionary of Organic Chemistry

Compounds of carbon with other elements in the periodic table

CH																	He
CLi	CBe											CB	CC	CN	CO	CF	Ne
CNa	CMg											CAl	CSi	CP	CS	CCl	Ar
CK	CCa	CSc	CTi	CV	CCr	CMn	CFe	CCo	CNi	CCu	CZn	CGa	CGe	CAs	CSe	CBr	CKr
CRb	CSr	CY	CZr	CNb	CMo	CTc	CRu	CRh	CPd	CAg	CCd	CIn	CSn	CSb	CTe	CI	CXe
CCs	CBa		CHf	CTa	CW	CRe	COs	CIr	CPt	CAu	CHg	CTl	CPb	CBi	CPo	CAt	Rn
Fr	CRa		Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Fl	Uup	Lv	Uus	Uuo
		↓															
		CLa	CCe	CPr	CNd	Pm	CSm	CEu	CGd	CTb	CDy	CHo	CEr	CTm	CYb	CLu	
		Ac	CTh	CPa	CU	CNp	CPu	CAm	CCm	CBk	Cf	CEs	Fm	Md	No	Lr	

Chemical bonds to carbon

Core organic chemistry	Many uses in chemistry
Academic research, but no widespread use	Bond unknown / not assessed

Abscisic acid (ABA) - A PLANT HORMONE once thought to be responsible for the shedding (abscission) of flowers and fruit and for the onset of dormancy in buds (hence its early name, dormin). The compound is associated with the closing of pores (stoma) in the leaves of plants deprived of water.

Absolute configuration - A particular molecular configuration of a CHIRAL molecule, as denoted by comparison with a reference molecule or by some sequence rule.There are two systems for expressing absolute configuration in common use: theD-L convention and the R-S convention.

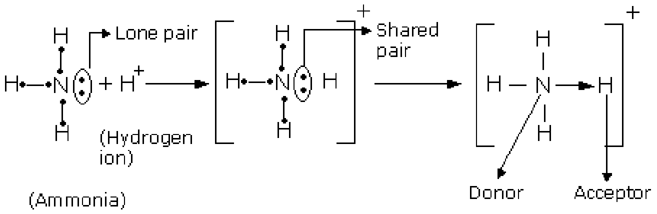
Absorption

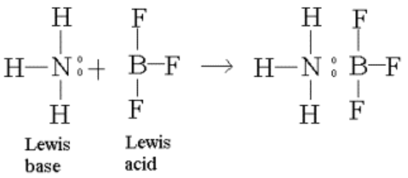
1. A process in which a gas is taken up by a liquid or solid, or in which a liquid is taken up by a solid. In absorption, the substance absorbed goes into the bulk of the material. Solids that absorb gases or liquids often have a porous structure. The absorption of gases in solids is sometimes called sorption. There is a distinction between absorption (in which one substance is assimilated into the bulk of another) and ADSORPTION (which involves attachment to the surface). Sometimes it is not obvious which process is occurring. For example, a porous solid, such as activated CHARCOAL may be said to absorb a large volume of gas, but the process may actually be adsorption on the high surface area of internal pores in the material.

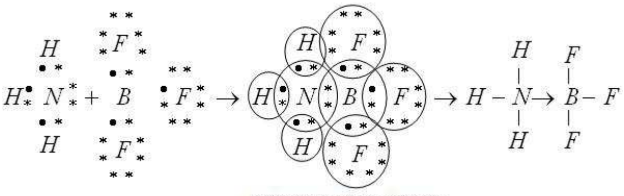
2. The process in which electromagnetic radiation, particles, or sound waves lose energy in passing through a medium. Absorption involves conversion of one form of energy into another.

Accelerator - A substance that increases the rate of a chemical reaction. In this sense the term is synonymous with CATALYST. It is common to refer to catalysts as ‘accelerators’ in certain industrial applications. For example, accelerators are used in the VULCANIZATION of rubber and in the polymerization of adhesives. Also, in the production of composite materials using polyester resins a distinction is sometimes made between the catalyst (which initiates the polymerization reaction) and the accelerator (which is an additional substance making the catalyst more effective). The terms promoter and activator are used in a similar way.

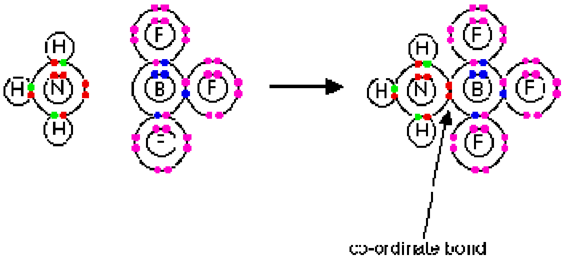
Acceptor - The atom or group to which a pair of electrons is donated in forming a COORDINATE BOND.



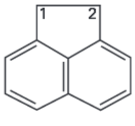




Formation of a co-ordinate bond between NH_3 and

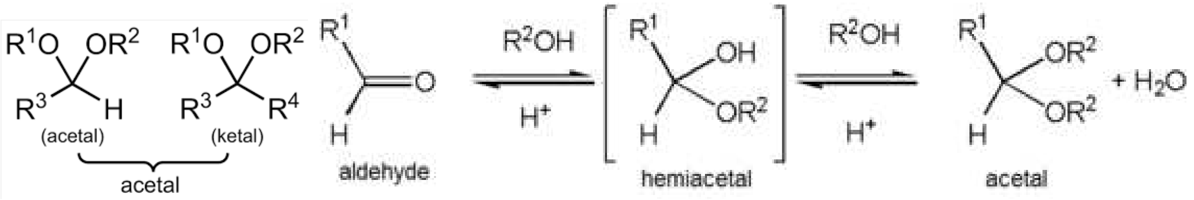


Acenaphthene - ($C_{12}H_{10}$) A colorless crystalline derivative of naphthalene, used in producing some dyes.



Acenaphthene

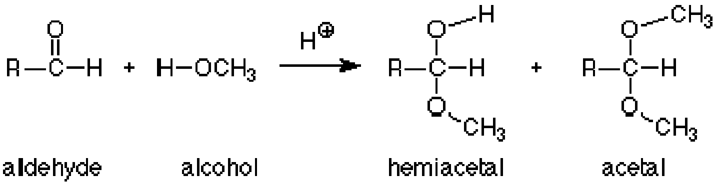
Acetal - A type of compound formed by reaction of an alcohol with either an aldehyde or a ketone. The first step in formation of an acetal is the formation of an intermediate, known as a hemiacetal. For example, ethanal (acetaldehyde; CH3CHO) reacts with ethanol (C2H5OH) as follows: $CH_3CHO + C_2H_5OH$ <==> $CH(OH)(CH_3)(C_2H_5O)$ The hemiacetal has a central carbon atom (from the aldehyde) attached to a hydrogen, a hydroxyl group, a hydrocarbon group (methyl - CH_3), and an alkoxy group (C_2H_5O). If a ketone is used rather than an aldehyde, the resulting hemiacetal contains two hydrocarbon groups. For example, reaction of the ketone R1COR2with the alcohol R3OH is:
R1COR2 + R3OH < == >CR1R2(OH)(OR3)



The formation of a hemiacetal is an example of NUCLEOPHILIC ADDITION to the carbonyl group of the aldehyde or ketone. The first step is attack of the lone pair on the O of the alcohol on the (positively charged) C of the carbonyl group. This is catalyzed by both acids and bases. Acid catalysis occurs by protonation of the O on the carbonyl, making the C more negative and more susceptible to nucleophilic attack. In base catalysis the OH^- ions from the base affect the -OH group of the alcohol, making it a more effective nucleophile.

In general, hemiacetals exist only in solution and cannot be isolated because they easily decompose back to the component alcohol and aldehyde or ketone. However, some cyclic hemiacetals are more stable. For example, cyclic forms of SUGAR molecules are hemiacetals.

Further reaction of hemiactals with another molecule of alcohol leads to a full acetal. For example:
$CH(OH)(CH3)(OC_2H_5) + C_2H_5OH$ <==>$CH(CH3)(OC_2H_5)_2$

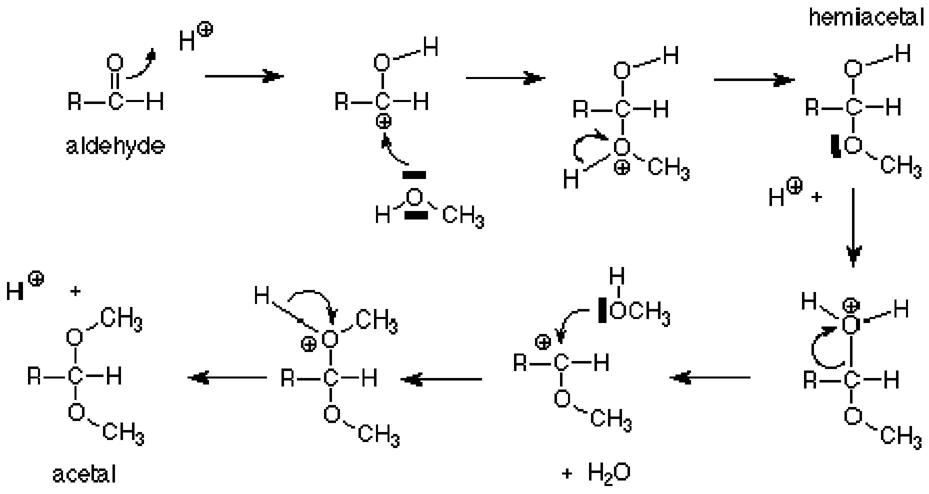


The overall reaction of an aldehyde or ketone with an alcohol to give an acetal can be written: R1COR2 + R3OH < == > CR1R2(OR3)2

It is also possible to have ‘mixed’ acetals with the general formula CR1R2(OR3)-(OR4). Note that if the acetal is derived from an aldehyde, then R1 and/or R2

May be a hydrogen atom. The mechanism of formation of an acetal from a hemiacetal is acid catalyzed. It involves protonation of the -OH group of the hemiacetal followed by loss of water to form an oxonium ion, which is attacked by the alcohol molecule.

Formerly it was conventional to use the terms ‘hemiacetal’ and ‘acetal’ for compounds formed by reaction between aldehydes and alcohols. Similar reactions between ketones and alcohols gave rise to compounds called hemiketals and ketals. Current nomenclature uses ‘hemiacetal’ and ‘acetal’ for compounds derived from either an aldehyde or a ketone, but reserves ‘hemiketal’ and ‘ketal’ for those derived from ketones. In other words, the ketals are a subclass of the acetals and the hemiketals are a subclass of the hemiacetals.

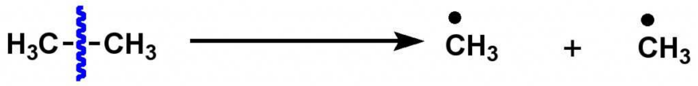


Achiral - Describing a molecule that does not have chiral properties; i.e. one that does not exhibit OPTICAL ACTIVITY.

Bond fission

A covalent bond is formed when electrons are shared between two atoms in the classical sense. A single bond (sigma bond) is thus made up of two electrons. Now a chemical reaction takes place when old bonds are broken and new ones are created. So how can one break a single bond—there are plainly two ways to go about breaking a bond as shown below.

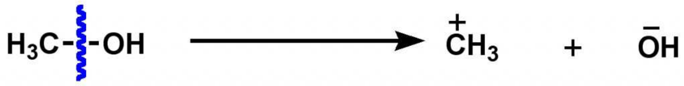
Homolytic fission



Homolysis

Homolytic fission is where each atom of the bond keeps an electron each resulting in species called free radicals. Radicals are important intermediates in organic chemistry and we will talk about them later. As the bond breaks to give two similar species each keeping an electron this form of bond breaking is called *Homolytic Fission.*

Heterolytic fission



Heterolysis

In this case we can see that one of the atoms carry a negative charge after bond cleavage indicating that it has both the electrons of the bond and the other has no electrons at all. Hence it is electron deficient thus positively charged. As the electrons are not divided equally after bond cleavage this is called *Heterolytic Fission*. In a case the C atom carries a positive charge it is called a carbocation and in the case it carries both the electrons of the broken bond and is negatively charged, it is called a Carbanion. Carbocation and Carbanions are the most important carbon intermediates in organic chemistry. Please note that both types of fissions are applicable to both homoatomic and heteroatomic bonds (bonds between two different atoms say C-N or C-O).

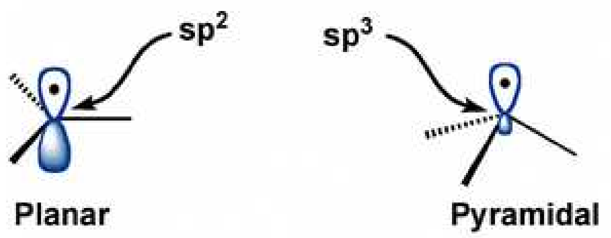
Free radicals

These are neutral intermediates, formed due to homolytic cleavage of a single bond. Some common bonds which cleave to give free radicals in organic chemistry are shown: C-O, C-Cl, C-Br, C-I, C-C, C-H. Carbon free radicals are mainly generated by:

i. Photolysis (action of light) like acetone alpha cleavage
ii. Other radical initiator like allylic bromination by N-Bromosuccinimide (NBS)

There has been a certain degree of debate as to what the shape and geometry of a free radical is like. Revisiting the theory of hybridization, there can be two basic shapes of these radicals.

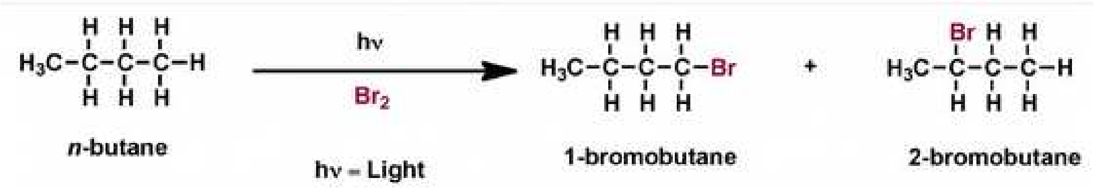
If the centre carbon atom of the radical is sp^3 hybridized (remember the one which was made of one s and three orbitals as in CH_4), the geometry will be tetrahedral.[1] But in the case of a radical there are only three groups attached to the sp^3 hybridized carbon atom so they we will have a shape of what resembles a pyramid—it's a tetrahedron with its head cut off. So sp^3 hybridized radicals are pyramidal in shape. The single electron of the radical would then be housed in a sp^3 orbital. The other option is sp^2 hybridization. In that case the C atom is sp^2 hybridized, so as discussed previously the shape would be planar with the single electron in the unhybridized p-orbital with the three substituents having sp^2 hybridized bonds.



Two different geometries of free radicals. The single electrons are shown as black dots.

So to summarize free radicals:

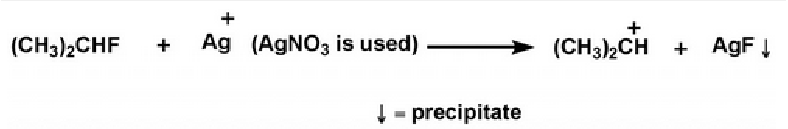
- Formed under activation by light or use of additional compounds called Radical Initiators.
- They are very reactive, because they have an unpaired electron which wants to get paired up.
- They are either pyramidal or planar with the lone electron in their sp^3 or p orbitals respectively.
- Because of their high reactivity, they tend to be less selective. In simple terms it means that it sometimes difficult to predict what products are formed in reactions which involve free radicals and we actually get several products from a single reaction.



This reaction shows the formation of two products with the Br atom attached to different carbons.

Carbocations

Carbocations are formed from the heterolytic cleavage of a carbon-heteroatom (meaning a non carbon atom in general) bond where the other atom is more electronegative than carbon like a C-O, C-N, C-X (X can be Cl, Br, I, etc) bond. This is quite logical as after the cleavage if a carbocation is to be formed the two electrons of the bond must go to the other atom. And this is favoured if that other atom is electronegative. Carbocations can be made in difficult conditions by using so-called superacids, developed by George Olah (Nobel Prize, 1994), which helps stabilize these intermediates substantially to be analyzed. Formation of carbocations can be assisted by using cations like Ag^+, with alkyl halides as substrates.



$$(CH_3)_2CHF + \overset{+}{Ag}\ (AgNO_3 \text{ is used}) \longrightarrow (CH_3)_2\overset{+}{C}H + AgF\downarrow$$

Formation of carbocation

The precipitating out of the silver salt forces the equilibrium to shift towards the forwards reaction.

A Similar reaction is Finkelstein reaction is particularly useful for preparing iodoalkanes. The iodoalkanes are obtained by heating chloro or bromoalkanes with a concentrated solution of sodium iodide in dry acetone.

In Acetone after Reflux (Heating in enclosed container) for a long time [Refluxing is a process of heating a liquid in a flask provided with a condenser. The vapours of the liquid which rise up get condensed back into the flask.]

RX + NaI ------ > RI + NaX where X = Cl, Br, and R is an Alkyl group

Sodium chloride and sodium bromide being less soluble in acetone get precipitated from the solution and can be removed by filtration. This also prevents the backward reaction to occur according to Le Chatelier's principle. The reaction gives the best results with primary halides.

With iodide, the solvate $NaI \cdot 3(CH_3)_2CO$ can be formed, in which each Na^+ is coordinated by 6 acetone ligands via oxygen lone pairs, and the I− ions fill in the gaps between these octahedral units, being surrounded by the methyl groups of the acetones. I− has a larger ionic radius compared to Cl− and Br− and thus fills the gap completely, thereby minimizing empty space in the lattice and its distance to Na+, which is, however, still larger than in solvent-free sodium iodide. This larger distance between the ions decreases the lattice energy, which in turn makes the compound readily soluble in acetone. Cl− and Br− are too small to fit in these gaps, and as the cation-anion distance becomes too large, the only way to minimize the lattice energy is the formation of a solvent-free structure NaX (X = Cl−, Br−). When the lattice energy is larger than the energy released by solvation of the ions, namely that of Na+ by complexation with acetone, the compound becomes insoluble and precipitates from the solution.

The positively charged carbon atom in carbocations is sp^2 hybridized, which means it's planar as we know by now. The three substituents of the carbocation lie in a plane leaving the unhybridized empty p orbital perpendicular to them.

These intermediates react with species which are electron rich (quite obvious) and being charged are stabilized in polar solvents. (Just as Na^+ is soluble and stable in polar water). Carbocations are important intermediates in most mechanisms along with carbanions as we shall see later.

To summarize carbocations:

- Formed due to heterolysis of a C-X bond (where X is more electronegative) and thus has a positive charge.
- Planar in shape (sp^2 hybridized carbon), with empty p orbital perpendicular to the plane of the molecule.
- Reactive towards electron rich species.

Carbanions

These are intermediates also formed as a result of heterolysis, but here the electron pair from the bond is kept by the carbon atom. From what we saw earlier the more electronegative atom keeps the electrons, so in this case carbon must the more electronegative of the two atoms making up the bond. Now there are only a few atoms (non-metals; metals are not usually part of organic chemistry) which are less electronegative, so the most common bond cleavage which yields carbanions is the C-H bond. The ease of breaking this bond and creating a carbanion is also a measure of the compound's acidity, because a H^+ is also generated with the carbanion, which makes the molecule an acid in the Bronsted sense.

Carbanions have three groups attached to each other and a lone pair of electrons which gives it its negative charge (similar to the ammonia molecule where the central N has 3 Hs and a lone pair of electrons). So its geometry is pyramidal (tetrahedral but since there is no fourth group again it's like a tetrahedral with head cut off) and the carbon atom is sp^3 hybridized.

Carbanions are also stable in polar solution (electrostatic stabilization).

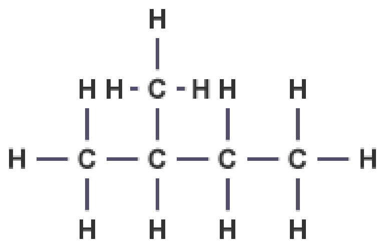
To summarize carbanions:

- Formed due to heterolysis of a C-X bond (where X is less electronegative) and thus has a negative charge.
- Pyramidal is shape (sp^3 hybridized) with the excess electrons placed in one sp^3 hybrid orbital.
- Reactive towards positively charged (electron deficient species).

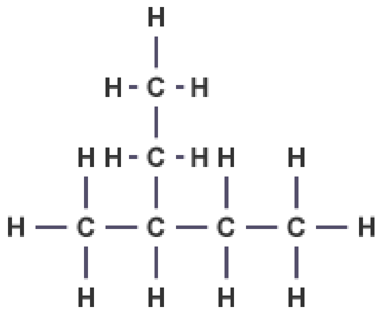
Stability of intermediates

Most organic reactions take place via formation of intermediates. So the study of different intermediates would help us predict the course of the reaction and the main aspect to look at would be their stability. No organic mechanism has been conclusively 'PROVEN', all the mechanism we see are the most plausible ones derived from many experiments, a major component of which is isolating and studying the intermediates. It is difficult to say that a certain mechanism is absolutely correct, but it is quite simple to point out an incorrect mechanism. One of the ways a chemist would confirm an incorrect mechanism is if it involves a very unstable intermediate. The good thing about this is that with a few empirical rules and principles in mind, it is quite simple to assign relative stability of intermediates like radicals, carbocations and carbanions. And what is even better is that we have already discussed these principles.

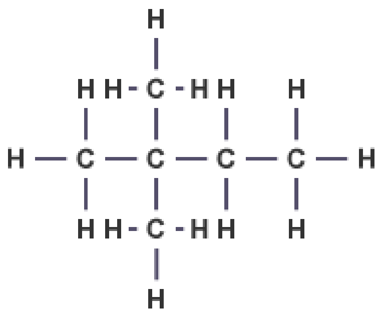
For carbocations and free radicals (both electron poor species), any group which donates electron density to the carbon centre would stabilize it and inversely electron withdrawing groups would increase electron deficiency on the carbon centre leading to destabilization. (Remember charge is not desirable, the most stable species are usually neutral). So following the same logic the effect should just be opposite in the case of carbanions as they are electron rich (negatively charged) instead of being electron deficient like the above two. So groups which pull away electrons from the charged carbon atom would have a stabilizing effect whereas electron donation would destabilize the intermediate as it loads more negative charge on an already negatively charged atom.



2-methylbutane

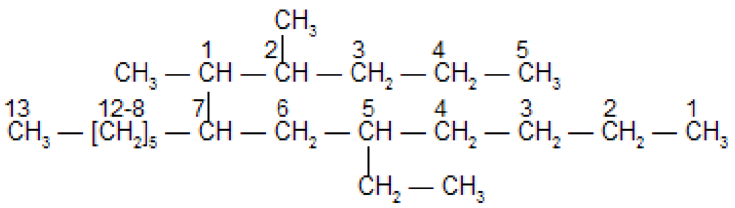


3-methylpentane (not 2-ethylbutane)



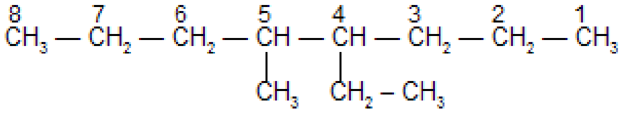
2,2-dimethylbutane

The name of a complex radical is considered to start with the first letter of its complete name.

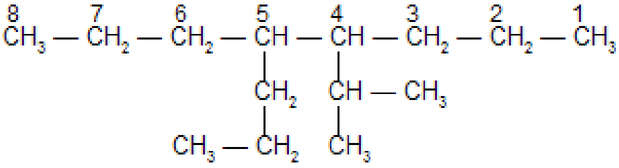


dimethylpentyl (as a complete single substituent) is alphabetized under " d ", thus 7-(1,2-Dimethylpentyl)-5-ethyltridecane

In cases where names of complex radicals are composed of identical words, preference for citation is given to that radical which contains the lowest locant at the first cited point of difference in the radical.

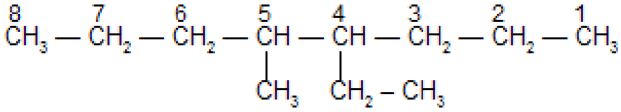


4-Ethyl-5-methyloctane

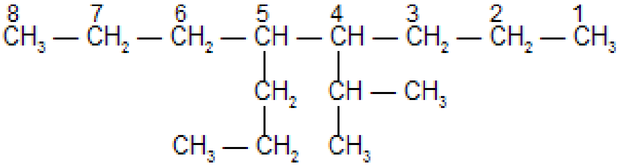


4-Isopropyl-5-propyloctane

If two or more side chains are in equivalent positions the one to be assigned the lower number is that cited first in the name.

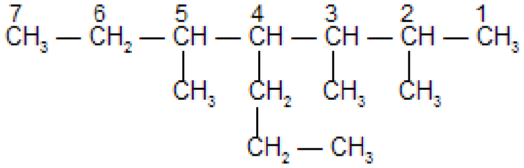


4-Ethyl-5-methyloctane

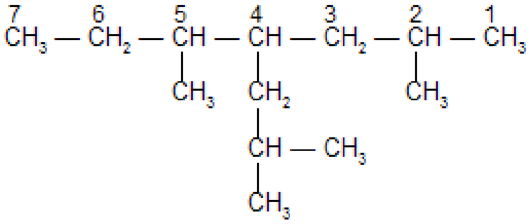


4-Isopropyl-5-propyloctane

The lowestsum rule.

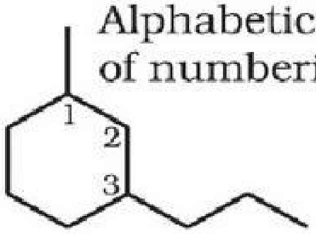


2,3,5-Trimethyl-4-propylheptane (not 3,5,6-Trimethyl-4-propylheptane)



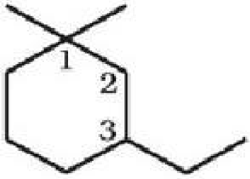
4-Isobutyl-2,5-dimethylheptane (not 4-Isobatyl-2, 6-dimethylheptane

Cyclopentane



Alphabetical order of numbering

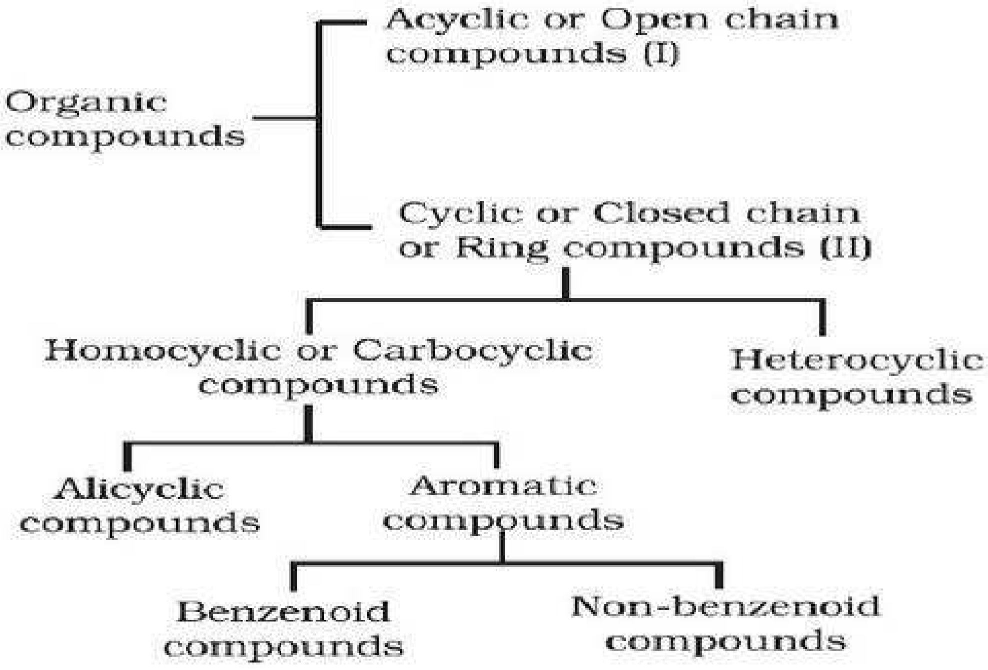
1-Methyl-3-propylcyclohexane



More branched carbon gets lower number

3-Ethyl-1,1-dimethylcyclohexane (not 1-ethyl-3,3-dimethylcyclohexane)

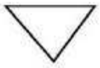
CLASSIFICATION OF ORGANIC COMPOUNDS



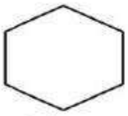
Alicyclic or closed chain or ring compounds

Alicyclic (aliphatic cyclic) compounds contain carbon atoms joined in the form of a ring (homocyclic). Sometimes atoms other than carbon are also present in the ring (heterocylic).

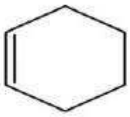
Some examples are



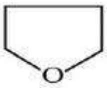
Cyclopropane



Cyclohexane



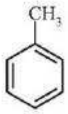
Cyclohexene



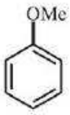
Tetrahydrofuran

Nomenclature of Substituted Benzene Compounds

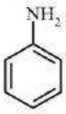
For IUPAC nomenclature of substituted benzene compounds, the substituent is placed as prefix to the word *benzene as* shown in the following examples.



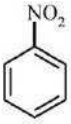
Methylbenzene
(Toluene)



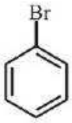
Methoxybenzene
(Anisole)



Aminobenzene
(Aniline)



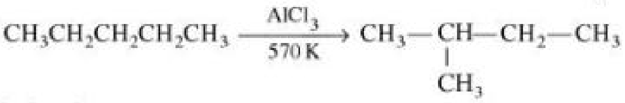
Nitrobenzene



Bromobenzene

How Isomers are made ?

(a) *Normal to branched chain in alkane* *n*-Alkane is heated with $AlCl_3$ at 570 K to give branched chain



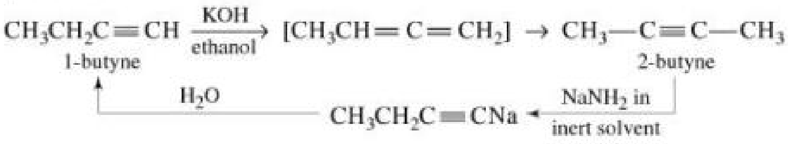
$$CH_3CH_2CH_2CH_2CH_3 \xrightarrow[570\,K]{AlCl_3} CH_3-\underset{\underset{CH_3}{|}}{CH}-CH_2-CH_3$$

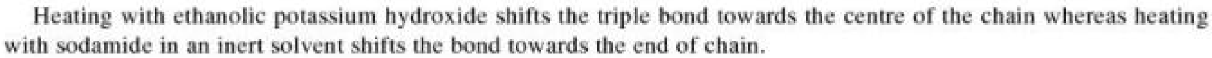
(b) *Shift of double bond*

In the presence of catalyst like $Al_2(SO_4)_3$, alkene undergo isomerization at high temperature (770 – 970 K).

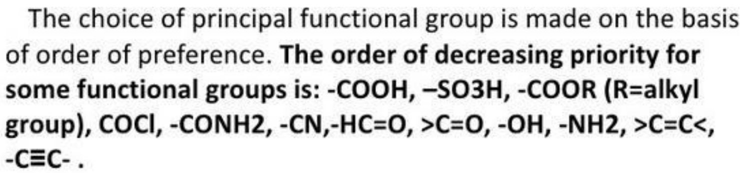
$$CH_3CH_2CH_2CH=CH_2 \rightarrow CH_3CH_2CH=CHCH_3$$

(c) *Shift of triple bond*



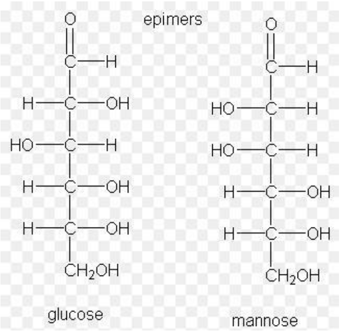


Heating with ethanolic potassium hydroxide shifts the triple bond towards the centre of the chain whereas heating with sodamide in an inert solvent shifts the bond towards the end of chain.

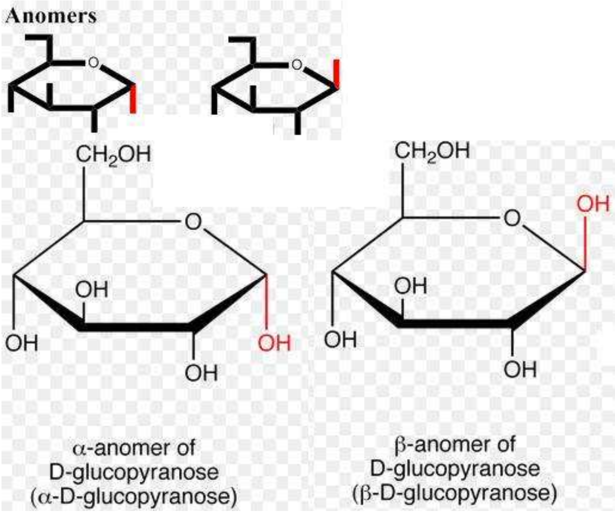


The choice of principal functional group is made on the basis of order of preference. **The order of decreasing priority for some functional groups is: -COOH, –SO3H, -COOR (R=alkyl group), COCl, -CONH2, -CN,-HC=O, >C=O, -OH, -NH2, >C=C<, -C≡C- .**

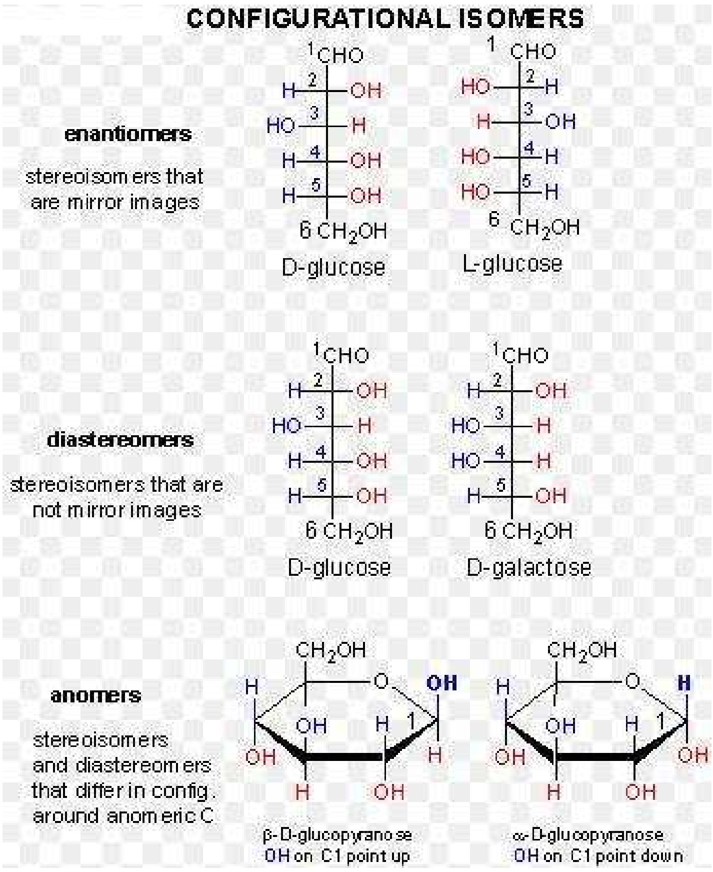
Epimers

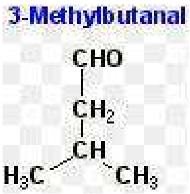


Anomers



Epimer Anomer comparison

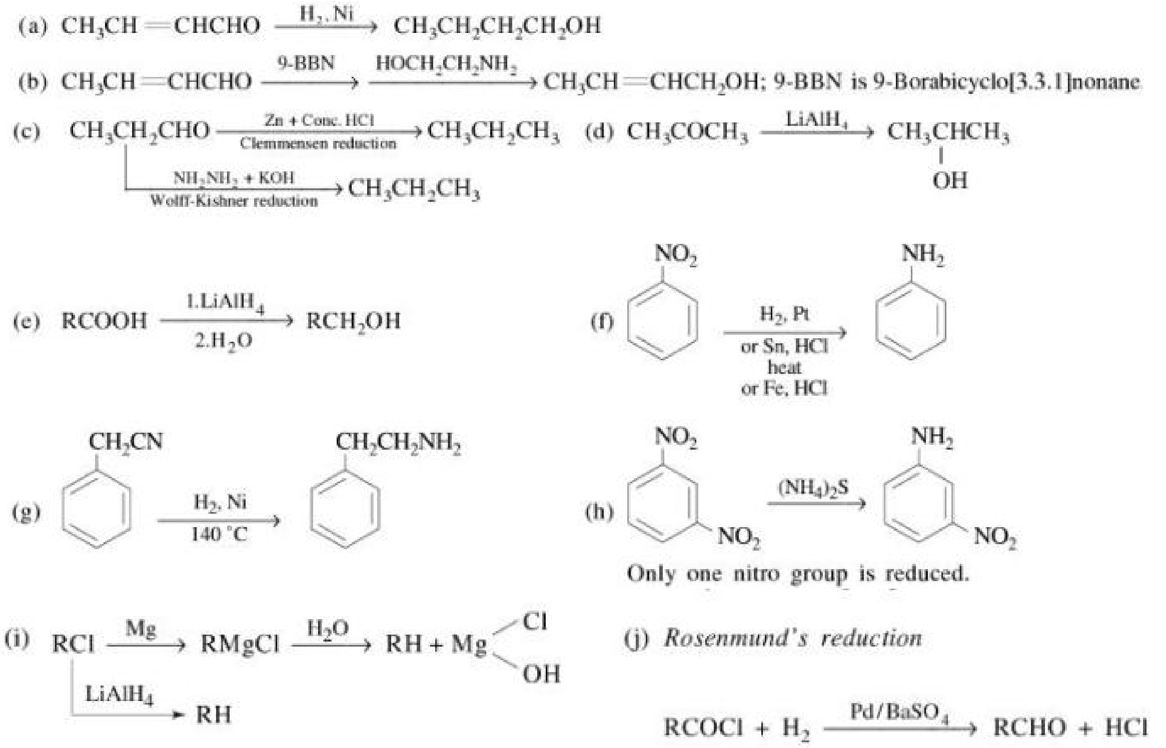




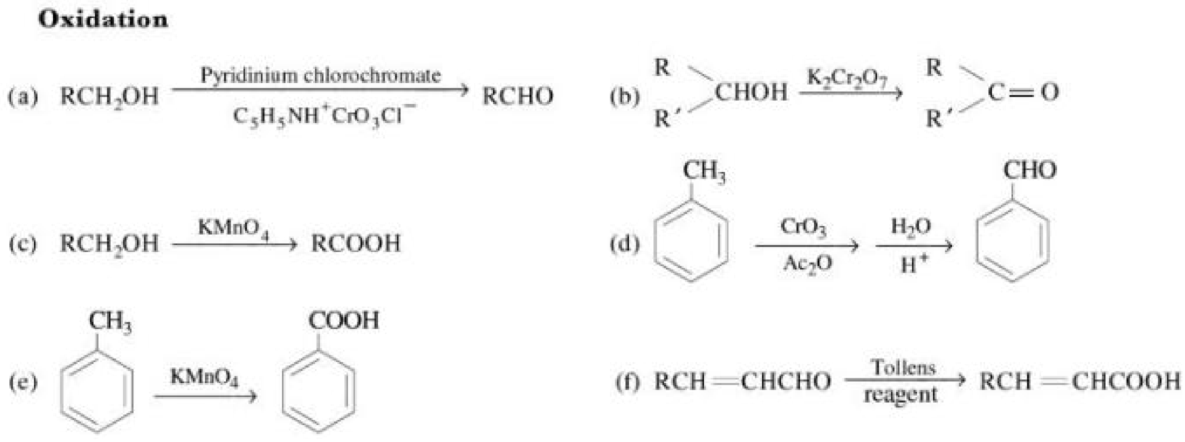
Epimers are monosaccharides that differ in chirality at only one carbon. In the straight-chain format, epimers will have H and OH-substituents switched at one backbone carbon, but not at any others.

Anomers are special epimers; in cyclic forms of one single monosaccharide, anomers differ in chirality at the anomeric (hemiacetal) carbon only. In the straight-chain format, anomers will have the same configuation.

Various Methods of Reduction used in Organic Chemistry

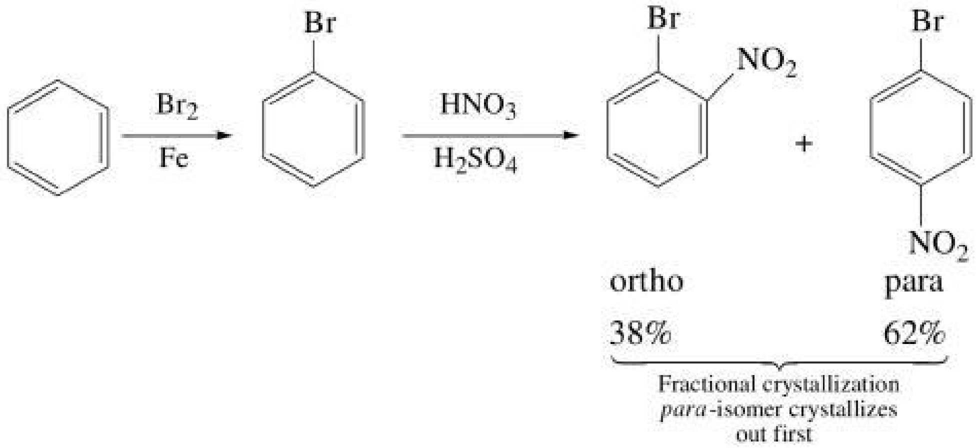


Various methods of Oxidation

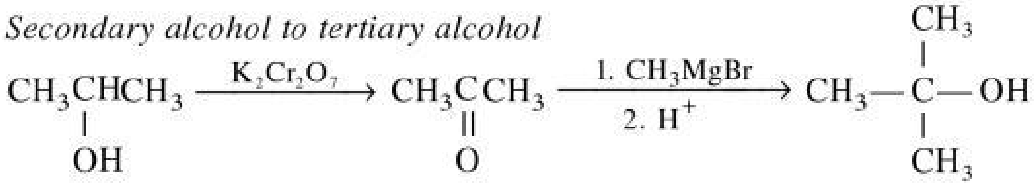


Synthesis of p-bromonitrobenzene in 2 steps

a synthesis of *p*-bromonitrobenzene from benzene in two steps.



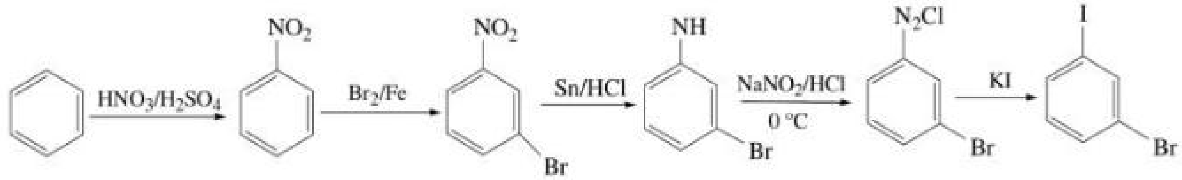
Secondary alcohol to tertiary alcohol



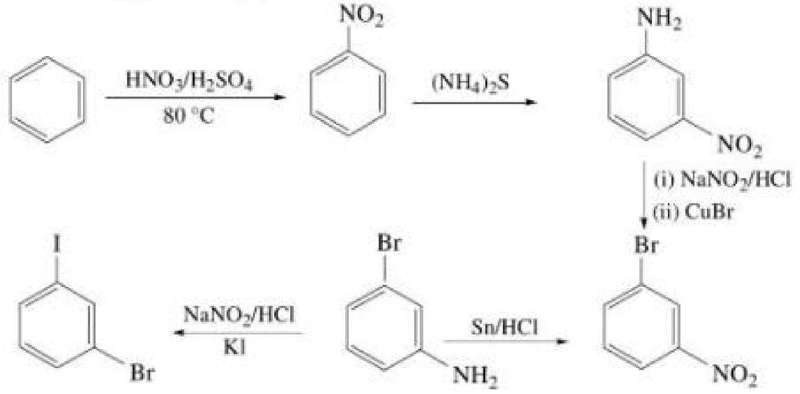
$$CH_3\underset{OH}{\underset{|}{C}H}CH_3 \xrightarrow{K_2Cr_2O_7} CH_3\underset{O}{\underset{\|}{C}}CH_3 \xrightarrow[2.\ H^+]{1.\ CH_3MgBr} CH_3-\overset{CH_3}{\overset{|}{\underset{CH_3}{\underset{|}{C}}}}-OH$$

Prepare m-bromoiodobenzene from benzene

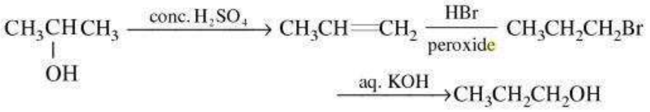
prepare *m*-bromoiodobenzene from benzene (in not more than 5-7 steps)?



Alternatively, we may proceed as follows:



Secondary alcohol to Primary alcohol

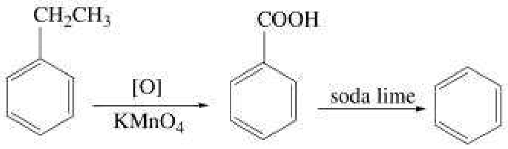


$$CH_3\underset{\displaystyle OH}{\underset{|}{C}H}CH_3 \xrightarrow{\text{conc. } H_2SO_4} CH_3CH{=}CH_2 \xrightarrow[\text{peroxide}]{HBr} CH_3CH_2CH_2Br \xrightarrow{\text{aq. KOH}} CH_3CH_2CH_2OH$$

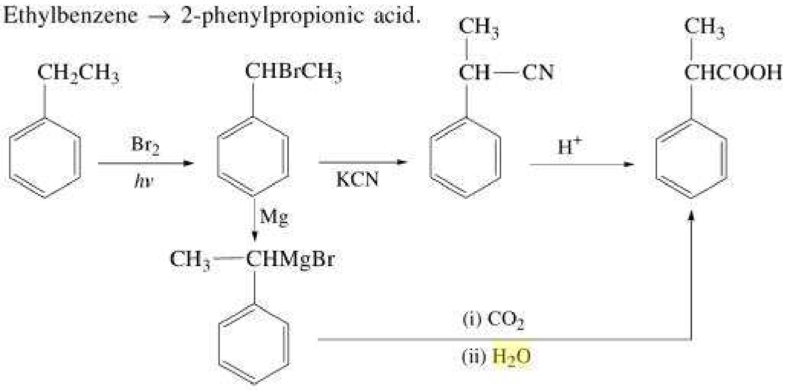
To convert Ethylebenzene to Benzene and Ethylebenzene to phenylpropionic acid

Show the steps to carry out the following transformations.

(i) Ethylbenzene → benzene



(ii) Ethylbenzene → 2-phenylpropionic acid.

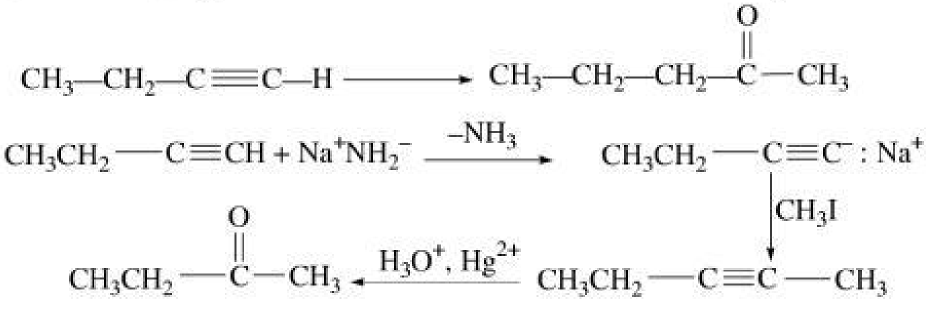


Alcohol to alcohol with one carbon more

$$CH_3OH \xrightarrow{PBr_3} CH_3Br \xrightarrow{KCN} CH_3CN \xrightarrow{[H]} CH_3CH_2NH_2 \xrightarrow{HNO_2} CH_3CH_2OH$$

Carryout the transformations 1-Butyne to ketone

Carry out the following transformation in not more than three steps.



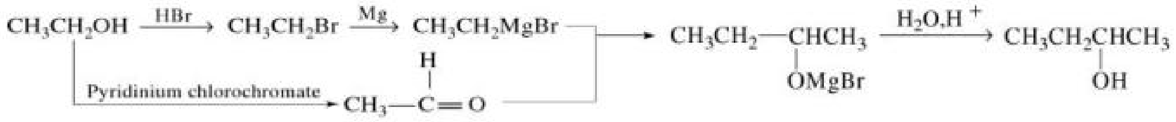
$$CH_3{-}CH_2{-}C{\equiv}C{-}H \longrightarrow CH_3{-}CH_2{-}CH_2{-}\overset{\displaystyle O}{\overset{\|}{C}}{-}CH_3$$

$$CH_3CH_2{-}C{\equiv}CH + Na^+NH_2^- \xrightarrow{-NH_3} CH_3CH_2{-}C{\equiv}C^-:Na^+ \xrightarrow{CH_3I} CH_3CH_2{-}C{\equiv}C{-}CH_3 \xrightarrow{H_3O^+,\ Hg^{2+}} CH_3CH_2{-}\overset{\displaystyle O}{\overset{\|}{C}}{-}CH_3$$

Alcohol to alcohol with one carbon less

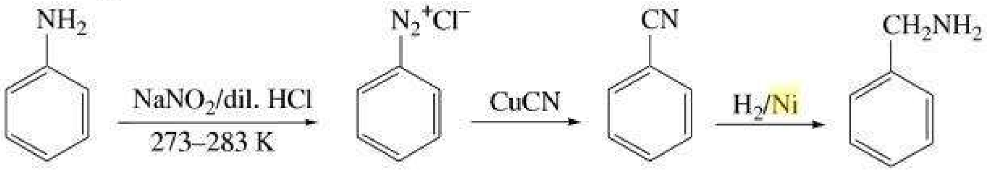
$$CH_3CH_2OH \xrightarrow[K_2Cr_2O_7]{[O]} CH_3COOH \xrightarrow{\text{soda lime}} CH_4 \xrightarrow[hv]{Cl_2} CH_3Cl \xrightarrow{\text{aq. KOH}} CH_3OH$$

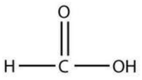
Simple alcohol to complicated alcohol



Conversion of Aniline to Benzylamine

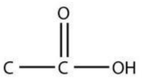
How would you bring about the following conversion (in 3 steps)?
Aniline → Benzylamine





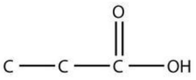
$H—\overset{\overset{O}{\|}}{C}—OH$

Methanoic acid



$C—\overset{\overset{O}{\|}}{C}—OH$

Ethanoic acid



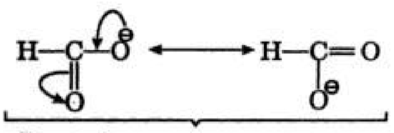
$C—C—\overset{\overset{O}{\|}}{C}—OH$

Propanoic acid

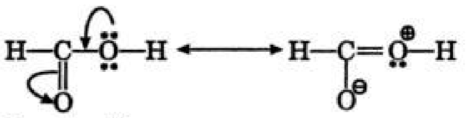
(i) Carbon-oxygen bond lengths in formic acid are different but are the same in sodium formate. Justify.

(ii) Phenate ion has more number of contributing structures than benzoate ion: but still benzoic acid is a stronger acid. Explain.

Solution. (*i*) In sodium formate, the contributing structures for the anion are equivalent while these are not the same in formic acid.



Formate ion



Formate acid

Thus, carbon-oxygen bond length in formate ion is the same for both the bonds while these are different in formic acid.

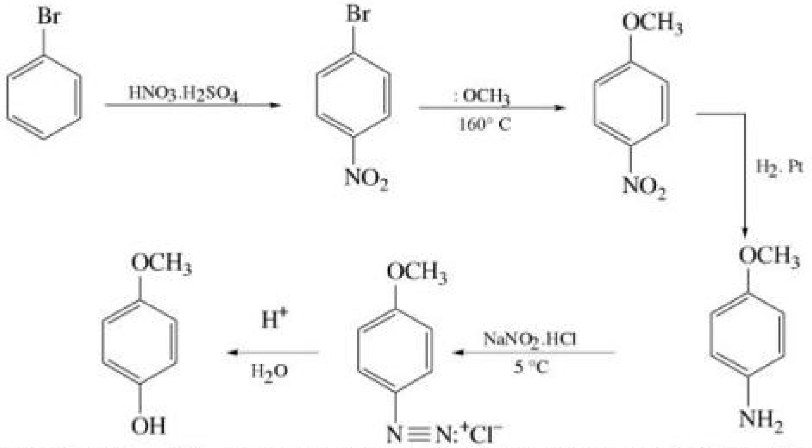
(*ii*) In phenate ion, the negative charge is dispersed only one electronegative oxygen atoms while there are two available oxygen atoms in benzoate ion to disperse the negative charge.

Amine to amine containing one carbon more

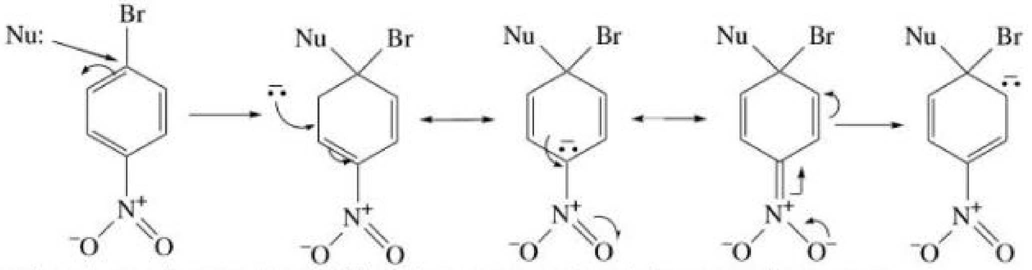
$$RCH_2NH_2 \xrightarrow[0\ ^\circ C]{HNO_2} RCH_2OH \xrightarrow{PBr_3} RCH_2Br \xrightarrow{NaCN} RCH_2CN \xrightarrow{H_2/Ni} RCH_2CH_2NH_2$$

Synthesis of 4-methoxyphenol from bromobenzene

How would you synthesise 4-methoxyphenol from bromobenzene is NOT more than five steps? State clearly the reagents used in each step and show the structure of the intermediate compounds in your synthetic scheme.

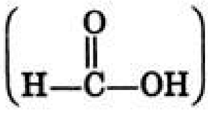
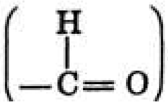


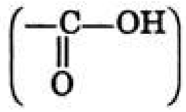
The introduction of —NO_2 group at *para* position in bromobenzene facilitates the nucleophilic replacement of —Br by —OCH_3.



The negative charge is strongly stabilised by resonance involving *para*-nitro group.

Write chemical tests to distinguish between:

Solution. Formic acid  $\left(H-\overset{\displaystyle O}{\overset{\|}{C}}-OH\right)$ contains both an aldehyde  $\left(\overset{\displaystyle H}{\overset{|}{-C}}=O\right)$ as well as carboxyl

group  $\left(-\underset{\underset{O}{\|}}{C}-OH\right)$ **but acetic acid contain only a carboxyl group. Formic acid behaves as reducing agent whereas acetic acid does not.**

(*a*) Formic acid reduces Tollen's reagent to metallic silver but acetic acid does not.

$$\underset{\text{Tollen's reagent}}{HCOOH + 2[Ag(NH_3)_2]^+ + 2(OH)^-} \longrightarrow \underset{\text{silver mirror}}{2Ag\downarrow} + CO_2\uparrow + 2H_2O + 4NH_3$$

No silver mirror is formed with acetic acid.

(*b*) Formic acid reduced Fehling solution to red ppt. of Cu_2O but acetic acid does not.

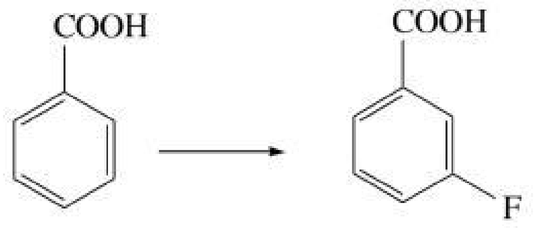
$$HCOOH + \underset{\text{Fehling solution}}{2Cu^{2+} + 4(OH)^-} \longrightarrow \underset{\substack{\text{red ppt}\\ \text{(cuprous oxide)}}}{Cu_2O\downarrow} + CO_2\uparrow + 3H_2O$$

Amine to amine containing one carbon less

$$RCH_2NH_2 \xrightarrow[0°C]{HNO_2} RCH_2OH \xrightarrow[KMnO_4]{[O]} RCOOH \xrightarrow{SOCl_2} RCOCl \xrightarrow{NH_3} RCONH_2 \xrightarrow{OBr^-} RNH_2$$

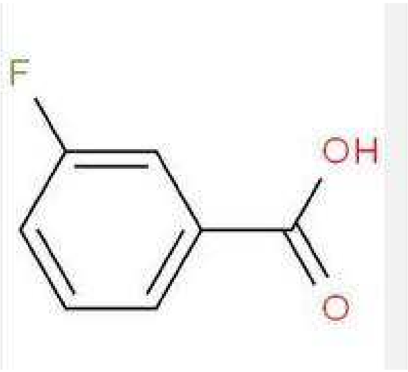
Convert Benzoic acid to meta fluorobenzoic acid (2-fluorobenzoic acid)

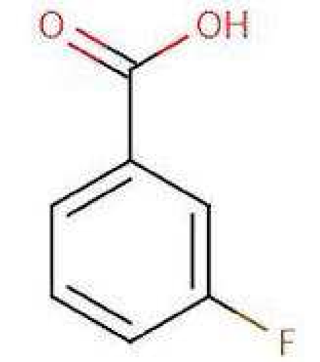
Convert

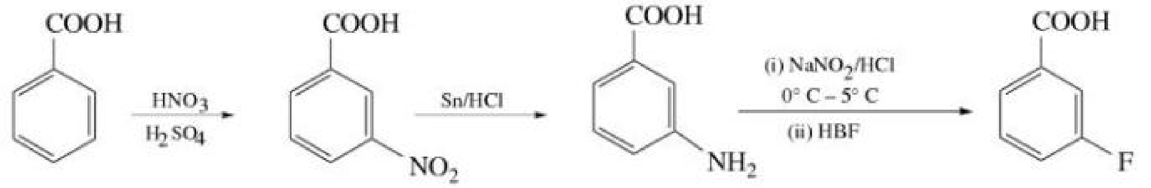


(not more than 3 steps)

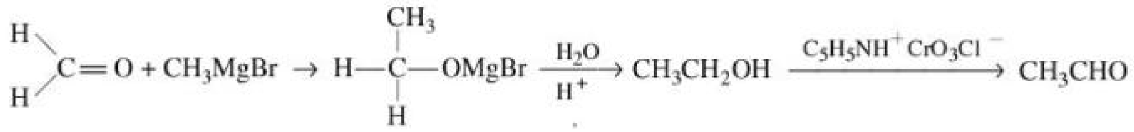
metafluoro benzoicacid





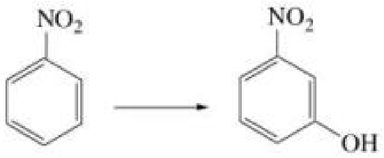


Aldehyde to aldehyde containing one carbon more



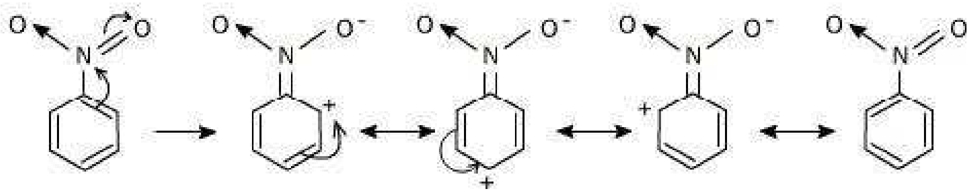
Convert Nitrobenzene to meta hydroxynitrobenzene (or 4-nitrophenol)

Convert

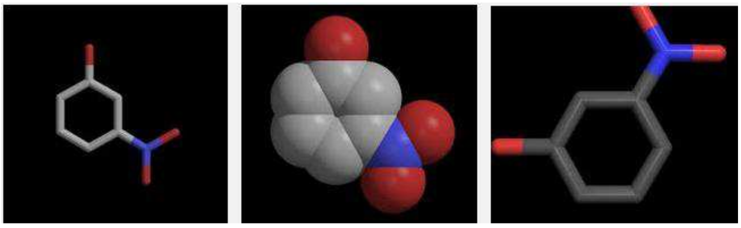


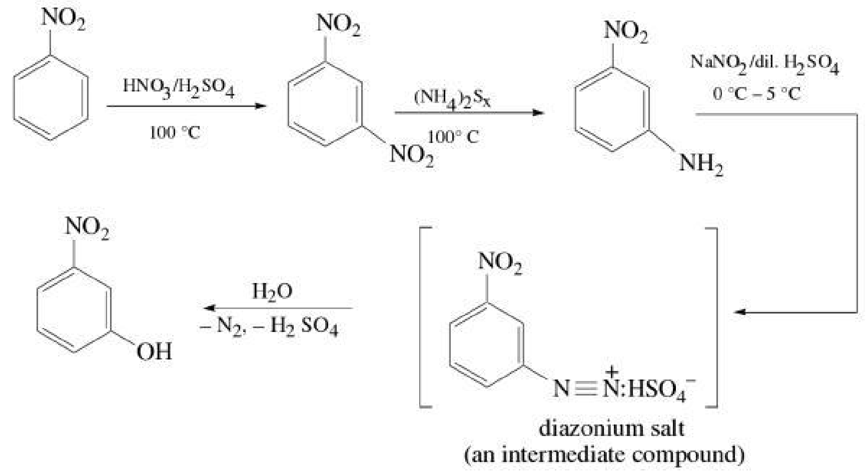
in not more than 4 steps mentioning the reagents, reaction conditions and intermediates in the above conversion.
The given conversion can be carried out as follows.

Nitrobenzene resonance structure

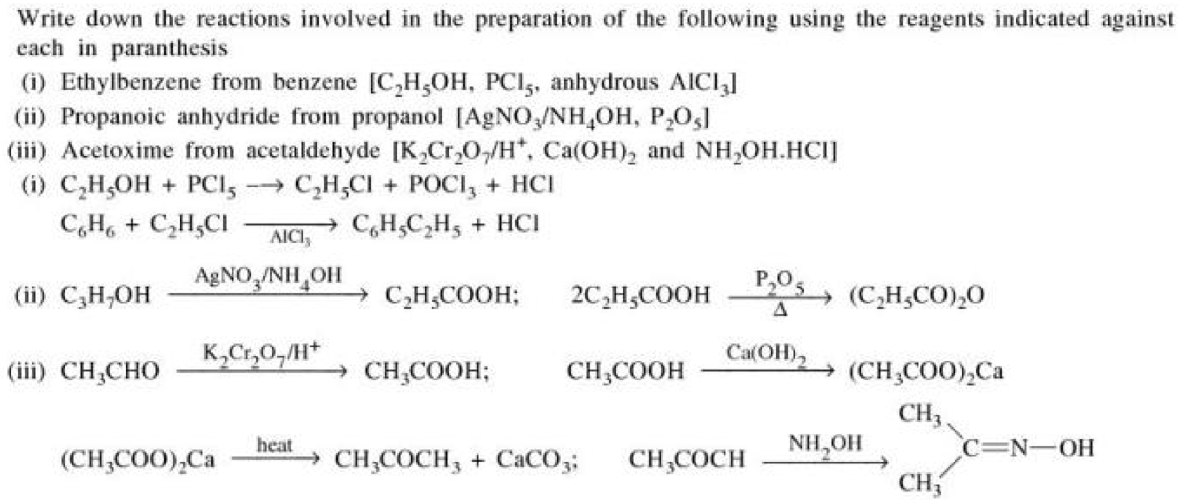


4 nitrophenol or metahydroxynitrobenzene





Many conversions or preparations



Write down the reactions involved in the preparation of the following using the reagents indicated against each in paranthesis

(i) Ethylbenzene from benzene [C_2H_5OH, PCl_5, anhydrous $AlCl_3$]

(ii) Propanoic anhydride from propanol [$AgNO_3/NH_4OH$, P_2O_5]

(iii) Acetoxime from acetaldehyde [$K_2Cr_2O_7/H^+$, $Ca(OH)_2$ and $NH_2OH.HCl$]

(i) $C_2H_5OH + PCl_5 \longrightarrow C_2H_5Cl + POCl_3 + HCl$

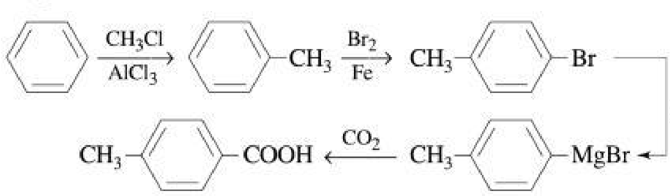
$C_6H_6 + C_2H_5Cl \xrightarrow[AlCl_3]{} C_6H_5C_2H_5 + HCl$

(ii) $C_3H_7OH \xrightarrow{AgNO_3/NH_4OH} C_2H_5COOH$; $2C_2H_5COOH \xrightarrow[\Delta]{P_2O_5} (C_2H_5CO)_2O$

(iii) $CH_3CHO \xrightarrow{K_2Cr_2O_7/H^+} CH_3COOH$; $CH_3COOH \xrightarrow{Ca(OH)_2} (CH_3COO)_2Ca$

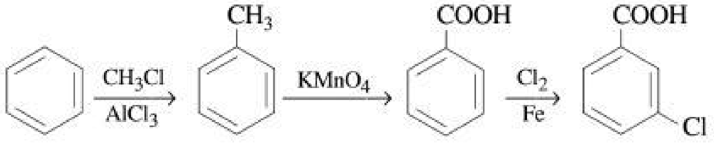
$(CH_3COO)_2Ca \xrightarrow{heat} CH_3COCH_3 + CaCO_3$; $CH_3COCH \xrightarrow{NH_2OH} \begin{matrix} CH_3 \\ CH_3 \end{matrix}\!\!>C{=}N{-}OH$

Benzene to p-toluic acid



Benzene to *p*-toluic acid

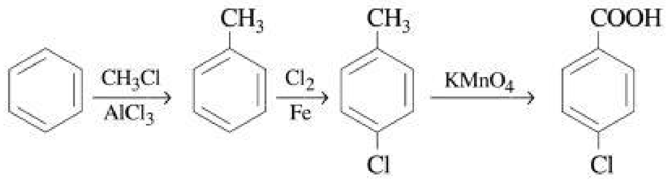
Benzene to m-chlorobenzoic acid



Benzene to *m*-chlorobenzoic acid

Benzene to p-chlorobenzoic acid

Benzene to *p*-chlorobenzoic acid



Acetic acid to Malonic acid

Acetic acid to malonic acid

$$CH_3COOH \xrightarrow{Br_2,\,P} BrCH_2COOH \xrightarrow{KCN} CNCH_2COOH \xrightarrow[H^+]{H_2O} HOOCCH_2COOH$$

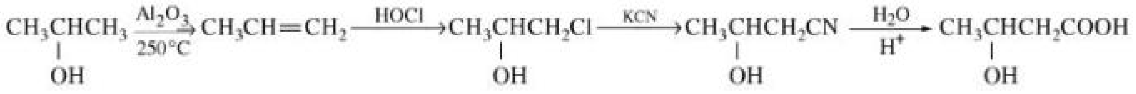
Ethyne to propanoic acid

Ethyne to propanoic acid

$$HC{\equiv}CH \xrightarrow[\text{catalyst}]{\text{Lindlar}} H_2C{=}CH_2 \xrightarrow{HCl} CH_3CH_2Cl \xrightarrow{KCN} C_2H_5CN \xrightarrow[H^+]{H_2O} C_2H_5COOH$$

Isopropyl alcohol to ß-hydroxybutyric acid

Isopropyl alcohol to β-hydroxybutyric acid



$$CH_3\underset{\substack{|\\OH}}{CH}CH_3 \xrightarrow[250^\circ C]{Al_2O_3} CH_3CH{=}CH_2 \xrightarrow{HOCl} CH_3\underset{\substack{|\\OH}}{CH}CH_2Cl \xrightarrow{KCN} CH_3\underset{\substack{|\\OH}}{CH}CH_2CN \xrightarrow[H^+]{H_2O} CH_3\underset{\substack{|\\OH}}{CH}CH_2COOH$$

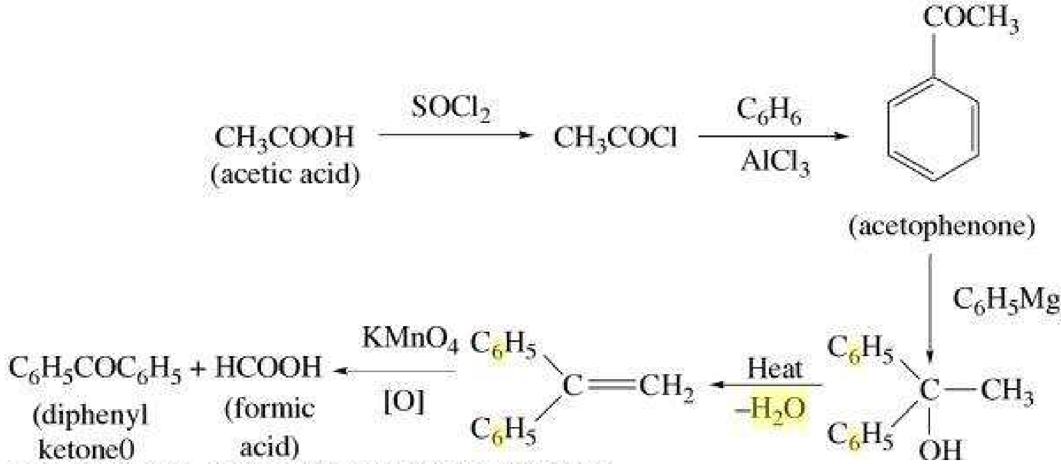
Acetic acid to Propanoic acid

Acetic acid to propanoic acid

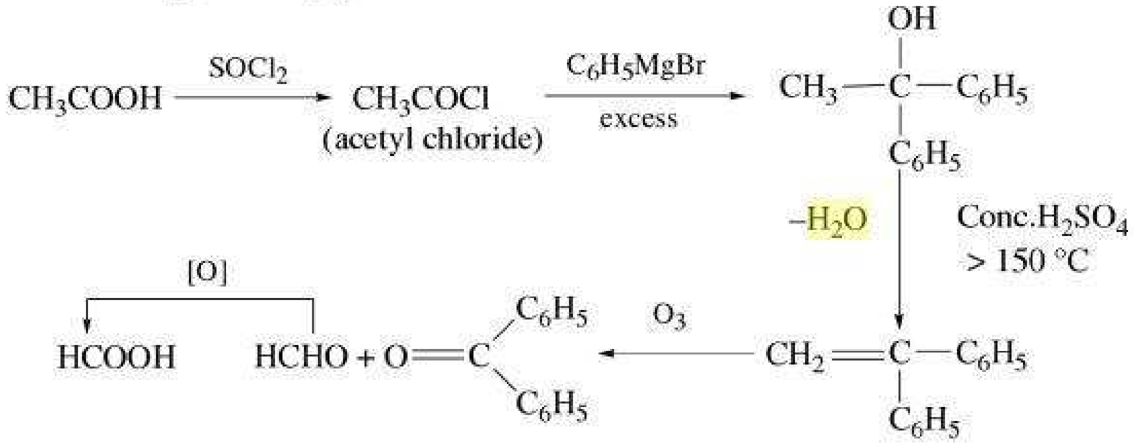
$$CH_3COOH \xrightarrow{LiAlH_4} CH_3CH_2OH \xrightarrow{PCl_5} CH_3CH_2Cl \xrightarrow{KCN} CH_3CH_2CN \xrightarrow[H^+]{H_2O} CH_3CH_2COOH$$

Ethanoic acid to a mixture of methanoic acid and diphenyl ketone

Ethanoic acid to a mixture of methanoic acid and diphenyl ketone



Alternatively, we may proceed as follows.

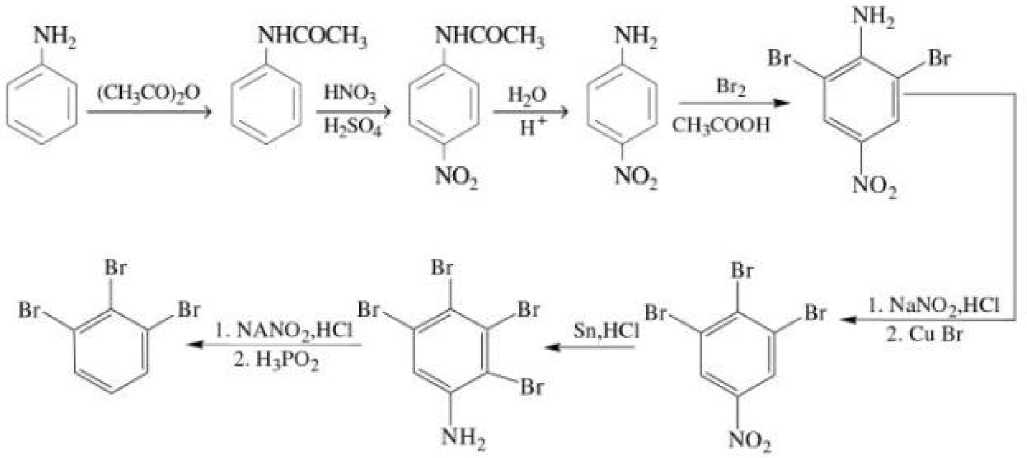


Propanoic acid to acetic acid

Propanoic acid to acetic acid

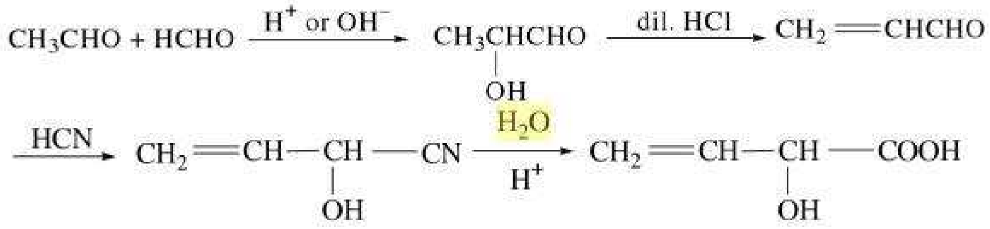
$$CH_3CH_2COOH \xrightarrow[heat]{NH_3} CH_3CH_2CONH_2 \xrightarrow{Br_2/KOH} CH_3CH_2NH_2 \xrightarrow{HNO_2} CH_3CH_2OH \xrightarrow[H^+]{Cr_2O_7^{2-}} CH_3COOH$$

Aniline to 1, 2, 3-tribromobenzene



Ethanal to 2-hydroxy-3-butenoic acid

Ethanal to 2-hydroxy-3-butenoic acid

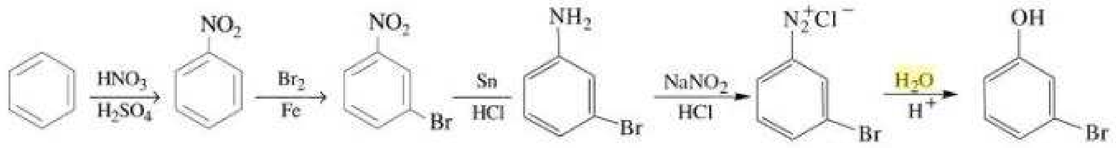


$$CH_3CHO + HCHO \xrightarrow{H^+ \text{ or } OH^-} CH_3\underset{\displaystyle OH}{\underset{|}{C}}HCHO \xrightarrow{\text{dil. HCl}} CH_2{=}CHCHO$$

$$\xrightarrow{HCN} CH_2{=}CH{-}\underset{\displaystyle OH}{\underset{|}{CH}}{-}CN \xrightarrow[H^+]{H_2O} CH_2{=}CH{-}\underset{\displaystyle OH}{\underset{|}{CH}}{-}COOH$$

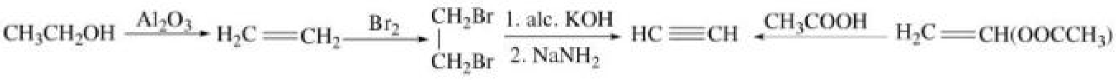
Benzene to m-bromophenol

Benzene to *m*-bromophenol



Ethyl alcohol to Vinyl alcohol

Ethyl alcohol to vinyl acetate



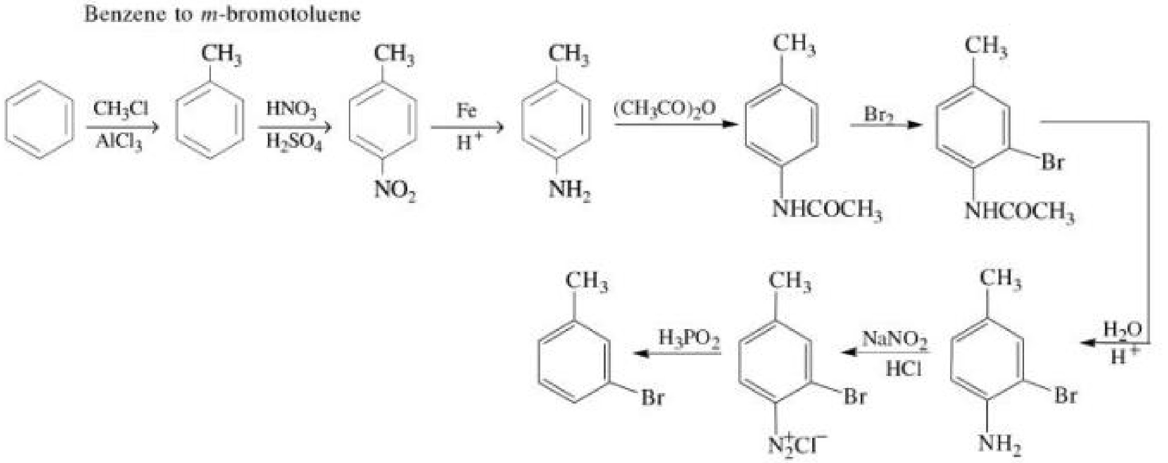
$$CH_3CH_2OH \xrightarrow{Al_2O_3} H_2C{=}CH_2 \xrightarrow{Br_2} \begin{matrix} CH_2Br \\ | \\ CH_2Br \end{matrix} \xrightarrow[\text{2. } NaNH_2]{\text{1. alc. KOH}} HC{\equiv}CH \xleftarrow{CH_3COOH} H_2C{=}CH(OOCCH_3)$$

Methylamine to ethylamine

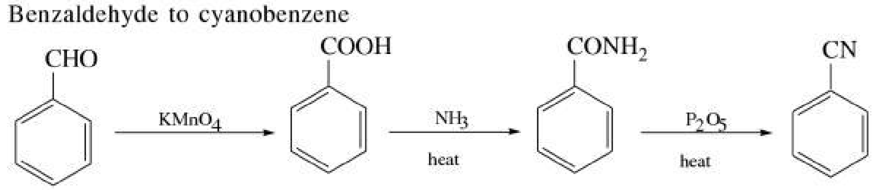
Methylamine to ethylamine

$$CH_3NH_2 \xrightarrow{HNO_2} CH_3OH \xrightarrow{PCl_3} CH_3Cl \xrightarrow{KCN} CH_3CN \xrightarrow{Sn/HCl} CH_3CH_2NH_2$$

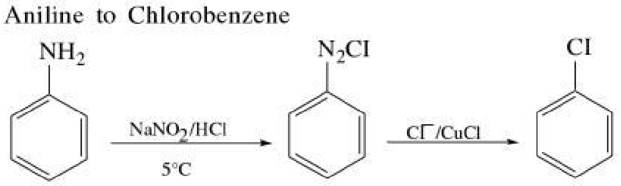
Benzene to m-bromotoluene



Benzaldehyde to cyanobenzene



Aniline to Chlorobenzene

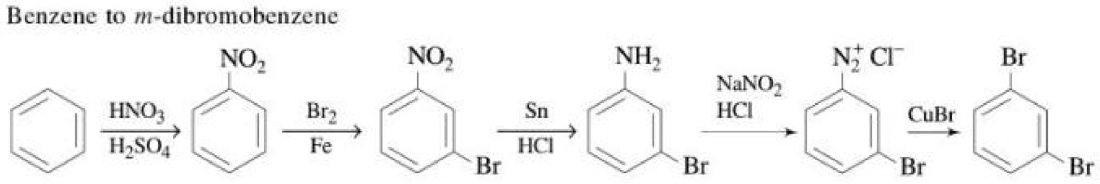


Ethylamine to Methylamine

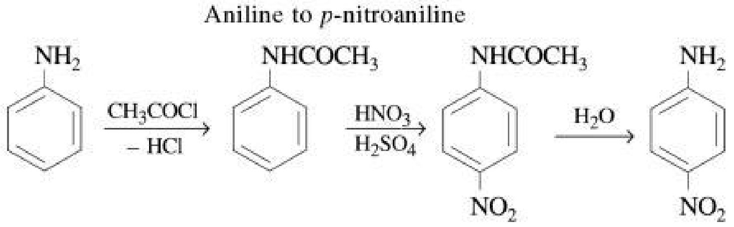
Ethylamine to methylamine

$$CH_3CH_2NH_2 \xrightarrow{HNO_2} CH_3CH_2OH \xrightarrow[H^+]{Cr_2O_7^{2-}} CH_3COOH \xrightarrow[heat]{NH_3} CH_3CONH_2 \xrightarrow{Br_2/KOH} CH_3NH_2$$

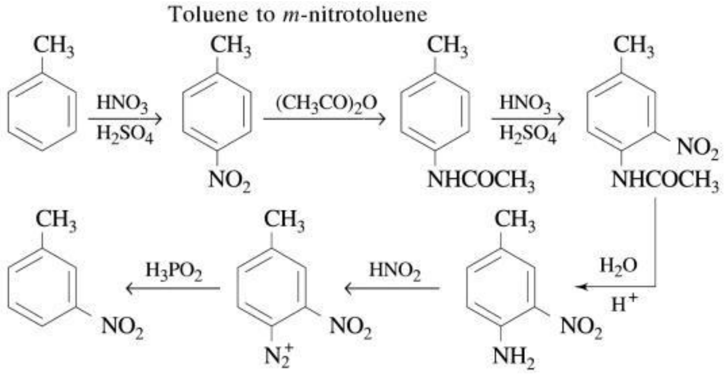
Benzene to m-dibromobenzene



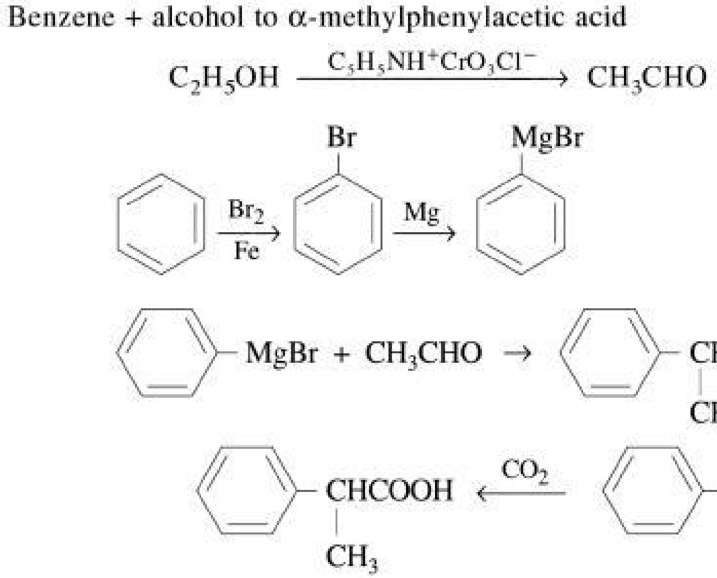
Aniline to p-nitroaniline

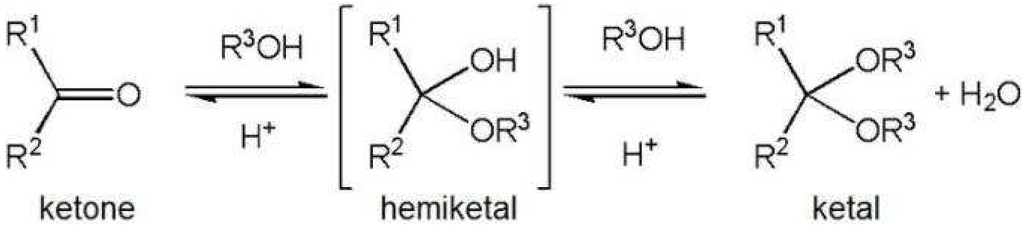


Toluene to m-nitrotoluene

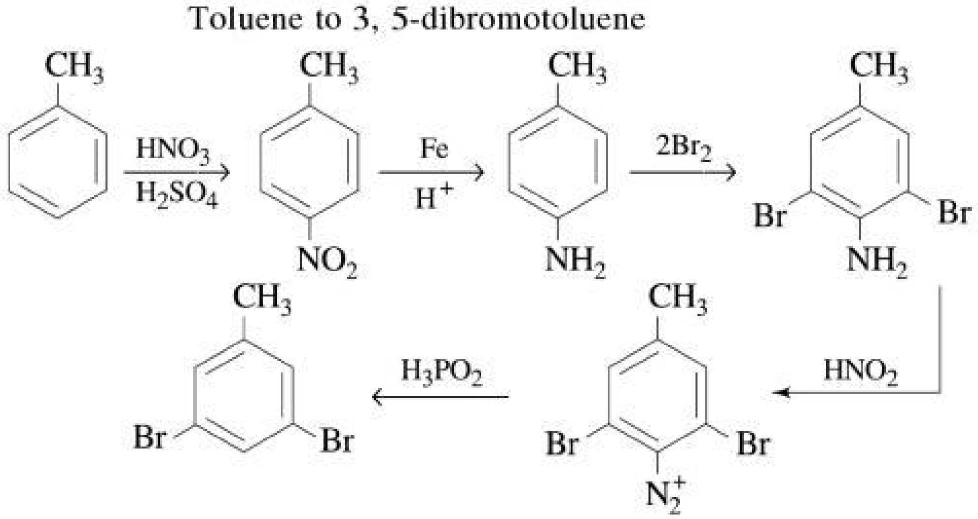


Convert Benzene + Alcohol to α-methylphenylacetic acid

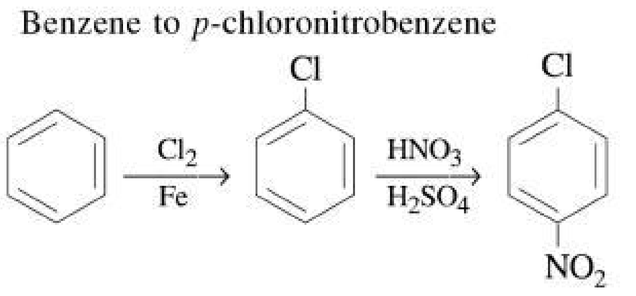

Benzene + alcohol to α-methylphenylacetic acid



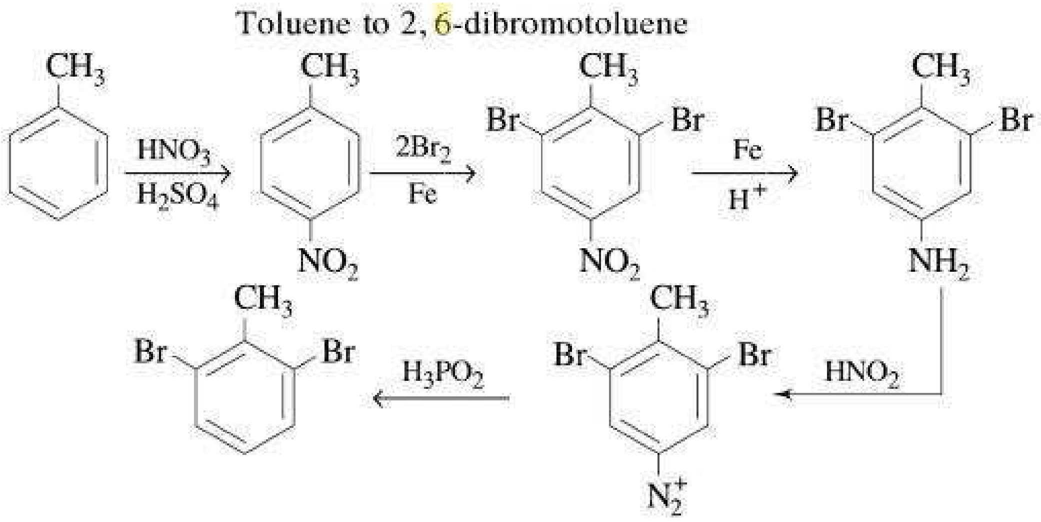
Toluene to 3,5 dibromotoluene


Toluene to 3, 5-dibromotoluene

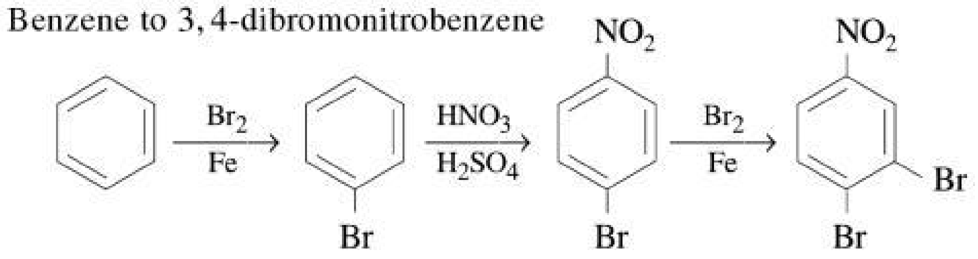
Benzene to p-chloronitrobenzene



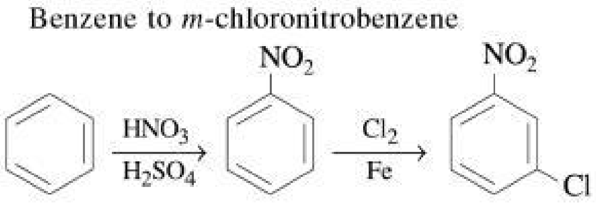
Toluene to 2,6 dibromotoluene



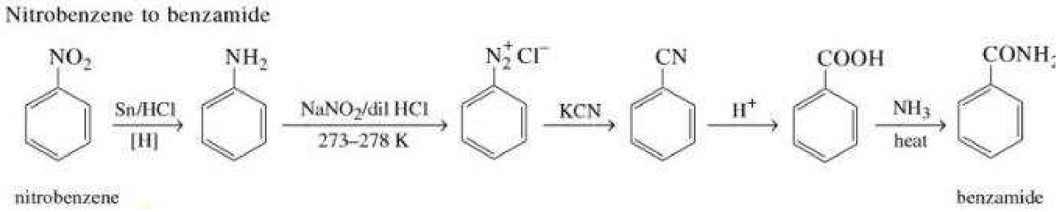
Benzene to 3,4-dibromonitrobenzene



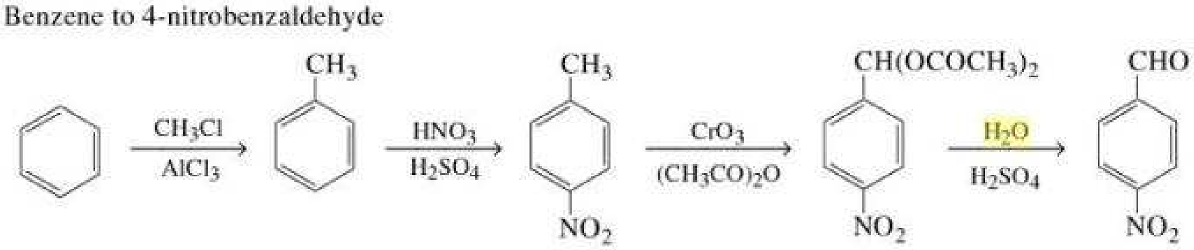
Benzene to m-chloronitrobenzene



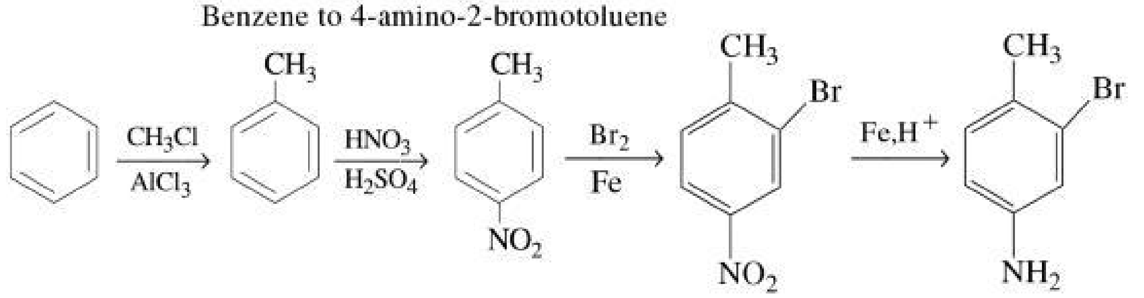
Nitrobenzene to benzamide

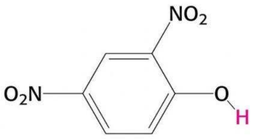


Benzene to 4-nitrobenzaldehyde



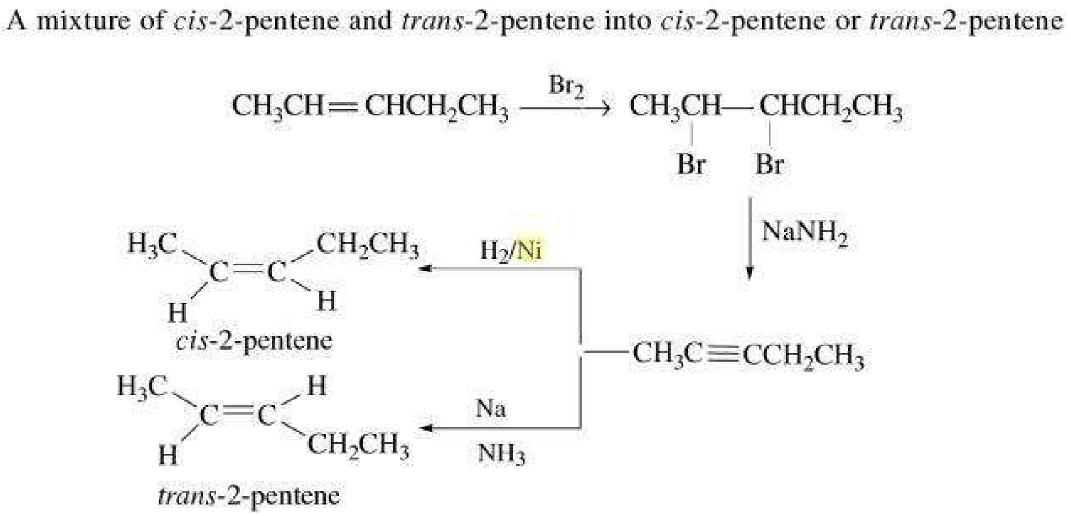
Benzene to 4-amino-2-bromotoluene

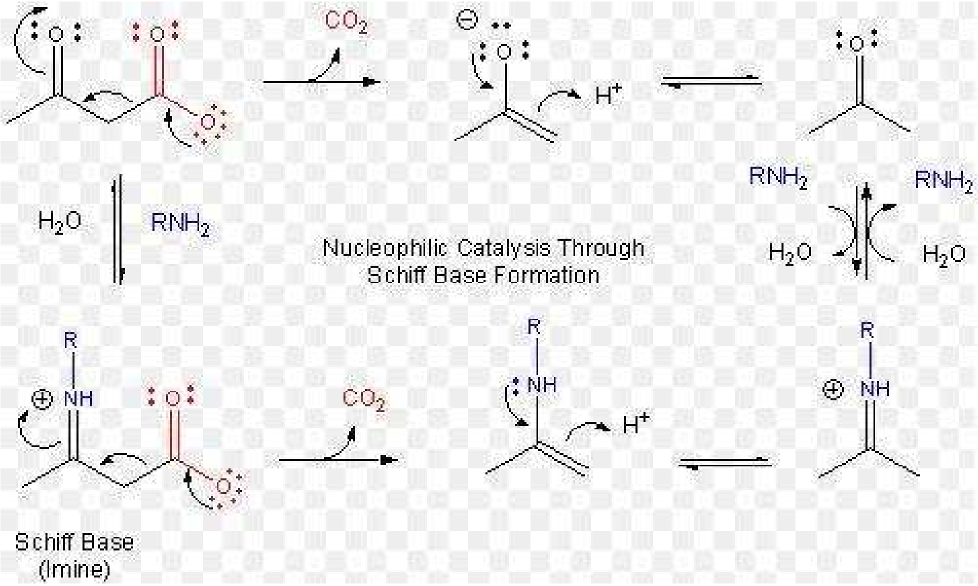




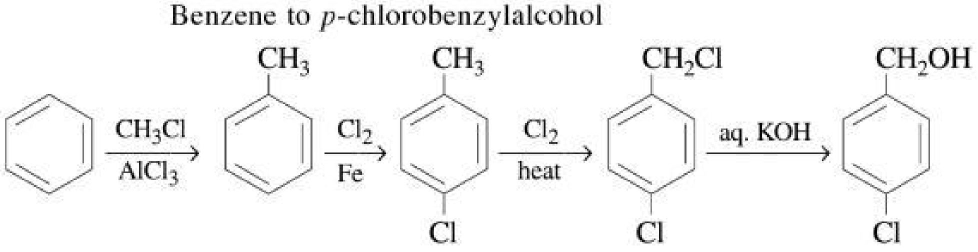
2,4-Dinitrophenol (DNP)

Convert a mixture of cis-2-pentene and trans-2-pentene into cis-2-pentene or trans-2-pentene



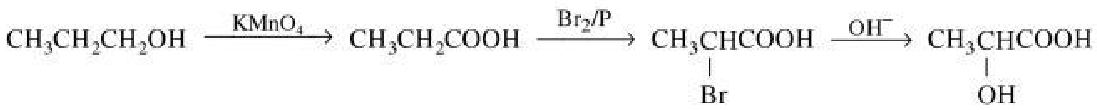


Benzene to p-chlorobenzylalcohol



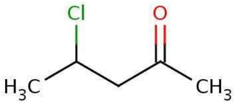
Convert Propanol to α-hydroxypropanoic acid

Propanol to α-hydroxypropionic acid

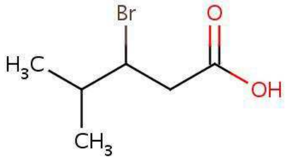


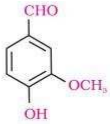
$$CH_3CH_2CH_2OH \xrightarrow{KMnO_4} CH_3CH_2COOH \xrightarrow{Br_2/P} CH_3\underset{\underset{Br}{|}}{C}HCOOH \xrightarrow{OH^-} CH_3\underset{\underset{OH}{|}}{C}HCOOH$$

4-chloropentane-2-one

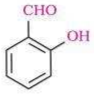


3-Bromo-4-methylpentanoic acid

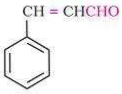




Vanillin

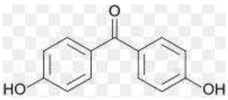


Salicylaldehyde

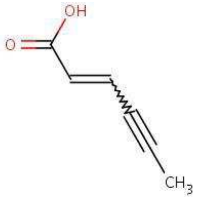


Cinnamaldehyde

p,p'- dihydroxybenzophenone

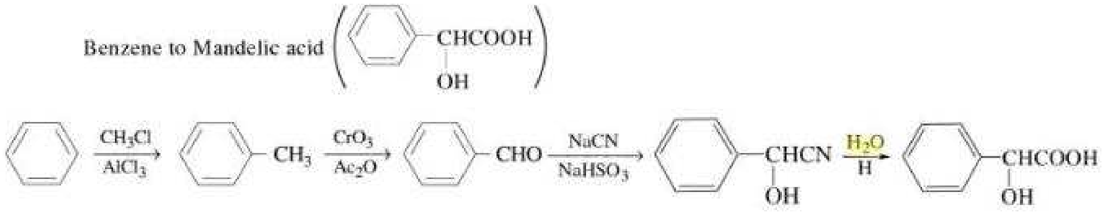


Hex-2-ene-4-ynoicacid

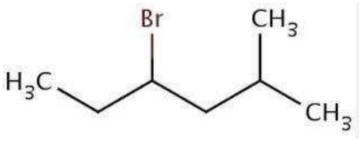


Heptan-2-one

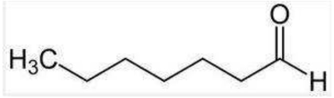
Benzene to Mandelic Acid



4-Bromo-2-methylhexane

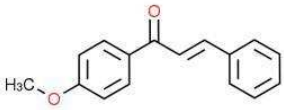


Heptanal

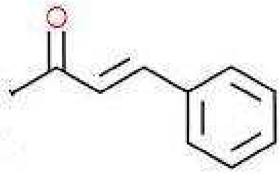


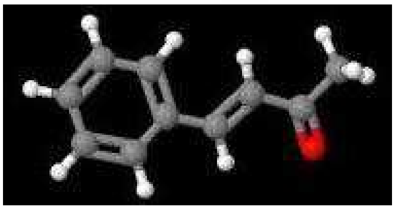
1-(4-Methoxyphenyl)-3-phenylprop-2-en-1-one

1-(4-Methoxyphenyl)-3-phenylprop-2-en-1-one

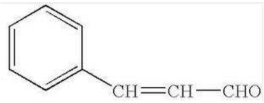


4-phenylprop-3-ene-2-one

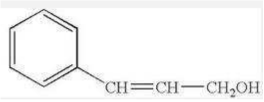




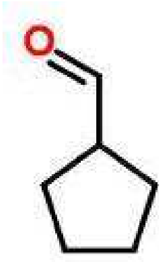
3-phenylprop-2-ene-1-al



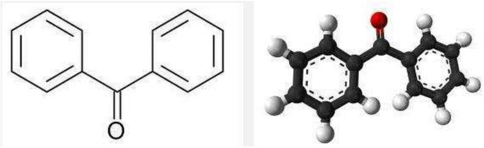
3-phenyl-prop-2-ene-1-ol



cyclopentancarbaldehyde



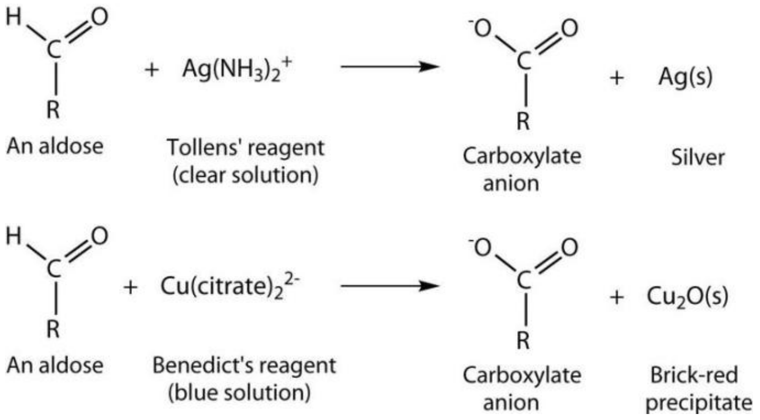
Benzophenone also diphenylmethanone



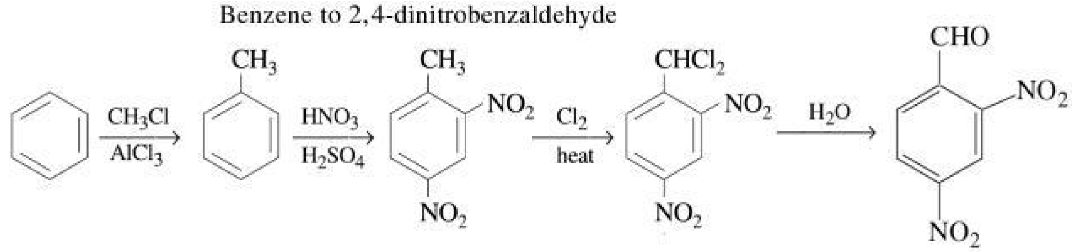
Tollen’s Reagent

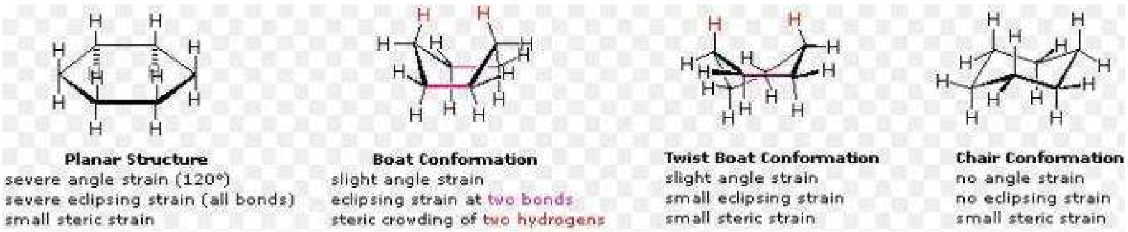
$$\underset{\text{An aldehyde}}{RCHO(aq)} + 2Ag(NH_3)_2^+(aq) + 3OH^-(aq) \longrightarrow$$

$$RCOO^-(aq) + \underset{\text{Free silver}}{2Ag(s)} + 4\,NH_3(aq) + 2H_2O$$

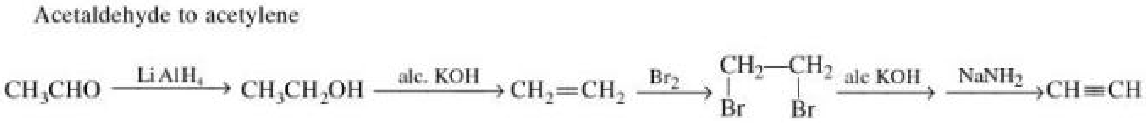


Benzene to 2,4-dinitrobenzaldehyde

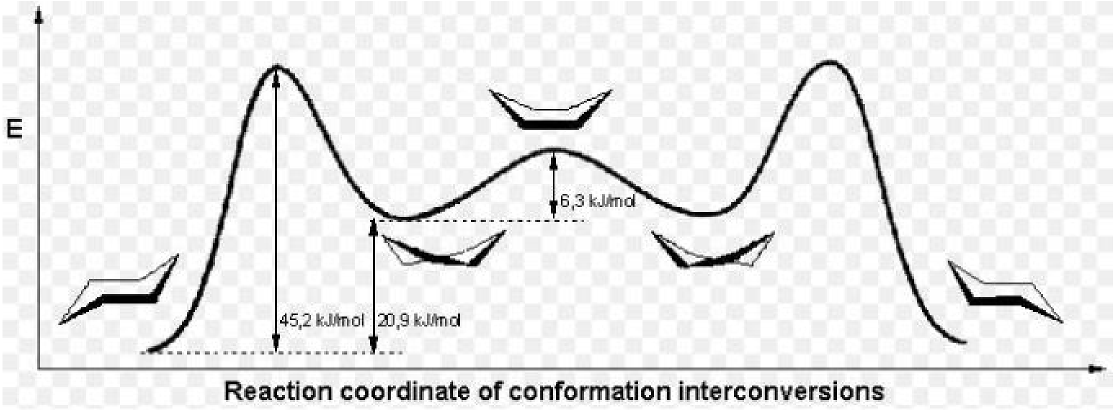




Acetaldehyde to Acetylene

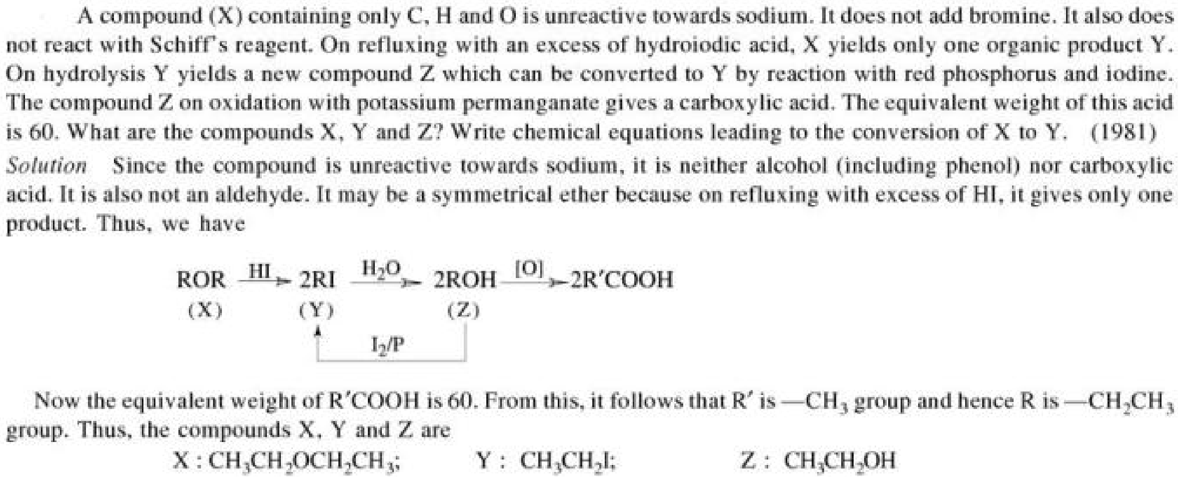


$$CH_3CHO \xrightarrow{LiAlH_4} CH_3CH_2OH \xrightarrow{alc.\ KOH} CH_2{=}CH_2 \xrightarrow{Br_2} \underset{Br}{CH_2}-\underset{Br}{CH_2} \xrightarrow{alc\ KOH} \xrightarrow{NaNH_2} CH{\equiv}CH$$



Identify the compounds

compound contains C, H and O unreactive towards Na (Sodium)



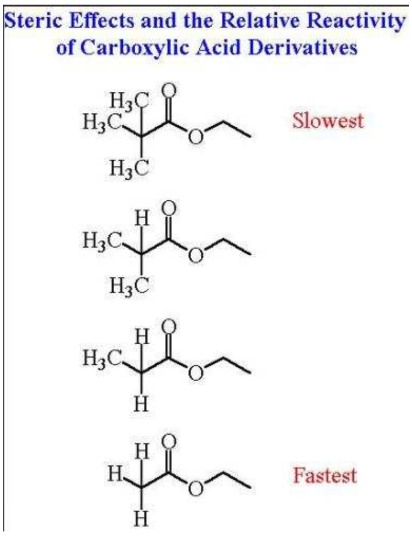
A compound (X) containing only C, H and O is unreactive towards sodium. It does not add bromine. It also does not react with Schiff's reagent. On refluxing with an excess of hydroiodic acid, X yields only one organic product Y. On hydrolysis Y yields a new compound Z which can be converted to Y by reaction with red phosphorus and iodine. The compound Z on oxidation with potassium permanganate gives a carboxylic acid. The equivalent weight of this acid is 60. What are the compounds X, Y and Z? Write chemical equations leading to the conversion of X to Y. (1981)

Solution Since the compound is unreactive towards sodium, it is neither alcohol (including phenol) nor carboxylic acid. It is also not an aldehyde. It may be a symmetrical ether because on refluxing with excess of HI, it gives only one product. Thus, we have

$$\begin{tikzcd} \underset{(X)}{ROR} \arrow[r, "HI"] & \underset{(Y)}{2RI} \arrow[r, "H_2O"] & \underset{(Z)}{2ROH} \arrow[r, "{[O]}"] \arrow[l, bend left, "I_2/P"] & 2R'COOH \end{tikzcd}$$

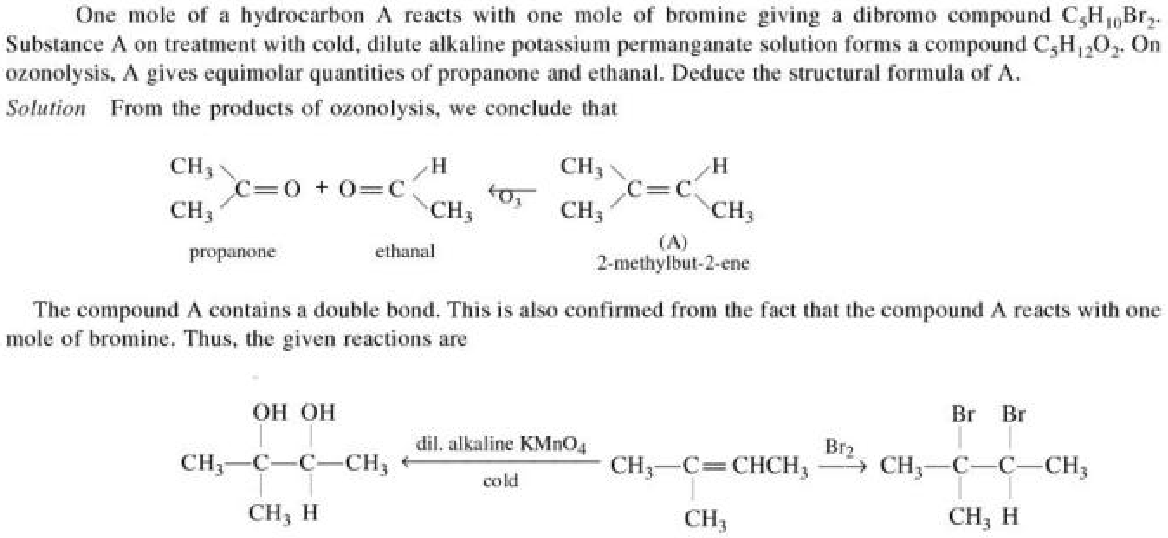
Now the equivalent weight of R′COOH is 60. From this, it follows that R′ is —CH_3 group and hence R is —CH_2CH_3 group. Thus, the compounds X, Y and Z are

X : $CH_3CH_2OCH_2CH_3$; Y : CH_3CH_2I; Z : CH_3CH_2OH



Identify the compound

A hydrocarbon reacts with Bromine to give dibromo compound



One mole of a hydrocarbon A reacts with one mole of bromine giving a dibromo compound $C_5H_{10}Br_2$. Substance A on treatment with cold, dilute alkaline potassium permanganate solution forms a compound $C_5H_{12}O_2$. On ozonolysis, A gives equimolar quantities of propanone and ethanal. Deduce the structural formula of A.

Solution From the products of ozonolysis, we conclude that

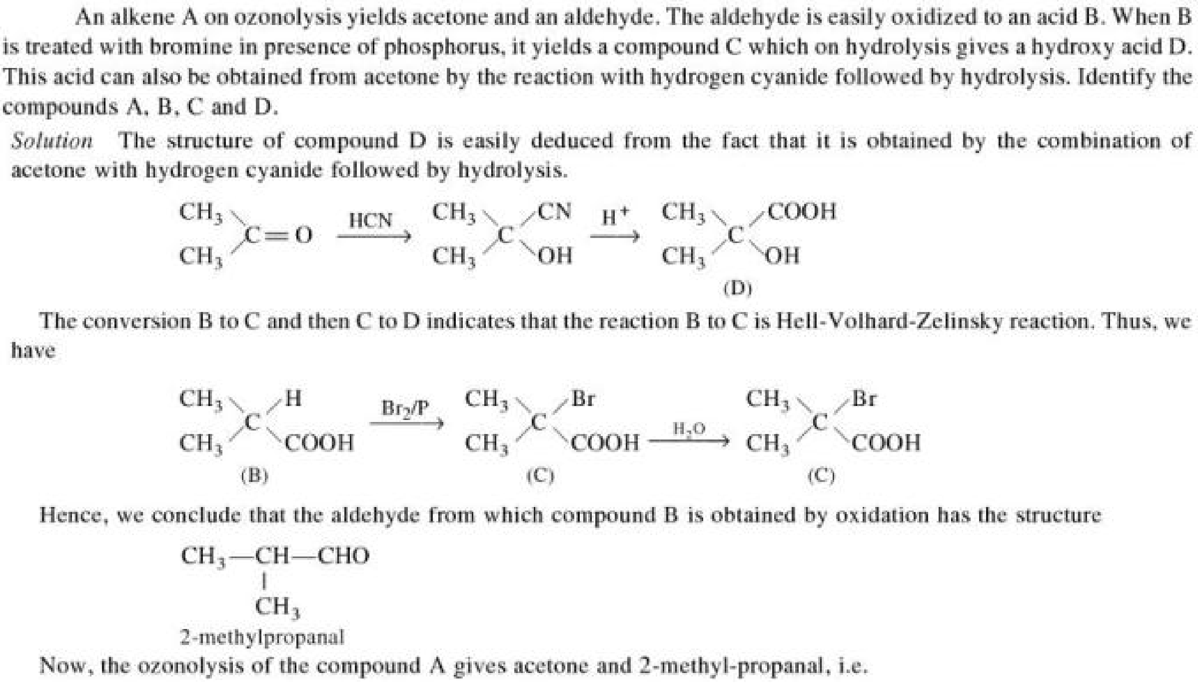
$$\underset{\text{propanone}}{\begin{matrix}CH_3\\CH_3\end{matrix}\!\!>C{=}O} + \underset{\text{ethanal}}{O{=}C\!<\!\!\begin{matrix}H\\CH_3\end{matrix}} \xleftarrow{O_3} \underset{\text{2-methylbut-2-ene}}{\underset{(A)}{\begin{matrix}CH_3\\CH_3\end{matrix}\!\!>C{=}C<\!\!\begin{matrix}H\\CH_3\end{matrix}}}$$

The compound A contains a double bond. This is also confirmed from the fact that the compound A reacts with one mole of bromine. Thus, the given reactions are

$$CH_3-\underset{CH_3}{\overset{OH}{C}}-\underset{H}{\overset{OH}{C}}-CH_3 \xleftarrow[\text{cold}]{\text{dil. alkaline } KMnO_4} CH_3-\underset{CH_3}{C}{=}CHCH_3 \xrightarrow{Br_2} CH_3-\underset{CH_3}{\overset{Br}{C}}-\underset{H}{\overset{Br}{C}}-CH_3$$

Identify the compound

An alkene A on ozonolysis yields acetone and an aldehyde



An alkene A on ozonolysis yields acetone and an aldehyde. The aldehyde is easily oxidized to an acid B. When B is treated with bromine in presence of phosphorus, it yields a compound C which on hydrolysis gives a hydroxy acid D. This acid can also be obtained from acetone by the reaction with hydrogen cyanide followed by hydrolysis. Identify the compounds A, B, C and D.

Solution The structure of compound D is easily deduced from the fact that it is obtained by the combination of acetone with hydrogen cyanide followed by hydrolysis.

$$\begin{matrix}CH_3\\CH_3\end{matrix}\!\!>C{=}O \xrightarrow{HCN} \begin{matrix}CH_3\\CH_3\end{matrix}\!\!>C<\!\!\begin{matrix}CN\\OH\end{matrix} \xrightarrow{H^+} \underset{(D)}{\begin{matrix}CH_3\\CH_3\end{matrix}\!\!>C<\!\!\begin{matrix}COOH\\OH\end{matrix}}$$

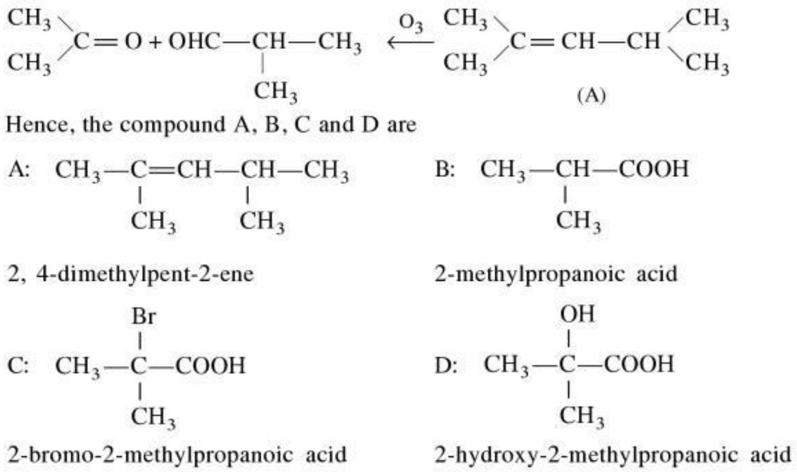
The conversion B to C and then C to D indicates that the reaction B to C is Hell-Volhard-Zelinsky reaction. Thus, we have

$$\underset{(B)}{\begin{matrix}CH_3\\CH_3\end{matrix}\!\!>C<\!\!\begin{matrix}H\\COOH\end{matrix}} \xrightarrow{Br_2/P} \underset{(C)}{\begin{matrix}CH_3\\CH_3\end{matrix}\!\!>C<\!\!\begin{matrix}Br\\COOH\end{matrix}} \xrightarrow{H_2O} \underset{(C)}{\begin{matrix}CH_3\\CH_3\end{matrix}\!\!>C<\!\!\begin{matrix}Br\\COOH\end{matrix}}$$

Hence, we conclude that the aldehyde from which compound B is obtained by oxidation has the structure

$$\underset{\text{2-methylpropanal}}{CH_3-\underset{CH_3}{CH}-CHO}$$

Now, the ozonolysis of the compound A gives acetone and 2-methyl-propanal, i.e.



$$\begin{matrix}CH_3 \\ CH_3\end{matrix}\!\!>C{=}O + OHC-\underset{\displaystyle CH_3}{\underset{|}{CH}}-CH_3 \xleftarrow{O_3} \begin{matrix}CH_3 \\ CH_3\end{matrix}\!\!>C{=}CH-CH\!\!<\!\!\begin{matrix}CH_3 \\ CH_3\end{matrix}$$

(A)

Hence, the compound A, B, C and D are

A: $CH_3-\underset{\displaystyle CH_3}{\underset{|}{C}}{=}CH-\underset{\displaystyle CH_3}{\underset{|}{CH}}-CH_3$

2, 4-dimethylpent-2-ene

B: $CH_3-\underset{\displaystyle CH_3}{\underset{|}{CH}}-COOH$

2-methylpropanoic acid

C: $CH_3-\overset{\displaystyle Br}{\overset{|}{\underset{\displaystyle CH_3}{\underset{|}{C}}}}-COOH$

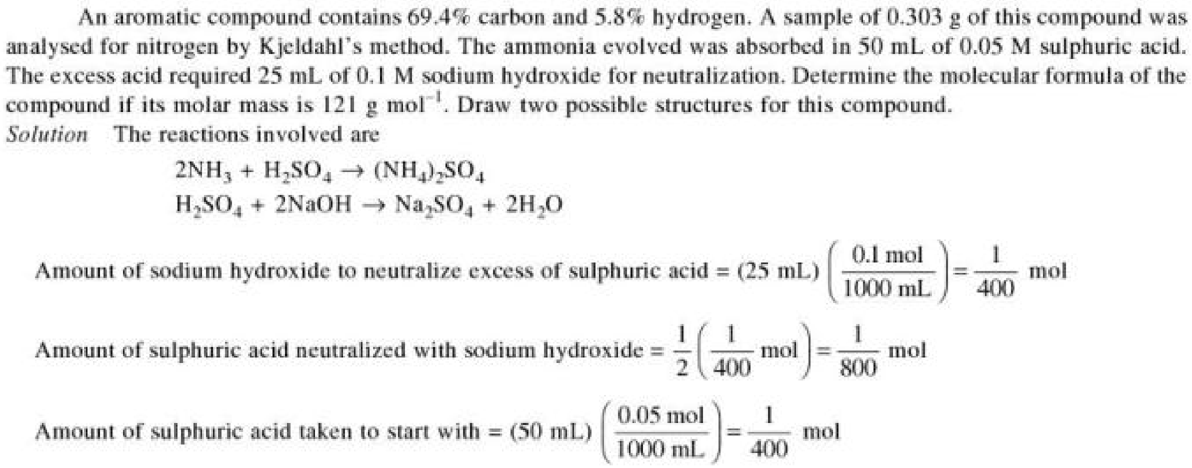
2-bromo-2-methylpropanoic acid

D: $CH_3-\overset{\displaystyle OH}{\overset{|}{\underset{\displaystyle CH_3}{\underset{|}{C}}}}-COOH$

2-hydroxy-2-methylpropanoic acid

Identify the compound

An aromatic comound contains 69.4% carbon, 5.8% hydrogen



An aromatic compound contains 69.4% carbon and 5.8% hydrogen. A sample of 0.303 g of this compound was analysed for nitrogen by Kjeldahl's method. The ammonia evolved was absorbed in 50 mL of 0.05 M sulphuric acid. The excess acid required 25 mL of 0.1 M sodium hydroxide for neutralization. Determine the molecular formula of the compound if its molar mass is 121 g mol^{-1}. Draw two possible structures for this compound.

Solution The reactions involved are

$$2NH_3 + H_2SO_4 \rightarrow (NH_4)_2SO_4$$
$$H_2SO_4 + 2NaOH \rightarrow Na_2SO_4 + 2H_2O$$

Amount of sodium hydroxide to neutralize excess of sulphuric acid = $(25\ \text{mL})\left(\frac{0.1\ \text{mol}}{1000\ \text{mL}}\right) = \frac{1}{400}$ mol

Amount of sulphuric acid neutralized with sodium hydroxide = $\frac{1}{2}\left(\frac{1}{400}\ \text{mol}\right) = \frac{1}{800}$ mol

Amount of sulphuric acid taken to start with = $(50\ \text{mL})\left(\frac{0.05\ \text{mol}}{1000\ \text{mL}}\right) = \frac{1}{400}$ mol

Amount of sulphuric acid neutralized by ammonia = $\left(\frac{1}{400}-\frac{1}{800}\right)$ mol = $\frac{1}{800}$ mol

Amount of ammonia evolved = $2\left(\frac{1}{800}\text{ mol}\right) = \frac{1}{400}$ mol

Mass of nitrogen in the evolved ammonia = $\left(\frac{1}{400}\text{ mol}\right)(14\text{ g mol}^{-1}) = \frac{7}{200}$ g

Per cent of nitrogen in the given aromatic compound = $\frac{(7/200)\text{ g}}{0.303\text{ g}} \times 100 = 11.6$

Per cent of oxygen in the given aromatic compound = 100 − (69.4 + 5.8 + 11.6) = 13.2
The ratios of atoms present in the given compound are

$$\text{C : H : N : O} \;::\; \frac{69.4}{12}:\frac{5.8}{1}:\frac{11.6}{14}:\frac{13.2}{16} \;::\; 5.8 : 5.8 : 0.83 : 0.83 \;::\; 7 : 7 : 1 : 1$$

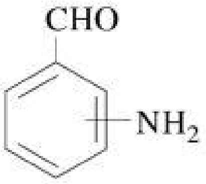
Hence, for the given compound, we have

Empirical formula = C_7H_7NO Empirical molar mass = 121 g mol^{-1}

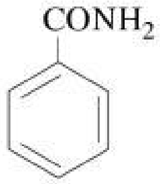
It is given that 121 g mol^{-1} is also the molar mass. Hence,

Molecular formula = C_7H_7NO

Since, the given compound is aromatic, the two possible structures are



and



Identify the compound

An Organic compound CxH2yOy was burnt with twice the amount of Oxygen

An organic compound $C_xH_{2y}O_y$ was burnt with twice the amount of oxygen needed for complete combustion to CO_2 and H_2O. The hot gases, when cooled to 0 °C and 1 atm pressure, measured 2.24 L. The water collected during cooling weighed 0.9 g. The vapour pressure of pure water at 20 °C is 17.5 mmHg and is lowered by 0.104 mmHg when 50 g of the organic compound is dissolved in 1000 g of water. Give the molecular formula of the organic compound.

Solution The combustion reaction is $C_xH_{2y}O_y + x\,O_2 \rightarrow x\,CO_2 + y\,H_2O$

To start with, the amount of O_2 taken is $2x$. Hence, after the combustion reaction, we will be left with the following amounts.

Amount of oxygen left unreacted = x; Amount of carbon dioxide = x; Amount of water = y

When this mixture is cooled to 0 °C and 1 atm, we will be left with oxygen and carbon dioxide. Hence, the amount $2x$ occupies the given volume of 2.24 L at STP. Hence,

$$\text{Amount } x = \frac{(2.24/2)\text{ L}}{22.4\text{ L mol}^{-1}} = 0.05\text{ mol}$$

Now, Mass of water collected = 0.9 g Amount of water collected, $y = \frac{0.9\text{ g}}{18\text{ g mol}^{-1}} = 0.05$ mol

Thus, the empirical formula of the compound is $C_{0.05}H_{2\times0.05}O_{0.05}$, i.e. CH_2O. Now, according to Raoult's law

$$-\frac{\Delta p}{p^*} = x_2 \quad \text{i.e.} \quad \frac{0.104\text{ mmHg}}{17.5\text{ mmHg}} = \frac{(50\text{ g}/M)}{(50\text{ g}/M) + (1000\text{ g}/18\text{ g mol}^{-1})}$$

Solving for M, we get $M = 150.5 \text{ g mol}^{-1}$

$$\text{Number of repeating units of CH}_2\text{O in the molecular formula} = \frac{150.5}{12+2+16} \approx 5$$

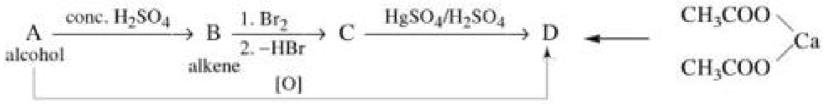
Hence, Molecular formula of the compound is $C_5H_{10}O_5$.

Identify the Compound

Alcohol when heated with concentrated H2SO4

An alcohol A, when heated with concentrated H_2SO_4 gives an alkene B. When B is bubbled through bromine water and the product obtained is dehydrohalogenated with excess of sodamide, a new compound C is obtained. The compound C gives D when treated with warm dilute H_2SO_4 in presence of $HgSO_4$. D can also be obtained either by oxidizing A with $KMnO_4$ or from acetic acid through its calcium salt. Identify A, B, C and D.

Solution The given reactions are as follows.



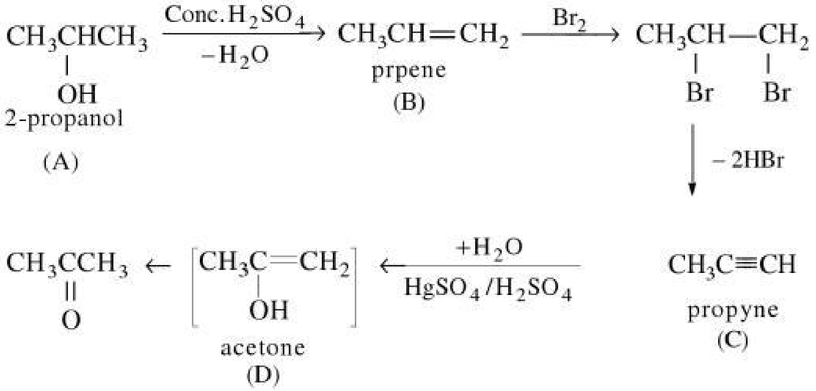
The reaction of obtaining D from calcium acetate is

$$(CH_3COO)_2Ca \rightarrow \underset{(D)}{CH_3COCH_3} + CaCO_3$$

Hence, D is acetone. Since the latter is also obtained from the oxidation of alcohol A, compound A will be

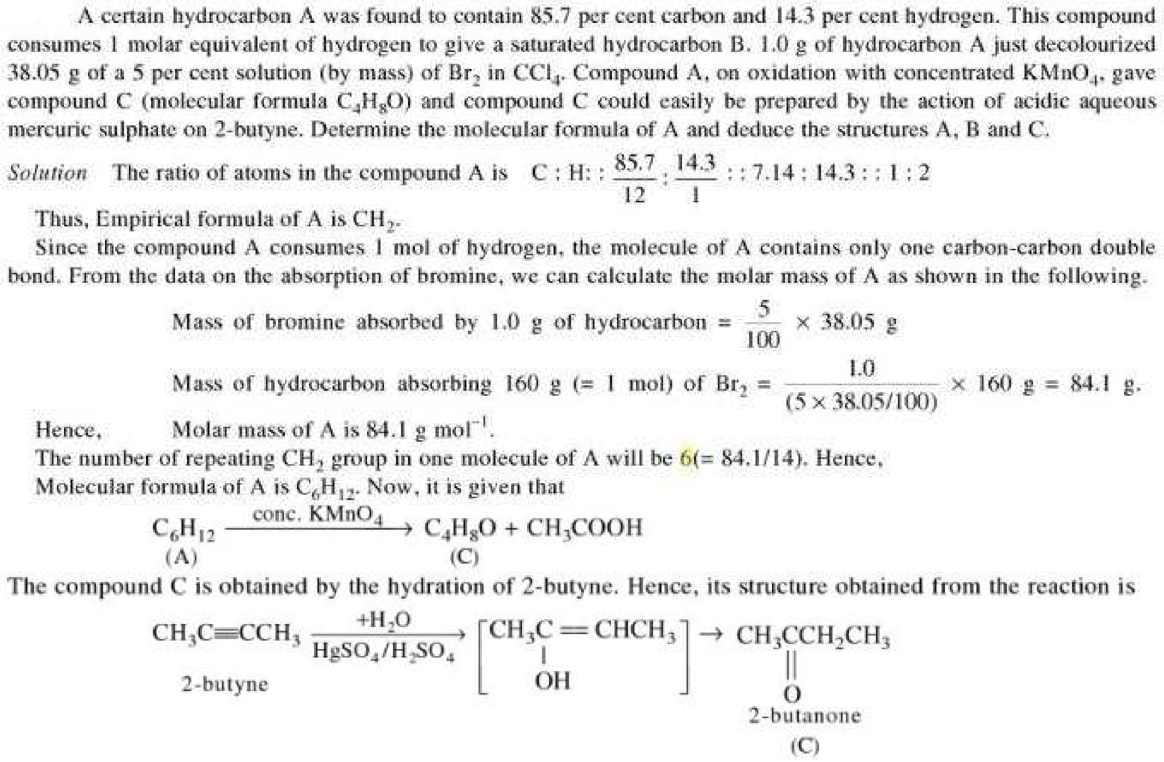
$$\underset{\displaystyle OH}{\overset{}{CH_3\underset{|}{C}HCH_3}}.$$

Hence, the reactions depicted above are as follows.



Identify the compound

Certain hydrocarbon was found to contain 85.7% carbon



A certain hydrocarbon A was found to contain 85.7 per cent carbon and 14.3 per cent hydrogen. This compound consumes 1 molar equivalent of hydrogen to give a saturated hydrocarbon B. 1.0 g of hydrocarbon A just decolourized 38.05 g of a 5 per cent solution (by mass) of Br_2 in CCl_4. Compound A, on oxidation with concentrated $KMnO_4$, gave compound C (molecular formula C_4H_8O) and compound C could easily be prepared by the action of acidic aqueous mercuric sulphate on 2-butyne. Determine the molecular formula of A and deduce the structures A, B and C.

Solution The ratio of atoms in the compound A is $C : H :: \frac{85.7}{12} : \frac{14.3}{1} :: 7.14 : 14.3 :: 1 : 2$

Thus, Empirical formula of A is CH_2.

Since the compound A consumes 1 mol of hydrogen, the molecule of A contains only one carbon-carbon double bond. From the data on the absorption of bromine, we can calculate the molar mass of A as shown in the following.

$$\text{Mass of bromine absorbed by 1.0 g of hydrocarbon} = \frac{5}{100} \times 38.05 \text{ g}$$

$$\text{Mass of hydrocarbon absorbing 160 g (= 1 mol) of } Br_2 = \frac{1.0}{(5 \times 38.05/100)} \times 160 \text{ g} = 84.1 \text{ g.}$$

Hence, Molar mass of A is 84.1 g mol^{-1}.

The number of repeating CH_2 group in one molecule of A will be 6(= 84.1/14). Hence,

Molecular formula of A is C_6H_{12}. Now, it is given that

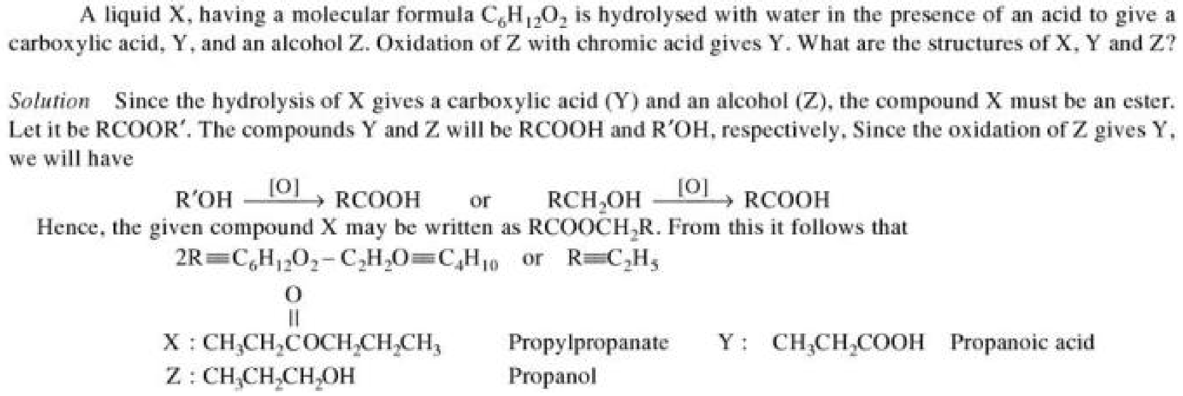
$$\underset{(A)}{C_6H_{12}} \xrightarrow{\text{conc. } KMnO_4} \underset{(C)}{C_4H_8O} + CH_3COOH$$

The compound C is obtained by the hydration of 2-butyne. Hence, its structure obtained from the reaction is

$$\underset{\text{2-butyne}}{CH_3C{\equiv}CCH_3} \xrightarrow[HgSO_4/H_2SO_4]{+H_2O} \left[\underset{\quad\;\; |\atop \quad\;\; OH}{CH_3C} = CHCH_3\right] \rightarrow \underset{\substack{\| \\ O \\ \text{2-butanone} \\ \text{(C)}}}{CH_3CCH_2CH_3}$$

Identify the compound

A liquid X having a molecular formula C6H12O2 is hydrolyzed with water



A liquid X, having a molecular formula $C_6H_{12}O_2$ is hydrolysed with water in the presence of an acid to give a carboxylic acid, Y, and an alcohol Z. Oxidation of Z with chromic acid gives Y. What are the structures of X, Y and Z?

Solution Since the hydrolysis of X gives a carboxylic acid (Y) and an alcohol (Z), the compound X must be an ester. Let it be RCOOR′. The compounds Y and Z will be RCOOH and R′OH, respectively. Since the oxidation of Z gives Y, we will have

$$R'OH \xrightarrow{[O]} RCOOH \quad \text{or} \quad RCH_2OH \xrightarrow{[O]} RCOOH$$

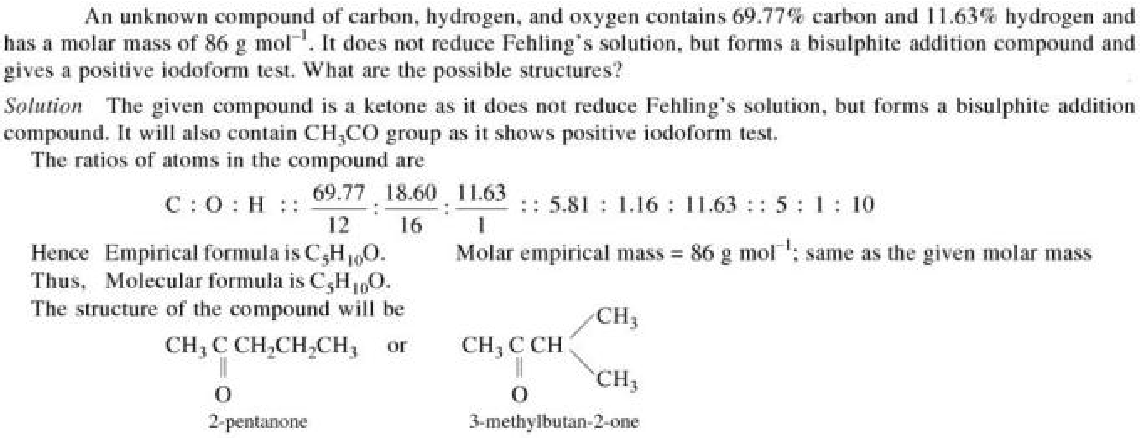
Hence, the given compound X may be written as $RCOOCH_2R$. From this it follows that

$$2R \equiv C_6H_{12}O_2 - C_2H_2O \equiv C_4H_{10} \quad \text{or} \quad R \equiv C_2H_5$$

$$X : CH_3CH_2\overset{\overset{\displaystyle O}{\|}}{C}OCH_2CH_2CH_3 \quad \text{Propylpropanate} \qquad Y : \; CH_3CH_2COOH \quad \text{Propanoic acid}$$

$$Z : CH_3CH_2CH_2OH \quad \text{Propanol}$$

Identify the compound



An unknown compound of carbon, hydrogen, and oxygen contains 69.77% carbon and 11.63% hydrogen and has a molar mass of 86 g mol^{-1}. It does not reduce Fehling's solution, but forms a bisulphite addition compound and gives a positive iodoform test. What are the possible structures?

Solution The given compound is a ketone as it does not reduce Fehling's solution, but forms a bisulphite addition compound. It will also contain CH_3CO group as it shows positive iodoform test.

The ratios of atoms in the compound are

$$C : O : H :: \frac{69.77}{12} : \frac{18.60}{16} : \frac{11.63}{1} :: 5.81 : 1.16 : 11.63 :: 5 : 1 : 10$$

Hence Empirical formula is $C_5H_{10}O$. Molar empirical mass = 86 g mol^{-1}; same as the given molar mass

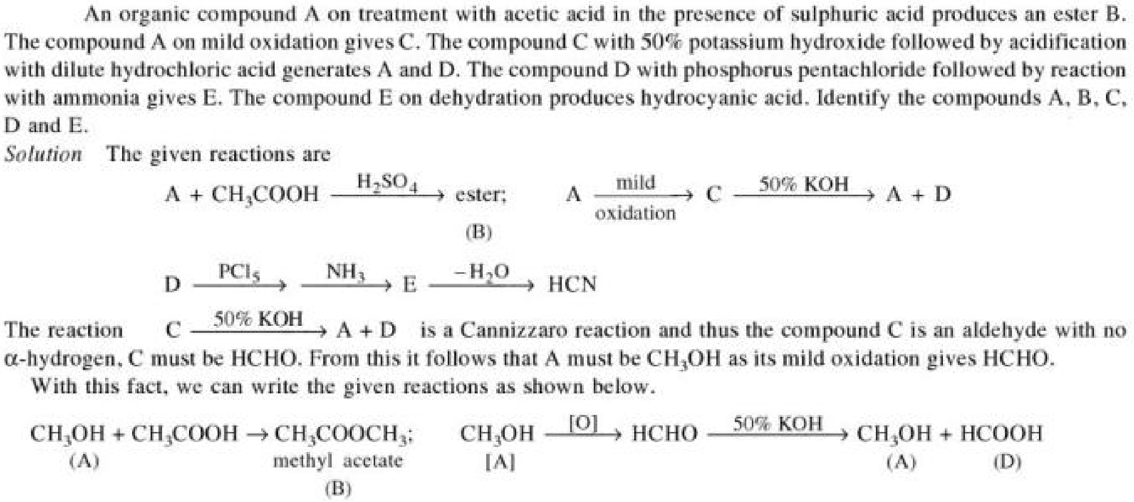
Thus, Molecular formula is $C_5H_{10}O$.

The structure of the compound will be

$$\underset{\text{2-pentanone}}{CH_3\overset{}{\underset{\underset{O}{\|}}{C}}CH_2CH_2CH_3} \quad \text{or} \quad \underset{\text{3-methylbutan-2-one}}{CH_3\underset{\underset{O}{\|}}{C}CH(CH_3)_2}$$

Identify the compound

An organic compound A on treatment with acetic acid in the presence of Sulphuric acid



An organic compound A on treatment with acetic acid in the presence of sulphuric acid produces an ester B. The compound A on mild oxidation gives C. The compound C with 50% potassium hydroxide followed by acidification with dilute hydrochloric acid generates A and D. The compound D with phosphorus pentachloride followed by reaction with ammonia gives E. The compound E on dehydration produces hydrocyanic acid. Identify the compounds A, B, C, D and E.

Solution The given reactions are

$$A + CH_3COOH \xrightarrow{H_2SO_4} \underset{(B)}{\text{ester}}; \qquad A \xrightarrow[\text{oxidation}]{\text{mild}} C \xrightarrow{50\%\ KOH} A + D$$

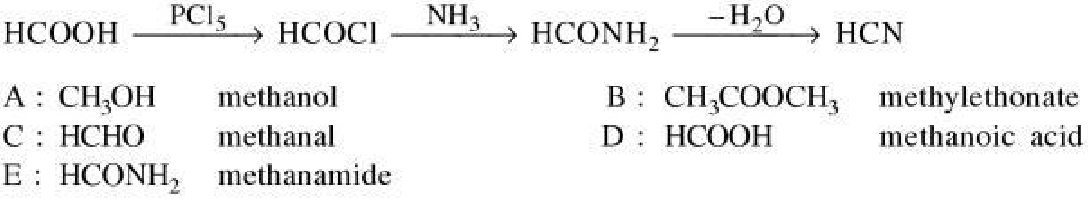
$$D \xrightarrow{PCl_5} \xrightarrow{NH_3} E \xrightarrow{-H_2O} HCN$$

The reaction $C \xrightarrow{50\%\ KOH} A + D$ is a Cannizzaro reaction and thus the compound C is an aldehyde with no α-hydrogen. C must be HCHO. From this it follows that A must be CH_3OH as its mild oxidation gives HCHO.

With this fact, we can write the given reactions as shown below.

$$\underset{(A)}{CH_3OH} + CH_3COOH \rightarrow \underset{\underset{(B)}{\text{methyl acetate}}}{CH_3COOCH_3}; \qquad \underset{[A]}{CH_3OH} \xrightarrow{[O]} HCHO \xrightarrow{50\%\ KOH} \underset{(A)}{CH_3OH} + \underset{(D)}{HCOOH}$$

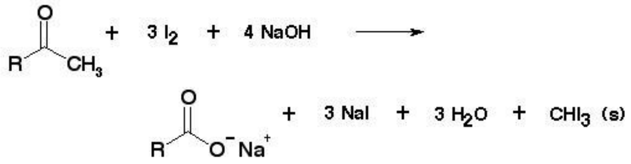
Hence

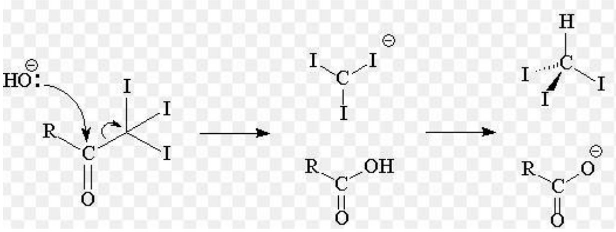


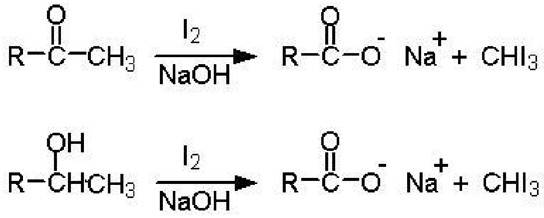
$$HCOOH \xrightarrow{PCl_5} HCOCl \xrightarrow{NH_3} HCONH_2 \xrightarrow{-H_2O} HCN$$

A :	CH_3OH	methanol	B :	CH_3COOCH_3	methylethonate
C :	HCHO	methanal	D :	HCOOH	methanoic acid
E :	$HCONH_2$	methanamide			

Iodoform Test

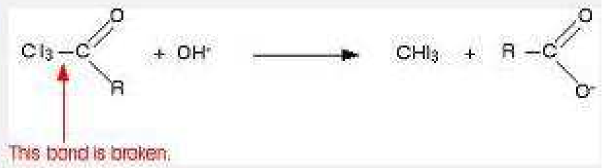




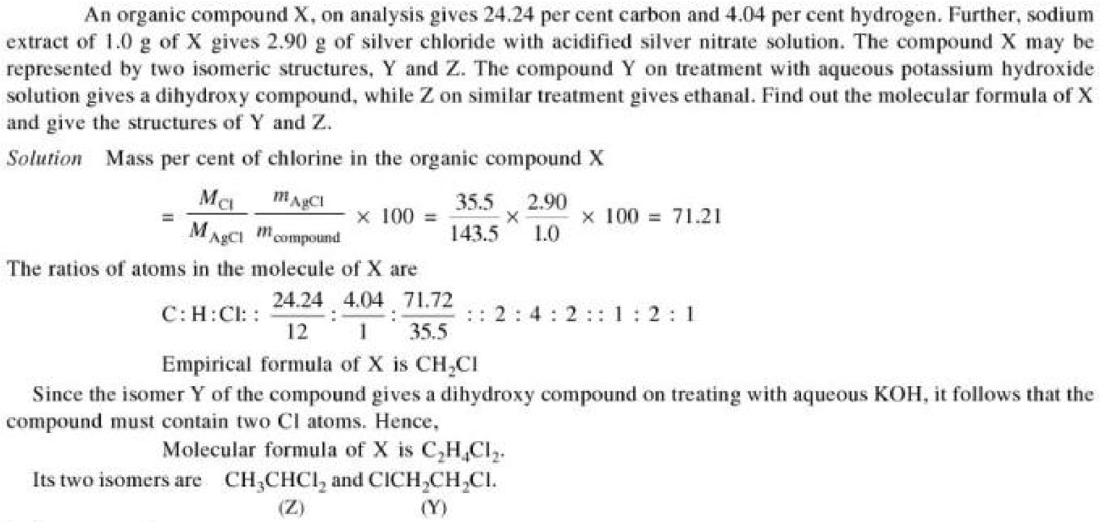


$$R-\overset{O}{\overset{\|}{C}}-CH_3 \xrightarrow[NaOH]{I_2} R-\overset{O}{\overset{\|}{C}}-O^- \; Na^+ + CHI_3$$

$$R-\overset{OH}{\overset{|}{C}H}CH_3 \xrightarrow[NaOH]{I_2} R-\overset{O}{\overset{\|}{C}}-O^- \; Na^+ + CHI_3$$



Identify the compound



An organic compound X, on analysis gives 24.24 per cent carbon and 4.04 per cent hydrogen. Further, sodium extract of 1.0 g of X gives 2.90 g of silver chloride with acidified silver nitrate solution. The compound X may be represented by two isomeric structures, Y and Z. The compound Y on treatment with aqueous potassium hydroxide solution gives a dihydroxy compound, while Z on similar treatment gives ethanal. Find out the molecular formula of X and give the structures of Y and Z.

Solution Mass per cent of chlorine in the organic compound X

$$= \frac{M_{Cl}}{M_{AgCl}} \frac{m_{AgCl}}{m_{compound}} \times 100 = \frac{35.5}{143.5} \times \frac{2.90}{1.0} \times 100 = 71.21$$

The ratios of atoms in the molecule of X are

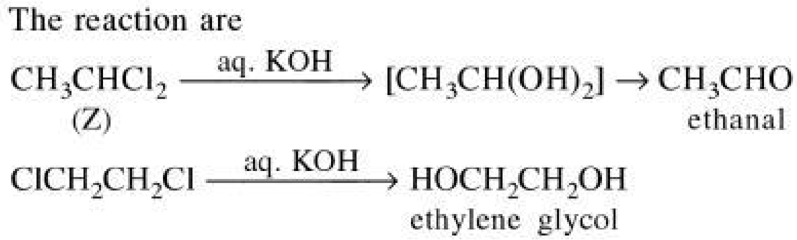
$$C:H:Cl: : \frac{24.24}{12} : \frac{4.04}{1} : \frac{71.72}{35.5} : : 2 : 4 : 2 : : 1 : 2 : 1$$

Empirical formula of X is CH_2Cl

Since the isomer Y of the compound gives a dihydroxy compound on treating with aqueous KOH, it follows that the compound must contain two Cl atoms. Hence,

Molecular formula of X is $C_2H_4Cl_2$.

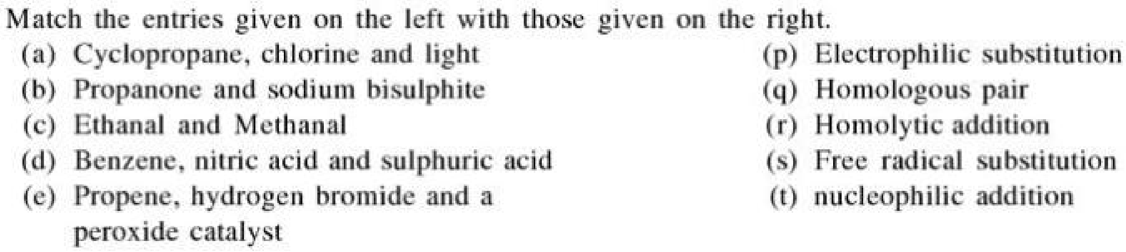
Its two isomers are $\underset{(Z)}{CH_3CHCl_2}$ and $\underset{(Y)}{ClCH_2CH_2Cl}$.



The reaction are

$$\underset{(Z)}{CH_3CHCl_2} \xrightarrow{\text{aq. KOH}} [CH_3CH(OH)_2] \rightarrow \underset{\text{ethanal}}{CH_3CHO}$$

$$ClCH_2CH_2Cl \xrightarrow{\text{aq. KOH}} \underset{\text{ethylene glycol}}{HOCH_2CH_2OH}$$



Match the entries given on the left with those given on the right.

(a) Cyclopropane, chlorine and light	(p) Electrophilic substitution
(b) Propanone and sodium bisulphite	(q) Homologous pair
(c) Ethanal and Methanal	(r) Homolytic addition
(d) Benzene, nitric acid and sulphuric acid	(s) Free radical substitution
(e) Propene, hydrogen bromide and a peroxide catalyst	(t) nucleophilic addition

Ans :

(a) – (s); (b) – (t); (c) – (q); (d) – (p); (e) – (r)

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Match the entries given on the left with those given on the right.

(a) Pyrolysis of alkanes	(p) Elimination reaction
(b) Benzene + Chloroethane (+ anhydrous $AlCl_3$)	(q) Saponification
(c) $CH_3COOC_2H_5$ + NaOH	(r) Wurtz reaction
(d) Preparation of alkanes	(s) Friedel-craft reaction
(e) Phenol + $CHCl_3$ (NaOH)	(t) Reimer-Tiemann reaction
(f) C_2H_5Br + alcKOH	(u) Cracking

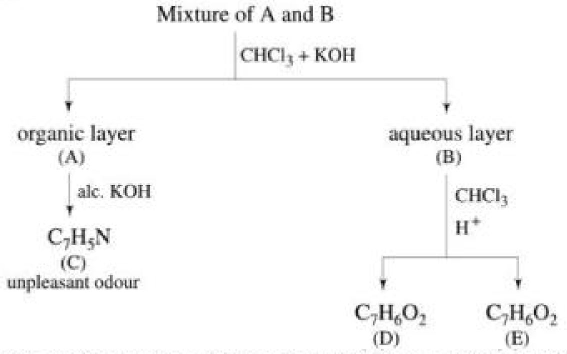
Ans :

(a) – (u); (b) – (s); (c) – (q); (d) – (r); (e) – (t);
(f) – (p)

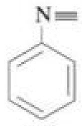
Identify the compound

A mixture of two aromatic compounds A and B was separated by dissolving in chloroform followed by extraction with aqueous KOH solution. The organic layer containing compound A, when heated with alcoholic solution of KOH produced a compound C (C_7H_5N) associated with an unpleasant odour. The alkaline aqueous layer on the other hand, when heated with chloroform and then acidified gave a mixture of two isomeric compounds D and E of molecular formula $C_7H_6O_2$. Identify the compounds A, B, C, D, and E and write their structures.

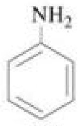
Solution We have



The compound C is an isocyanide which is produced from a primary amine. Hence, the structures of C and A are

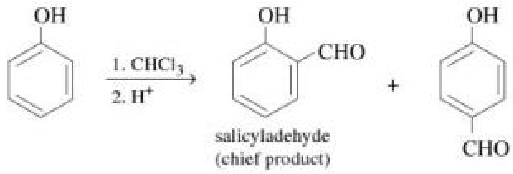


(C)

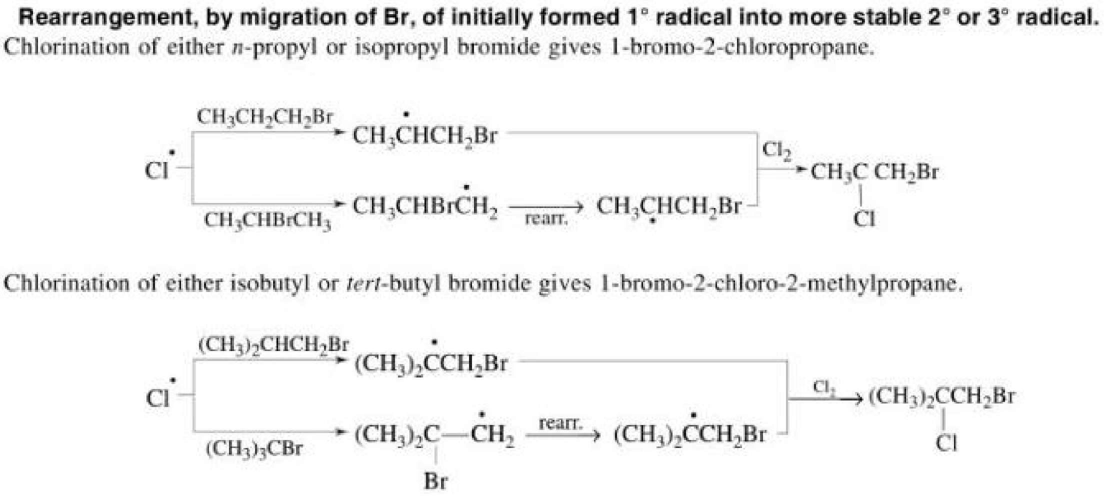


(A)

The compound B exhibits Reimer-Tiemann reaction. Thus, the compound B must be phenol. The reaction is



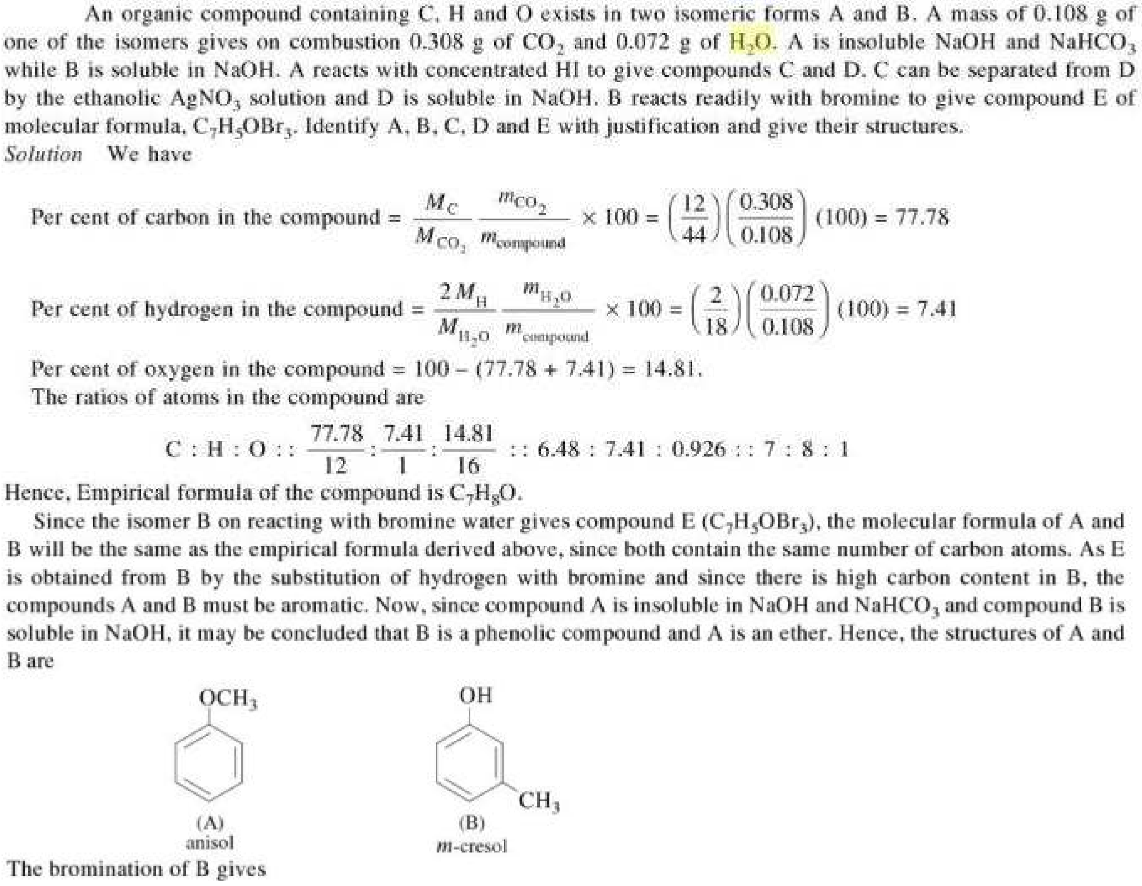
Rearrangement by migration of Bromine



Rearrangement, by migration of Br, of initially formed 1° radical into more stable 2° or 3° radical.
Chlorination of either *n*-propyl or isopropyl bromide gives 1-bromo-2-chloropropane.

Chlorination of either isobutyl or *tert*-butyl bromide gives 1-bromo-2-chloro-2-methylpropane.

Identify the Compound



An organic compound containing C, H and O exists in two isomeric forms A and B. A mass of 0.108 g of one of the isomers gives on combustion 0.308 g of CO_2 and 0.072 g of H_2O. A is insoluble NaOH and $NaHCO_3$ while B is soluble in NaOH. A reacts with concentrated HI to give compounds C and D. C can be separated from D by the ethanolic $AgNO_3$ solution and D is soluble in NaOH. B reacts readily with bromine to give compound E of molecular formula, $C_7H_5OBr_3$. Identify A, B, C, D and E with justification and give their structures.

Solution We have

$$\text{Per cent of carbon in the compound} = \frac{M_C}{M_{CO_2}}\frac{m_{CO_2}}{m_{\text{compound}}} \times 100 = \left(\frac{12}{44}\right)\left(\frac{0.308}{0.108}\right)(100) = 77.78$$

$$\text{Per cent of hydrogen in the compound} = \frac{2M_H}{M_{H_2O}}\frac{m_{H_2O}}{m_{\text{compound}}} \times 100 = \left(\frac{2}{18}\right)\left(\frac{0.072}{0.108}\right)(100) = 7.41$$

Per cent of oxygen in the compound = 100 − (77.78 + 7.41) = 14.81.
The ratios of atoms in the compound are

$$C : H : O :: \frac{77.78}{12} : \frac{7.41}{1} : \frac{14.81}{16} :: 6.48 : 7.41 : 0.926 :: 7 : 8 : 1$$

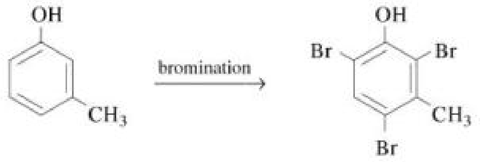
Hence, Empirical formula of the compound is C_7H_8O.

Since the isomer B on reacting with bromine water gives compound E ($C_7H_5OBr_3$), the molecular formula of A and B will be the same as the empirical formula derived above, since both contain the same number of carbon atoms. As E is obtained from B by the substitution of hydrogen with bromine and since there is high carbon content in B, the compounds A and B must be aromatic. Now, since compound A is insoluble in NaOH and $NaHCO_3$ and compound B is soluble in NaOH, it may be concluded that B is a phenolic compound and A is an ether. Hence, the structures of A and B are

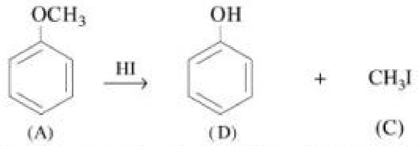
(A)
anisol

(B)
m-cresol

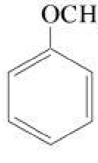
The bromination of B gives



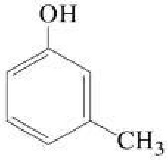
The reaction of compound A with HI is



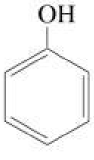
The compound C can be separated from D by use of ethanolic $AgNO_3$ solution as it is soluble in it whereas D will remain insoluble. The compound D will be soluble in NaOH as it is phenol. Hence, the structures of A, B, C, D and E are


(A)
anisol

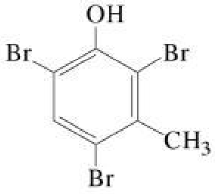
;


(B)
m-cresol

; CH_3I ;
(C)


(D)

;


(E)

Match the entries given on the left with those given on the right.

(a) Decarboxylation	(p) Addition reaction
(b) Ozonolysis	(q) Sodalime
(c) Williamson's synthesis	(r) Structure of alkene
(d) Dichloroethylene	(s) Ether

Ans :

(a) – (q); (b) – (r); (c) – (s); (d) – (p)

Identify the compound

Basic volatile Nitrogen compound

A basic volatile nitrogen compound gave a foul smelling gas when treated with chloroform and alcoholic potash. A 0.295 g sample of the substance, dissolved in aqueous HCl, and treated with $NaNO_2$ solution at 0 °C liberated a colourless, odourless gas whose volume corresponded to 112 mL at STP. After the evolution of the gas was complete, the aqueous solution was distilled to give an organic liquid which did not contain nitrogen and which on warming with alkali and iodine gave a yellow precipitate. Identify the original substance. Assume that it contains one N atom per molecule.

Solution Since the compound gives a foul smelling gas on treating with $CHCl_3$ and alcoholic KOH, the compound must be a primary amine.

$$RNH_2 + CHCl_3 + 3KOH \rightarrow \underset{\substack{\text{alkyl isocyanide}\\ \text{(foul smelling gas)}}}{RNC} + 3KCl + 3H_2O \qquad (1)$$

Since the compound on treating with $NaNO_2/HCl$ at 0 °C produces a colourless gas, the compound must be an aliphatic primary amine.

$$RNH_2 + HNO_2 \rightarrow ROH + N_2 + H_2O \qquad (2)$$

Thus, the gas produced is nitrogen.

$$\text{Amount of gas liberated} = \frac{112\ \text{mL}}{22400\ \text{mL mol}^{-1}} = \frac{1}{200}\ \text{mol}$$

From the above equation, it is obvious that

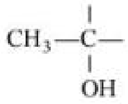
$$\text{Amount of compound } RNH_2 = \frac{1}{200} \text{ mol}$$

If M is the molar mass of RNH_2, then

$$\frac{2.295 \text{ g}}{M} = \frac{1}{200} \text{ mol} \quad \text{or} \quad M = 0.295 \times 200 \text{ g mol}^{-1} = 59 \text{ g mol}^{-1}.$$

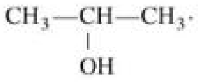
Thus, the molar mass of alkyl group R is (59 – 16) g mol^{-1}, i.e. 43 g mol^{-1}. Hence, R must be C_3H_7.

From Eq. (2), it is obvious that the liquid obtained after distillation is ROH. Since this gives yellow precipitate with alkali and iodine (iodoform test), it must contain



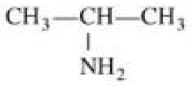
$CH_3-\overset{|}{\underset{OH}{\underset{|}{C}}}-$

group. Hence, it is concluded that ROH is



$CH_3-\underset{OH}{\underset{|}{CH}}-CH_3$.

Thus, the original compound is



$CH_3-\underset{NH_2}{\underset{|}{CH}}-CH_3$

isopropylamine

Match the entries given on the left with those given on the right.

(a) Lucas test	(p) Phenol
(b) Neutral $FeCl_3$	(q) Glucose
(c) Dye test	(r) Tertiary alcohol
(d) Tollens test	(s) Aniline

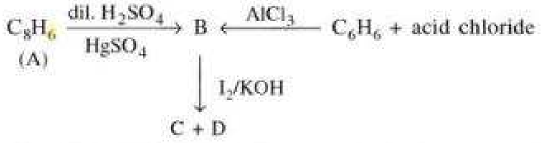
Ans :

(a) – (r); (b) – (p); (c) – (s); (d) – (q)

Identify the compound

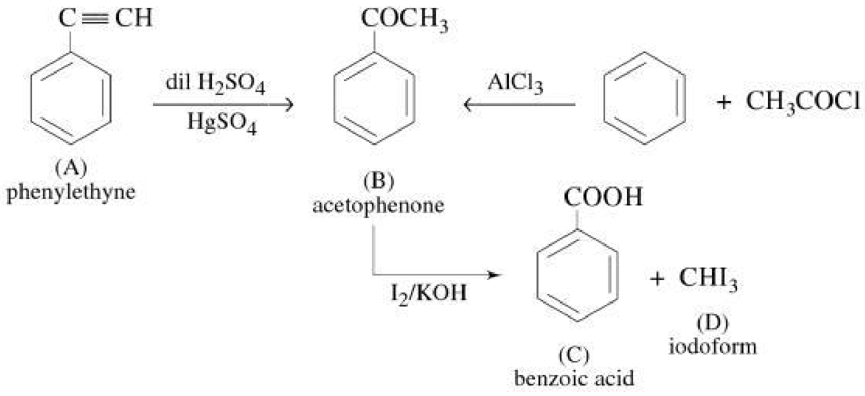
An organic compound A, C_8H_6, on treatment with dilute sulphuric acid containing mercuric sulphate gives a compound B, which can also be obtained from a reaction of benzene with an acid chloride in the presence of anhydrous aluminium chloride. The compound B, when treated with iodine in aqueous KOH, yields C and a yellow compound D. Identify A, B, C and D with justification. Show how B is formed from A.

Solution The given reactions are

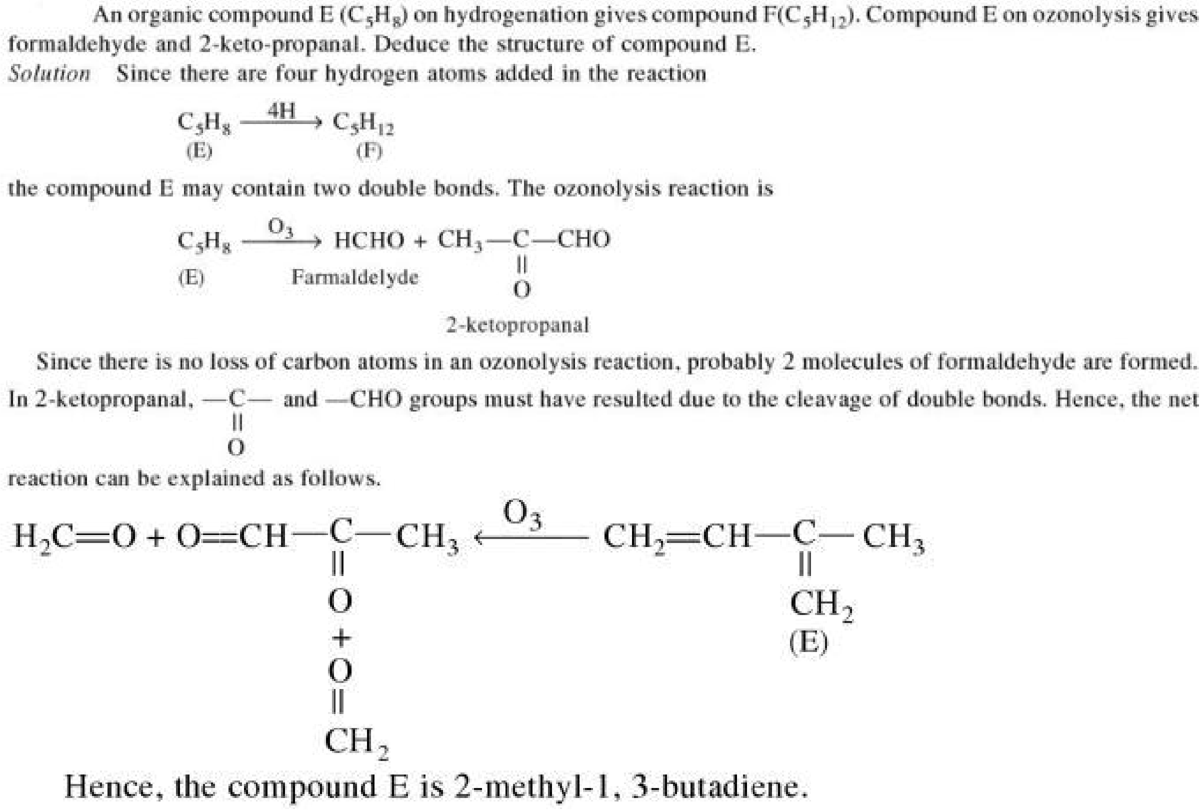


The reaction of B with I_2 in KOH is iodoform reaction. The compound D is iodoform, CHI_3. The compound B must contain $—COCH_3$ group so as to exhibit iodoform reaction. Since D is obtained from benzene by Friedel-Crafts reaction, it is an aromatic ketone ($C_6H_5COCH_3$). The compound C must be an acid.

The compound A may be represented as $C_6H_5C_2H$. Since it gives $C_6H_5COCH_3$ on treating with dilute H_2SO_4 and $HgSO_4$, it must contain a triple bond ($—C{\equiv}CH$) in the side chain. Hence, the given reactions may be represented as



Identify the compound



An organic compound E (C_5H_8) on hydrogenation gives compound F(C_5H_{12}). Compound E on ozonolysis gives formaldehyde and 2-keto-propanal. Deduce the structure of compound E.

Solution Since there are four hydrogen atoms added in the reaction

$$\underset{(E)}{C_5H_8} \xrightarrow{4H} \underset{(F)}{C_5H_{12}}$$

the compound E may contain two double bonds. The ozonolysis reaction is

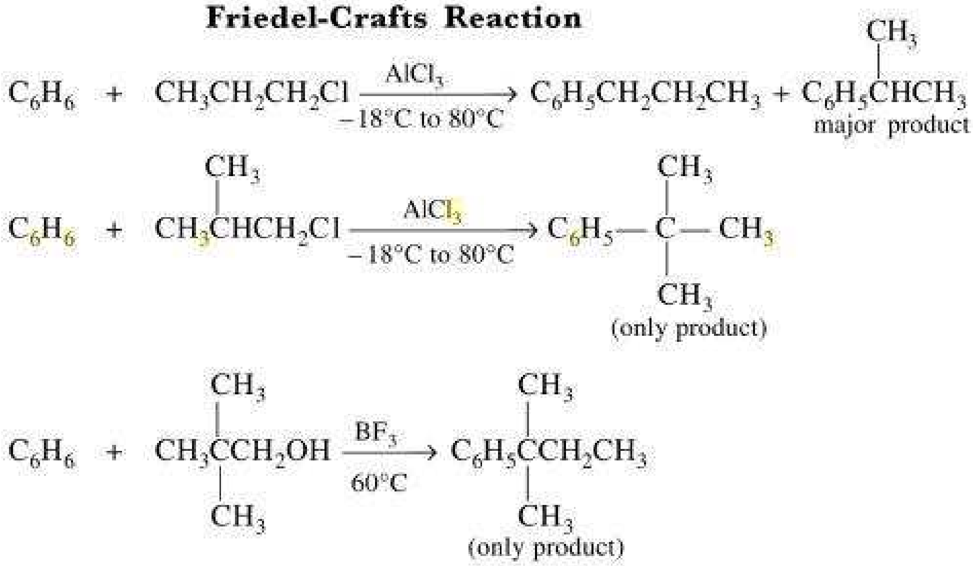
$$\underset{(E)}{C_5H_8} \xrightarrow{O_3} \underset{\text{Farmaldelyde}}{HCHO} + \underset{\text{2-ketopropanal}}{CH_3-\underset{\underset{O}{\|}}{C}-CHO}$$

Since there is no loss of carbon atoms in an ozonolysis reaction, probably 2 molecules of formaldehyde are formed. In 2-ketopropanal, $-\underset{\underset{O}{\|}}{C}-$ and —CHO groups must have resulted due to the cleavage of double bonds. Hence, the net reaction can be explained as follows.

$$H_2C{=}O + O{=}CH-\underset{\underset{\underset{\underset{\underset{CH_2}{\|}}{O}}{+}}{O}}{\|}}{C}-CH_3 \xleftarrow{O_3} \underset{(E)}{CH_2{=}CH-\underset{\underset{CH_2}{\|}}{C}-CH_3}$$

Hence, the compound E is 2-methyl-1, 3-butadiene.

Friedel Crafts Reaction



Friedel-Crafts Reaction

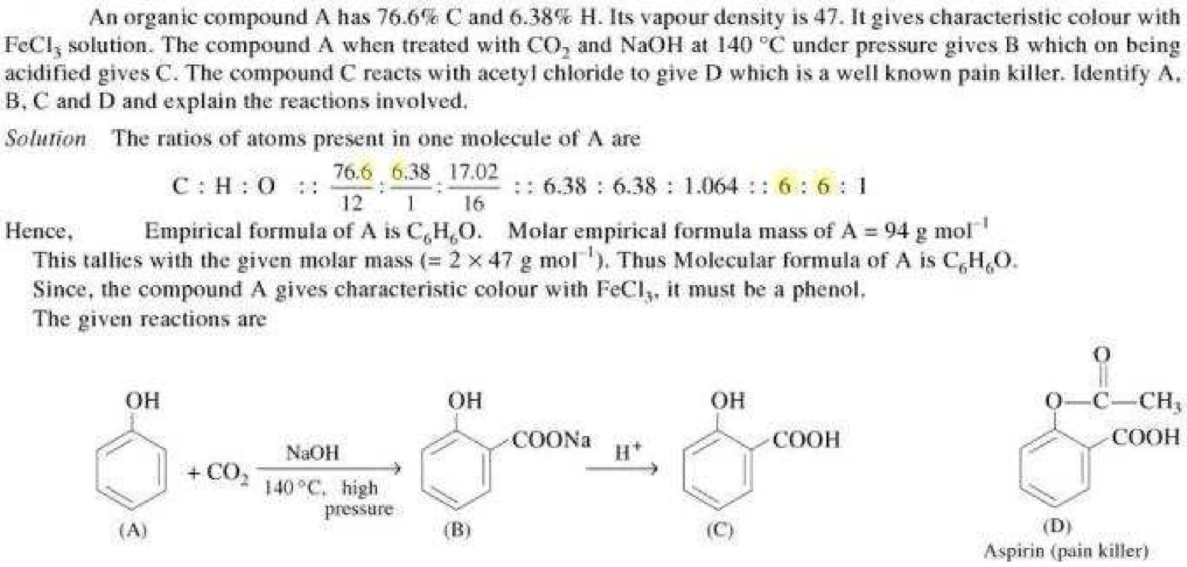
$$C_6H_6 + CH_3CH_2CH_2Cl \xrightarrow[-18°C \text{ to } 80°C]{AlCl_3} C_6H_5CH_2CH_2CH_3 + \underset{\text{major product}}{C_6H_5\overset{CH_3}{\overset{|}{C}}HCH_3}$$

$$C_6H_6 + CH_3\overset{CH_3}{\overset{|}{C}}HCH_2Cl \xrightarrow[-18°C \text{ to } 80°C]{AlCl_3} \underset{\text{(only product)}}{C_6H_5-\overset{CH_3}{\overset{|}{\underset{CH_3}{\underset{|}{C}}}}-CH_3}$$

$$C_6H_6 + CH_3\overset{CH_3}{\overset{|}{\underset{CH_3}{\underset{|}{C}}}}CH_2OH \xrightarrow[60°C]{BF_3} \underset{\text{(only product)}}{C_6H_5\overset{CH_3}{\overset{|}{\underset{CH_3}{\underset{|}{C}}}}CH_2CH_3}$$

Identify the compound

An Organic compound A has 76.6% C and 6.38% H (Aspirin)



An organic compound A has 76.6% C and 6.38% H. Its vapour density is 47. It gives characteristic colour with $FeCl_3$ solution. The compound A when treated with CO_2 and NaOH at 140 °C under pressure gives B which on being acidified gives C. The compound C reacts with acetyl chloride to give D which is a well known pain killer. Identify A, B, C and D and explain the reactions involved.

Solution The ratios of atoms present in one molecule of A are

$$C : H : O :: \frac{76.6}{12} : \frac{6.38}{1} : \frac{17.02}{16} :: 6.38 : 6.38 : 1.064 :: 6 : 6 : 1$$

Hence, Empirical formula of A is C_6H_6O. Molar empirical formula mass of A = 94 g mol^{-1}

This tallies with the given molar mass (= 2 × 47 g mol^{-1}). Thus Molecular formula of A is C_6H_6O.

Since, the compound A gives characteristic colour with $FeCl_3$, it must be a phenol.

The given reactions are

Match the entries given on the left with those given on the right.

Friedel-Crafts	oil	alkenes
Fermentation	Lewis acid	soap
Dehydrohalogenation	cuprous chloride	anhydrous $AlCl_3$
Sandmeyer	yeast	chlorobenzene
Saponification	alcoholic alkali	ethanol

Ans :

Friedel-Crafts;	Lewis acid;	anhydrous $AlCl_3$
Fermentation;	yeast;	ethanol
Dehydrohalogenation;	alcoholic alkali;	alkenes
Sandmeyer;	cuprous chloride;	chlorobenzene
Saponification;	oil;	soap

Identify the compound

An organic compound A of molar mass 140.5 gm per mole has 68.32% Hydrogen

An organic compound A of molar mass 140.5 g mol^{-1} has 68.32% C, 6.4% H and 25.26% Cl. Hydrolysis of A with dilute acid gives compound B, $C_8H_{10}O$. Compound B can be oxidized under mild conditions to compound C, C_8H_8O. Compound C forms a phenylhydrazone D with $PhNHNH_2$ and gives a positive iodoform test. Deduce the structures of compounds A to D with proper reasoning.

Solution The ratios of atoms in a molecule of A are

$$C:H:Cl::\frac{68.32}{12}:\frac{6.4}{1}:\frac{25.26}{35.5}::5.69:6.4:0.71::8:9:1$$

Hence, Empirical formula of A is C_8H_9Cl. Molar empirical formula mass of A = 140.5 g mol^{-1}

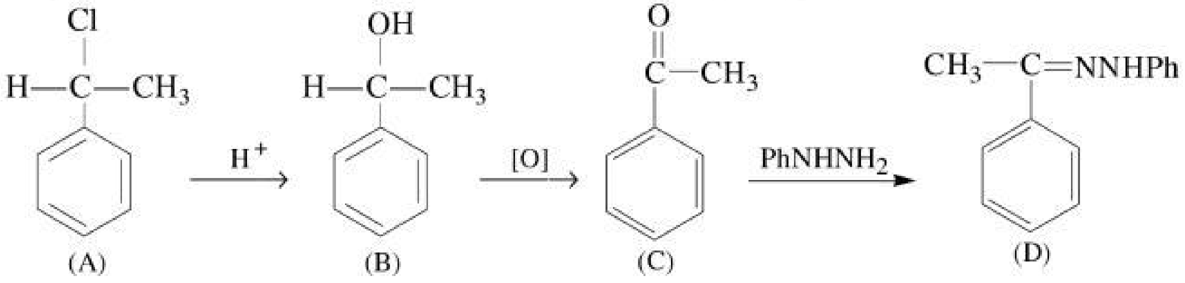
This tallies with the given molar mass. Thus Molecular formula of A is C_8H_9Cl.

The high content of carbon indicates the presence of aromatic (or benzene) ring in the molecule.

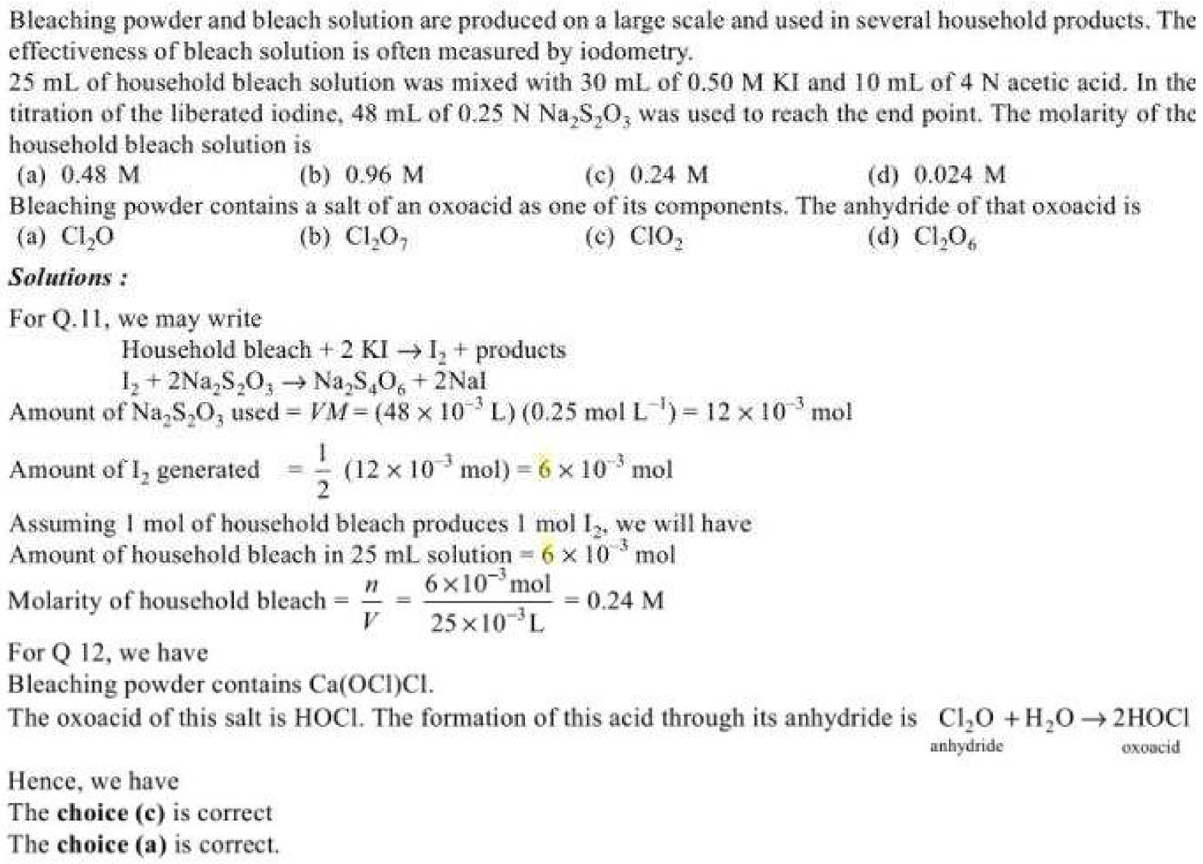
Since Cl atom in the compound A is easily replaced during hydrolysis of A with dilute acid, the Cl atom must be attached to the aliphatic portion of the molecule.

Since the compound C gives a positive iodoform test, it must contain—$COCH_3$ group.

Since the compound C is obtained by mild oxidation of B, the latter must be a secondary alcohol. Hence, the compound A also contains secondary chlorine atom. Consistent with this, the given reactions are as follows.



Question



Bleaching powder and bleach solution are produced on a large scale and used in several household products. The effectiveness of bleach solution is often measured by iodometry.

25 mL of household bleach solution was mixed with 30 mL of 0.50 M KI and 10 mL of 4 N acetic acid. In the titration of the liberated iodine, 48 mL of 0.25 N $Na_2S_2O_3$ was used to reach the end point. The molarity of the household bleach solution is

(a) 0.48 M (b) 0.96 M (c) 0.24 M (d) 0.024 M

Bleaching powder contains a salt of an oxoacid as one of its components. The anhydride of that oxoacid is

(a) Cl_2O (b) Cl_2O_7 (c) ClO_2 (d) Cl_2O_6

Solutions :

For Q.11, we may write

$$\text{Household bleach} + 2\,KI \rightarrow I_2 + \text{products}$$

$$I_2 + 2Na_2S_2O_3 \rightarrow Na_2S_4O_6 + 2NaI$$

Amount of $Na_2S_2O_3$ used $= VM = (48 \times 10^{-3}\ L)(0.25\ mol\ L^{-1}) = 12 \times 10^{-3}$ mol

Amount of I_2 generated $= \frac{1}{2}(12 \times 10^{-3}\ mol) = 6 \times 10^{-3}$ mol

Assuming 1 mol of household bleach produces 1 mol I_2, we will have

Amount of household bleach in 25 mL solution $= 6 \times 10^{-3}$ mol

Molarity of household bleach $= \frac{n}{V} = \frac{6 \times 10^{-3}\,mol}{25 \times 10^{-3}\,L} = 0.24$ M

For Q 12, we have

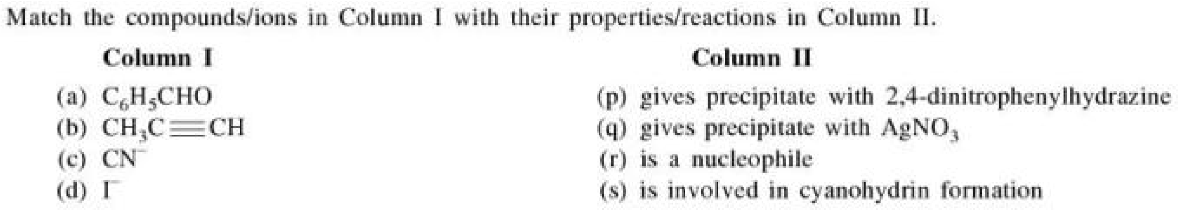
Bleaching powder contains Ca(OCl)Cl.

The oxoacid of this salt is HOCl. The formation of this acid through its anhydride is $\underset{\text{anhydride}}{Cl_2O} + H_2O \rightarrow \underset{\text{oxoacid}}{2HOCl}$

Hence, we have

The **choice (c)** is correct

The **choice (a)** is correct.



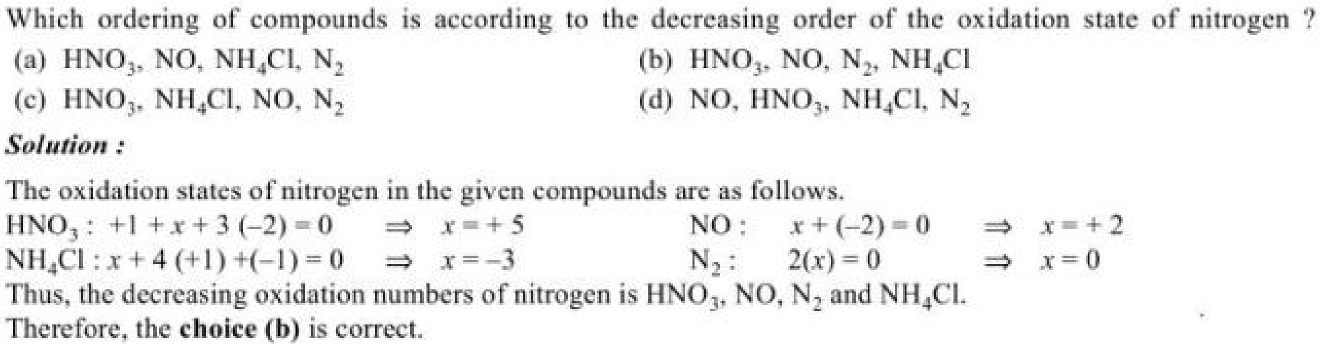
Match the compounds/ions in Column I with their properties/reactions in Column II.

Column I	Column II
(a) C_6H_5CHO	(p) gives precipitate with 2,4-dinitrophenylhydrazine
(b) $CH_3C\equiv CH$	(q) gives precipitate with $AgNO_3$
(c) CN^-	(r) is a nucleophile
(d) I^-	(s) is involved in cyanohydrin formation

Ans :

(a) – (p), (q), (s); (b) – (q); (c) – (q), (r), (s); (d) – (q)

In decreasing order of Oxidation state of Nitrogen



Which ordering of compounds is according to the decreasing order of the oxidation state of nitrogen ?

(a) HNO_3, NO, NH_4Cl, N_2 (b) HNO_3, NO, N_2, NH_4Cl

(c) HNO_3, NH_4Cl, NO, N_2 (d) NO, HNO_3, NH_4Cl, N_2

Solution :

The oxidation states of nitrogen in the given compounds are as follows.

HNO_3 : $+1 + x + 3(-2) = 0 \quad \Rightarrow \quad x = +5$ NO : $x + (-2) = 0 \quad \Rightarrow \quad x = +2$

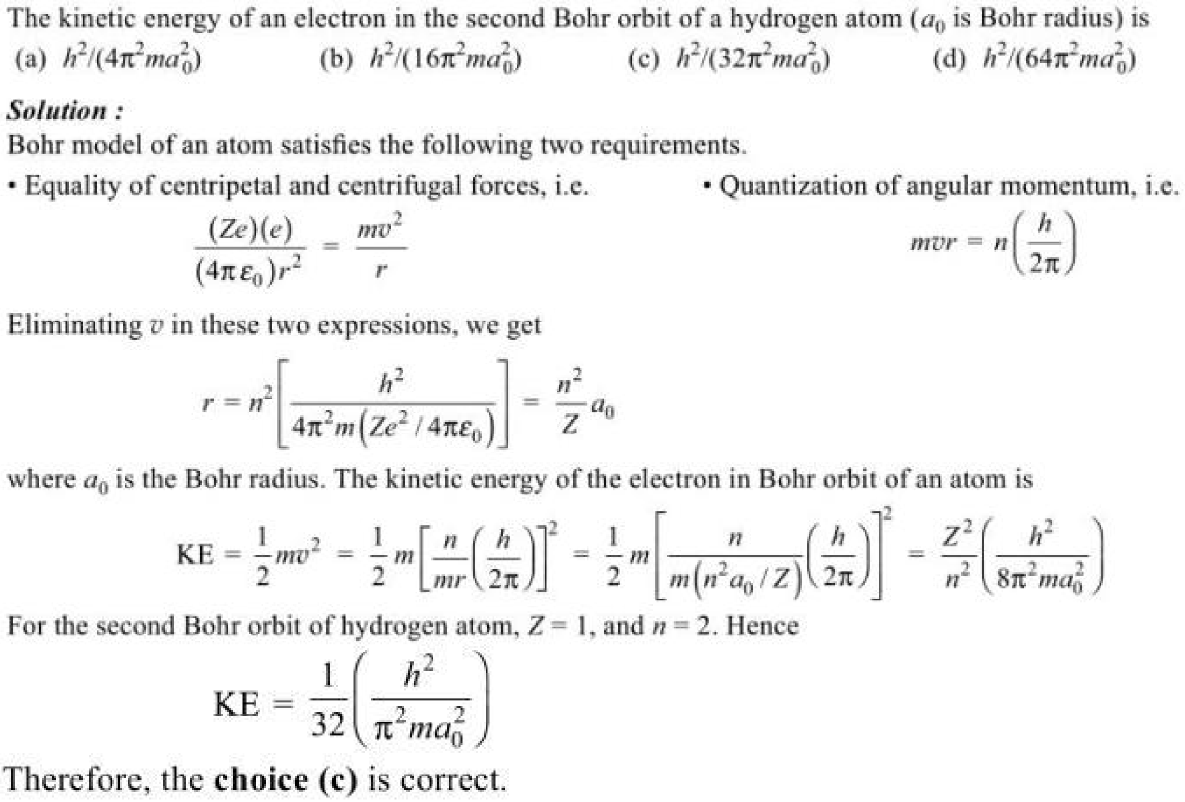
NH_4Cl : $x + 4(+1) + (-1) = 0 \quad \Rightarrow \quad x = -3$ N_2 : $2(x) = 0 \quad \Rightarrow \quad x = 0$

Thus, the decreasing oxidation numbers of nitrogen is HNO_3, NO, N_2 and NH_4Cl.

Therefore, the **choice (b)** is correct.

Gyan Question

The Kinetic energy of an electron in the second Bohr orbit



The kinetic energy of an electron in the second Bohr orbit of a hydrogen atom (a_0 is Bohr radius) is

(a) $h^2/(4\pi^2 ma_0^2)$ (b) $h^2/(16\pi^2 ma_0^2)$ (c) $h^2/(32\pi^2 ma_0^2)$ (d) $h^2/(64\pi^2 ma_0^2)$

Solution :

Bohr model of an atom satisfies the following two requirements.

- Equality of centripetal and centrifugal forces, i.e.

$$\frac{(Ze)(e)}{(4\pi\varepsilon_0)r^2} = \frac{mv^2}{r}$$

- Quantization of angular momentum, i.e.

$$mvr = n\left(\frac{h}{2\pi}\right)$$

Eliminating v in these two expressions, we get

$$r = n^2\left[\frac{h^2}{4\pi^2 m\left(Ze^2/4\pi\varepsilon_0\right)}\right] = \frac{n^2}{Z}a_0$$

where a_0 is the Bohr radius. The kinetic energy of the electron in Bohr orbit of an atom is

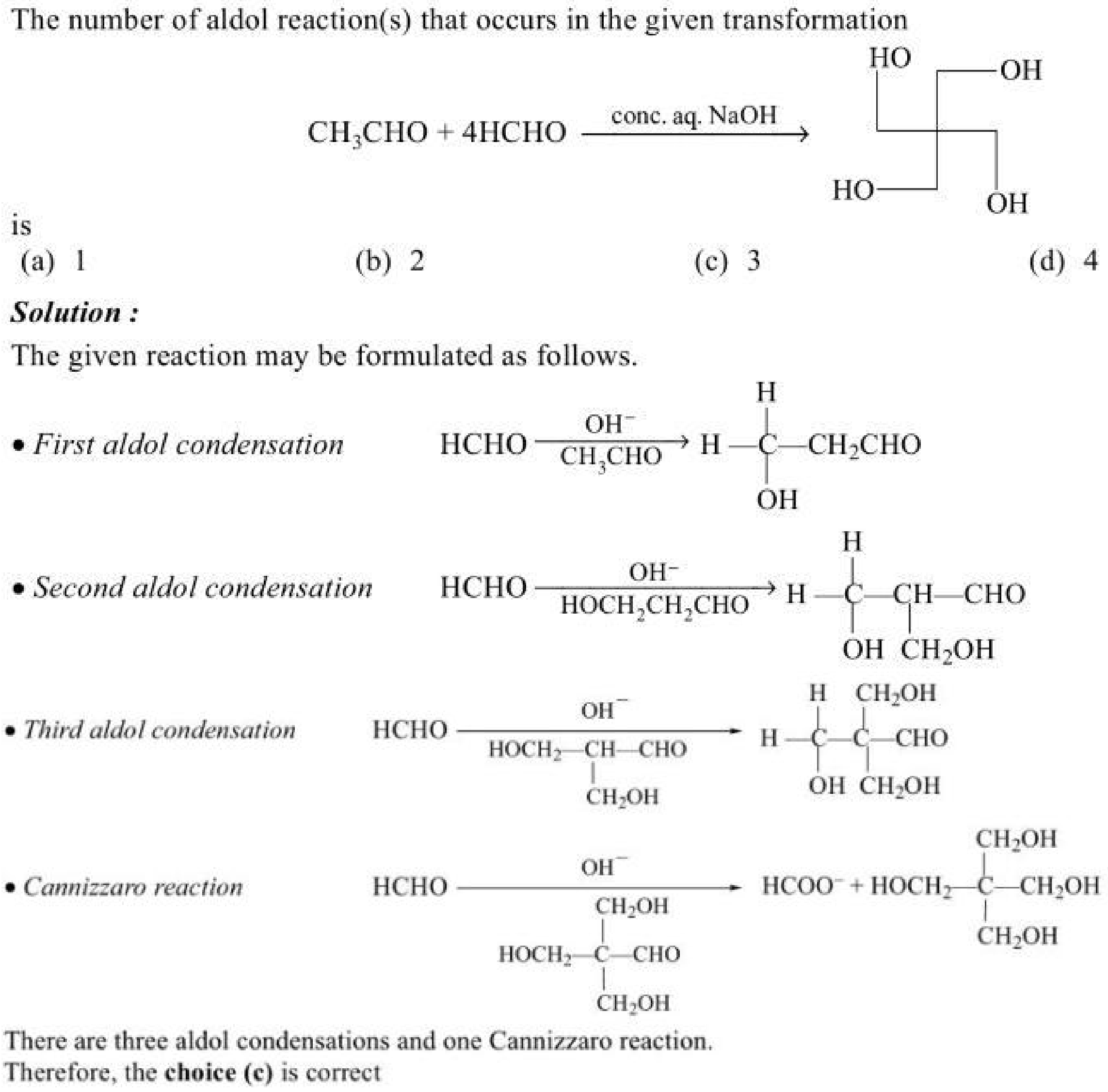
$$\text{KE} = \frac{1}{2}mv^2 = \frac{1}{2}m\left[\frac{n}{mr}\left(\frac{h}{2\pi}\right)\right]^2 = \frac{1}{2}m\left[\frac{n}{m\left(n^2a_0/Z\right)}\left(\frac{h}{2\pi}\right)\right]^2 = \frac{Z^2}{n^2}\left(\frac{h^2}{8\pi^2 ma_0^2}\right)$$

For the second Bohr orbit of hydrogen atom, $Z = 1$, and $n = 2$. Hence

$$\text{KE} = \frac{1}{32}\left(\frac{h^2}{\pi^2 ma_0^2}\right)$$

Therefore, the **choice (c)** is correct.

Question on Aldol reaction



The number of aldol reaction(s) that occurs in the given transformation

$$CH_3CHO + 4HCHO \xrightarrow{\text{conc. aq. NaOH}}$$

is

(a) 1 (b) 2 (c) 3 (d) 4

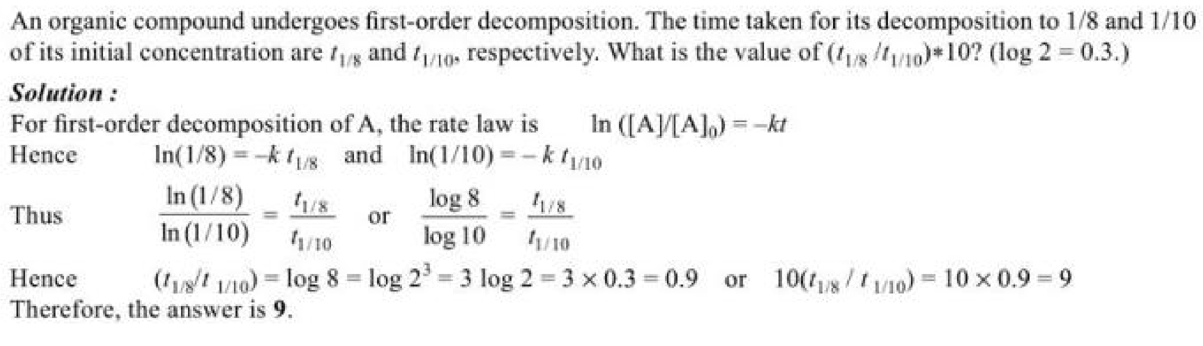
Solution :

The given reaction may be formulated as follows.

- *First aldol condensation* $HCHO \xrightarrow[CH_3CHO]{OH^-} H-\underset{OH}{\overset{H}{C}}-CH_2CHO$
- *Second aldol condensation* $HCHO \xrightarrow[HOCH_2CH_2CHO]{OH^-} H-\underset{OH}{\overset{H}{C}}-\underset{CH_2OH}{CH}-CHO$
- *Third aldol condensation* $HCHO \xrightarrow[HOCH_2-\underset{CH_2OH}{CH}-CHO]{OH^-} H-\underset{OH}{\overset{H}{C}}-\underset{CH_2OH}{\overset{CH_2OH}{C}}-CHO$
- *Cannizzaro reaction* $HCHO \xrightarrow[HOCH_2-\underset{CH_2OH}{\overset{CH_2OH}{C}}-CHO]{OH^-} HCOO^- + HOCH_2-\underset{CH_2OH}{\overset{CH_2OH}{C}}-CH_2OH$

There are three aldol condensations and one Cannizzaro reaction.
Therefore, the **choice (c)** is correct

An organic compound undergoes first order decomposition



An organic compound undergoes first-order decomposition. The time taken for its decomposition to 1/8 and 1/10 of its initial concentration are $t_{1/8}$ and $t_{1/10}$, respectively. What is the value of $(t_{1/8}/t_{1/10})*10$? (log 2 = 0.3.)

Solution :

For first-order decomposition of A, the rate law is $\ln([A]/[A]_0) = -kt$

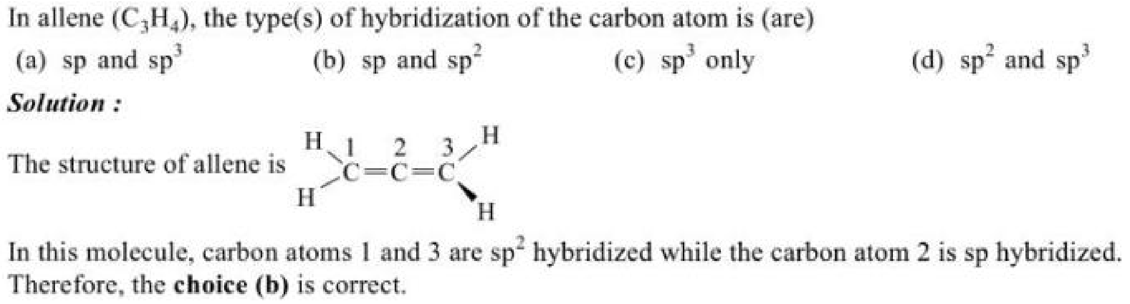
Hence $\ln(1/8) = -k\,t_{1/8}$ and $\ln(1/10) = -k\,t_{1/10}$

Thus $\dfrac{\ln(1/8)}{\ln(1/10)} = \dfrac{t_{1/8}}{t_{1/10}}$ or $\dfrac{\log 8}{\log 10} = \dfrac{t_{1/8}}{t_{1/10}}$

Hence $(t_{1/8}/t_{1/10}) = \log 8 = \log 2^3 = 3 \log 2 = 3 \times 0.3 = 0.9$ or $10(t_{1/8}/t_{1/10}) = 10 \times 0.9 = 9$

Therefore, the answer is **9**.

Allene has sp hybridization



In allene (C_3H_4), the type(s) of hybridization of the carbon atom is (are)

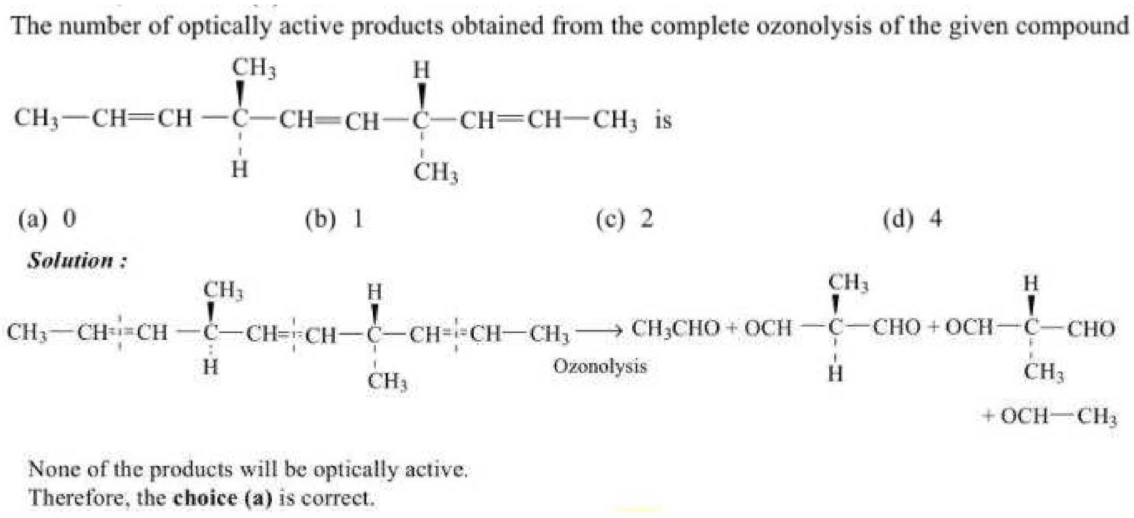
(a) sp and sp^3 (b) sp and sp^2 (c) sp^3 only (d) sp^2 and sp^3

Solution :

The structure of allene is $H_2\overset{1}{C}=\overset{2}{C}=\overset{3}{C}H_2$

In this molecule, carbon atoms 1 and 3 are sp^2 hybridized while the carbon atom 2 is sp hybridized. Therefore, the **choice (b)** is correct.

Number of chiral carbon and optically active products



The number of optically active products obtained from the complete ozonolysis of the given compound

$$CH_3-CH=CH-\underset{H}{\overset{CH_3}{C}}-CH=CH-\underset{CH_3}{\overset{H}{C}}-CH=CH-CH_3$$ is

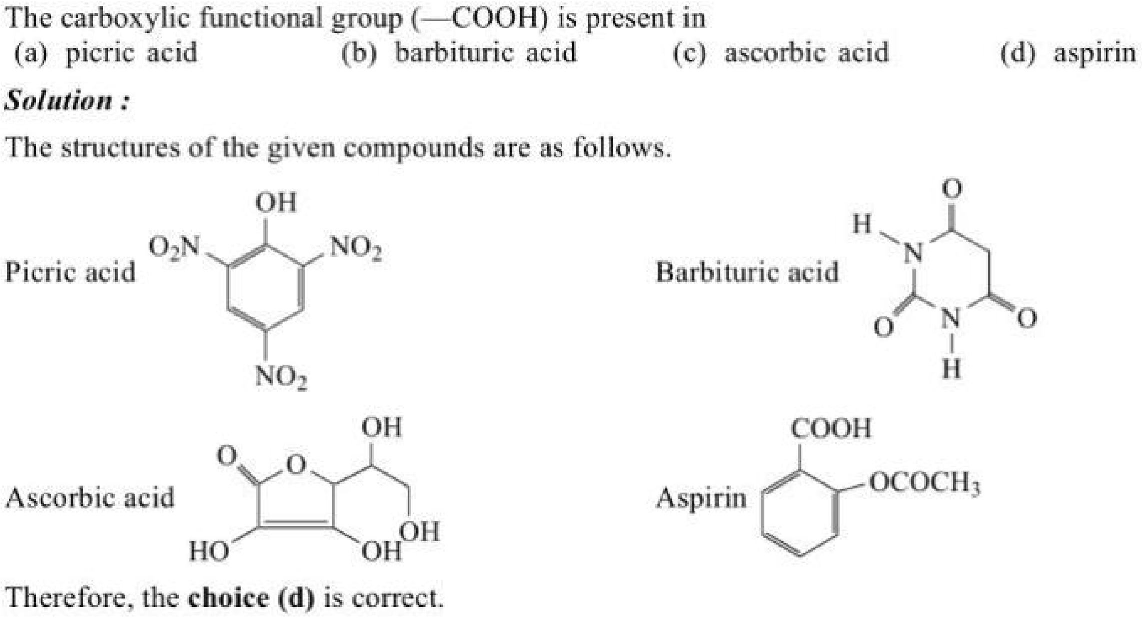
(a) 0 (b) 1 (c) 2 (d) 4

Solution :

$$CH_3-CH=CH-\underset{H}{\overset{CH_3}{C}}-CH=CH-\underset{CH_3}{\overset{H}{C}}-CH=CH-CH_3 \xrightarrow[\text{Ozonolysis}]{} CH_3CHO + OCH-\underset{H}{\overset{CH_3}{C}}-CHO + OCH-\underset{CH_3}{\overset{H}{C}}-CHO + OCH-CH_3$$

None of the products will be optically active.
Therefore, the **choice (a)** is correct.

Carboxylic functional group is present in aspirin



The carboxylic functional group (—COOH) is present in
(a) picric acid (b) barbituric acid (c) ascorbic acid (d) aspirin

Solution :

The structures of the given compounds are as follows.

Picric acid

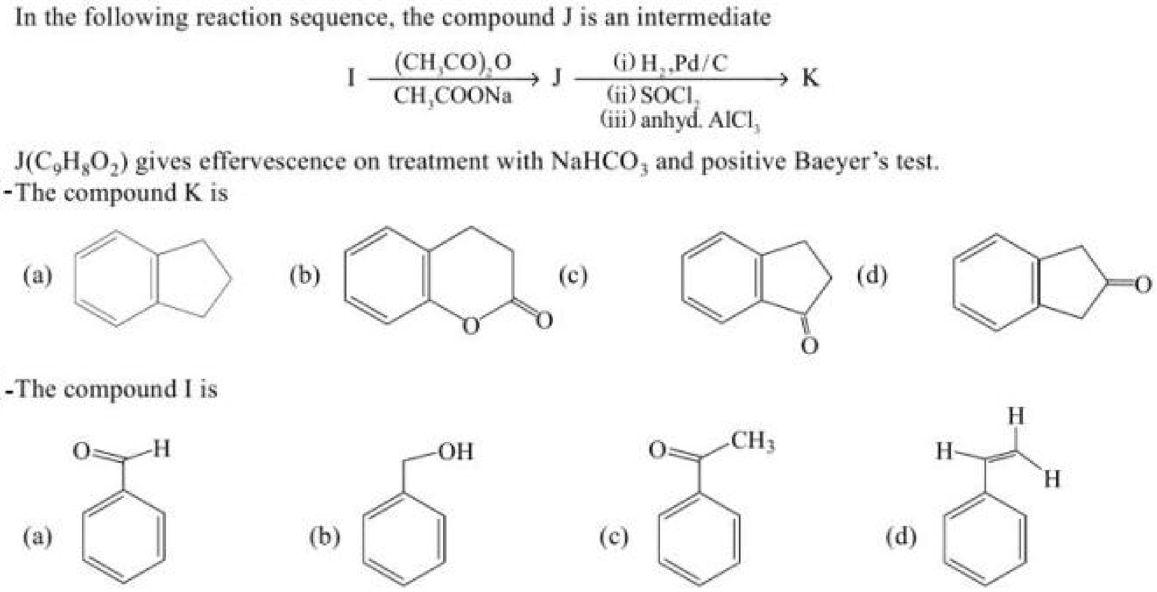
Barbituric acid

Ascorbic acid

Aspirin

Therefore, the **choice (d)** is correct.

Identify the compound



In the following reaction sequence, the compound J is an intermediate

$$I \xrightarrow[CH_3COONa]{(CH_3CO)_2O} J \xrightarrow[\text{(ii) } SOCl_2 \text{ (iii) anhyd. } AlCl_3]{\text{(i) } H_2,Pd/C} K$$

$J(C_9H_8O_2)$ gives effervescence on treatment with $NaHCO_3$ and positive Baeyer's test.

Q-The compound K is

(a) (b) (c) (d)

Q-The compound I is

(a) (b) (c) (d)

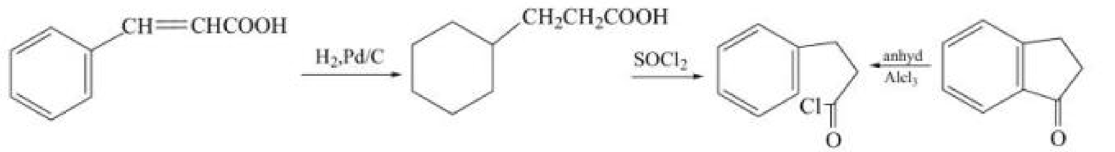
Perkin Condensation

Solutions :

The compound J must contain a —COOH group as it gives effervescence with $NaHCO_3$. Also, it is an unsaturated compound as it gives positive Baeyer's test (decolourization of pink colour of alkaline $KMnO_4$). Since the compound J is obtained by treating the compound I with $(CH_3CO)_2O$ and CH_3COONa, and more over the compound J seems to be aromatic (high carbon content), the compound I amongst the choices given in Q.14 seems to be benzaldehyde as it shows Perkin condensation shown in the following

α, β- unsaturated
carboxylic acid

The conversion J to K is as follows.

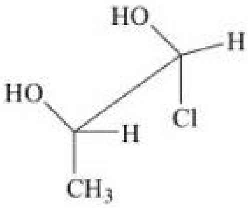


choice (c) of 2.13

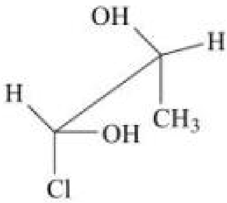
Hence, we have
The **choice (c)** is correct.
The **choice (a)** is correct.

Question on isomers

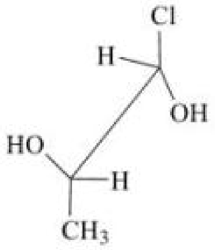
Which of the given statement(s) about N, O, P and Q with respect to M is(are) correct?



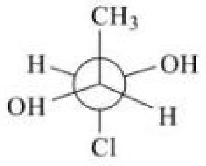
(M)



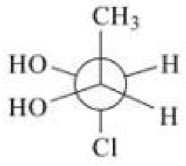
(N)



(O)



(P)

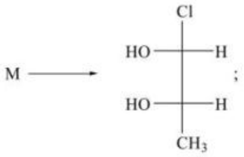


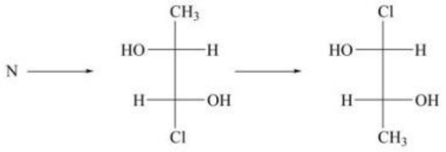
(Q)

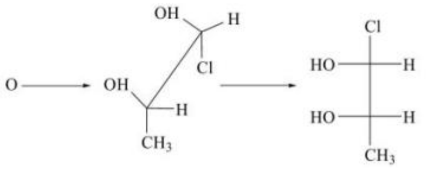
(a) M and N are non-mirror image stereoisomers. (b) M and O are identical
(c) M and P are enantiomers (d) M and Q are identical

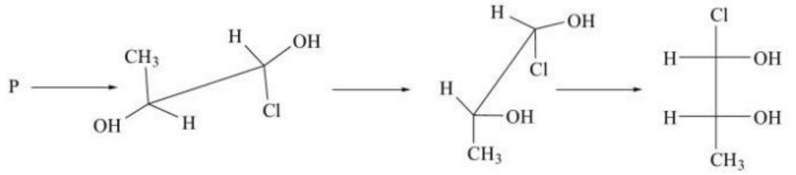
Solution :

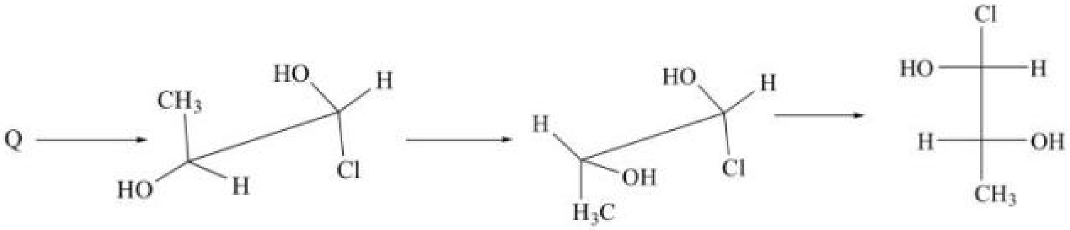
First of all, we may draw Fischer projections of the given compounds











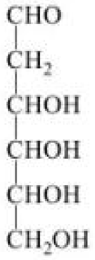
From the Fischer projections, it follows that
M and N are non-mirror images.
M and P are enantiomers.

M and O are identical.
M and Q are not identical.

Therefore the **chioces (a) ,(b) and (c)** are correct.

Aldohexose exists in D configuration

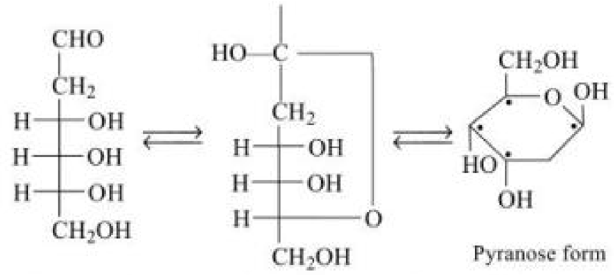
When the following aldohexose exists in D–configuration, the total number of stereoisomers in its pyranose form is



$$\begin{array}{l} CHO \\ | \\ CH_2 \\ | \\ CHOH \\ | \\ CHOH \\ | \\ CHOH \\ | \\ CH_2OH \end{array}$$

Solution:

In D-configuration, the hydroxyl group at the assymmetric carbon placed at the bottom of Fischer projection lies to the right of the straight skeleton. Thus, we will have

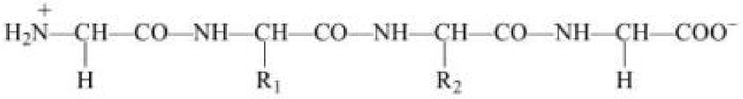


There are four chiral carbon atoms in the pyranose form of the given aldohexose. Thus, there will be 16 ($= 2^4$) stereoisomers, out of which 8 are of D-configurations (CH_2OH group above the ring) and 8 are of L-configurations (CH_2OH group below the ring).

Therefore, the correct answer is **8**.

Peptides

The substituents R_1 and R_2 for nine peptides are listed in the table given below. How many these peptides are positively charged at pH = 7.0 ?



$$\overset{+}{H_2N}-\underset{H}{CH}-CO-NH-\underset{R_1}{CH}-CO-NH-\underset{R_2}{CH}-CO-NH-\underset{H}{CH}-COO^-$$

Peptide	R_1	R_2
I	H	H
II	H	CH_3
III	CH_2COOH	H
IV	CH_2CONH_2	$(CH_2)_4NH_2$
V	CH_2CONH_2	CH_2CONH_2
VI	$(CH_2)_4NH_2$	$(CH_2)_4NH_2$
VII	CH_2COOH	CH_2CONH_2
VIII	CH_2OH	$(CH_2)_4NH_2$
IX	$(CH_2)_4NH_2$	CH_3

Solution

In general, the pH of the solution at which amino acids exist as zwitterion follows the order :

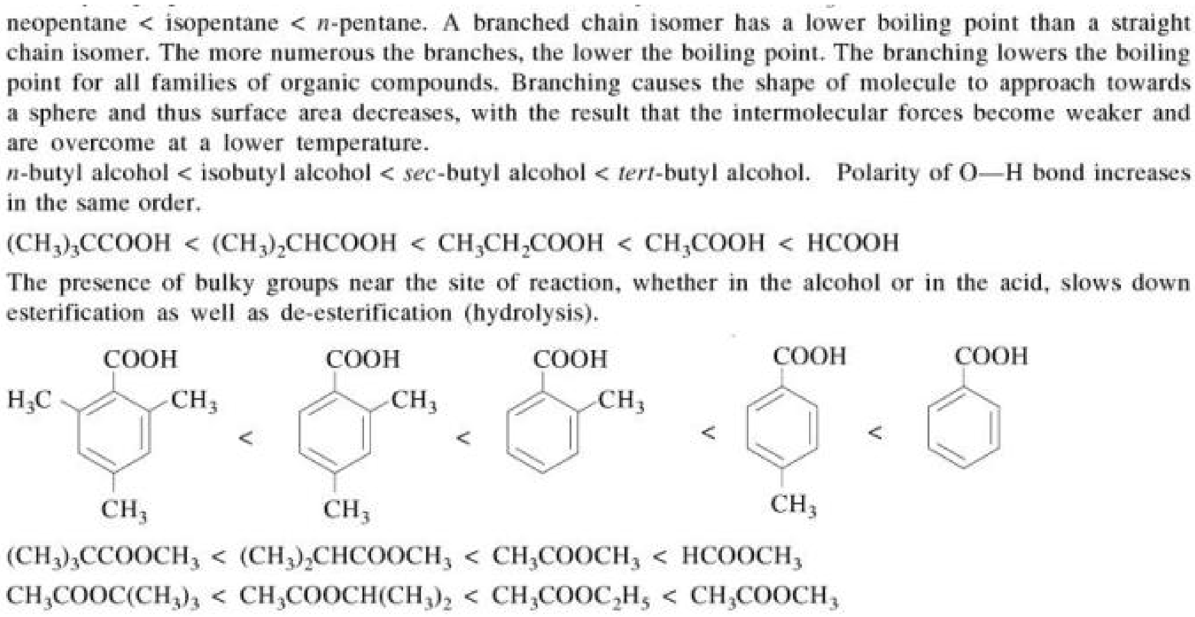
Acidic side chain < neutral chain < basic side chain
(pH = 3) (pH = 5.5 –6.0) (pH= 9–10)

This pH is known as isoelectric point.

At pH = 7, the acidic side chain (for which isoelectric point < 7) exists as a negatively–charged species (e.g. $—COO^-$) and the basic side chain (for which isoelectric point > 7) exists as a positively charged species (e.g. $—NH_3^+$). Since peptides IV, VI, VIII and IX contain $—NH_2$ group in R_1 or/and R_2 group(s), these are expected to exist as positively–charged species.

Therefore, the correct answer is **4**.

Try to figure out what sequence is depicted



neopentane < isopentane < *n*-pentane. A branched chain isomer has a lower boiling point than a straight chain isomer. The more numerous the branches, the lower the boiling point. The branching lowers the boiling point for all families of organic compounds. Branching causes the shape of molecule to approach towards a sphere and thus surface area decreases, with the result that the intermolecular forces become weaker and are overcome at a lower temperature.

n-butyl alcohol < isobutyl alcohol < *sec*-butyl alcohol < *tert*-butyl alcohol. Polarity of O—H bond increases in the same order.

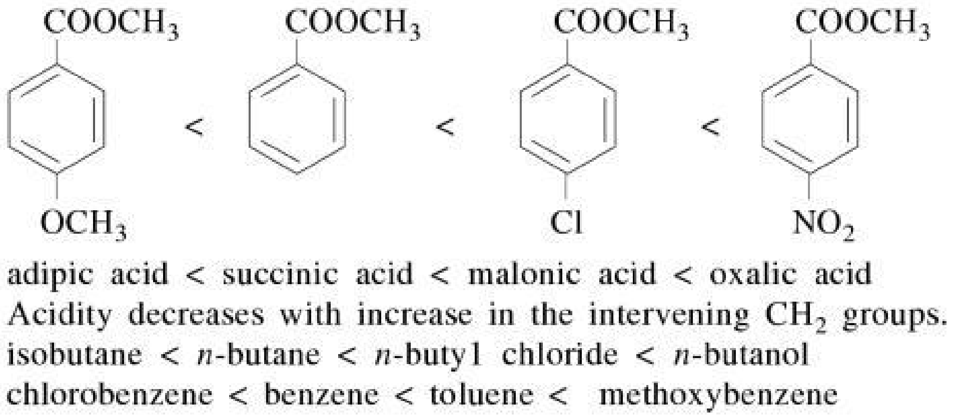
$(CH_3)_3CCOOH < (CH_3)_2CHCOOH < CH_3CH_2COOH < CH_3COOH < HCOOH$

The presence of bulky groups near the site of reaction, whether in the alcohol or in the acid, slows down esterification as well as de-esterification (hydrolysis).

$(CH_3)_3CCOOCH_3 < (CH_3)_2CHCOOCH_3 < CH_3COOCH_3 < HCOOCH_3$

$CH_3COOC(CH_3)_3 < CH_3COOCH(CH_3)_2 < CH_3COOC_2H_5 < CH_3COOCH_3$

Try to figure out the sequences

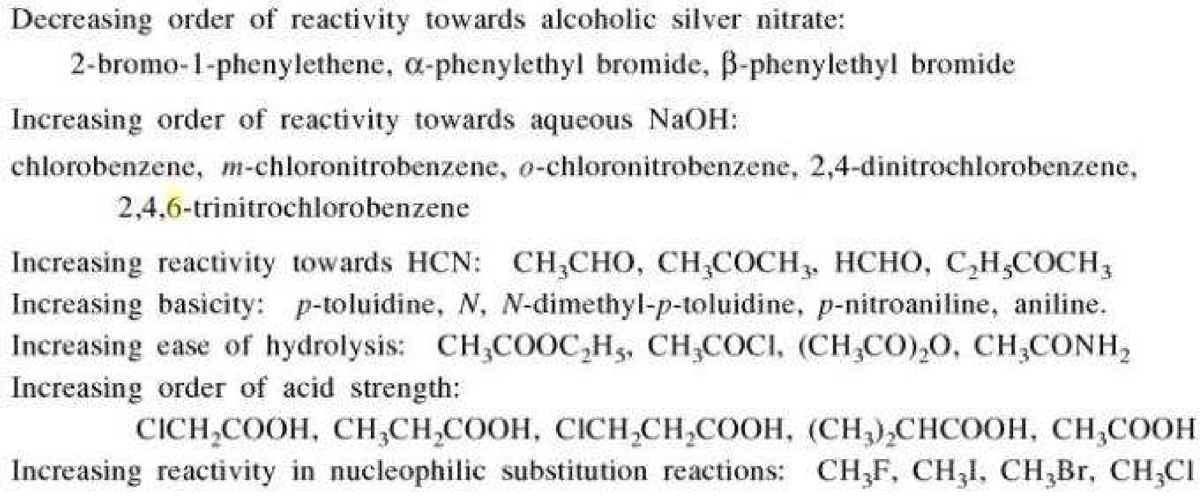


adipic acid < succinic acid < malonic acid < oxalic acid

Acidity decreases with increase in the intervening CH_2 groups.

isobutane < *n*-butane < *n*-butyl chloride < *n*-butanol

chlorobenzene < benzene < toluene < methoxybenzene



Decreasing order of reactivity towards alcoholic silver nitrate:

2-bromo-1-phenylethene, α-phenylethyl bromide, β-phenylethyl bromide

Increasing order of reactivity towards aqueous NaOH:

chlorobenzene, *m*-chloronitrobenzene, *o*-chloronitrobenzene, 2,4-dinitrochlorobenzene, 2,4,6-trinitrochlorobenzene

Increasing reactivity towards HCN: CH_3CHO, CH_3COCH_3, $HCHO$, $C_2H_5COCH_3$

Increasing basicity: *p*-toluidine, *N*, *N*-dimethyl-*p*-toluidine, *p*-nitroaniline, aniline.

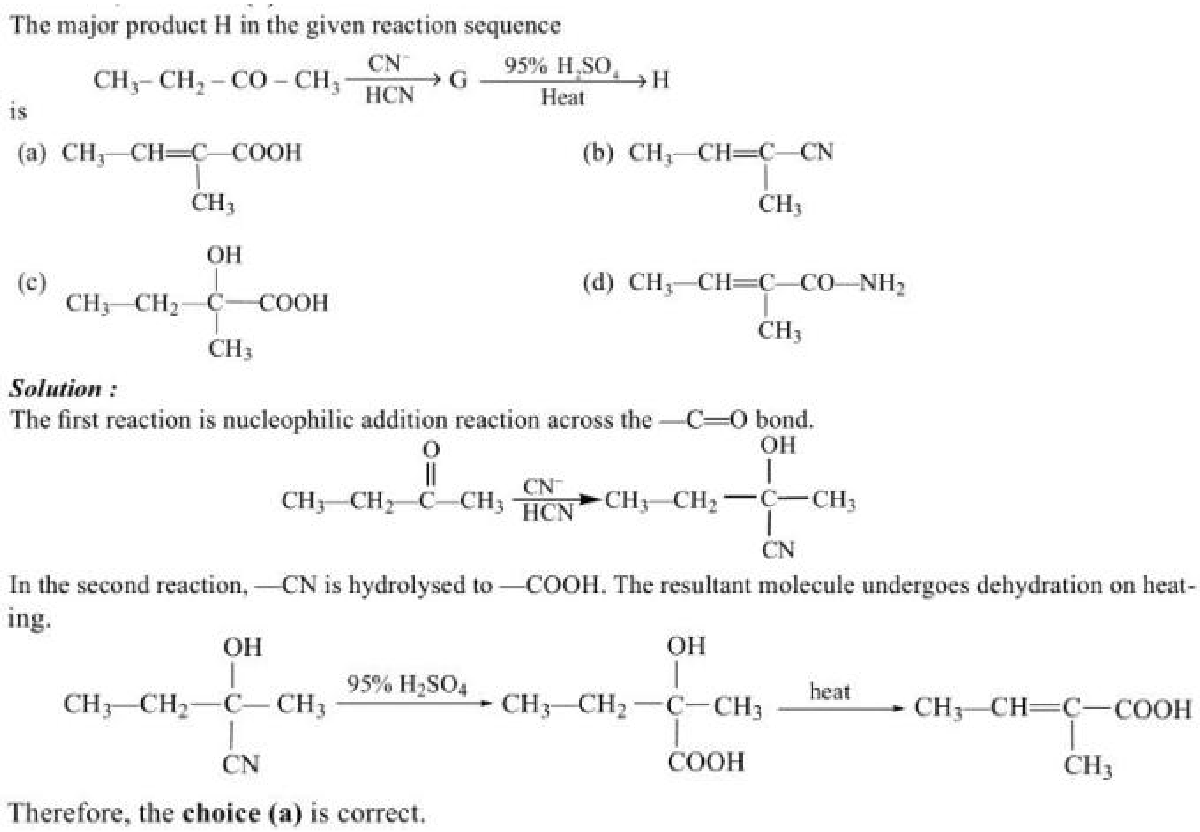
Increasing ease of hydrolysis: $CH_3COOC_2H_5$, CH_3COCl, $(CH_3CO)_2O$, CH_3CONH_2

Increasing order of acid strength:

$ClCH_2COOH$, CH_3CH_2COOH, $ClCH_2CH_2COOH$, $(CH_3)_2CHCOOH$, CH_3COOH

Increasing reactivity in nucleophilic substitution reactions: CH_3F, CH_3I, CH_3Br, CH_3Cl

major product H in the given reaction sequence



The major product H in the given reaction sequence

$$CH_3-CH_2-CO-CH_3 \xrightarrow[HCN]{CN^-} G \xrightarrow[Heat]{95\%\ H_2SO_4} H$$

is

(a) $CH_3-CH=\underset{\underset{CH_3}{|}}{C}-COOH$

(b) $CH_3-CH=\underset{\underset{CH_3}{|}}{C}-CN$

(c) $CH_3-CH_2-\overset{\overset{OH}{|}}{\underset{\underset{CH_3}{|}}{C}}-COOH$

(d) $CH_3-CH=\underset{\underset{CH_3}{|}}{C}-CO-NH_2$

Solution :

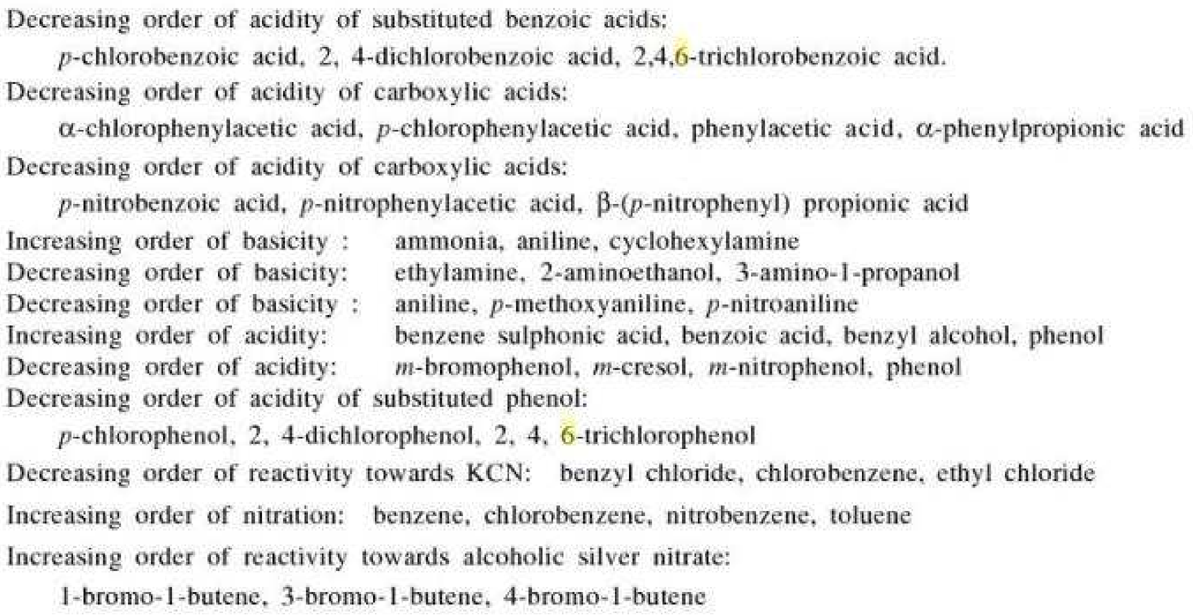
The first reaction is nucleophilic addition reaction across the —C=O bond.

$$CH_3-CH_2-\overset{\overset{O}{\|}}{C}-CH_3 \xrightarrow[HCN]{CN^-} CH_3-CH_2-\overset{\overset{OH}{|}}{\underset{\underset{CN}{|}}{C}}-CH_3$$

In the second reaction, —CN is hydrolysed to —COOH. The resultant molecule undergoes dehydration on heating.

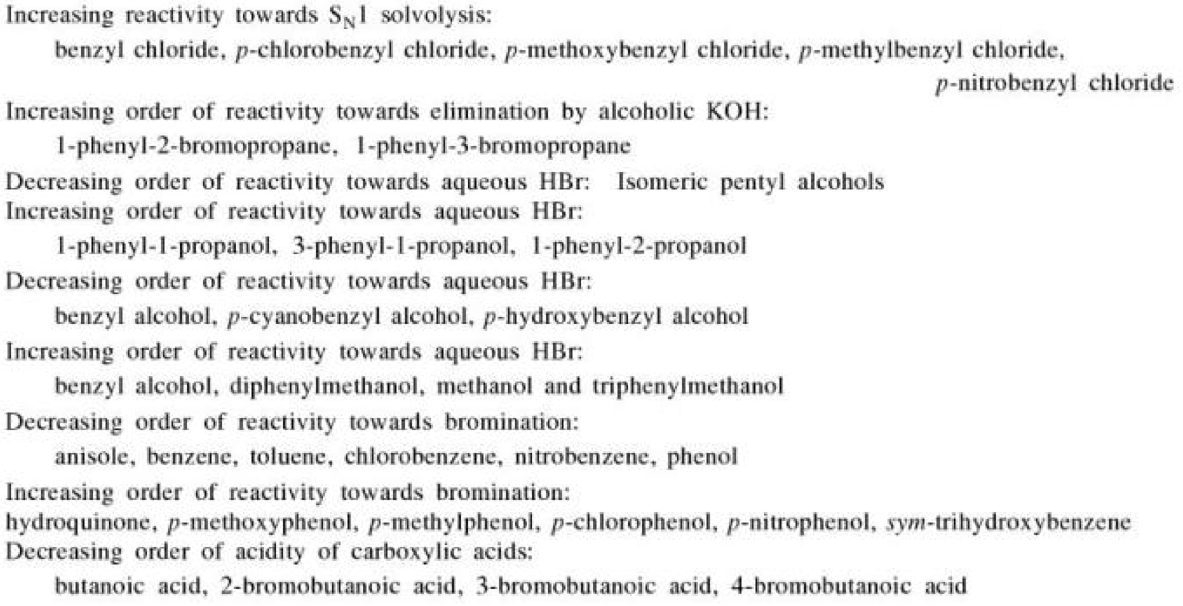
$$CH_3-CH_2-\overset{\overset{OH}{|}}{\underset{\underset{CN}{|}}{C}}-CH_3 \xrightarrow{95\%\ H_2SO_4} CH_3-CH_2-\overset{\overset{OH}{|}}{\underset{\underset{COOH}{|}}{C}}-CH_3 \xrightarrow{heat} CH_3-CH=\underset{\underset{CH_3}{|}}{C}-COOH$$

Therefore, the **choice (a)** is correct.



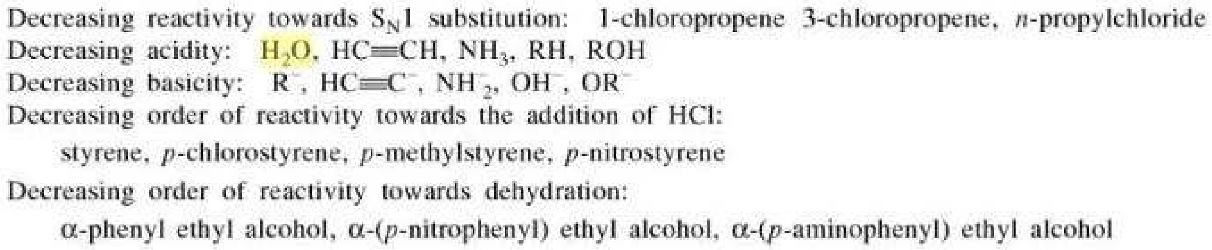
Decreasing order of acidity of substituted benzoic acids:
p-chlorobenzoic acid, 2, 4-dichlorobenzoic acid, 2,4,6-trichlorobenzoic acid.
Decreasing order of acidity of carboxylic acids:
α-chlorophenylacetic acid, *p*-chlorophenylacetic acid, phenylacetic acid, α-phenylpropionic acid
Decreasing order of acidity of carboxylic acids:
p-nitrobenzoic acid, *p*-nitrophenylacetic acid, β-(*p*-nitrophenyl) propionic acid
Increasing order of basicity : ammonia, aniline, cyclohexylamine
Decreasing order of basicity: ethylamine, 2-aminoethanol, 3-amino-1-propanol
Decreasing order of basicity : aniline, *p*-methoxyaniline, *p*-nitroaniline
Increasing order of acidity: benzene sulphonic acid, benzoic acid, benzyl alcohol, phenol
Decreasing order of acidity: *m*-bromophenol, *m*-cresol, *m*-nitrophenol, phenol
Decreasing order of acidity of substituted phenol:
p-chlorophenol, 2, 4-dichlorophenol, 2, 4, 6-trichlorophenol
Decreasing order of reactivity towards KCN: benzyl chloride, chlorobenzene, ethyl chloride
Increasing order of nitration: benzene, chlorobenzene, nitrobenzene, toluene
Increasing order of reactivity towards alcoholic silver nitrate:
1-bromo-1-butene, 3-bromo-1-butene, 4-bromo-1-butene

Explain these by solving in right order -



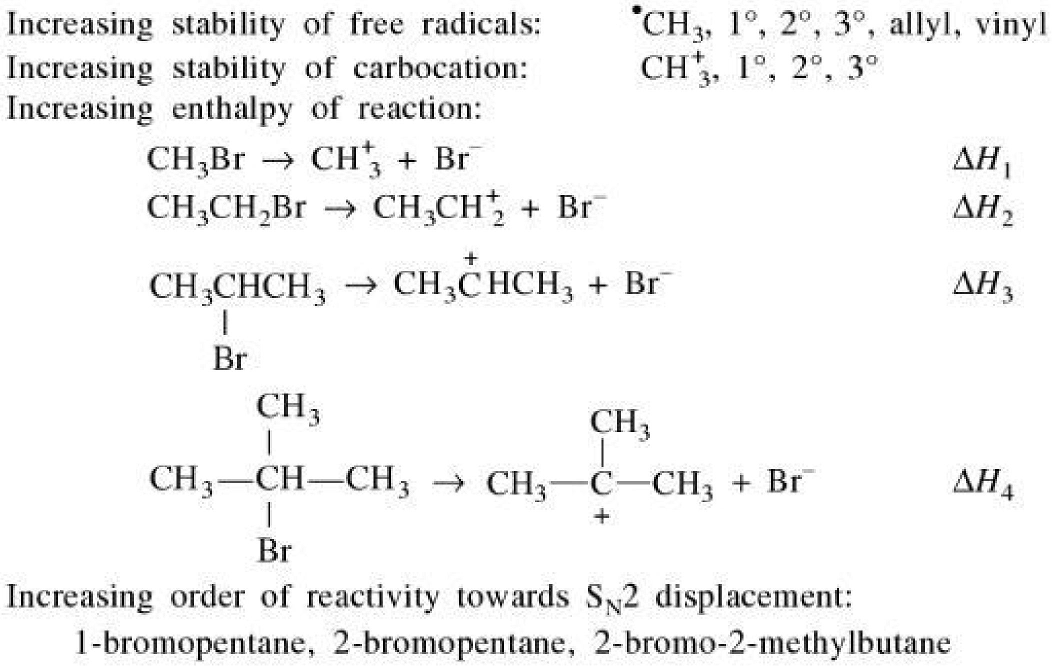
Increasing reactivity towards S_N1 solvolysis:
benzyl chloride, *p*-chlorobenzyl chloride, *p*-methoxybenzyl chloride, *p*-methylbenzyl chloride, *p*-nitrobenzyl chloride
Increasing order of reactivity towards elimination by alcoholic KOH:
1-phenyl-2-bromopropane, 1-phenyl-3-bromopropane
Decreasing order of reactivity towards aqueous HBr: Isomeric pentyl alcohols
Increasing order of reactivity towards aqueous HBr:
1-phenyl-1-propanol, 3-phenyl-1-propanol, 1-phenyl-2-propanol
Decreasing order of reactivity towards aqueous HBr:
benzyl alcohol, *p*-cyanobenzyl alcohol, *p*-hydroxybenzyl alcohol
Increasing order of reactivity towards aqueous HBr:
benzyl alcohol, diphenylmethanol, methanol and triphenylmethanol
Decreasing order of reactivity towards bromination:
anisole, benzene, toluene, chlorobenzene, nitrobenzene, phenol
Increasing order of reactivity towards bromination:
hydroquinone, *p*-methoxyphenol, *p*-methylphenol, *p*-chlorophenol, *p*-nitrophenol, *sym*-trihydroxybenzene
Decreasing order of acidity of carboxylic acids:
butanoic acid, 2-bromobutanoic acid, 3-bromobutanoic acid, 4-bromobutanoic acid

-



Decreasing reactivity towards S_N1 substitution: 1-chloropropene 3-chloropropene, *n*-propylchloride
Decreasing acidity: H_2O, $HC\equiv CH$, NH_3, RH, ROH
Decreasing basicity: R^-, $HC\equiv C^-$, NH^-_2, OH^-, OR^-
Decreasing order of reactivity towards the addition of HCl:
styrene, *p*-chlorostyrene, *p*-methylstyrene, *p*-nitrostyrene
Decreasing order of reactivity towards dehydration:
α-phenyl ethyl alcohol, α-(*p*-nitrophenyl) ethyl alcohol, α-(*p*-aminophenyl) ethyl alcohol

Various Increasing Properties



Increasing stability of free radicals: $^{\bullet}CH_3$, 1°, 2°, 3°, allyl, vinyl
Increasing stability of carbocation: CH_3^+, 1°, 2°, 3°
Increasing enthalpy of reaction:

$$CH_3Br \rightarrow CH_3^+ + Br^- \qquad \Delta H_1$$

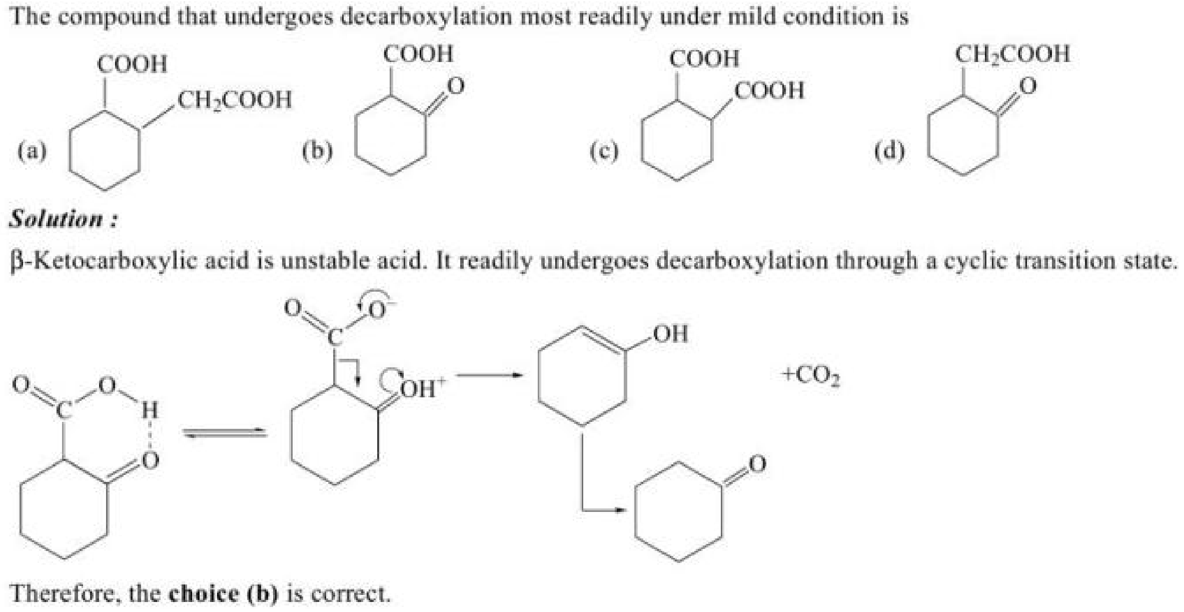
$$CH_3CH_2Br \rightarrow CH_3CH_2^+ + Br^- \qquad \Delta H_2$$

$$CH_3\underset{\displaystyle Br}{\underset{|}{C}}HCH_3 \rightarrow CH_3\overset{+}{C}HCH_3 + Br^- \qquad \Delta H_3$$

$$CH_3-\underset{\displaystyle Br}{\underset{|}{\overset{\displaystyle CH_3}{\overset{|}{C}}}}H-CH_3 \rightarrow CH_3-\underset{+}{\overset{\displaystyle CH_3}{\overset{|}{C}}}-CH_3 + Br^- \qquad \Delta H_4$$

Increasing order of reactivity towards S_N2 displacement:
1-bromopentane, 2-bromopentane, 2-bromo-2-methylbutane

compound undergoes decarboxylation most readily

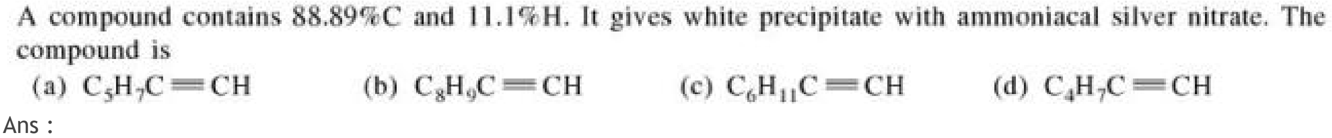


The compound that undergoes decarboxylation most readily under mild condition is

Solution :

β-Ketocarboxylic acid is unstable acid. It readily undergoes decarboxylation through a cyclic transition state.

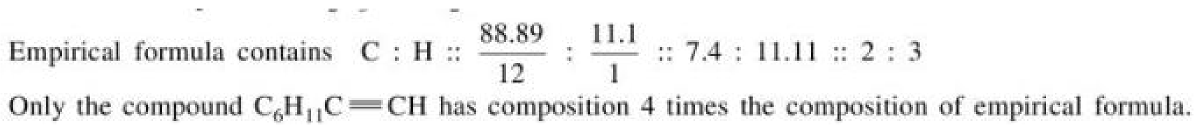
Therefore, the **choice (b)** is correct.



A compound contains 88.89%C and 11.1%H. It gives white precipitate with ammoniacal silver nitrate. The compound is

(a) $C_5H_7C{\equiv}CH$ (b) $C_8H_9C{\equiv}CH$ (c) $C_6H_{11}C{\equiv}CH$ (d) $C_4H_7C{\equiv}CH$

Ans :



Empirical formula contains C : H :: $\frac{88.89}{12}$: $\frac{11.1}{1}$ :: 7.4 : 11.11 :: 2 : 3

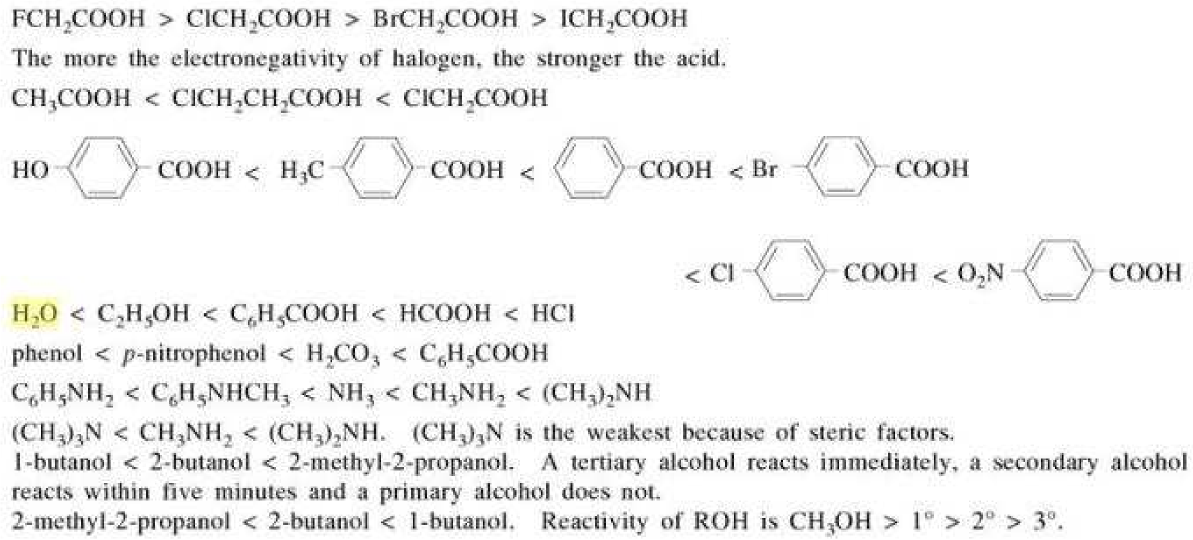
Only the compound $C_6H_{11}C{\equiv}CH$ has composition 4 times the composition of empirical formula.

More the number of electron withdrawing group stronger is the acid

$Cl_3CCOOH > Cl_2CHCOOH > ClCH_2COOH > CH_3COOH$

The more the number of electron-withdrawing group, the stronger the acid.

Try to figure out what sequence is depicted



$FCH_2COOH > ClCH_2COOH > BrCH_2COOH > ICH_2COOH$

The more the electronegativity of halogen, the stronger the acid.

$CH_3COOH < ClCH_2CH_2COOH < ClCH_2COOH$

$H_2O < C_2H_5OH < C_6H_5COOH < HCOOH < HCl$

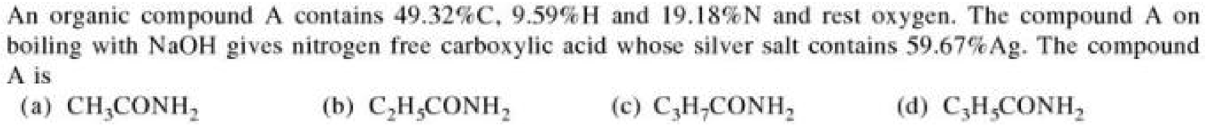
phenol < *p*-nitrophenol < $H_2CO_3 < C_6H_5COOH$

$C_6H_5NH_2 < C_6H_5NHCH_3 < NH_3 < CH_3NH_2 < (CH_3)_2NH$

$(CH_3)_3N < CH_3NH_2 < (CH_3)_2NH$. $(CH_3)_3N$ is the weakest because of steric factors.

1-butanol < 2-butanol < 2-methyl-2-propanol. A tertiary alcohol reacts immediately, a secondary alcohol reacts within five minutes and a primary alcohol does not.

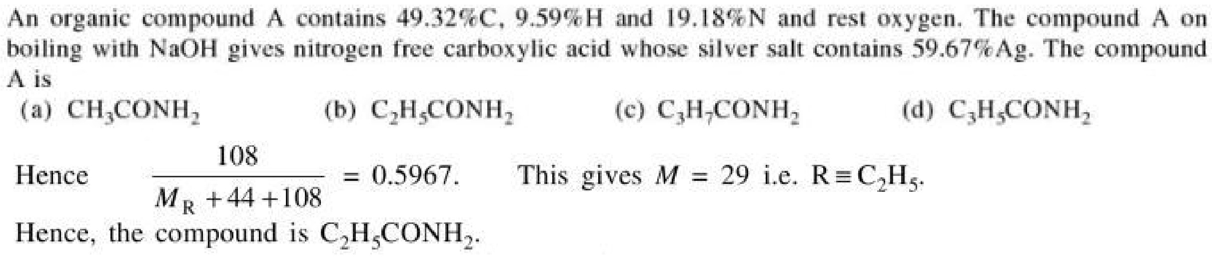
2-methyl-2-propanol < 2-butanol < 1-butanol. Reactivity of ROH is CH_3OH > 1° > 2° > 3°.



An organic compound A contains 49.32%C, 9.59%H and 19.18%N and rest oxygen. The compound A on boiling with NaOH gives nitrogen free carboxylic acid whose silver salt contains 59.67%Ag. The compound A is

(a) CH_3CONH_2 (b) $C_2H_5CONH_2$ (c) $C_3H_7CONH_2$ (d) $C_3H_5CONH_2$

Ans :



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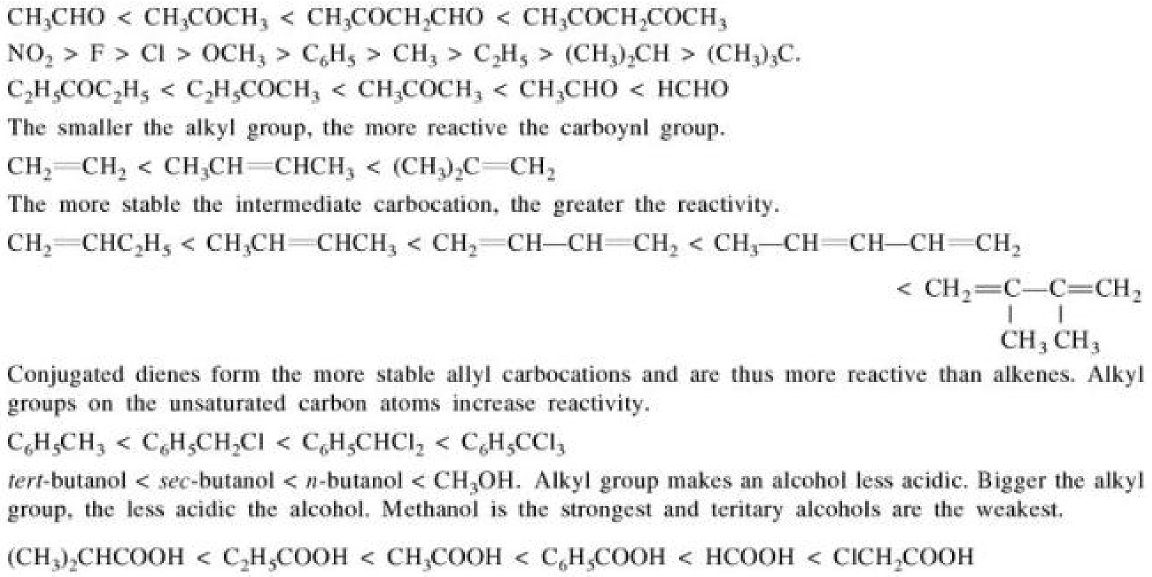
Hence $\frac{108}{M_R + 44 + 108} = 0.5967$. This gives $M = 29$ i.e. $R \equiv C_2H_5$.

Hence, the compound is $C_2H_5CONH_2$.

Try to figure out what sequence is depicted



$CH_3CHO < CH_3COCH_3 < CH_3COCH_2CHO < CH_3COCH_2COCH_3$

$NO_2 > F > Cl > OCH_3 > C_6H_5 > CH_3 > C_2H_5 > (CH_3)_2CH > (CH_3)_3C.$

$C_2H_5COC_2H_5 < C_2H_5COCH_3 < CH_3COCH_3 < CH_3CHO < HCHO$

The smaller the alkyl group, the more reactive the carboynl group.

$CH_2{=}CH_2 < CH_3CH{=}CHCH_3 < (CH_3)_2C{=}CH_2$

The more stable the intermediate carbocation, the greater the reactivity.

$$CH_2{=}CHC_2H_5 < CH_3CH{=}CHCH_3 < CH_2{=}CH{-}CH{=}CH_2 < CH_3{-}CH{=}CH{-}CH{=}CH_2 < CH_2{=}\underset{\large CH_3}{\underset{|}{C}}{-}\underset{\large CH_3}{\underset{|}{C}}{=}CH_2$$

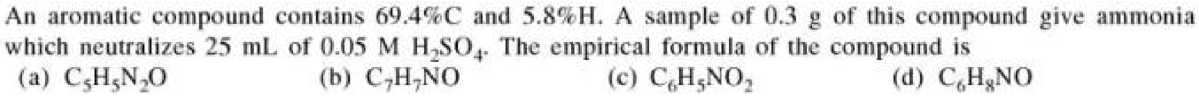
Conjugated dienes form the more stable allyl carbocations and are thus more reactive than alkenes. Alkyl groups on the unsaturated carbon atoms increase reactivity.

$C_6H_5CH_3 < C_6H_5CH_2Cl < C_6H_5CHCl_2 < C_6H_5CCl_3$

tert-butanol < *sec*-butanol < *n*-butanol < CH_3OH. Alkyl group makes an alcohol less acidic. Bigger the alkyl group, the less acidic the alcohol. Methanol is the strongest and teritary alcohols are the weakest.

$(CH_3)_2CHCOOH < C_2H_5COOH < CH_3COOH < C_6H_5COOH < HCOOH < ClCH_2COOH$

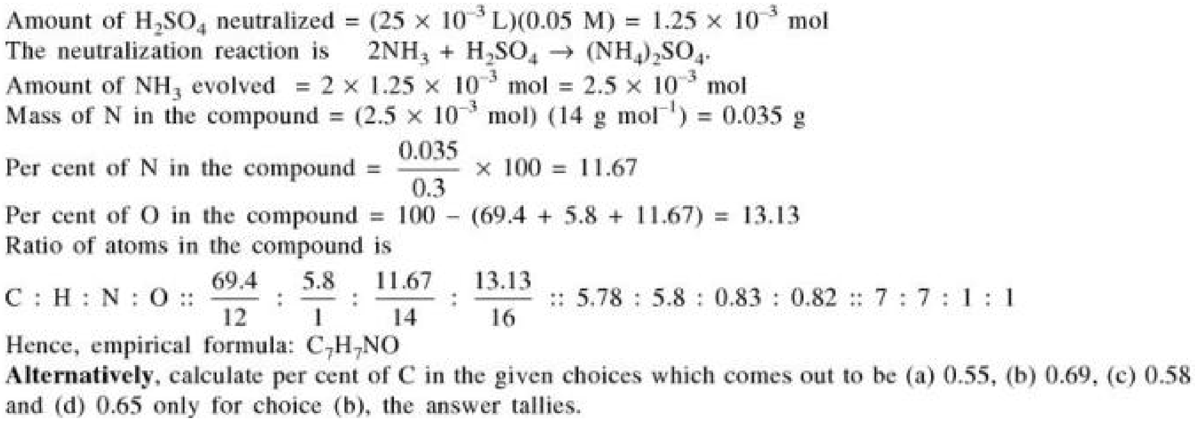
An aromatic compound contains 69.4% C and 5.8% H



An aromatic compound contains 69.4%C and 5.8%H. A sample of 0.3 g of this compound give ammonia which neutralizes 25 mL of 0.05 M H_2SO_4. The empirical formula of the compound is

(a) $C_5H_5N_2O$ (b) C_7H_7NO (c) $C_6H_5NO_2$ (d) C_6H_8NO

Ans :



Amount of H_2SO_4 neutralized = $(25 \times 10^{-3}\,L)(0.05\,M) = 1.25 \times 10^{-3}$ mol

The neutralization reaction is $2NH_3 + H_2SO_4 \rightarrow (NH_4)_2SO_4$.

Amount of NH_3 evolved $= 2 \times 1.25 \times 10^{-3}\ \text{mol} = 2.5 \times 10^{-3}$ mol

Mass of N in the compound $= (2.5 \times 10^{-3}\ \text{mol})\ (14\ \text{g mol}^{-1}) = 0.035$ g

$$\text{Per cent of N in the compound} = \frac{0.035}{0.3} \times 100 = 11.67$$

Per cent of O in the compound $= 100 - (69.4 + 5.8 + 11.67) = 13.13$

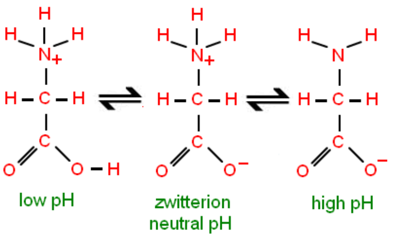
Ratio of atoms in the compound is

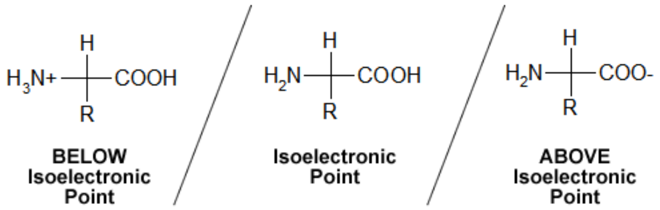
$$C : H : N : O :: \frac{69.4}{12} : \frac{5.8}{1} : \frac{11.67}{14} : \frac{13.13}{16} :: 5.78 : 5.8 : 0.83 : 0.82 :: 7 : 7 : 1 : 1$$

Hence, empirical formula: C_7H_7NO

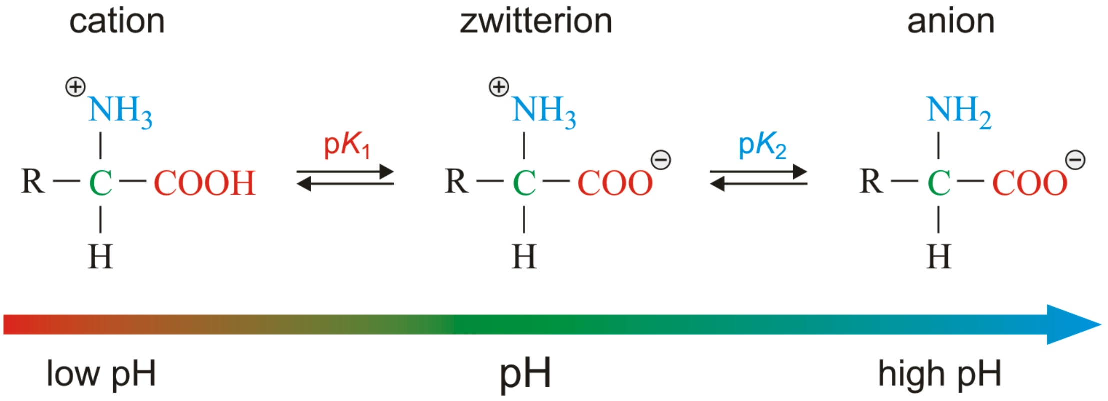
Alternatively, calculate per cent of C in the given choices which comes out to be (a) 0.55, (b) 0.69, (c) 0.58 and (d) 0.65 only for choice (b), the answer tallies.

Zwitter Ions - The Hydrogen atom (rather ion) moves to different positions depending on the PH (So can have a positive charge at some position, or a negative charge at some other position, depending on the pH)



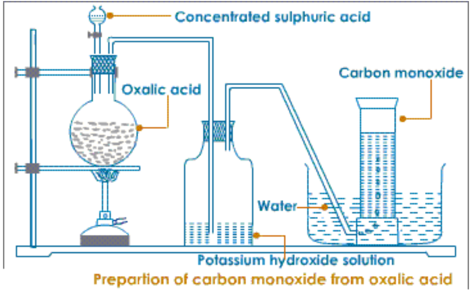


Electrophoresis - is a method to confirm if a positive charge is present or a negative charge is present in the ion.



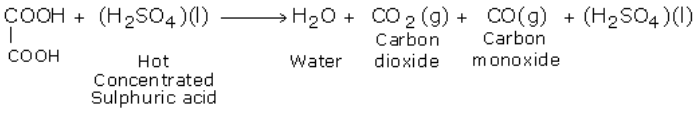
Iso -Electric point - the pH at which the ion (molecule) does not migrate to any electrodes.

What happens if oxalic acid is heated with conc sulphuric acid ?



Prepartion of carbon monoxide from oxalic acid

We get Carbonmonoxide



$$\underset{\text{COOH}}{\overset{\text{COOH}}{|}} + \underset{\text{Hot Concentrated Sulphuric acid}}{(H_2SO_4)(l)} \longrightarrow \underset{\text{Water}}{H_2O} + \underset{\text{Carbon dioxide}}{CO_2(g)} + \underset{\text{Carbon monoxide}}{CO(g)} + (H_2SO_4)(l)$$

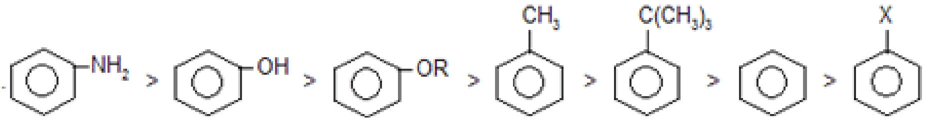
IMPORTANT ORDER AND FACTS OF ORGANIC CHEMISTRY

1. RCOCl > RCOOCOR > RCOOR > $RCONH_2$ Nucleophilic substitution reaction.

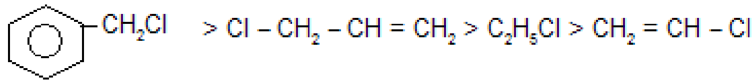
2. HI > HBr > HCl > RCOOH > C_6H_5OH > H_2O > $CH \equiv CH$ > NH_3 (Acidic nature).

3. $CCl_3CHO > HCHO > CH_3CHO > CH_3COCH_3$ Nucleophilic addition reaction.

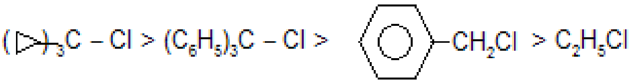
4. $CH_2 = CH_2 > CH \equiv CH > C_6H_6$ Electrophilic addition reaction.

5. 

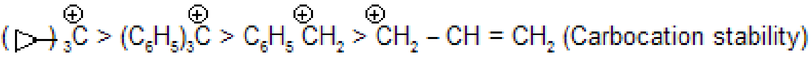
Electrophilic Substitution Reaction

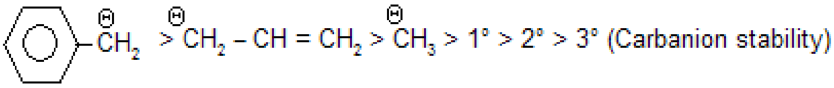
6. 

Nucleophilic Substitution Reaction

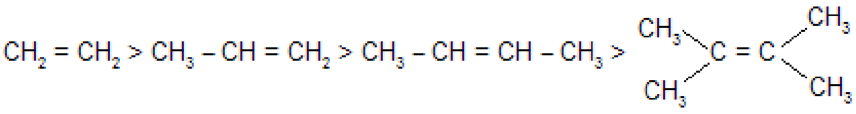
7. 

Nucleophilic Substitution Reaction

8. 
$(\triangleright\!\!-)_3\overset{\oplus}{C} > (C_6H_5)_3\overset{\oplus}{C} > C_6H_5\overset{\oplus}{C}H_2 > \overset{\oplus}{C}H_2 - CH = CH_2$ (Carbocation stability)

9. 

10. $(CH_3)_2C = C(CH_3)_2 > CH_3 - CH = C -(CH_3)_2 > CH_3 - CH = CH - CH_3 > CH_3 - CH = CH_2$ (Stability)

11. 

(Heat of Hydration)

12. $\overset{\ominus}{N}H_2 > \overset{\ominus}{O}H > NH_3 > H_2O$ (Basic strength)

13. $NI_3 > NBr_3 > NCl_3 > NF_3$ (Basic strength)

14. $Br_2 > Cl_2 > I_2$ (Selectivity for halogenation)

15. Halogenation of alkenes by cyclic halonium state, so anti attack takes place.

16. Hydroboration followed by oxidation is always anti markownikoff's addition due to steric effect.

17. Oximercuration - demercuration is markownikoff's addition of water because some carbocation character in cyclic mercurium state.

18. $CHCl_3$ in the presence of strong bases forms biradical : CCl_2 which undergo addition with double or triple bonds.

19. When conjugated diene reacts with alkene or alkyne it is known as diel's elder synthesis.

20. Ozonolysis of cyclo alkene forms one mole dialdehyde while ozonolysis of cyclo alkadiene forms two moles of dialdehyde.

21. Ozonolysis with $(CH_3)_2S$ is known as reductive ozonolysis.

22. Hydration of alkyne occur's in $HgSO_4$ and dil H_2SO_4.

23. 1-alkynes forms ppt with $Ag(NH_3)_2^{\oplus}$ and $Cu(NH_3)_4^{+}$.

24. Cis-2-butene reacts with Br2 to forms dl(±) pair of enantiomers of 2,3-dibromobutane while in case of trans-2-butene forms meso-2,3-dibromo butane due to anti addition always.

25. Haloform test given by species with CH_3CO-group but not in case of A.A.E. and tert. Butyl alcohol.

26. Chloral reacts with chloro benzene in con. H_2SO_4 to form insectiside DDT.

27. NBS is used for free radical allylation.

28. Rate for S_N^1 reaction is $3^\circ > 2^\circ > 1^\circ$ in protic polar solvent.

29. Rate for S_N^2 reaction is $1^\circ > 2^\circ > 3^\circ$ in polar aprotic solvent like DMSO, DMF, HMPT.

30. Chemical reactions like Hoffmann carbylamine and Reimer Tiemann's reaction active species is biradical CCl_2.

31. If cyclo 1,3-penta diene reacts with $CHCl_3$ and potassium tert. butoxide to form chlorobenzene.

32. Alkyl halides reacts with AgCN to form isocyanides due to ambident nature of nucleophile, other ambident nucleophiles are $\overset{\ominus}{N}O_2$and SO_3^{-2}.

33. In dehydration of alcohols active species is carbocation so rearrangement occurs like hydride shift or alkyl shift.

34. Dehydration of cyclobutyl methyl alcohol ring expansion takes place, formation of cyclo pentene occurs.

35. In esterification where acid reacts with alcohol to form ester, - OH given by acid while - H by alcohol

$$CH_3CO\overset{18}{O}H + C_2H_5\overset{16}{O}H \xrightarrow{\overset{\oplus}{H}} CH_3COOC_2H_5 + H_2\overset{18}{O}$$

36.Ether's reacts with HI to form alcohol and halide where fission of lower ether by S_N^2 mechanism while higher ethers like ter. butyl methyl ether or alkyl methyl ether by S_N^1 mechanism.

37. Quantitative estimation of ethers is done by ziesal's method.

38. If unsym. cyclic ether undergo fission it depends upon medium weather it is acid or basic like in acidic medium some character of carbocation so nucleophile goes to carbon where more alkyl groups are there while reverse in basic medium due to steric factor.

39. Aldehydes are reducing agent while ketones are not.

40. Aldehydes and ketones are separated by tollen's reagent.

41. Carbonyl and noncarbonyl are seprated by sodium bi sulphite and bradye's reagent.

42. As the size of alkyl group increases steric hindrance comes into play, reactivity towards nucleophilic addition decreases.

43. Aldehydes with $\propto$hydrogen atom in the presence of dil base undergo enolization and form $\propto$

carbonian to give aldol product.

44. If there is two-CHO group with ∝hydrogen atom to form cyclic intra aldol product.

45. Aldehydes without -hydrogen atom in the presence of con. alkali to form each molecule of acid and alcohol by hydride active species.

46. Glyoxal reacts with con. KOH to form glycolate ion by Intra Cannizaro's.

47. Pinacol pinacolone type reactions involve protonation, deprotonation and alkyl shift.

48. Aldol, Perkin, reformatsky and knovengel's reactions are classified as carbanian active reaction.

49. In Beckmann's rearrangement migration of group which is anti to-OH group takes place.

50. Beckmann's rearrangent is a reaction of oximes in the presence of H_2SO_4 or PCl_5 to produce N-alkyl amide derivatives.

51. Migratory attitude of alkyl group in Pinacol-pinacolone, beckmann's and bayer villegar oxidation is -C_6H_5 > $(CH_3)_3C$- > $(CH_3)_2CH$- > C_2H_5- > CH_3-

52. Cyclo hexanoneoxime on beckmann's reaction gives caprolectum which on reaction with $\overset{\oplus}{H}$ to give polymer nylon-6

53. 2-methyl propanal even contains ∝-hydrogen atom but does not give aldol reaction.

54. Diphenyl glyoxal reacts with con. KOH to form salt of benzillic acid.

55. Propanone in the presence of dry HCl gas by enolic intermediate to form diacetone alcohol undergo heating form mesityl oxide. If this again reacts with propanone to form phoron.

56. Benzaldehyde reacts with alc. KCN to form Benzoin which on oxidation form benzil.

57. By wittig reaction carbonyl compounds are converted into E-Z form of alkene.

58. Benzaldehyde do not respond to benedict's and fehling's solution due to less redusing power of aromatic aldehyde.

59. HCOOH respond to oxidising agent due to presence of - CHO group.

60. $HI > H_2SO_4 > HNO_3 > RCOOH > H_2CO_3 > C_6H_5OH > H_2O > CH \equiv CH > NH_3 > CH_3 - C \equiv CH > CH_2 = CH_2 > H_2$ this is decreasing acidic nature.

61. Acids with ∝-hydrogen atom when reacts with halogen in the presence of P to form ∝-haloacid (HVZ).

62. Carboxyllic acid on reaction with PCl_5, CH_2N_2 and H_2O to form higher acid or next homologue.

63. For reactivity of acid derivatives use funda weaker the base better the leaving group.

64. Acid amide on reaction with PCl_5 to form alkane nitrile.

65. Anhydride on reaction with carbonyl compound in the presence of base (carbanian) forms ∝, β-unsaturated carboxyllic acid (perkin reaction)

66. Amide on reaction with Br_2 and alkali to form primary amine of lower homologue. Intermediate species is nitrene which undergo intra rearrangement to form RNH_2.

67. Ester's with $\propto$-hydrogen atom in the presence of strong base to form carbanian undergo nucleophilic substitution reaction forms β-keto ester for example ethyl acetate in the presence of pot. ter. Butoxide form aceto acetic ester (AAE). reaction is known as clesen's ester condensation of four types

(a) Simple clesen's ester condensation.
(b) CROSS clesen's ester condensation.
(c) Intra CEC (Dieckmann's condensation).
(d) mixed clesen ester condensation.

68. For a compound to be aromatic it must be cyclic, planar and obey $(4n + 2)\pi$ electron rule.

69. Cyclooctatetraene is non aromatic compound while pyrrole, pyredene, furan, cyclopentadieneylanion all are aromatic.

70. In aromatic electrophilic substitution reaction there is no hydrogen isotopic effect except sulphonation and iodination.

71. m- directing groups like nitrobenzene and benzaldehyde cannot undergo fridal craft reaction.

72. Phenol is less acidic than general carboxylic acid, cannot react with $NaHCO_3$.

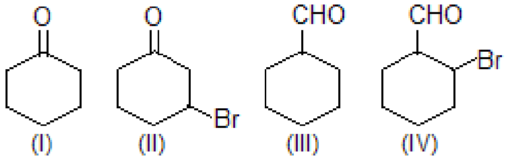
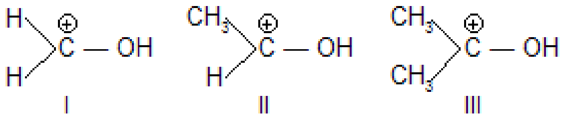
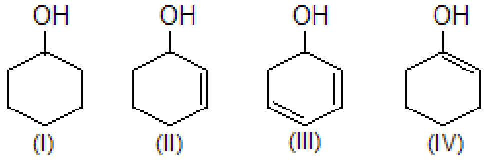
73. Anilene is more reactive than phenol towards electrophilic substitution reaction because less energy difference between nitrogen and carbon.

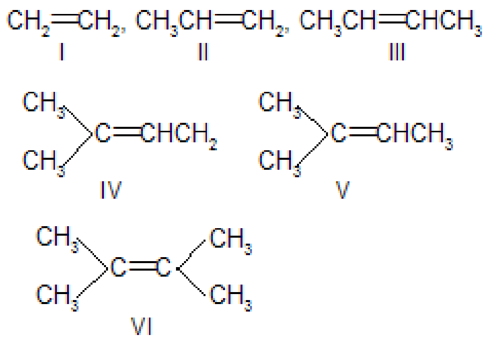
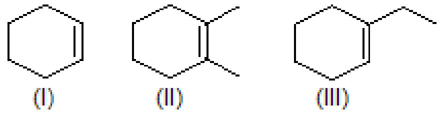
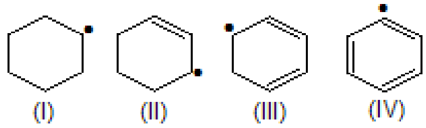
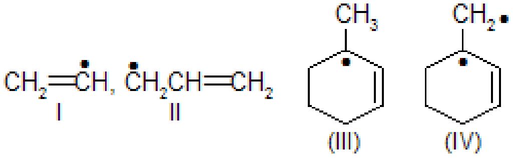
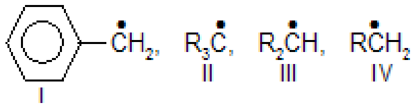
74. Benzene diazonium chloride reacts with phenol or anilene to form azo compounds.

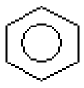
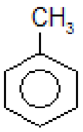
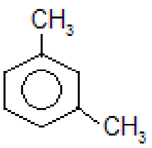
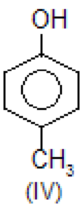
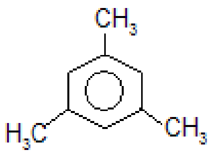
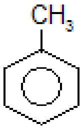
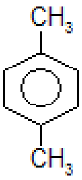
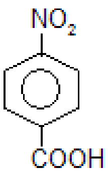
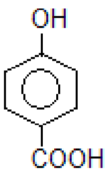
75. $C_6H_5NH_2 > C_6H_5OH > C_6H_5OR > C_6H_5Br > C_6H_5NO_2$ (Reactivity towards electrophilic substitution reaction

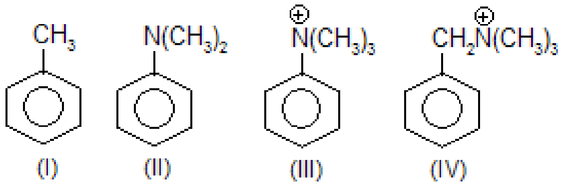
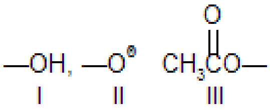
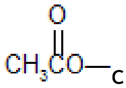
General Series - Chemistry Fact Sheet - 3

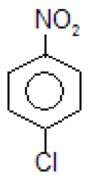
General series	**Order**	**Why ?**	
1.	**B.P. of** $\underset{I}{CH_3CH_2OH}$, $\underset{II}{CH_3OCH_3}$, $\underset{III}{CH_3CH_3}$	I > II > III	There is intermolecular H-bonding I. III has weak force of attraction and is most volatile.
2.	**B.P. of *o*, *m*, *p*-nitro phenol**	o < m < p	Intramolecular H-bonding in o-isomer makes it more volatile.

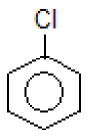
3.	**Reactivity of … with Tollen's reagent** $HCHO$, CH_3CHO, CH_3COCH_3, C_6H_5CHO I, II, III, IV	I > II > IV > III	—CHO group is easily oxidised compared to keto group due to redusing hydrogen.
4.	**Reactivity of … with Fehling's solution** I II III IV (above)	I > II > IV > III	—do—
5.	**Extent of hydration of** 	I < II < III < IV	Aldehydes are more hydrated than ketones. Halide makes C of carbonyl group more electropositive.
6.	**Electrophilic nature of …….. for nucleophilic attack** 	I > II > III	CH_3 group decreases +ve charge on C hence nucleophilic attack.
7.	**Reactivity of isomeric 1°, 2°, 3° butyl halide towards elimination (E1 or E2)**	3° < 2° < 1°	due to stability of intermediate carbocation
8.	**Dehydration of** 	IV < I < II < III	Alcohol leading to increase in conjugation due to dehydration is more easily dehydration is more easily dehydrated. IV is vinylic, hence least.
9.	**Stability of**		

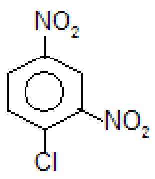
		I < II < III < IV < V < VI	Substituted alkenes are more stable.More the alkyl groups are attached to the doubly bonded carbon atom more is the stability.
10.	**Stability of** 	I < III < II	II is more substituted than III (More hyperconjugation more stability)
11.	**Stability of** 	III > II > I > IV	IV is vinylic while in conjugative, II allylic.
12.	**Stability of** 	I < IV < II < III	III is 3° allylic and II is 1° allylic
13.	**Dehydration of** 1°, 2°, 3° isomeric butyl alcohol	3° < 2° < 1°	More the stability of intermediate, greater the reactivity of chemical reaction.
14.	**Boiling points of** *n*-butyl amine, *n*-butyl alcohol, *n*-pentane I II III	II > I > III	I, II have H-bonding but electronegativity of O > N hence H-bonding in II > I
15.	**Formation of** 	I > II > III > IV (easiest I)	greater the stability, easier the formation of perticular species.
16.	**Reactivity of C—H bond (abstraction of H)**		

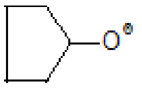
	$CH_2{=}CH{-}\mathbf{H}$, vinylic (I) $CH_3{-}\mathbf{H}$, methyl (II) 1° $CH_3CH_2{-}\mathbf{H}$ (III) $(CH_3)_2CH{-}\mathbf{H}$ 2° (IV) $(CH_3)_3C{-}\mathbf{H}$, 3° (V) $CH_2{=}CHCH_2{-}\mathbf{H}$ (allylic) (VI)	I < II < III < IV < V < VI	Vinyl < methyl 1° < 2° < 3° < allylic
17.	**Leaving nature (tendency) of … in S_N reaction.** H^-, R^-, MeO^-, OH^-, CN^-, CH_3COO^-, I II III IV V VI $ROSO_2^-$, $ArSO_3^-$ VII VIII	I < II < III ~ IV < V < VI < VII < VIII	If acid is strong, its conjugate base is weak and greater the leaving tendency.
18.	**Rate of esterification of the following acids with MeOH** $MeCH_2COOH$, $Me_2CHCOOH$, $MeCCOOH$, I II III Et_3CCOOH $(i\text{-}Pr)_2CHCOOH$ IV V	I > II > III > IV > V	As the size of the substituents on the $\alpha{-}C$ increases, the tetrahedrally bonded interme. diate becomes more crowded and these slower the rate.
19.	**Relative reactivity of … with electrophile in S_E reaction**  (I)  (II)  (III)  (IV)  (V)	I > II > IV > III > V	$-CH_3$ is *o*-, *p*-directing and responsible for activation.
20.	**Relative reactivity of these compunds with electrophile inS_E reaction**  (I)  (II)  (III)  (IV)	II > I > III > IV	$-CH_3$ is *o*-, *p*-directing due to activation while $-COOH$ is m-directing and deactivating group.
21.	**Relative reactivity of … with electrophile in S_E reaction.**	II > I > IV > II	As the number of sp^3 hybridised C

			atoms separating the ring from the positively charged substituent increases, deactivating effect decreases due to less electronegativity.
22.	**Activating effects of the following o, p-directors.**  $-OH, -O^{\ominus}, CH_3\overset{O}{\overset{\|}{C}}O-$ I II III	II > I > III	$-O^{\ominus}$ is best able to donate electrons there by giving a very stable uncharged intermediate. In  $CH_3\overset{O}{\overset{\|}{C}}O-$ cross conjugation diminished its ability to donate electrons to an arenium ion.
23.	**Relative reactivity of ... towards S_N1 reaction** benzyl chloride, *p* - methoxy benzyl chloride I II and *p* - nitro benzyl chloride III	II > I > III	Intermediates are benzylic cations. So CH_3O(electron repelling) gives greater stability through delocalisation while NO_2 (electron attracting) decreases stability.
24.	**Relative reactivity of ... towards S_N1 and S_N2 reaction** $PhCH_2Cl$ $PhCHClMe$ $PhCClMe_2$ (I) (II) (III) (Ph stands for phenyl, C_6H_5)	**S_N1 :** III > II > I **S_N2 :** II < II < I	**S_N1 :** 1° < 2° < 3° alkyl halide **S_N2 :** 3° < 2° < 1° alkyl halide
25.	**Relative reactivity of ... with E^+ (electrophile) in S_E reaction.**	II > I > III	$-NO_2$ deactivates benzene ring for S_E

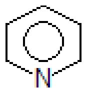
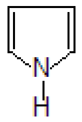
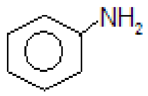
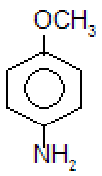
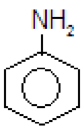
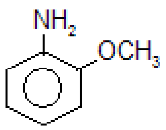
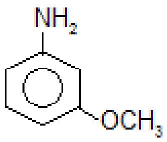
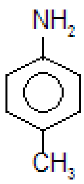
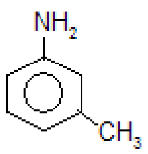
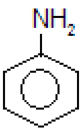
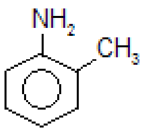

(I)

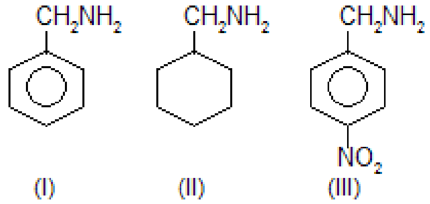
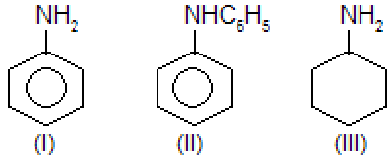
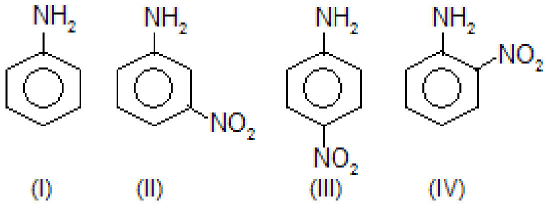
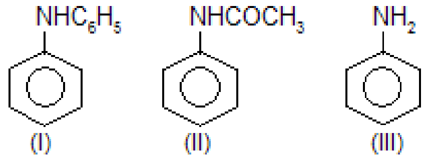
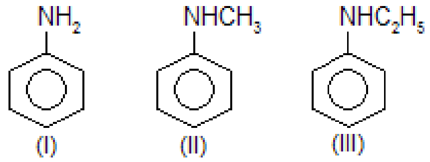

(II)

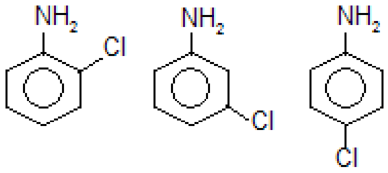
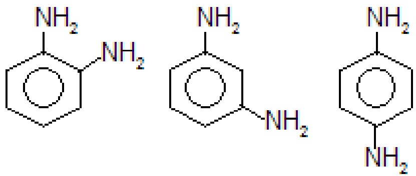

(III)

26.	**Order of S_N2 reactivity of alkoxide nucleophiles** Me_3CO^- (I), MeO^- (II), $MeCH_2O^-$ (III), Me_2CHO^- (IV)  V	I < IV < V < III < II	S_N2 reactivity is suseptible to steric hindrance by the nucleophile as well as by the size of alkyl group.

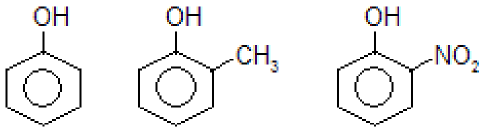
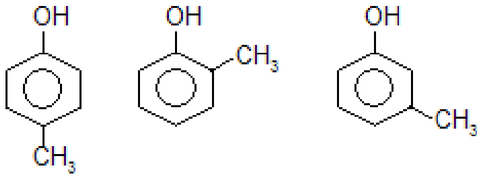
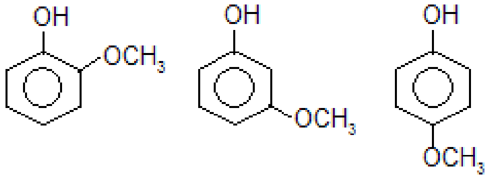
Basic power - Chemistry Fact Sheet - 2

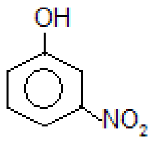
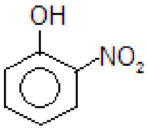
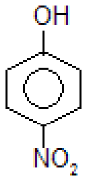
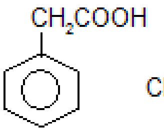
Basic power	**Order**	**Why ?**	
1.	 (I)  (II)  (III)	I > III > II	lone pair on N is not used in resonance of π - electrons in I. In II lone pair of the ring is itself used in delocalisation while that of outside ring in III.
2.	 (I)  (II)  (III)  (IV)	I > II > III > IV	$-OCH_3$ is strong electron donating group. This is due to **ortho** effect, all the aniline are less basic than *p*-substituted aniline due to **steric hindrance.**
3.	 (I)  (II)  (III)  (IV)	I > II > III > IV	I (hyper conugation and induction) II (induction) IV (ortho effect), ortho effect normally decreases basic nature.

4.		II > I > III	In II there is *sp*^{3} hybridised C, In I, *sp*^{2}. NO_2 is electron withdrawing.
5.		III > I > II	lone pair on N is used in delocalisation of π-electrons in aromatic amines while cyclohexyl is electron repelling (III); in II, lone pair on N is used by two benzene ring.
6.		I > II > III > IV	NO_2 is electron-withdrawing, thus nitro-anilines are less basic than aniline. IV is less basic than III because $-NO_2$ is closer and exerts a stronger inductive effect.
7.		III > I > II	phenyl and $-COCH_3$ are electronwithdrawing and $-C_6H_5$ < $COCH_3$
8.		I < II < III	Electron donating nature of $C_2H_5 > CH_3$ So more basic strength.

9.	 (I) (II) (III)	I < II < III	ortho effect in I.
10.	 (I) (II) (III)	I < II < III	ortho effect in (I)

Acidic Powers & Their Orders - Chemistry Fact Sheet

Acidic power	Order	Why ?	
1.	1- , 2- , 3- chlorobutanoic acid I II III	III < II < I	Farther the (-I) group (Cl), lesser the acidic strength
2.	1-, 2-, 3- methyl pentanoic acid I II III	I < II < III	Farther the (+I) group, greater the acidic power
3.	 (I) (II) (III)	II < I < III	$-CH_3$ is electron donating and $-NO_2$ is electron attracting
4.	 (I) (II) (III)	II < I < III	$-CH_3$ is electron repelling; decreases acidic strength of phenol
5.	 (I) (II) (III)	III < I < II	$-OCH_3$ group contains +M effect and decreases acidic poer.

6.	 (I)  (II)  (III)	I < II < III	— NO_2 is electron attracting; III is more resonance stabilised than I and also than II. In I, only inductive effect is operative.
7.	 (I) $CH_2{=}CHCH_2COOH$ (II) CH_3CH_2COOH (III)	I > II > III	sp^2 hybridised carbon of I, II are more electronegative hence acid strength is increased. Benzylic ($C_6H_5CH_2$) is more stabilised than allylic ($CH_2{=}{=}CHCH_2$).
8.	oxalic acid, succinic acid, malonic acid, adipic acid I II III adipic acid (all dibasic) IV	I > II > III > IV	Effect of one —COOH on the other decreases as its distance between them increases, $(COOH)_2$ is maximum acidic.
9.	*o* - *m* - *p* - nitrobenzoic acid I II III	II < III < I	—NO_2 is electron attracting (-I effect)
10.	*o* - *m* - *p* - hydroxy benzoic acid I II III	III < II < I	—OH shows electron withdrawing nature at *o* - and *m* - and electron repelling at *p* -, *o* - isomer due to intramolecular bonding in salicylate ion is stronger than *m* - isomer

11.	*o*- *m*- *p*- methoxy benzoic acid I II III	III < II < I	—do—
12.	*o*- *m*- *p*- amino benzoic acid I II III	I < III < II	—NH_2 is electron donating.

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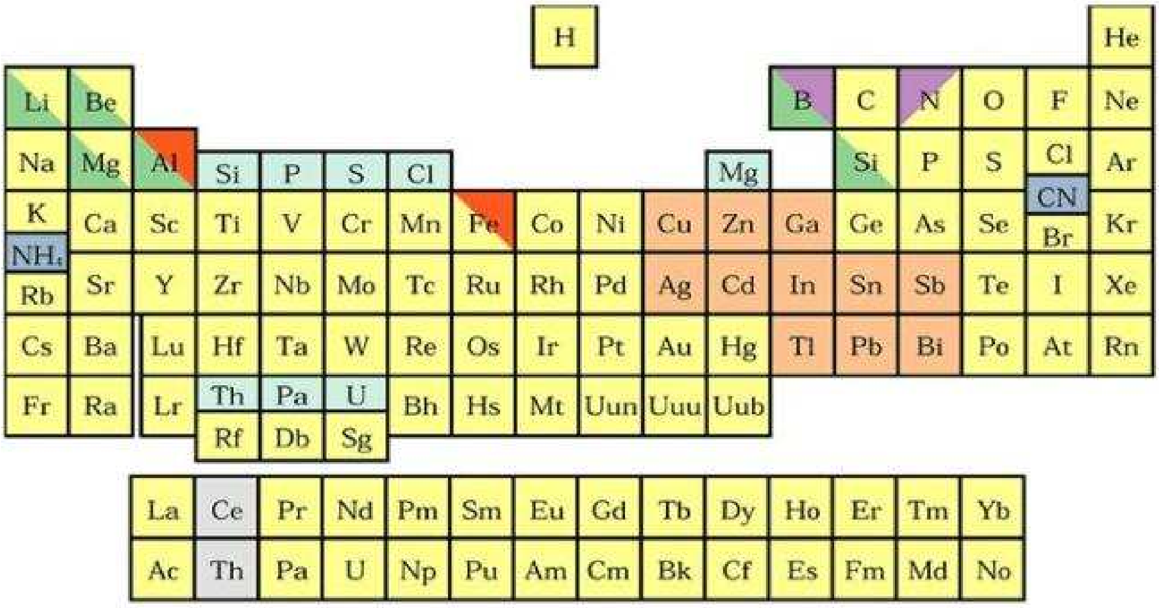
Topic wise grouping of information kind of Dictionary of Inorganic Chemistry



Given two Equations

(1) $B_2O_3 + 6NaOH \rightarrow (X) + 3H_2O$

(2) $B_2O_3 + 2NaOH \rightarrow (Y) + H_2O$

What are (X) and (Y) ?

Ans : X is Sodium Borate and Y is Sodium meta borate

$B_2O_3 + 6NaOH \rightarrow 2Na_3BO_3$ (Sodium Borate) $+ 3H_2O$

$B_2O_3 + 2NaOH \rightarrow 2NaBO_2$ (Sodium metaborate) $+ H_2O$

-

Silica is soluble in which of the following acids ?

(1) HF (2) HCl (3) HBr (4) HI

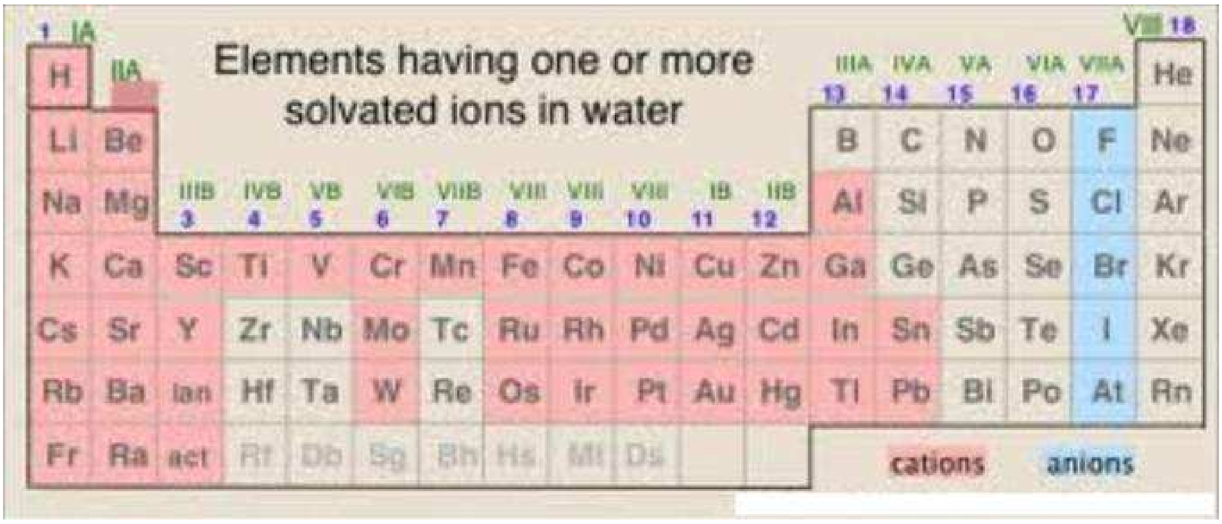
Solution : $SiO_2 + 4HF \rightarrow SiF_4 + 2H_2O$ and again $SiF_4 + 2HF \rightarrow H_2SiF_6$ (Hydroflurosilicic acid)

-

$K_3Co(NO_2)_6$ is Fischer's Salt

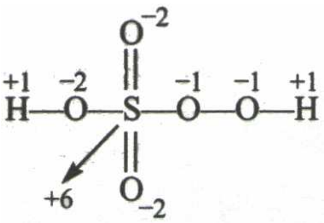
The brown ring test for NO_3^- is due to formation of the complex $[Fe(H_2O)_5NO]^{2+}$

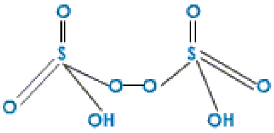
Merck's perhydrol is 30.4% H_2O_2


Elements having one or more solvated ions in water

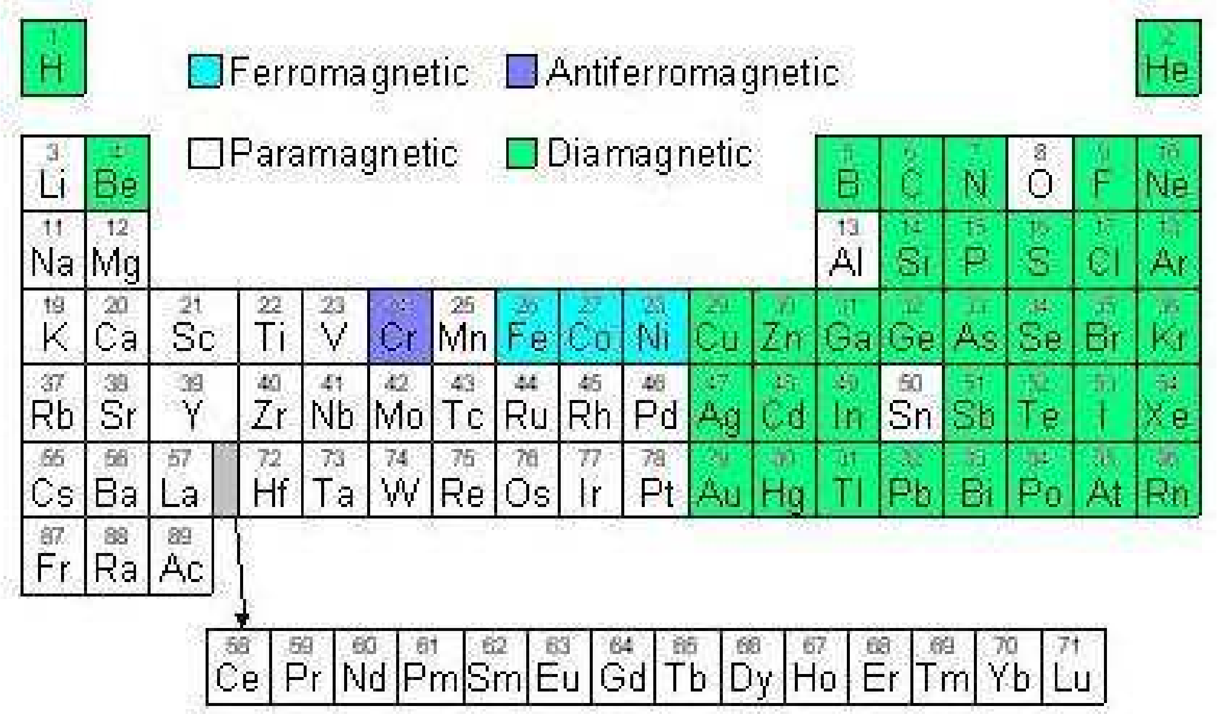
1 IA																	VIII 18
H	IIA											IIIA 13	IVA 14	VA 15	VIA 16	VIIA 17	He
Li	Be											B	C	N	O	F	Ne
Na	Mg	IIIB 3	IVB 4	VB 5	VIB 6	VIIB 7	VIII 8	VIII 9	VIII 10	IB 11	IIB 12	Al	Si	P	S	Cl	Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Cs	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
Rb	Ba	lan	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ra	act	Rf	Db	Sg	Bh	Hs	Mt	Ds					cations		anions	

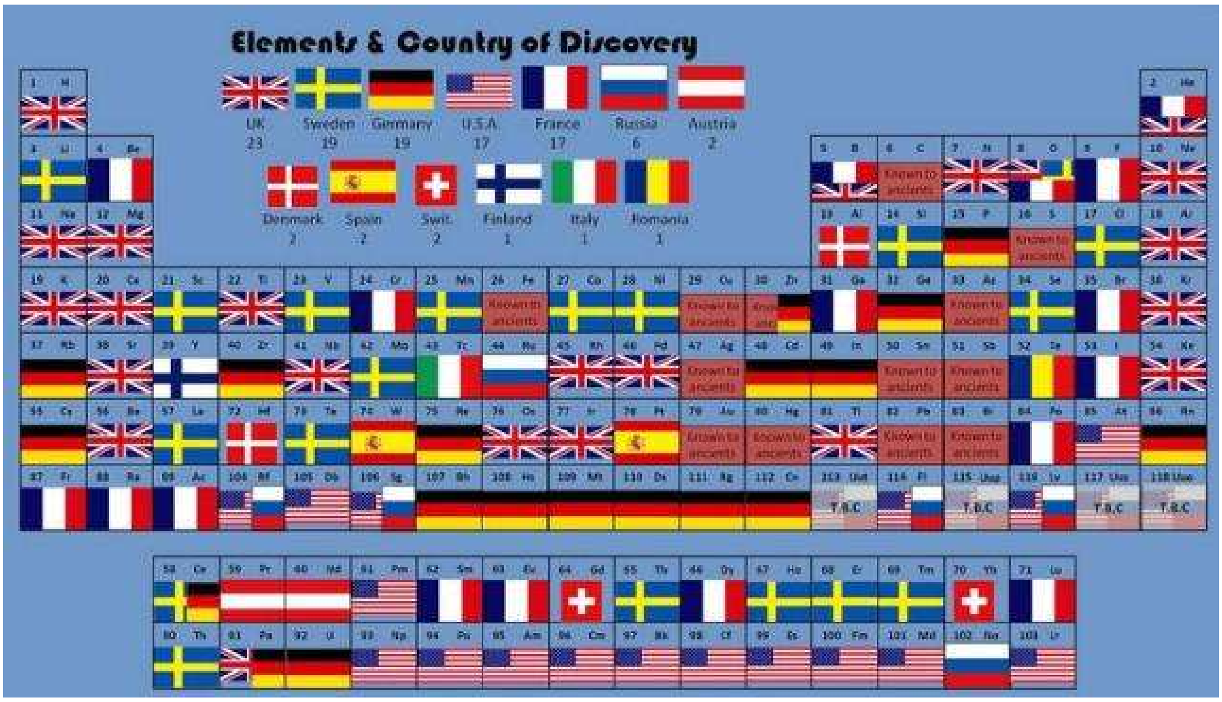
H_2SO_5 , $H_2S_2O_8$ have peroxide linkages





$H_2S_2O_8$(+6)
Peroxydisulphuric acid
(or Marshall's acid)





Bromine reacting with NaOH in cold and hot give different mix of products. Specify

Br_2 + 2NaOH (cold) -> NaBr + NaOBr + H_2O

$3Br_2$ + 6 NaOH (Hot) -> 5NaBr + $NaBrO_3$ + $3H_2O$

Malachite decomposes to give A + CO_2 + H_2O and compound A on reduction with Carbon gives CO + B

Identify A and B

Solution : $CuCO_3.Cu(OH)_2$ -> 2CuO (A) + CO_2 + H_2O

CuO + C -> Cu (B) + CO

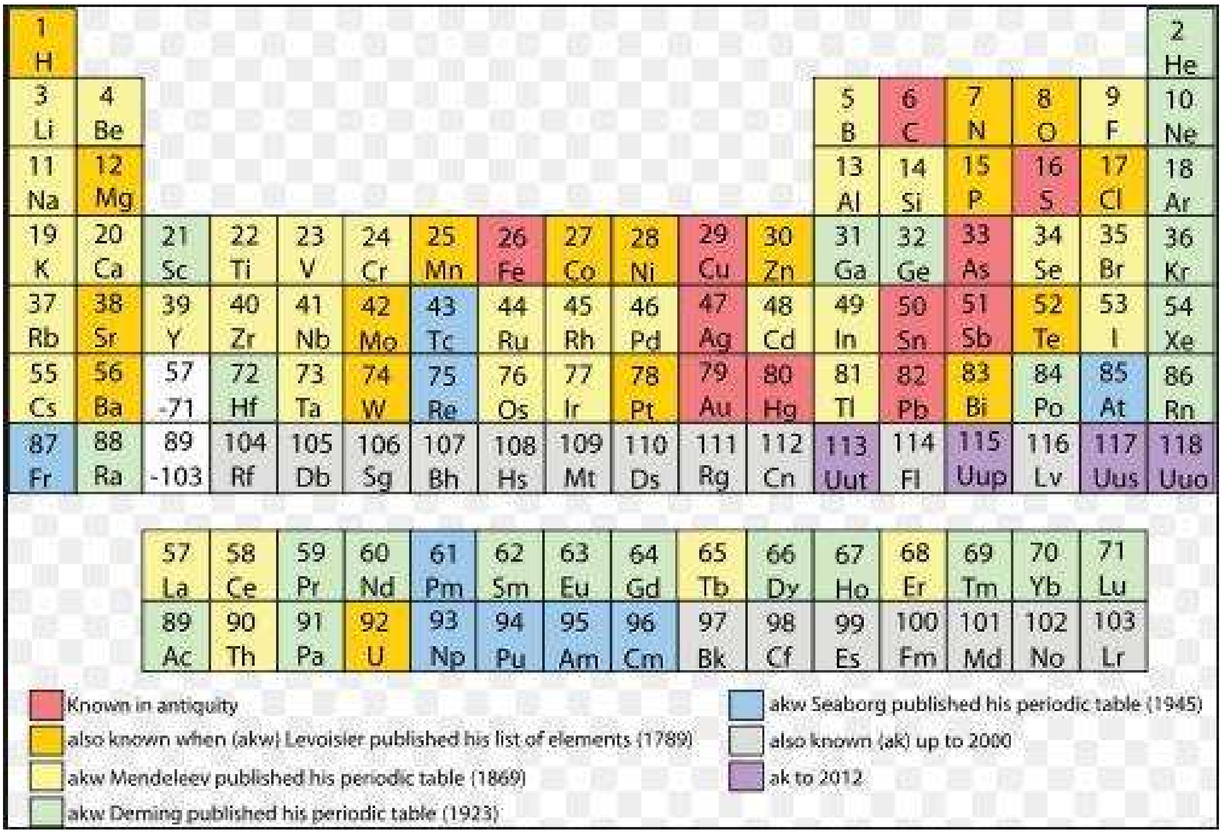
-

Some examples of Complex Anions

K_2PtCl_6 < == > $2K^+$ + $[PtCl_6]^{2-}$

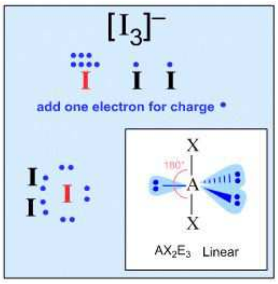
Na_3AlF_6 < == > $3Na^+$ + AlF_6^{-3}

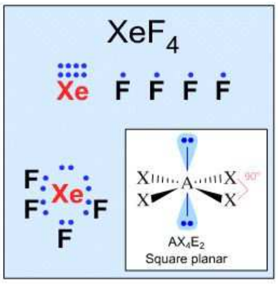
K_2SiF_6 < == > $2K^+$ + $SiF6^{-2}$



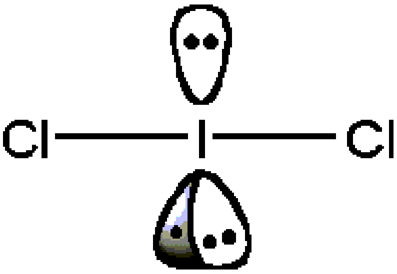
1 H																	2 He
3 Li	4 Be											5 B	6 C	7 N	8 O	9 F	10 Ne
11 Na	12 Mg											13 Al	14 Si	15 P	16 S	17 Cl	18 Ar
19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr
37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe
55 Cs	56 Ba	57 -71	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn
87 Fr	88 Ra	89 -103	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Uut	114 Fl	115 Uup	116 Lv	117 Uus	118 Uuo
		57 La	58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu	
		89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr	

🟥 Known in antiquity
🟧 also known when (akw) Levoisier published his list of elements (1789)
🟨 akw Mendeleev published his periodic table (1869)
🟩 akw Deming published his periodic table (1923)
🟦 akw Seaborg published his periodic table (1945)
⬜ also known (ak) up to 2000
🟪 ak to 2012

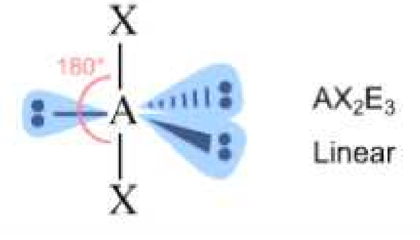




and ICl_2^-



Sp^{3}d hybridization



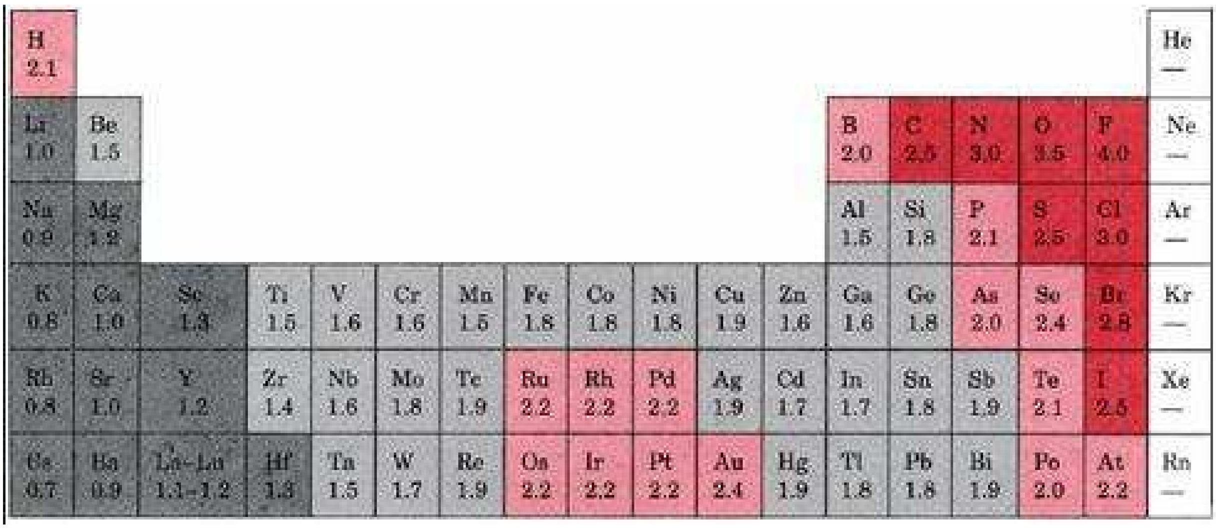
for I_3^- and ICl_2^-

The ionization potential of Pb is higher than Sn due to poor shielding by 14f electrons present in Pb

Selenous acid H_2SeO_3 (Oxidation Number of Se is +4)

Telluric acid H_6TeO_6 (Oxidation Number of Te is +6)

Electronegativity Values (in Pauling Scale)



H 2.1																	He —
Li 1.0	Be 1.5											B 2.0	C 2.5	N 3.0	O 3.5	F 4.0	Ne —
Na 0.9	Mg 1.2											Al 1.5	Si 1.8	P 2.1	S 2.5	Cl 3.0	Ar —
K 0.8	Ca 1.0	Sc 1.3	Ti 1.5	V 1.6	Cr 1.6	Mn 1.5	Fe 1.8	Co 1.8	Ni 1.8	Cu 1.9	Zn 1.6	Ga 1.6	Ge 1.8	As 2.0	Se 2.4	Br 2.8	Kr —
Rb 0.8	Sr 1.0	Y 1.2	Zr 1.4	Nb 1.6	Mo 1.8	Tc 1.9	Ru 2.2	Rh 2.2	Pd 2.2	Ag 1.9	Cd 1.7	In 1.7	Sn 1.8	Sb 1.9	Te 2.1	I 2.5	Xe —
Cs 0.7	Ba 0.9	La–Lu 1.1–1.2	Hf 1.3	Ta 1.5	W 1.7	Re 1.9	Os 2.2	Ir 2.2	Pt 2.2	Au 2.4	Hg 1.9	Tl 1.8	Pb 1.8	Bi 1.9	Po 2.0	At 2.2	Rn —

Microcosmic Salt is $Na(NH_4)HPO_4$ Sodium ammonium hydrogen phosphate

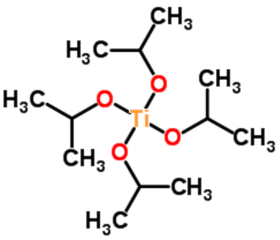
Which of the following electrolyte will be most effective in coagulation of gold sol ?

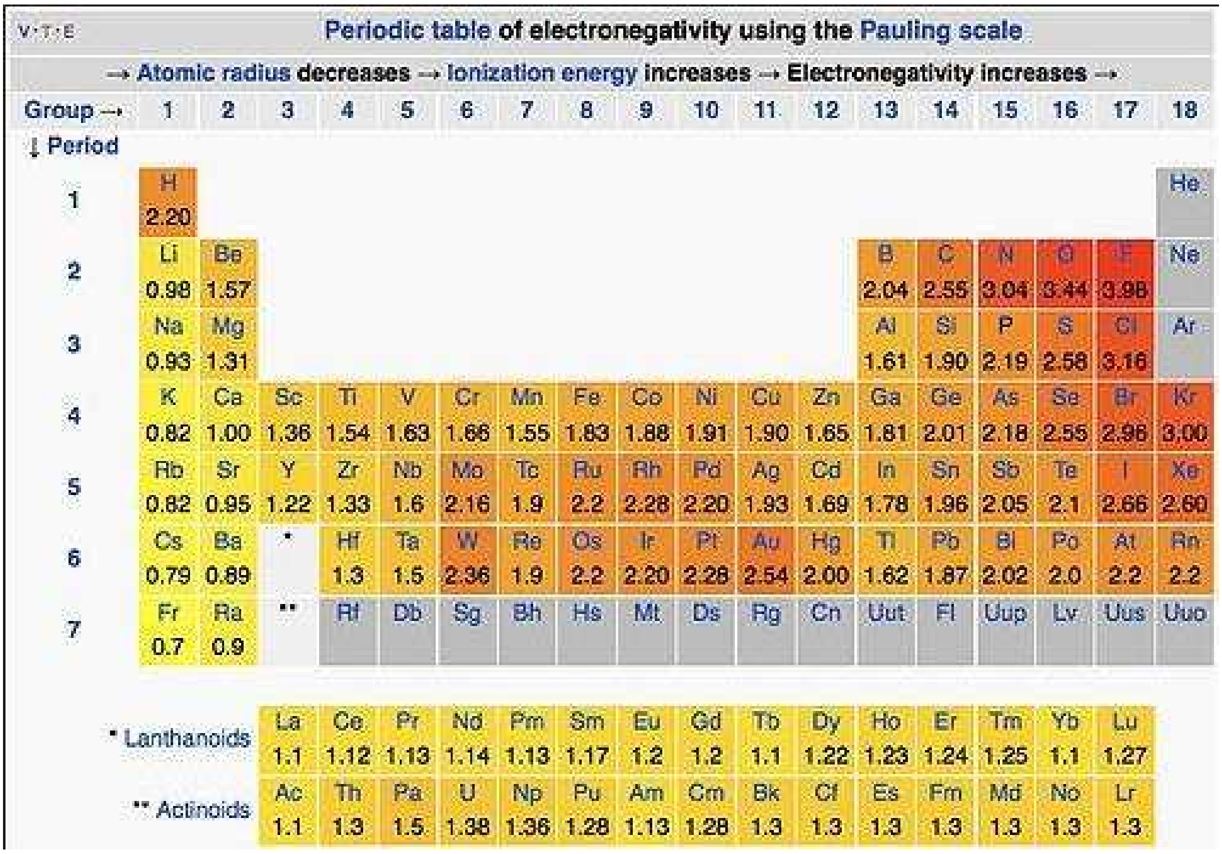
(1) $NaNO_3$ (2) $K_4Fe(CN)_6$ (3) Na_3PO_4 (4) $MgCl_2$

Gold sol is a negatively charged coagulation which is carried out by positively charged ion furnished by electrolyte

Coagulating power depends on the valency of cation of electrolyte. (Schulze Hardy Law) Mg^{2+} ion has highest valency so $MgCl_2$ is most effective.

Organometallic compounds are Metal atoms directly linked with Carbon. So $Ti(OC_3H_7)_4$ is not an organometallic compound.





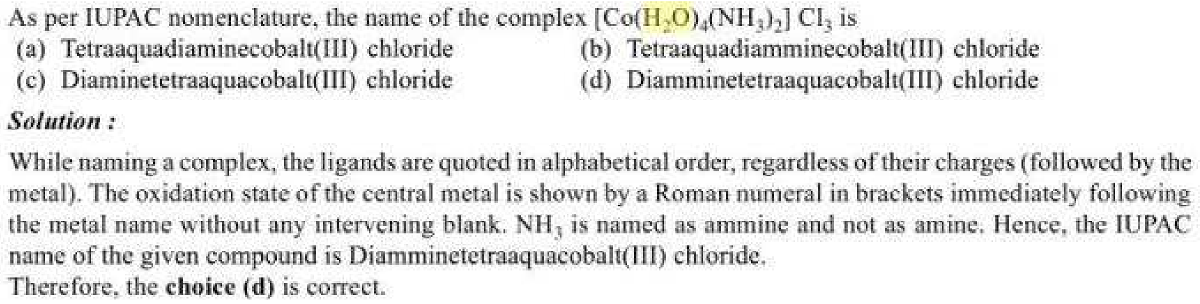
V·T·E

Periodic table of electronegativity using the Pauling scale

→ Atomic radius decreases → Ionization energy increases → Electronegativity increases →

Group →	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
↓ Period																		
1	H 2.20																	He
2	Li 0.98	Be 1.57											B 2.04	C 2.55	N 3.04	O 3.44	F 3.98	Ne
3	Na 0.93	Mg 1.31											Al 1.61	Si 1.90	P 2.19	S 2.58	Cl 3.16	Ar
4	K 0.82	Ca 1.00	Sc 1.36	Ti 1.54	V 1.63	Cr 1.66	Mn 1.55	Fe 1.83	Co 1.88	Ni 1.91	Cu 1.90	Zn 1.65	Ga 1.81	Ge 2.01	As 2.18	Se 2.55	Br 2.96	Kr 3.00
5	Rb 0.82	Sr 0.95	Y 1.22	Zr 1.33	Nb 1.6	Mo 2.16	Tc 1.9	Ru 2.2	Rh 2.28	Pd 2.20	Ag 1.93	Cd 1.69	In 1.78	Sn 1.96	Sb 2.05	Te 2.1	I 2.66	Xe 2.60
6	Cs 0.79	Ba 0.89	*	Hf 1.3	Ta 1.5	W 2.36	Re 1.9	Os 2.2	Ir 2.20	Pt 2.28	Au 2.54	Hg 2.00	Tl 1.62	Pb 1.87	Bi 2.02	Po 2.0	At 2.2	Rn 2.2
7	Fr 0.7	Ra 0.9	**	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Fl	Uup	Lv	Uus	Uuo
* Lanthanoids			La 1.1	Ce 1.12	Pr 1.13	Nd 1.14	Pm 1.13	Sm 1.17	Eu 1.2	Gd 1.2	Tb 1.1	Dy 1.22	Ho 1.23	Er 1.24	Tm 1.25	Yb 1.1	Lu 1.27	
** Actinoids			Ac 1.1	Th 1.3	Pa 1.5	U 1.38	Np 1.36	Pu 1.28	Am 1.13	Cm 1.28	Bk 1.3	Cf 1.3	Es 1.3	Fm 1.3	Md 1.3	No 1.3	Lr 1.3	

IUPAC name of Diamminetetraaquacobalt III chloride



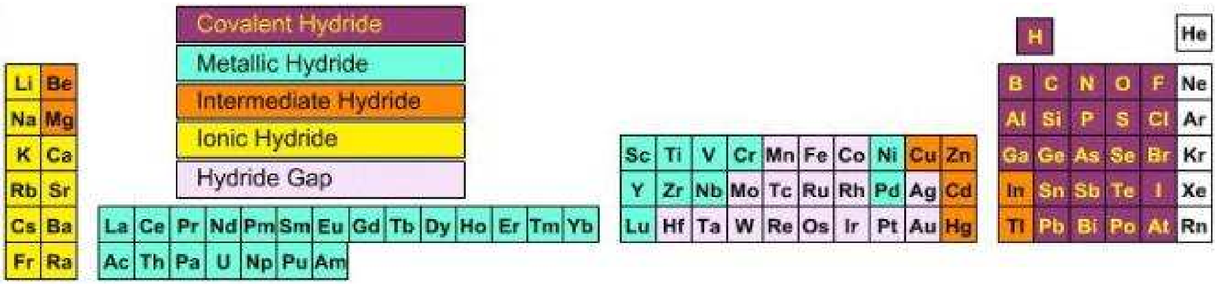
As per IUPAC nomenclature, the name of the complex $[Co(H_2O)_4(NH_3)_2]\,Cl_3$ is

(a) Tetraaquadiaminecobalt(III) chloride (b) Tetraaquadiamminecobalt(III) chloride

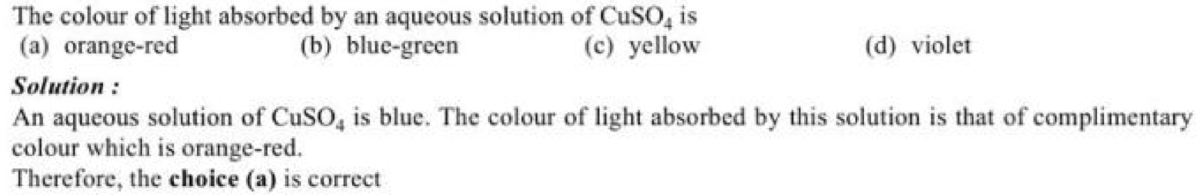
(c) Diaminetetraaquacobalt(III) chloride (d) Diamminetetraaquacobalt(III) chloride

Solution :

While naming a complex, the ligands are quoted in alphabetical order, regardless of their charges (followed by the metal). The oxidation state of the central metal is shown by a Roman numeral in brackets immediately following the metal name without any intervening blank. NH_3 is named as ammine and not as amine. Hence, the IUPAC name of the given compound is Diamminetetraaquacobalt(III) chloride.
Therefore, the **choice (d)** is correct.



Color of light absorbed by aqueous solution of CuSO4

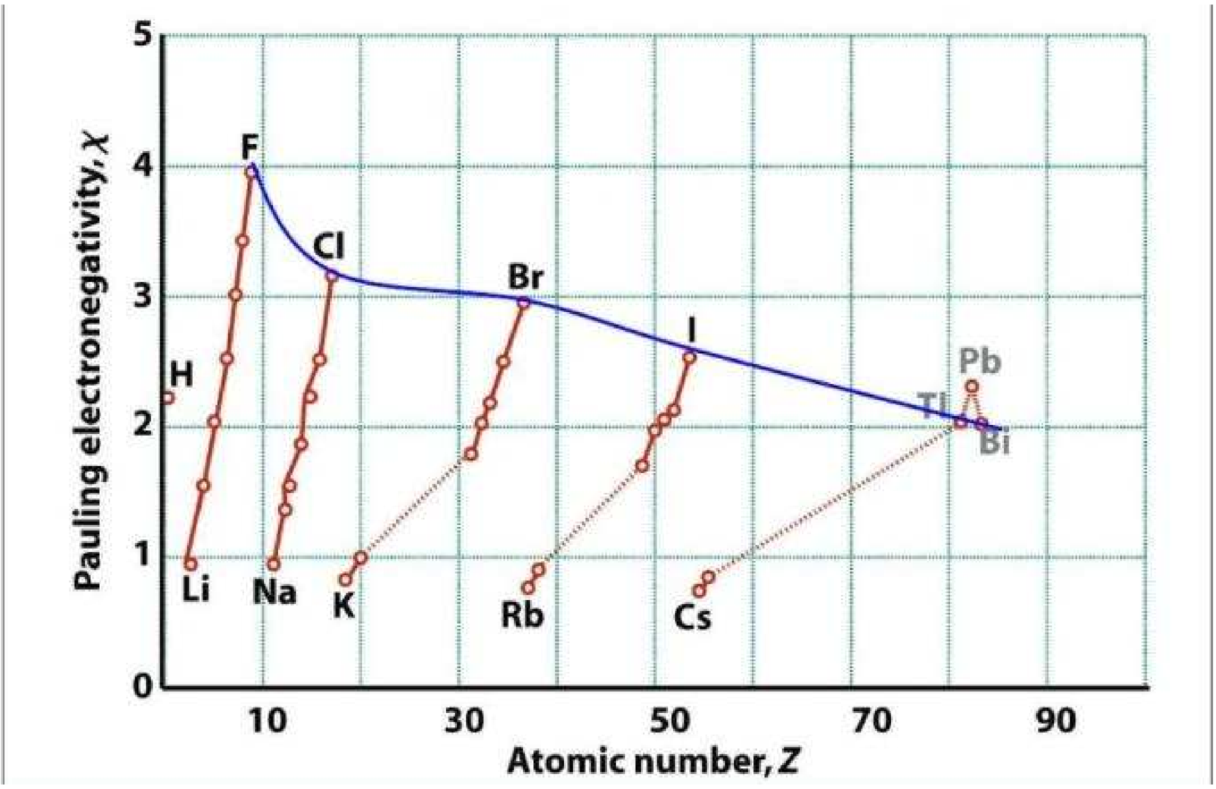


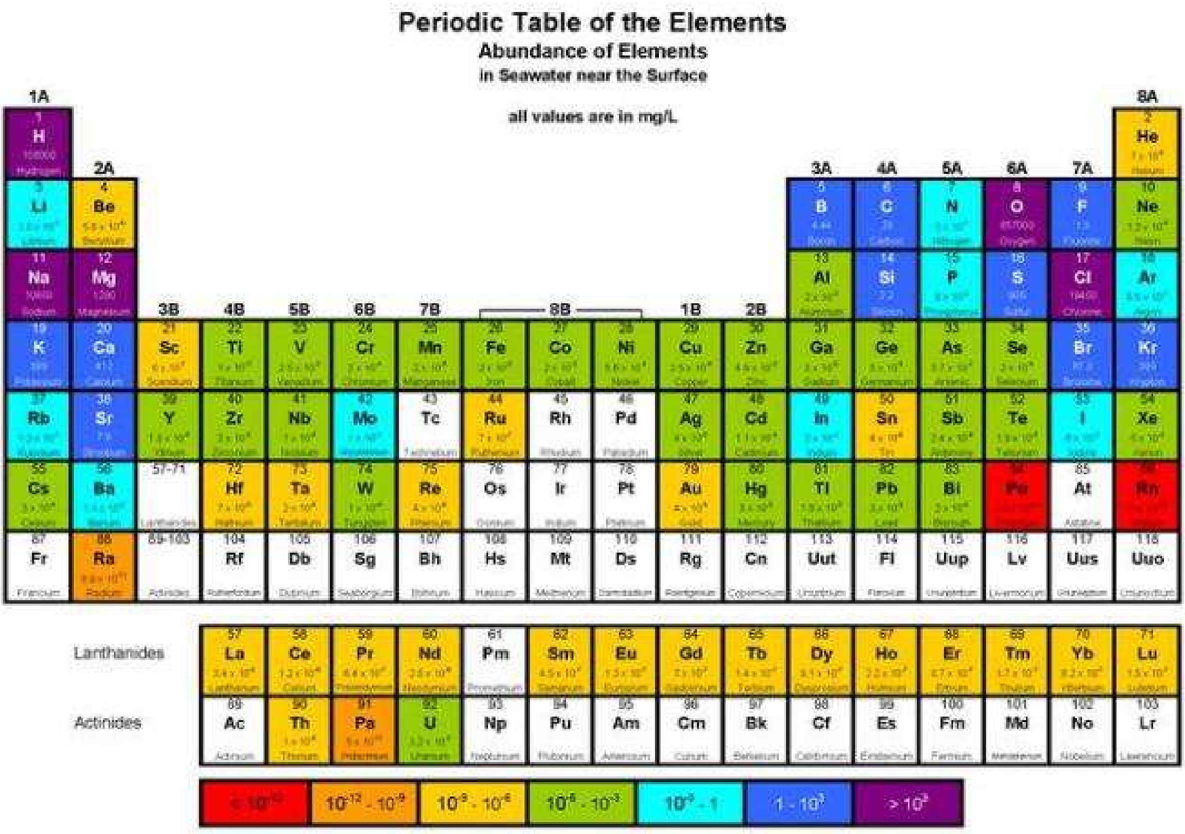
The colour of light absorbed by an aqueous solution of $CuSO_4$ is

(a) orange-red (b) blue-green (c) yellow (d) violet

Solution :

An aqueous solution of $CuSO_4$ is blue. The colour of light absorbed by this solution is that of complimentary colour which is orange-red.
Therefore, the **choice (a)** is correct

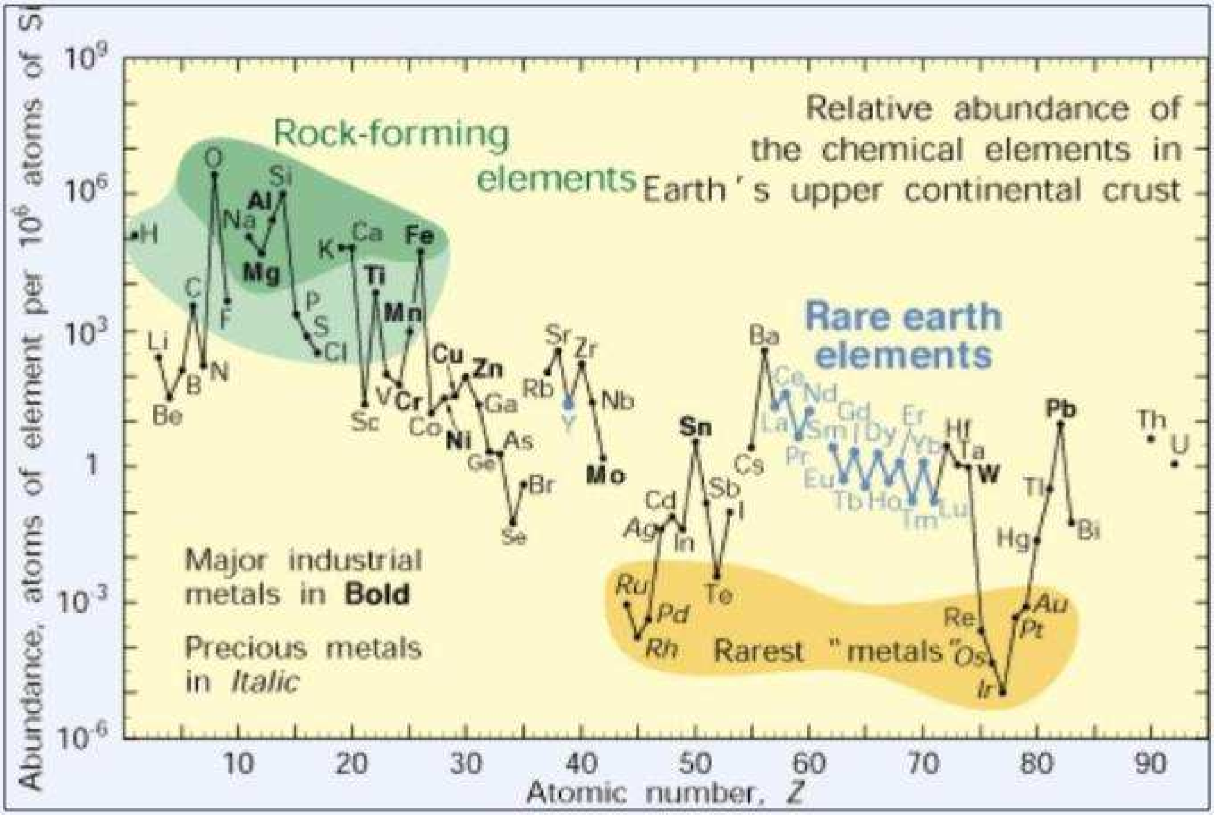


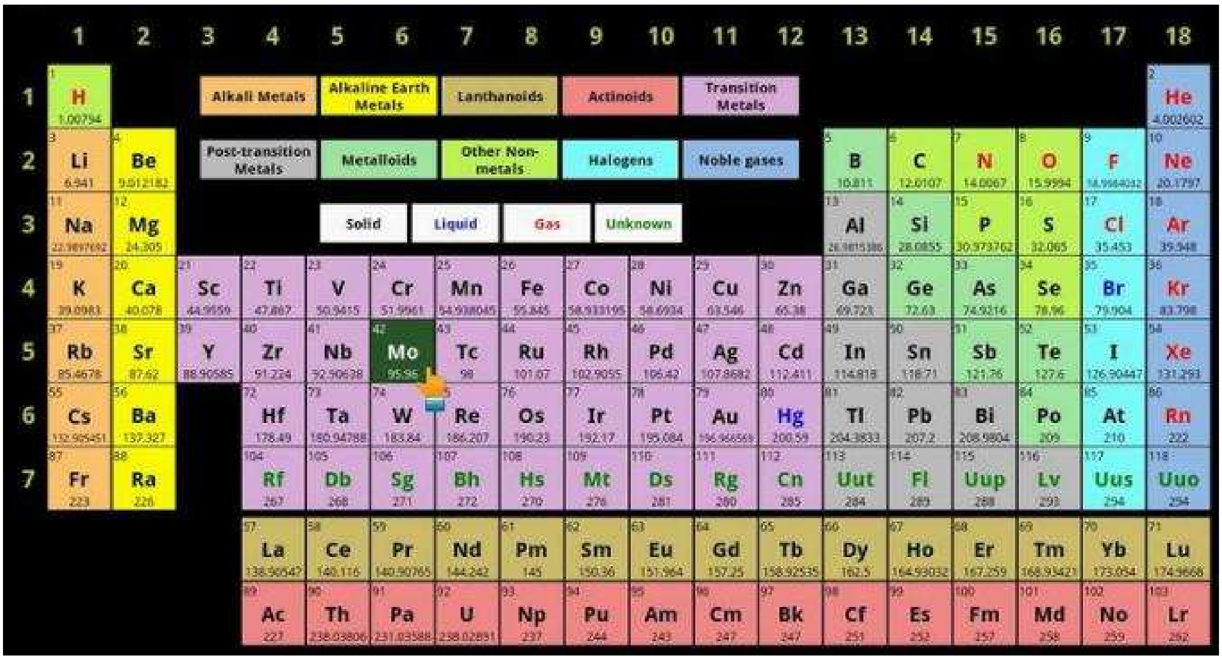


Periodic Table of the Elements
Abundance of Elements
in Seawater near the Surface

all values are in mg/L

🟥	🟧	🟨	🟩	🟦	🟦	🟪
$< 10^{-12}$	$10^{-12} - 10^{-9}$	$10^{-9} - 10^{-6}$	$10^{-6} - 10^{-3}$	$10^{-3} - 1$	$1 - 10^{3}$	$> 10^{3}$

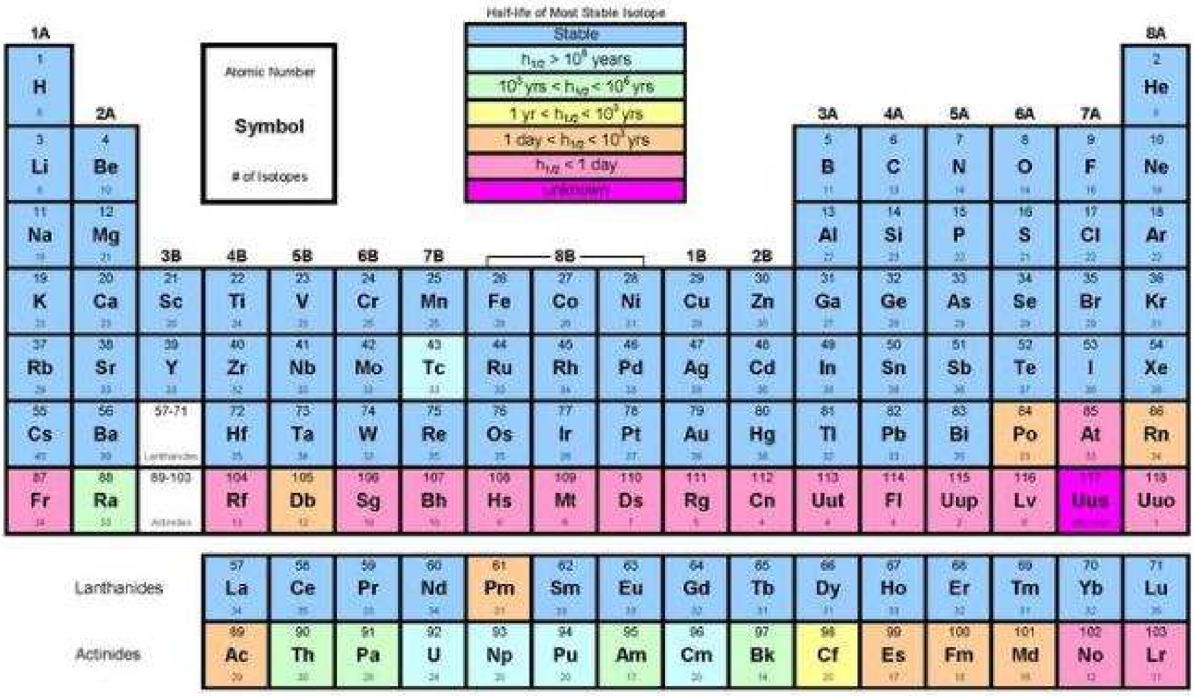




	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1	1 H 1.00794																	2 He 4.002602
2	3 Li 6.941	4 Be 9.012182											5 B 10.811	6 C 12.0107	7 N 14.0067	8 O 15.9994	9 F 18.9984032	10 Ne 20.1797
3	11 Na 22.9897692	12 Mg 24.305											13 Al 26.9815386	14 Si 28.0855	15 P 30.973762	16 S 32.065	17 Cl 35.453	18 Ar 39.948
4	19 K 39.0983	20 Ca 40.078	21 Sc 44.9559	22 Ti 47.867	23 V 50.9415	24 Cr 51.9961	25 Mn 54.938045	26 Fe 55.845	27 Co 58.933195	28 Ni 58.6934	29 Cu 63.546	30 Zn 65.38	31 Ga 69.723	32 Ge 72.63	33 As 74.9216	34 Se 78.96	35 Br 79.904	36 Kr 83.798
5	37 Rb 85.4678	38 Sr 87.62	39 Y 88.90585	40 Zr 91.224	41 Nb 92.90638	42 Mo 95.96	43 Tc 98	44 Ru 101.07	45 Rh 102.9055	46 Pd 106.42	47 Ag 107.8682	48 Cd 112.411	49 In 114.818	50 Sn 118.71	51 Sb 121.76	52 Te 127.6	53 I 126.90447	54 Xe 131.293
6	55 Cs 132.905451	56 Ba 137.327		72 Hf 178.49	73 Ta 180.94788	74 W 183.84	75 Re 186.207	76 Os 190.23	77 Ir 192.17	78 Pt 195.084	79 Au 196.966569	80 Hg 200.59	81 Tl 204.3833	82 Pb 207.2	83 Bi 208.9804	84 Po 209	85 At 210	86 Rn 222
7	87 Fr 223	88 Ra 226		104 Rf 267	105 Db 268	106 Sg 271	107 Bh 272	108 Hs 270	109 Mt 276	110 Ds 281	111 Rg 280	112 Cn 285	113 Uut 284	114 Fl 289	115 Uup 288	116 Lv 293	117 Uus 294	118 Uuo 294
				57 La 138.90547	58 Ce 140.116	59 Pr 140.90765	60 Nd 144.242	61 Pm 145	62 Sm 150.36	63 Eu 151.964	64 Gd 157.25	65 Tb 158.92535	66 Dy 162.5	67 Ho 164.93032	68 Er 167.259	69 Tm 168.93421	70 Yb 173.054	71 Lu 174.9668
				89 Ac 227	90 Th 232.03806	91 Pa 231.03588	92 U 238.02891	93 Np 237	94 Pu 244	95 Am 243	96 Cm 247	97 Bk 247	98 Cf 251	99 Es 252	100 Fm 257	101 Md 258	102 No 259	103 Lr 262

Alkali Metals | Alkaline Earth Metals | Lanthanoids | Actinoids | Transition Metals
Post-transition Metals | Metalloids | Other Non-metals | Halogens | Noble gases
Solid | Liquid | Gas | Unknown

Periodic Table of the Radioactive Elements



1A	2A	3B	4B	5B	6B	7B	8B	8B	8B	1B	2B	3A	4A	5A	6A	7A	8A
1 H																	2 He
3 Li	4 Be											5 B	6 C	7 N	8 O	9 F	10 Ne
11 Na	12 Mg											13 Al	14 Si	15 P	16 S	17 Cl	18 Ar
19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr
37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe
55 Cs	56 Ba	57-71 Lanthanides	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn
87 Fr	88 Ra	89-103 Actinides	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Uut	114 Fl	115 Uup	116 Lv	117 Uus	118 Uuo

Lanthanides	57 La	58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu
Actinides	89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr

Temperature dependent Magnetic behaviour

$NiCl_2\{P(C_2H_5)_2(C_6H_5)\}_2$ exhibits temperature dependent magnetic behaviour (paramagnetic/diamagnetic). The coordination geometries of Ni^{2+} in the paramagnetic and diamagnetic states are respectively

(a) tetrahedral and tetrahedral
(b) square planar and square planar
(c) tetrahedral and square planar
(d) square planar and tetrahedral

Solution :

In the paramagnetic state, the complex will have unpaired 3d electrons in Ni^{2+} ion, whereas no such electrons (i.e. all the 3d elctrons will be paired) are present in the diamagnetic state.

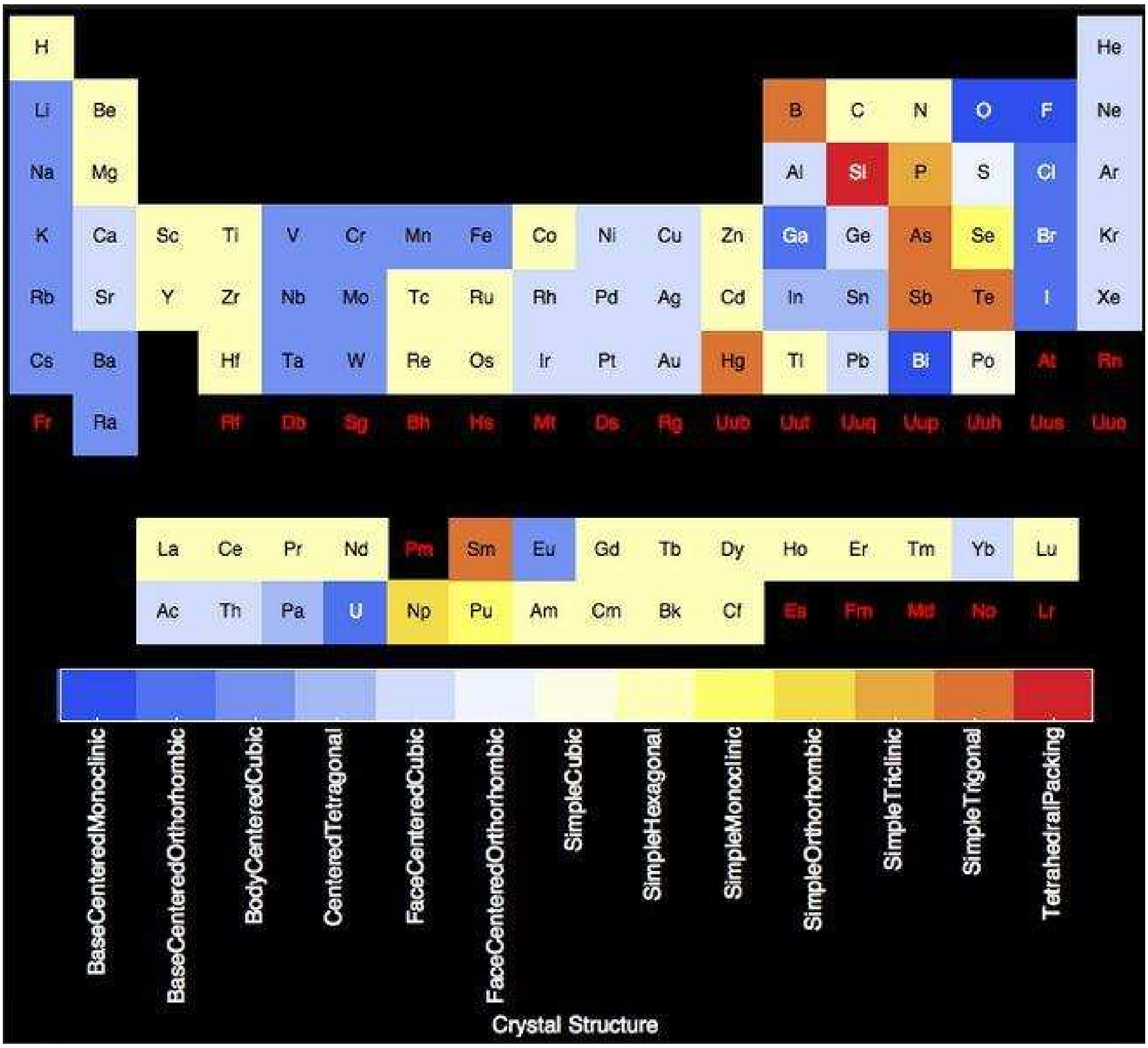
The electronic configuration of $_{28}Ni^{2+}$ is

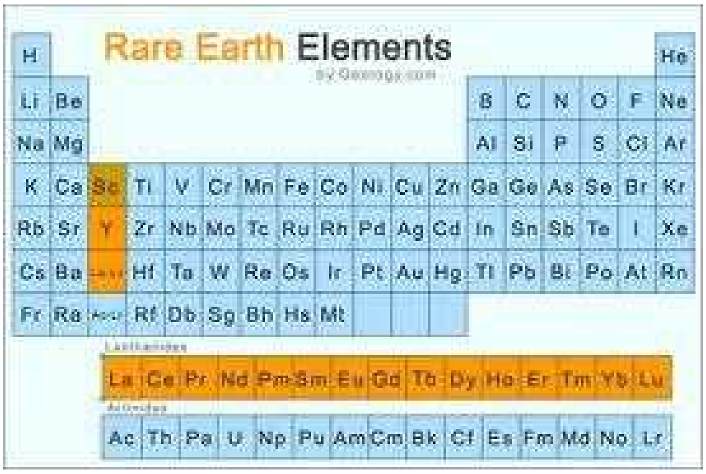
3d					4s	4p		
↑↓	↑↓	↑↓	↑	↑				

If the two unpaired electrons remain present in the ion, it will undergo sp^3 hybridization to accommodate ligands and the geometry of the complex ion will be tetrahedral.

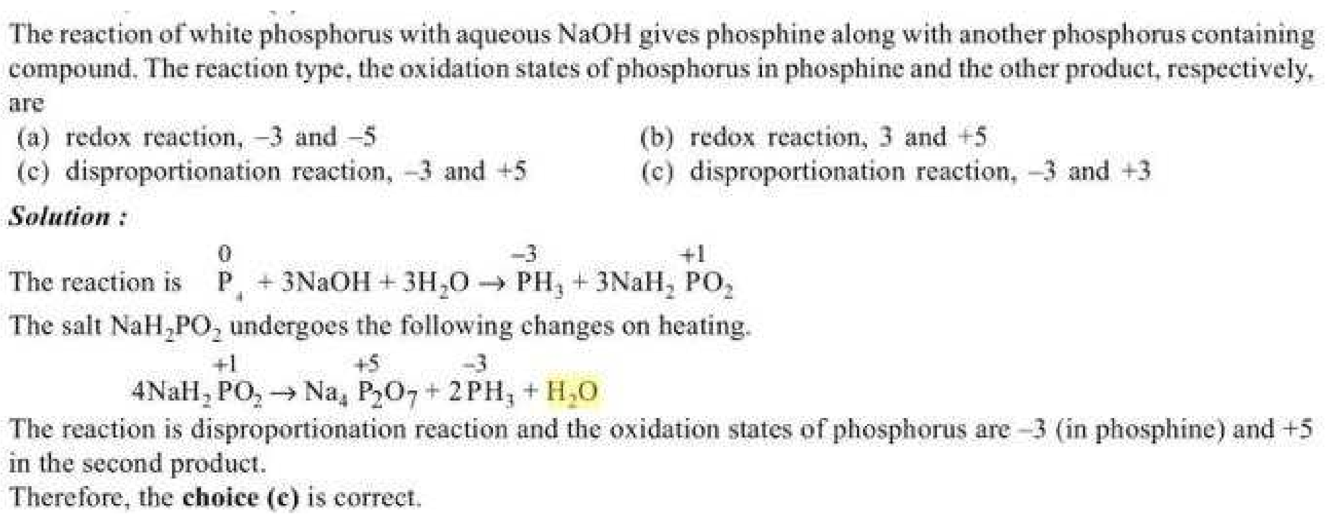
If the two unpaired electrons are coupled, the ion will be diamagnetic and it can undergo dsp^2 hybridization to accommodate ligands leading to the square planar geometry to the complex ion.

Therefore, the **choice (c)** is correct.


H He
Li Be B C N O F Ne
Na Mg Al Si P S Cl Ar
K Ca Sc Ti V Cr Mn Fe Co Ni Cu Zn Ga Ge As Se Br Kr
Rb Sr Y Zr Nb Mo Tc Ru Rh Pd Ag Cd In Sn Sb Te I Xe
Cs Ba Hf Ta W Re Os Ir Pt Au Hg Tl Pb Bi Po At Rn
Fr Ra Rf Db Sg Bh Hs Mt Ds Rg Uub Uut Uuq Uup Uuh Uus Uuo
La Ce Pr Nd Pm Sm Eu Gd Tb Dy Ho Er Tm Yb Lu
Ac Th Pa U Np Pu Am Cm Bk Cf Es Fm Md No Lr
BaseCenteredMonoclinic
BaseCenteredOrthorhombic
BodyCenteredCubic
CenteredTetragonal
FaceCenteredCubic
FaceCenteredOrthorhombic
SimpleCubic
SimpleHexagonal
SimpleMonoclinic
SimpleOrthorhombic
SimpleTriclinic
SimpleTrigonal
TetrahedralPacking
Crystal Structure


Rare Earth Elements
by Geology.com
H He
Li Be B C N O F Ne
Na Mg Al Si P S Cl Ar
K Ca Sc Ti V Cr Mn Fe Co Ni Cu Zn Ga Ge As Se Br Kr
Rb Sr Y Zr Nb Mo Tc Ru Rh Pd Ag Cd In Sn Sb Te I Xe
Cs Ba Hf Ta W Re Os Ir Pt Au Hg Tl Pb Bi Po At Rn
Fr Ra Rf Db Sg Bh Hs Mt
Lanthanides
La Ce Pr Nd Pm Sm Eu Gd Tb Dy Ho Er Tm Yb Lu
Actinides
Ac Th Pa U Np Pu Am Cm Bk Cf Es Fm Md No Lr

Reaction of white phosphorus with aqueous NaOH



The reaction of white phosphorus with aqueous NaOH gives phosphine along with another phosphorus containing compound. The reaction type, the oxidation states of phosphorus in phosphine and the other product, respectively, are

(a) redox reaction, –3 and –5
(b) redox reaction, 3 and +5
(c) disproportionation reaction, –3 and +5
(c) disproportionation reaction, –3 and +3

Solution :

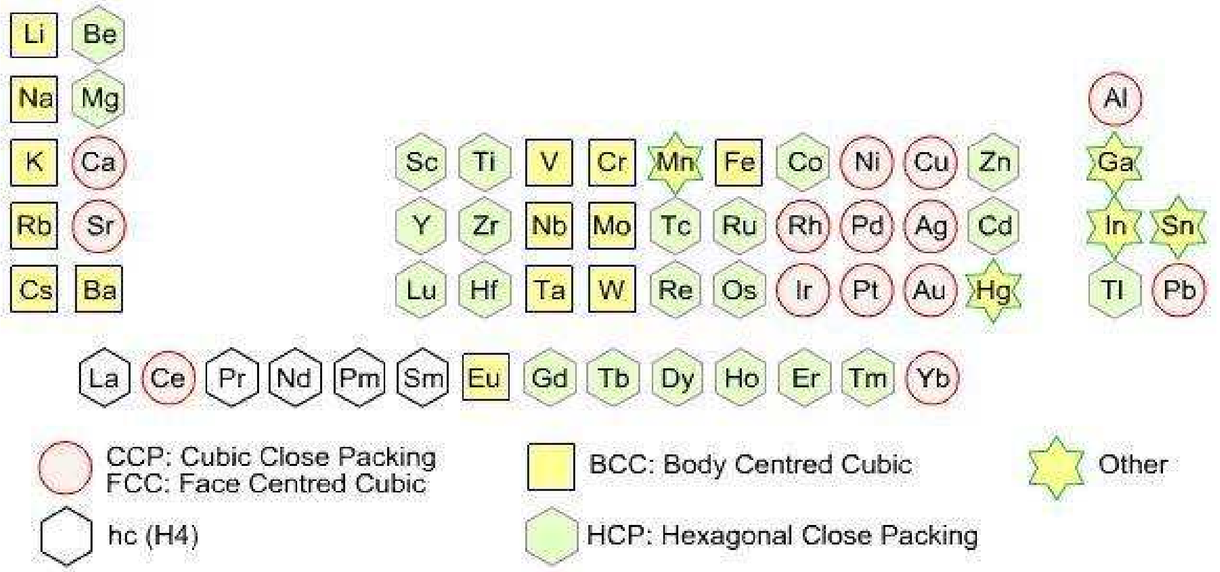
The reaction is $\overset{0}{P}_4 + 3NaOH + 3H_2O \rightarrow \overset{-3}{P}H_3 + 3NaH_2\overset{+1}{P}O_2$

The salt NaH_2PO_2 undergoes the following changes on heating.

$$4NaH_2\overset{+1}{P}O_2 \rightarrow Na_4\overset{+5}{P}_2O_7 + 2\overset{-3}{P}H_3 + H_2O$$

The reaction is disproportionation reaction and the oxidation states of phosphorus are –3 (in phosphine) and +5 in the second product.

Therefore, the **choice (c)** is correct.

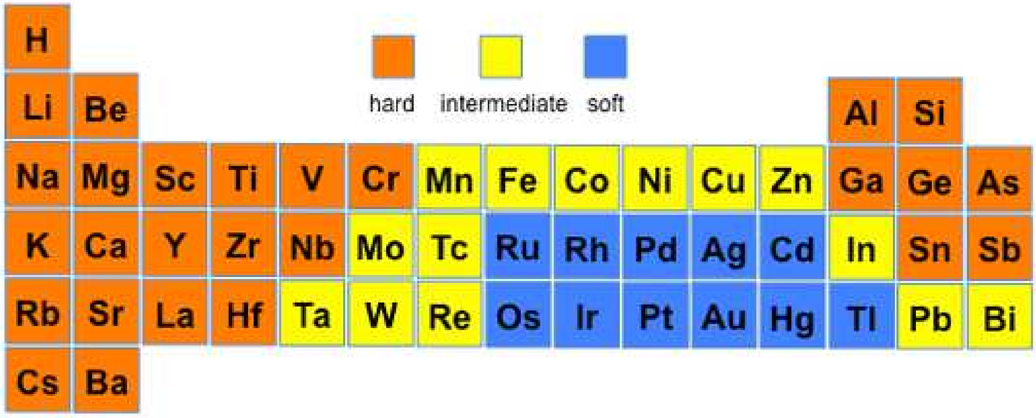


Generally the stabilities of complexes decreases with increasing atomic number for the electropositive metals e.g. group IIA or hard Lewis acids, and increase with increasing atomic number for the noble metals (soft acids), following the general trend of ionization energy.

(a) For electropositive metals (hard acids) the order of stabilities of the Halide complexes F > Cl > Br > I but for highly polarizing soft acid metal ions such as Hg^{+2} , we see the reverse order.

(b) The most electropositive metals (hardest acids) show a greater tendency of forming complexes, with hard ligands such as F^- or Oxygen containing ligands.

Class A



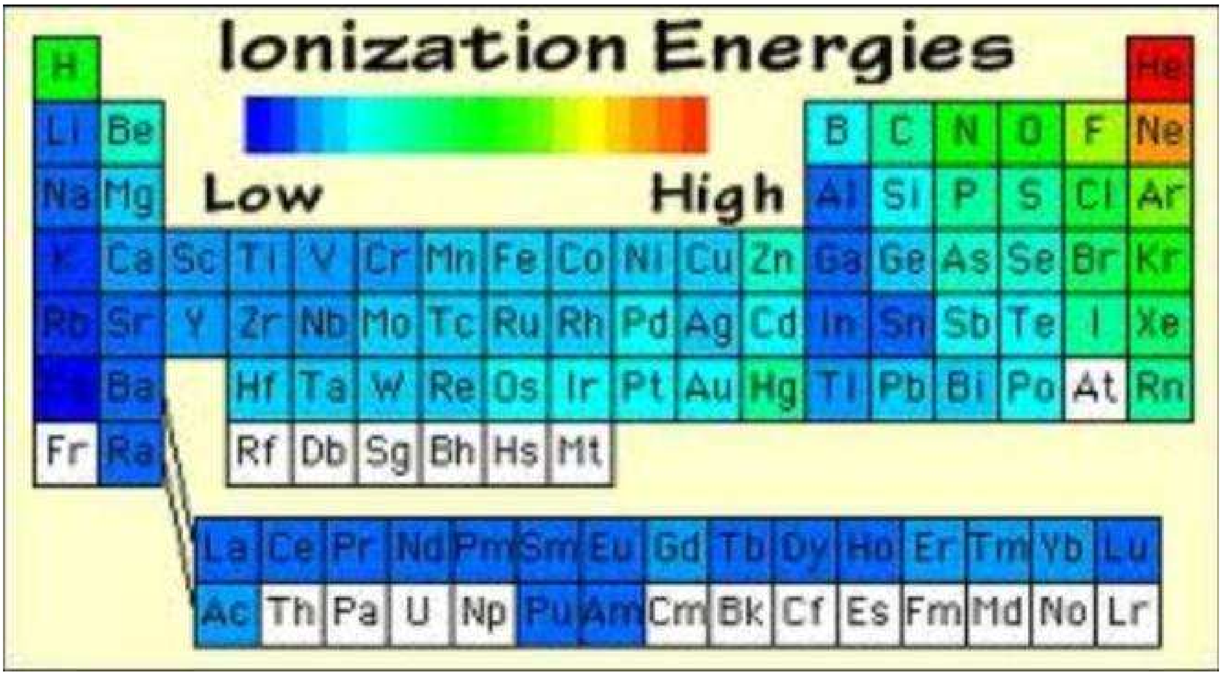
Class A **Class C(Yellow)** **Class B** **Class C (Yellow)**

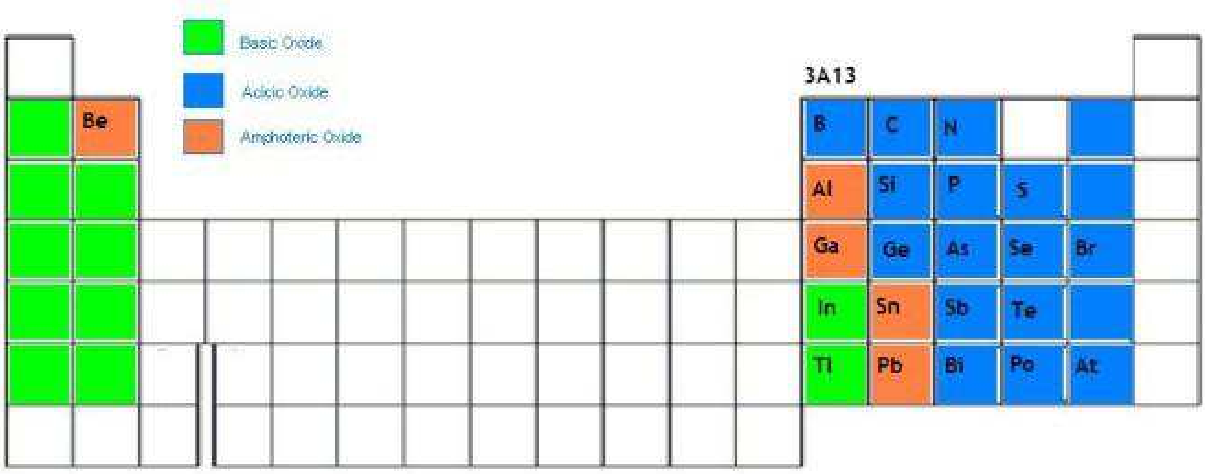
Metals are classified according to their acceptor properties.

Class A - Hard Acids. Show affinities to ligands whose basicity is proportional to protons.

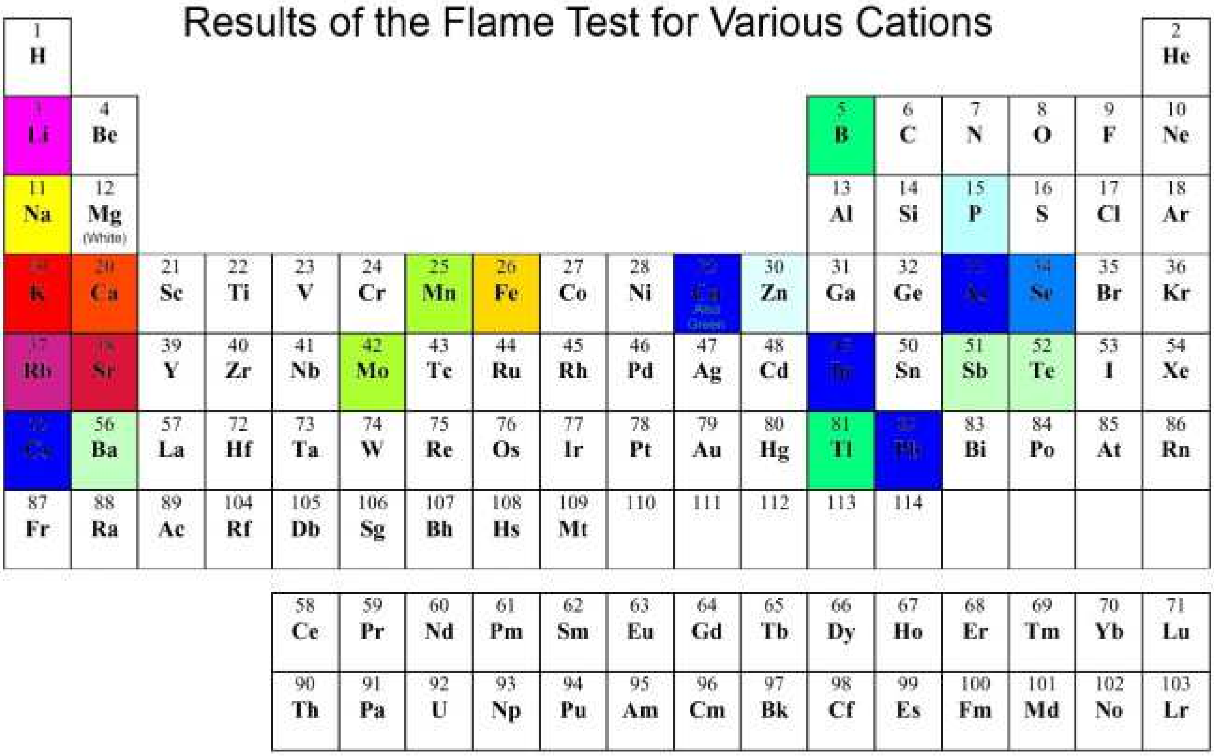
Class B - Soft acids form stable olefin complexes.

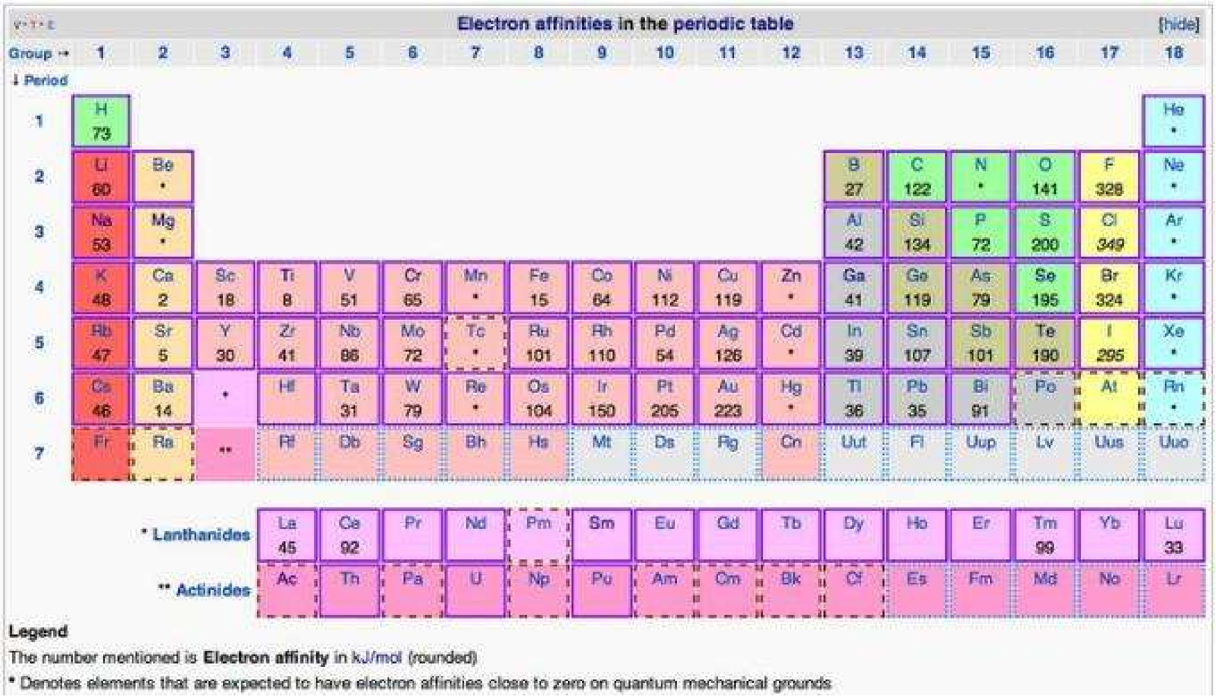
Class C - Borderline metals





Results of the Flame Test for Various Cations

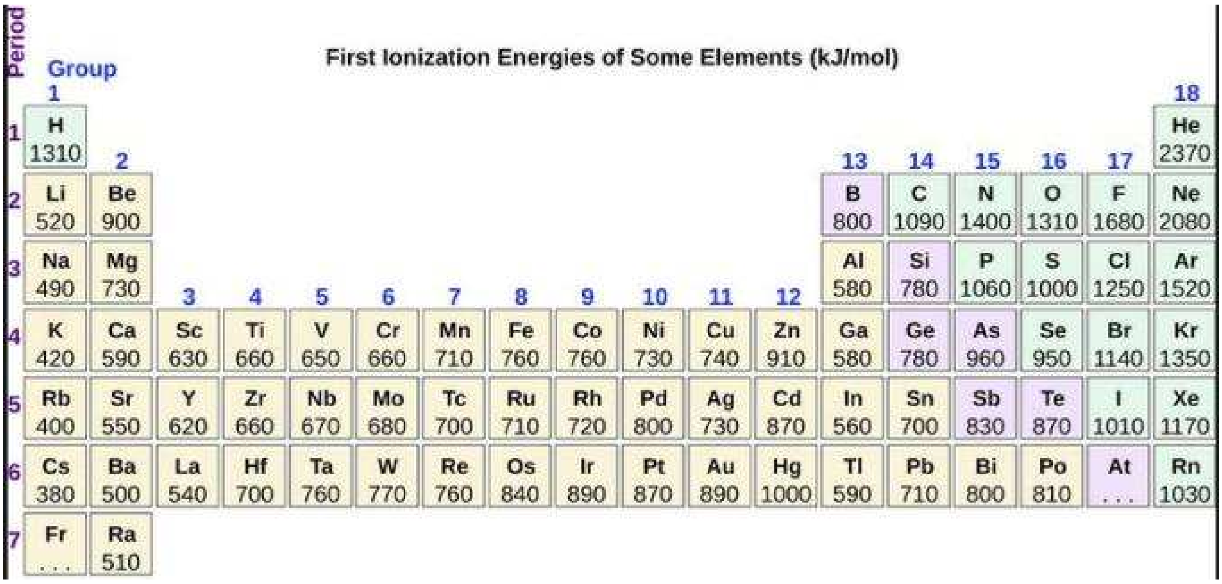




v · T · E **Electron affinities in the periodic table** [hide]

Group → ↓ Period	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1	H 73																	He •
2	Li 60	Be •											B 27	C 122	N •	O 141	F 328	Ne •
3	Na 53	Mg •											Al 42	Si 134	P 72	S 200	Cl 349	Ar •
4	K 48	Ca 2	Sc 18	Ti 8	V 51	Cr 65	Mn •	Fe 15	Co 64	Ni 112	Cu 119	Zn •	Ga 41	Ge 119	As 79	Se 195	Br 324	Kr •
5	Rb 47	Sr 5	Y 30	Zr 41	Nb 86	Mo 72	Tc •	Ru 101	Rh 110	Pd 54	Ag 126	Cd •	In 39	Sn 107	Sb 101	Te 190	I 295	Xe •
6	Cs 46	Ba 14	*	Hf	Ta 31	W 79	Re •	Os 104	Ir 150	Pt 205	Au 223	Hg •	Tl 36	Pb 35	Bi 91	Po	At	Rn •
7	Fr	Ra	**	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Fl	Uup	Lv	Uus	Uuo
* Lanthanides				La 45	Ce 92	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm 99	Yb	Lu 33
** Actinides				Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr

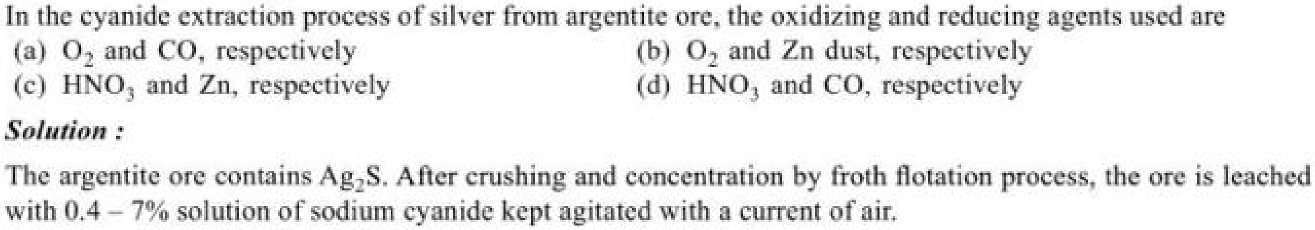
Legend
The number mentioned is **Electron affinity** in kJ/mol (rounded)
* Denotes elements that are expected to have electron affinities close to zero on quantum mechanical grounds



First Ionization Energies of Some Elements (kJ/mol)

Period \ Group	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1	H 1310																	He 2370
2	Li 520	Be 900											B 800	C 1090	N 1400	O 1310	F 1680	Ne 2080
3	Na 490	Mg 730											Al 580	Si 780	P 1060	S 1000	Cl 1250	Ar 1520
4	K 420	Ca 590	Sc 630	Ti 660	V 650	Cr 660	Mn 710	Fe 760	Co 760	Ni 730	Cu 740	Zn 910	Ga 580	Ge 780	As 960	Se 950	Br 1140	Kr 1350
5	Rb 400	Sr 550	Y 620	Zr 660	Nb 670	Mo 680	Tc 700	Ru 710	Rh 720	Pd 800	Ag 730	Cd 870	In 560	Sn 700	Sb 830	Te 870	I 1010	Xe 1170
6	Cs 380	Ba 500	La 540	Hf 700	Ta 760	W 770	Re 760	Os 840	Ir 890	Pt 870	Au 890	Hg 1000	Tl 590	Pb 710	Bi 800	Po 810	At . . .	Rn 1030
7	Fr . . .	Ra 510																

cyanide extraction process of silver from argentite ore

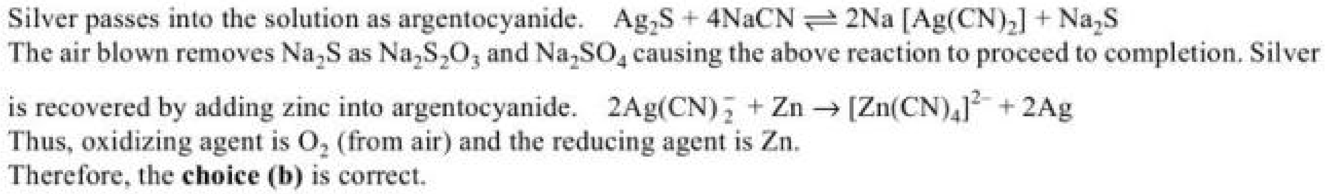


In the cyanide extraction process of silver from argentite ore, the oxidizing and reducing agents used are

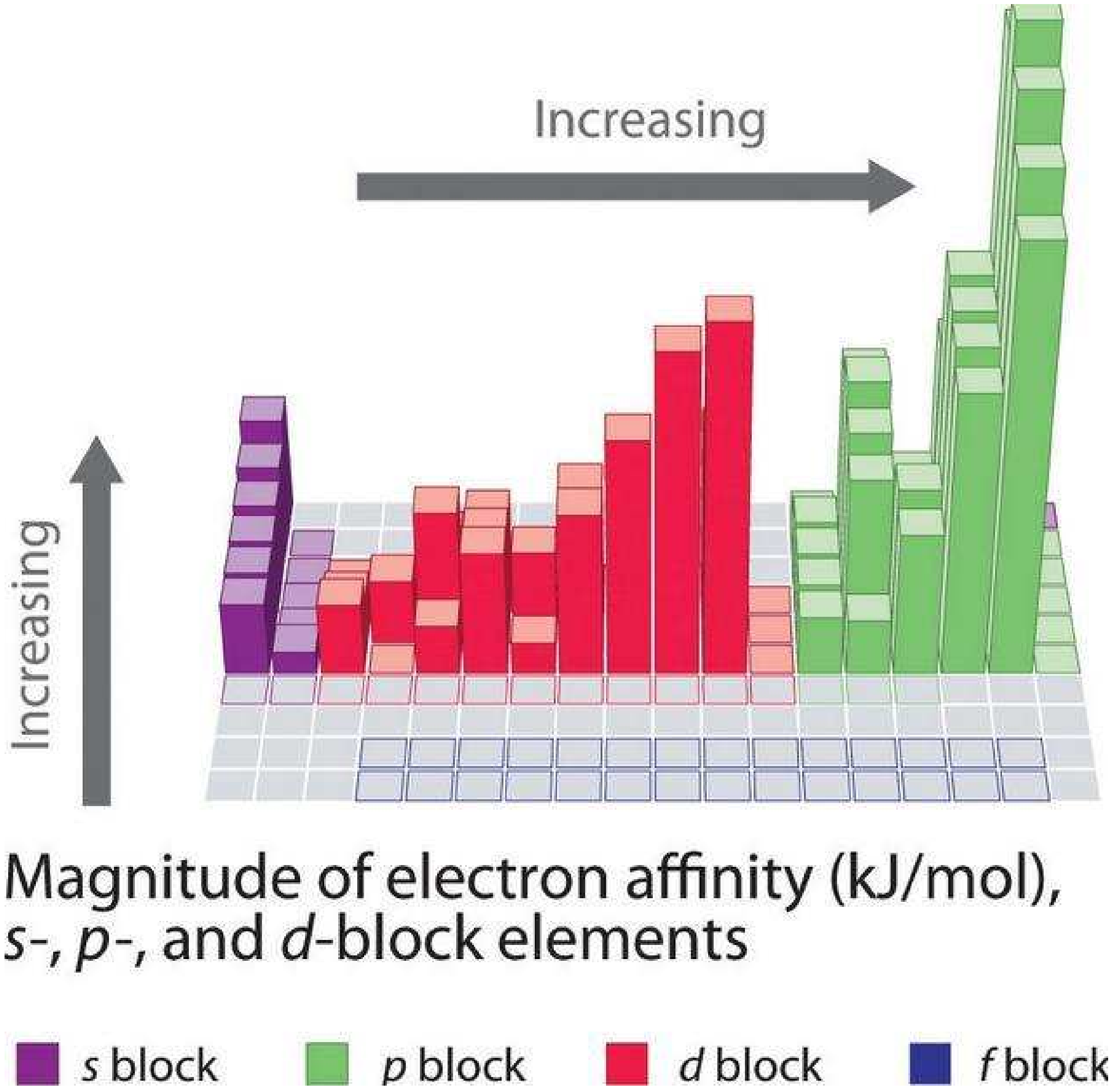
(a) O_2 and CO, respectively
(b) O_2 and Zn dust, respectively
(c) HNO_3 and Zn, respectively
(d) HNO_3 and CO, respectively

Solution :

The argentite ore contains Ag_2S. After crushing and concentration by froth flotation process, the ore is leached with 0.4 – 7% solution of sodium cyanide kept agitated with a current of air.



Silver passes into the solution as argentocyanide. $Ag_2S + 4NaCN \rightleftharpoons 2Na[Ag(CN)_2] + Na_2S$
The air blown removes Na_2S as $Na_2S_2O_3$ and Na_2SO_4 causing the above reaction to proceed to completion. Silver is recovered by adding zinc into argentocyanide. $2Ag(CN)_2^- + Zn \rightarrow [Zn(CN)_4]^{2-} + 2Ag$
Thus, oxidizing agent is O_2 (from air) and the reducing agent is Zn.
Therefore, the **choice (b)** is correct.



Magnitude of electron affinity (kJ/mol), *s*-, *p*-, and *d*-block elements

Common Name	**Chemical Name**
acetone	dimethyl ketone; 2-propanone (usually known as acetone)
acid potassium sulfate	potassium bisulfate
acid of sugar	oxalic acid
ackey	nitric acid
alcali volatil	ammonium hydroxide
alcohol, grain	ethyl alcohol
alcohol sulfuris	carbon disulfide
alcohol, wood	methyl alcohol
alum	aluminum potassium sulfate
alumina	aluminum oxide
antichlor	sodium thiosulfate
antimony black	antimony trisulfide
antimony bloom	antimony trioxide
antimony glance	antimony trisulfide
antimony red (vermillion)	antimony oxysulfide
aqua ammonia	aqueous solution of ammonium hydroxide
aqua fortis	nitric acid
aqua regia	nitrohydrochloric acid
aromatic spirit of ammonia	ammonia in alcohol
arsenic glass	arsenic trioxide
azurite	mineral form of basic copper carbonate
asbestos	magnesium silicate
aspirin	acetylsalicylic acid
baking soda	sodium bicarbonate
banana oil (artificial)	isoamyl acetate
barium white	barium sulfate
benzol	benzene
bicarbonate of soda	sodium hydrogen carbonate or sodium bicarbonate
bichloride of mercury	mercuric chloride
bichrome	potassium dichromate
bitter salt	magnesium sulfate
black ash	crude form of sodium carbonate

black copper oxide	cupric oxide
black lead	graphite (carbon)
blanc-fixe	barium sulfate
bleaching powder	chlorinated lime; calcium hypochlorite
blue copperas	copper sulfate (crystals)
blue lead	lead sulfate
blue salts	nickel sulfate
blue stone	copper sulfate (crystals)
blue vitriol	copper sulfate
bluestone	copper sulfate
bone ash	crude calcium phosphate
bone black	crude animal charcoal
boracic acid	boric acid
borax	sodium borate; sodium tetraborate
bremen blue	basic copper carbonate
brimstone	sulfur
burnt alum	anhydrous potassium aluminum sulfate
burnt lime	calcium oxide
burnt ochre	ferric oxide
burnt ore	ferric oxide
brine	aqueous sodium chloride solution
butter of antimony	antimony trichloride
butter of tin	anhydrous stannic chloride
butter of zinc	zinc chloride
calomel	mercury chloride; mercurous chloride
carbolic acid	phenol
carbonic acid gas	carbon dioxide
caustic lime	calcium hydroxide
caustic potash	potassium hydroxide
caustic soda	sodium hydroxide
chalk	calcium carbonate
Chile saltpeter	sodium nitrate
Chile nitre	sodium nitrate

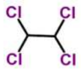
Chinese red	basic lead chromate
Chinese white	zinc oxide
chloride of soda	sodium hypochlorite
chloride of lime	calcium hypochlorite
chrome alum	chromic potassium sulfate
chrome green	chromium oxide
chrome yellow	lead (VI) chromate
chromic acid	chromium trioxide
copperas	ferrous sulfate
corrosive sublimate	mercury (II) chloride
corundum (ruby, sapphire)	chiefly aluminum oxide
cream of tartar	potassium bitartrate
crocus powder	ferric oxide
crystal carbonate	sodium carbonate
dechlor	sodium thiophosphate
diamond	carbon crystal
emery powder	impure aluminum oxide
epsom salts	magnesium sulfate
ethanol	ethyl alcohol
farina	starch
ferro prussiate	potassium ferricyanide
ferrum	iron
flores martis	anhydride iron (III) chloride
fluorspar	natural calcium fluoride
fixed white	barium sulfate
flowers of sulfur	sulfur
'flowers of' any metal	oxide of the metal
formalin	aqueous formaldehyde solution
French chalk	natural magnesium silicate
French vergidris	basic copper acetate
galena	natural lead sulfide
Glauber's salt	sodium sulfate
green verditer	basic copper carbonate

green vitriol	ferrous sulfate crystals
gypsum	natural calcium sulfate
hard oil	boiled linseed oil
heavy spar	barium sulfate
hydrocyanic acid	hydrogen cynanide
hypo (photography)	sodium thiosulfate solution
Indian red	ferric oxide
Isinglass	agar-agar gelatin
jeweler's rouge	ferric oxide
killed spirits	zinc chloride
lampblack	crude form of carbon; charcoal
laughing gas	nitrous oxide
lead peroxide	lead dioxide
lead protoxide	lead oxide
lime	calcium oxide
lime, slaked	calcium hydroxide
limewater	aqueous solution of calcium hydroxide
liquor ammonia	ammonium hydroxide solution
litharge	lead monoxide
lunar caustic	silver nitrate
liver of sulfur	sufurated potash
lye or soda lye	sodium hydroxide
magnesia	magnesium oxide
manganese black	manganese dioxide
marble	mainly calcium carbonate
mercury oxide, black	mercurous oxide
methanol	methyl alcohol
methylated spirits	methyl alcohol
milk of lime	calcium hydroxide
milk of magnesium	magnesium hydroxide
milk of sulfur	precipitated sulfur
"muriate" of a metal	chloride of the metal
muriatic acid	hydrochloric acid

natron	sodium carbonate
nitre	potassium nitrate
nordhausen acid	fuming sulfuric acid
oil of mars	deliquescent anhydrous iron (III) chloride
oil of vitriol	sulfuric acid
oil of wintergreen (artificial)	methyl salicylate
orthophosphoric acid	phosphoric acid
Paris blue	ferric ferrocyanide
Paris green	copper acetoarsenite
Paris white	powdered calcium carbonate
pear oil (artificial)	isoamyl acetate
pearl ash	potassium carbonate
permanent white	barium sulfate
plaster of Paris	calcium sulfate
plumbago	graphite
potash	potassium carbonate
potassa	potassium hydroxide
precipitated chalk	calcium carbonate
Prussic acid	hydrogen cyanide
pyro	tetrasodium pyrophosphate
quicklime	calcium oxide
quicksilver	mercury
red lead	lead tetraoxide
red liquor	aluminum acetate solution
red prussiate of potash	potassium ferrocyanide
red prussiate of soda	sodium ferrocyanide
Rochelle salt	potassium sodium tartrate
rock salt	sodium chloride
rouge, jeweler's	ferric oxide
rubbing alcohol	isopropyl alcohol
sal ammoniac	ammonium chloride
sal soda	sodium carbonate
salt, table	sodium chloride

salt of lemon	potassium binoxalate
salt of tartar	potassium carbonate
saltpeter	potassium nitrate
silica	silicon dioxide
slaked lime	calcium hydroxide
soda ash	sodium carbonate
soda nitre	sodium nitrate
soda lye	sodium hydroxide
soluble glass	sodium silicate
sour water	dilute sulfuric acid
spirit of hartshorn	ammonium hydroxide solution
spirit of salt	hydrochloric acid
spirit of wine	ethyl alcohol
spirits of nitrous ether	ethyl nitrate
sugar, table	sucrose
sugar of lead	lead acetate
sulfuric ether	ethyl ether
talc or talcum	magnesium silicate
tin crystals	stannous chloride
trona	natural sodium carbonate
unslaked lime	calcium oxide
Venetian red	ferric oxide
verdigris	basic copper acetate
Vienna lime	calcium carbonate
vinegar	impure dilute acetic acid
vitamin C	ascorbic acid
vitriol	sulfuric acid
washing soda	sodium carbonate
water glass	sodium silicate
white caustic	sodium hydroxide
white lead	basic lead carbonate
white vitriol	zinc sulfate crystals
yellow prussiate of potash	potassium ferrocyanide

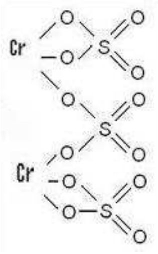
yellow prussiate of soda	sodium ferrocyanide
zinc vitriol	zinc sulfate
zinc white	zinc oxide

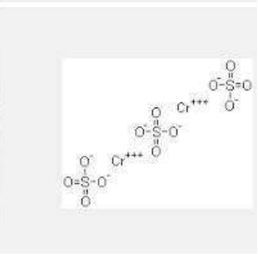
Westron - A refrigerant. 

$C_2H_2Cl_4$

Colours of Compounds

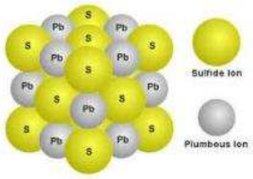
Cr2(SO4)3 - green



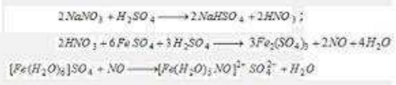



PbS - Black





[Fe(H2O)5NO]SO4 - brown



$2NaNO_3 + H_2SO_4 \longrightarrow 2NaHSO_4 + 2HNO_3$;

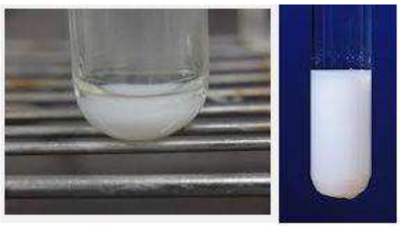
$2HNO_3 + 6FeSO_4 + 3H_2SO_4 \longrightarrow 3Fe_2(SO_4)_3 + 2NO + 4H_2O$

$[Fe(H_2O)_6]SO_4 + NO \longrightarrow [Fe(H_2O)_5NO]^{2+} SO_4^{2-} + H_2O$

(CH3COO)3Fe - blood red

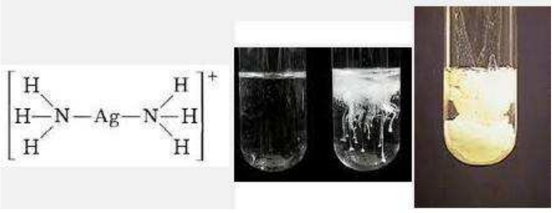
NH4Cl - white fumes

AgCl - curdy white precipitate



AgCl - curdy white precipitate

AgCl + NH4OH -> Ag(NH3)2Cl +H2O



AgCl + NH4OH --> Ag(NH3)2Cl +H2O

CrO3 + HCl -> CrO2Cl2 (deep red fumes) + H2O

CrO_2Cl_2 (deep red fumes)

$NaOH + CrO_2Cl_2 \rightarrow Na_2CrO_4$ (yellow) + $NaCl + H_2O$

Na_2CrO_4 (yellow)

Sodium Chromate

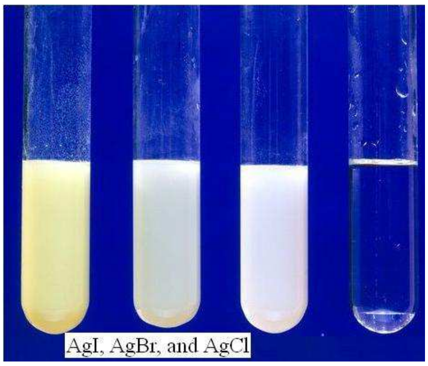
Sodium chromate is used in the industry to manufacture chromium compounds. Sodium chromate (Na_2CrO_4) is a yellow solid chemical compound. It is used in many industries like petroleum industry, textile industry. Sodium chromate is used in the industry as an important source for manufacture of chromium compounds. Sodium Chromate makes a protective film on metal surfaces and is used is corrosion control applications such as coolants containing alcohols and water cooling systems.

$Na_2CrO_4 + Pb(CH_3COO)_2 \rightarrow PbCrO_4$ (yellow) + CH_3COONa

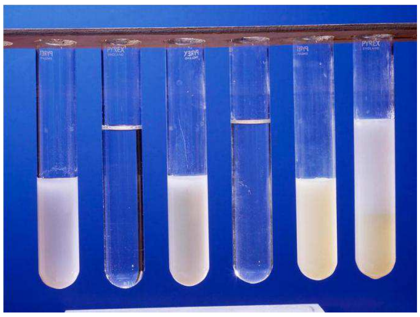
$PbCrO_4$ (yellow)

$NaBr + AgNO_3 \rightarrow AgBr$ (pale yellow) + $NaNO_3$

AgI, AgBr, AgCl

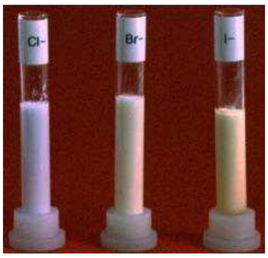


Ammonia test for silver halides. Test tubes and the results of using concentrated ammonia to test for the presence of three silver halides. Each one of the silver halides is next to a fuller test tube that shows the result of adding concentrated ammonia. Halides are compounds containing halogens such as here, from left to right silver chloride, silver bromide and silver iodide. Adding dilute ammonia dissolves the chloride, while concentrated ammonia is needed to dissolve the bromide, but it fails to dissolve the iodide. The silver halides are themselves the results of standard tests using silver nitrate to test for the halide ions.



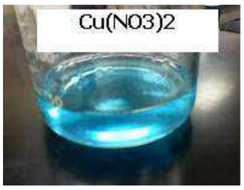
BrCl - pale yellow

$NaI + AgNO_3 \rightarrow AgI$ (yellow ppt) $+ NaNO_3$

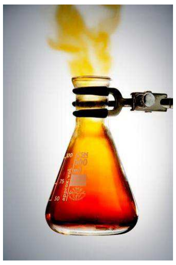


$Cu + HNO_3 \rightarrow Cu(NO_3)_2 + NO_2 + H_2O$

$Cu(NO_3)_2$



NO_2



$H_3BO_3 + C_2H_5OH \rightarrow (C_2H_5)_3BO_3 + H_2O$ (greenish flame)

$H_3PO_4 + 12(NH_4)_2MoO_4 + HNO_3 \rightarrow (NH_4)_3.PO_4.12MoO_3$ (canary yellow) $+ 21\ NH_4NO_3 + 12\ H_2O$

In the presence of Arsenic, we get yellow precipitate of $(NH_4)_3.AsO_4.12MoO_3$ (ammonium arsenomolybdate)

$Mo_3O_8.xH_2O$ - molybdenum blue

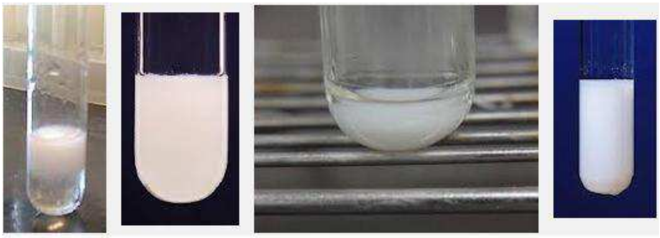
$SiF_4 + H_2O \rightarrow H_4SiO_4$ (silicic acid - white) + H_2SiO_4

$Na_2S + Na_2[Fe(CN)_5NO] \rightarrow Na_4[Fe(CN)_5NOS]$ - violet colour

$Na_2S + CdCO_3 \rightarrow CdS$ (yellow) + Na_2CO_3

$Na_3[Co(NO_2)_6] + KCl \rightarrow K_3[Co(NO_2)_6]$ (potassium cobaltnitrate - yellow ppt) + $NaCl$

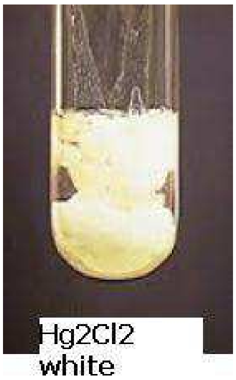
$AgCl$, $PbCl_2$, Hg_2Cl_2 - white



AgCl White

PbCl2 is white

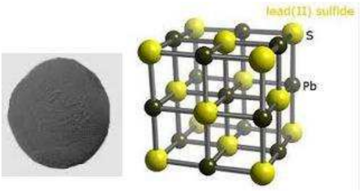
Hg2Cl2 is white



H2S, PbS, Bi2S3, CuS - black

H2S is Black

PbS is Black



PbS Lead Sulfide is Black

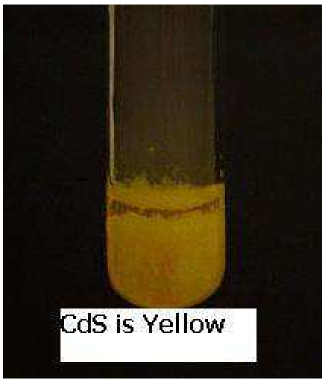
Bi2S3 is Black

CuS - black

Copper Sulfide Powder

CdS, As2S3, SnS2 - Yellow

CdS is Yellow

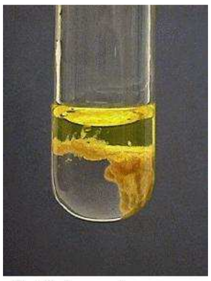


As2S3 is Yellow

SnS2 is Yellow

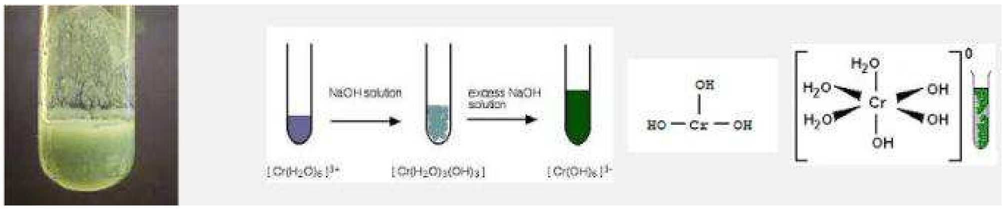
Sb2S3 is orange

SnS is Brown



SnS is Brown is your teacher says so. SnS is Yellow if your teacher says so :-)

Cr(OH)3 is green

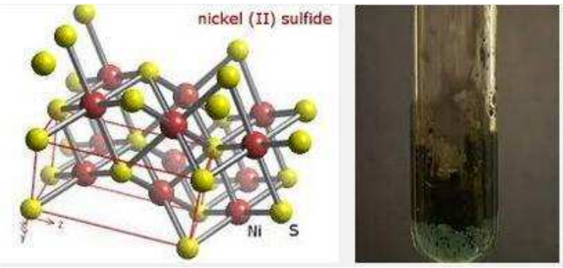


Chromium (III) Oxide is Green

CoS, NiS - black

CoS is black

NiS is Black Mineral Millerite



Cobalt Oxide is Black

Cobalt Sulphate is deep red brown

Cobalt Nitrate is deep red brown

Cobalt Chloride is deep Brown red

Cobalt Acetate is deep Brown

Azulene - **Deep Bue**

Iodoform - CHI_3 Yellow

Iron Phenol complex - **Violet** - $[Fe (OC_6H_5)_6]^{-3}$

$6 C_6H_5OH + FeCl_3 \text{ - -> } [Fe (OC_6H_5)_6]^{-3} + 3 H^+ + Cl^-$

Ammonium Phosphomolybdate - Yellow - $(NH_4)_3PO_4.12MoO_3$ (Yellow)

Prusian Blue - $Fe_4 [Fe(CN)_6]_3.xH_2O$

$[Fe(CN)_3 NOS]^{-4}$ - **Violet**

$[Fe(SCN)]^{2+}$ **Blood Red**

$[Co(NH_3)_5 Br]^{+2} SO_4^{-2}$ - **Violet** - Pentaaminebromocobalt(III)sulphate

$[Co(NH_3)_5SO_4]^+ Br^-$ -> **Red**

Test for Halogen:

Halogens present in an organic compound forms sodium halide on fusion with sodium metal. Sodium halide extracted with water can be easily identified by adding silver nitrate solution after acidifying with dil. HNO_3.

If chlorine is present, a white curdy precipitate soluble in ammonium hydroxide solution is formed.

$$Na + Cl \rightarrow NaCl$$

$$NaCl + AgNO_3 \rightarrow AgCl + NaNO_3$$

If bromine is present, an yellowish white precipitate sparingly soluble in ammonium hydroxide solution is formed.

$$Na + Br \rightarrow NaBr$$

$$NaBr + AgNO_3 \rightarrow AgBr + NaNO_3$$

If iodine is present, an yellow precipitate insoluble in ammonium hydroxide solution is formed.

$$Na + I \rightarrow NaI$$

$$NaI + AgNO_3 \rightarrow AgI + NaNO_3$$

Test for Nitrogen:

The carbon and nitrogen present in the organic compound on fusion with sodium metal gives sodium cyanide (NaCN) soluble in water. This is converted in to sodium ferrocyanide by the addition of sufficient quantity of ferrous sulphate. Ferric ions generated during the process react with ferrocyanide to form prussian blue precipitate of ferric ferrocyanide.

$$Na + C + N \rightarrow NaCN$$

$$6NaCN + FeSO_4 \rightarrow Na_4[Fe(CN)_6] + Na_2SO_4$$

Sodium ferrocyanide

$$Na_4[Fe(CN)_6] + Fe^{3+} \rightarrow Fe_4[Fe(CN)_6]_3$$

Ferric ferrocyanide

$FeCl_3 + K_4[Fe(CN)_6] \rightarrow Fe_4[Fe(CN)_6]_3$ (prussian blue) + KCl

Test for Sulphur:

If sulphur is present in the organic compound, sodium fusion will convert it into sodium sulphide. Sulphide ions are readily identified using sodium nitroprusside.

$$Na + S \rightarrow Na_2S$$

$$Na_2S + Na_2[Fe(CN)_5NO] \rightarrow Na_4[Fe(CN)_5NOS]$$

Sodium nitroprusside violet colour

Test for both Nitrogen and Sulphur:

If both nitrogen and sulphur are present in an organic compound, sodium fusion will convert it into sodium thiocyanate which then react with Fe^{3+} to form blood colour complex $[Fe(SCN)]^{2+}$

$$Na + C + N + S \rightarrow NaSCN$$

$$Fe^{3+} + 2Na \rightarrow [Fe(SCN)]^{2+}$$

Note: If the Lassaigne's extract containing excess of sodium metal, sodium cyanide and sulphides are formed instead of sodium thiocyanate.

$$NaSCN + 2Na \rightarrow NaCN + Na_2S$$

Here in this type of cases, both sulphur and nitrogen are to be identified in separate tests.

$BaCO_3$, $SrCO_3$, $CaCO_3$, $Mg(NH_4)PO_4$ - white

$BaCO_3$ is white

$BaCO_3$ powder

$SrCO_3$ is white

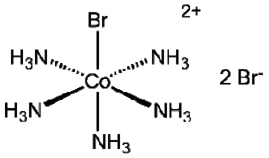
$CaCO_3$ powder is white

$CaCO_3$ Calcium Carbonate Powder

$Mg(NH_4)PO_4$ is white

Pentaamminechlorocobalt(III)Bromide

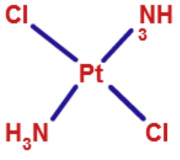
$[Co(NH_3)_5Cl]\ Cl_2$ Pentaamminechlorocobalt(III)chloride



Pentaamminesulphatocobalt(III)Bromide

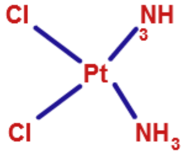
-

Dark Yellow

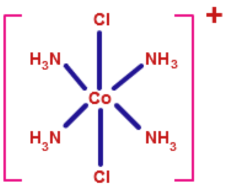


trans- $Pt(NH_3)_2Cl_2$

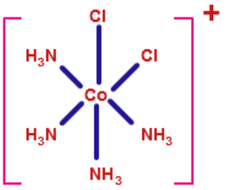
Pale Yellow



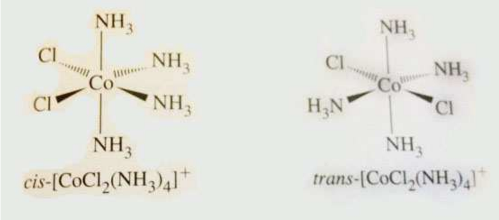
cis- $Pt(NH_3)_2Cl_2$



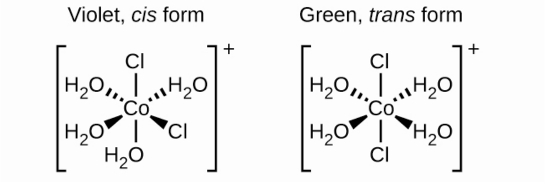
Trans - $[Co(NH_3)_4Cl_2]^+$

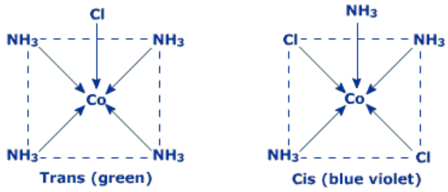


Cis - $[Co(NH_3)_4Cl_2]^+$



Cis is Blue Violet trans is green

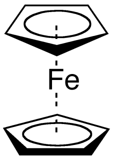


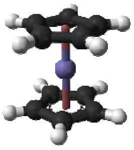


Eta bonds - Also known as Hapticity - Hapticity is the coordination of a ligand to a metal center via an uninterrupted and contiguous series of atoms.

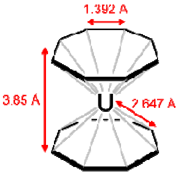
The hapticity of a ligand is described with the Greek letter η ('eta'). For example, η2 describes a ligand that coordinates through 2 contiguous atoms. In general the η-notation only applies when multiple atoms are coordinated (otherwise the κ-notation is used). In addition, if the ligand coordinates through multiple atoms that are not contiguous then this is considered denticity (not hapticity), and the κ-notation is used once again. Lastly, bridging ligands are described with the μ ('mu') notation.

Ferrocene - bis(η5-cyclopentadienyl)iron





Uranocene - bis(η8-1,3,5,7-cyclooctatetraene)uranium



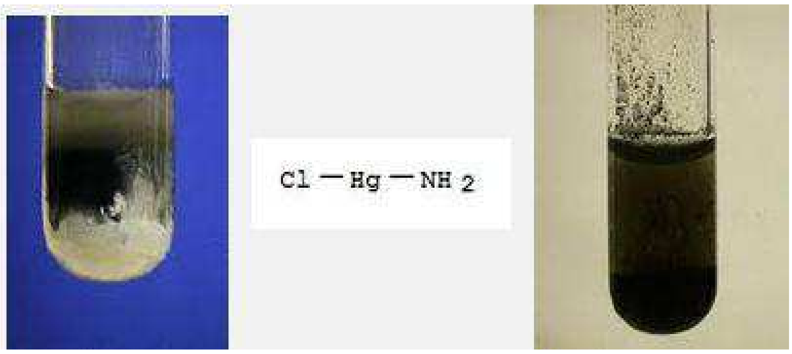
$W(CO)_3(PPr^i_3)_2(\eta^2\text{-}H_2)$ - the first compound to be synthesized with a dihydrogen ligand (also known as Dihydrogen Complexes)
$IrCl(CO)[P(C_6H_5)_3]_2(\eta^2\text{-}O_2)$ - the dioxygen derivative which forms reversibly upon oxygenation of Vaska's complex.

-

Chromium (III) chloride is purple

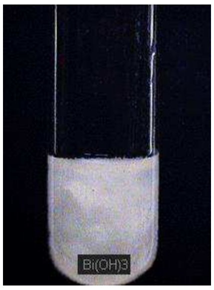
$HgCl_2 + NH_4OH \rightarrow (Hg + Hg(NH_2)Cl) + NH_4Cl + H_2O$

black ppt

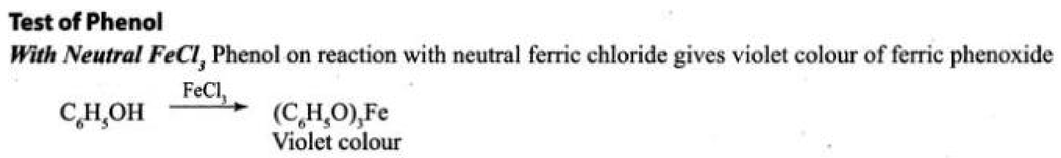


$BiCl_3 + H_2O \rightarrow BiOCl$ (white) $+ HCl$

$Bi(OH)_3$ Bismuth Hydroxide is white



Oxidation of Phenol



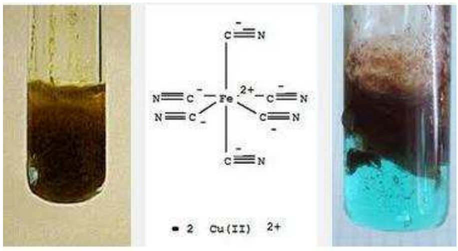
Test of Phenol

With Neutral $FeCl_3$ Phenol on reaction with neutral ferric chloride gives violet colour of ferric phenoxide

$$C_6H_5OH \xrightarrow{FeCl_3} (C_6H_5O)_3Fe$$

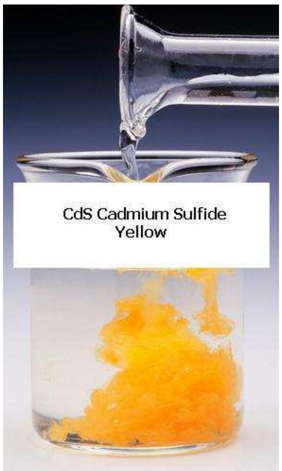
Violet colour

Cu(NO3)2 + K4[Fe(CN)6] -> Cu2[Fe(CN)6] (Chocolate brown) +KNO3



Cu2[Fe(CN)6]

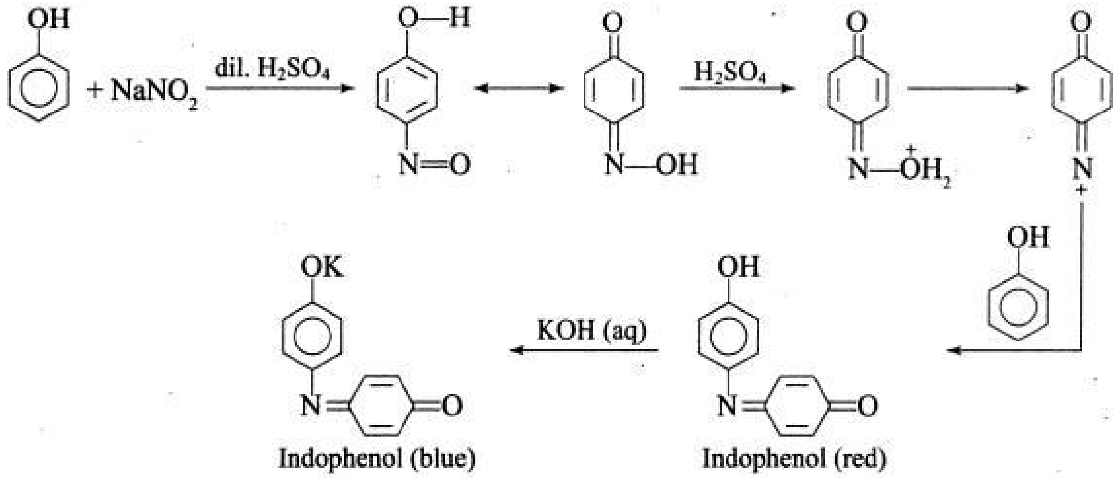
Cd(NH3)4(NO3)2 + H2S -> CdS (yellow) + NH4NO3 + NH3



$HgCl_2 + SnCl_2 \rightarrow Hg_2Cl_2$ (white) $+ SnCl_4$

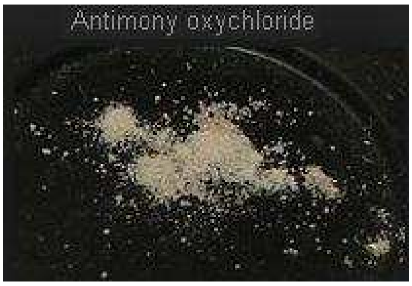
Liebermann's Nitroso Reaction

$$C_6H_5OH \text{ In } H_2SO_4 \xrightarrow[\text{in excess of } H_2O]{NaNO_2} \text{Red colouration} \xrightarrow[\text{excess}]{NaOH} \text{Blue colouration}$$



$SbCl_3 + H_2O \rightarrow SbOCl$ (white) $+ HCl$

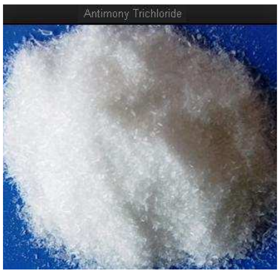
SbOCl Antimony oxychloride



Antimony Trioxide is white

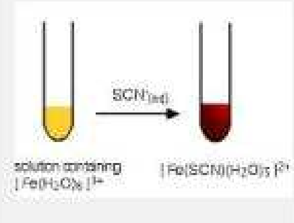
Antimony Pentoxide is gray

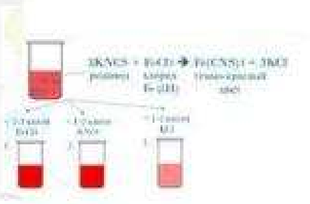
Antimony Trichloride is white

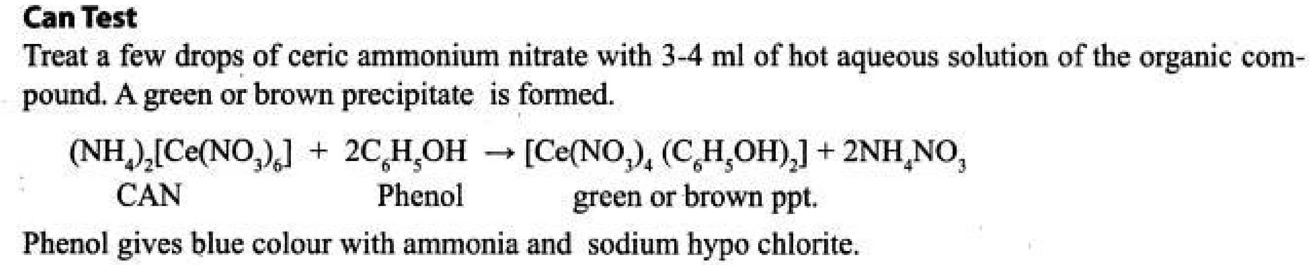


Potassium antimonate white

$FeCl_3$ + KSCN -> $Fe(CNS)_3$ (blood red) + KCN







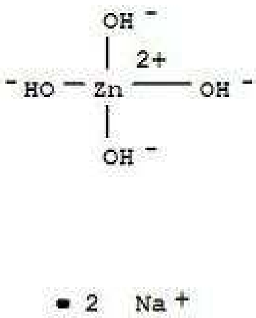
Can Test

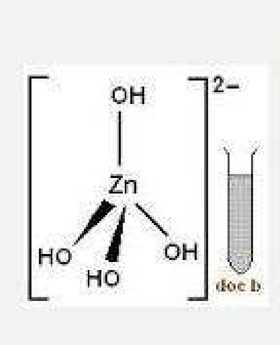
Treat a few drops of ceric ammonium nitrate with 3-4 ml of hot aqueous solution of the organic compound. A green or brown precipitate is formed.

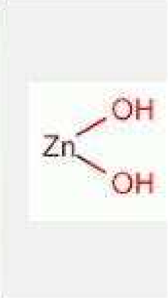
$$\underset{\text{CAN}}{(NH_4)_2[Ce(NO_3)_6]} + \underset{\text{Phenol}}{2C_6H_5OH} \longrightarrow \underset{\text{green or brown ppt.}}{[Ce(NO_3)_4\,(C_6H_5OH)_2]} + 2NH_4NO_3$$

Phenol gives blue colour with ammonia and sodium hypo chlorite.

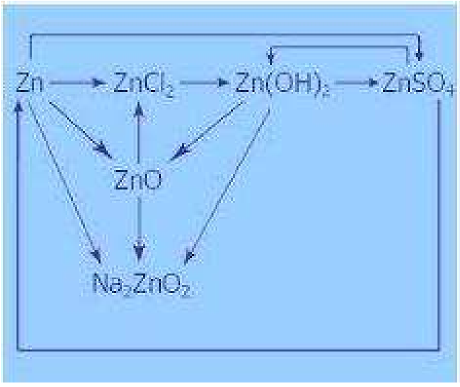
$ZnCl_2$ + 2NaOH -> $Zn(OH)_2$ (white ppt) + 2NaCl

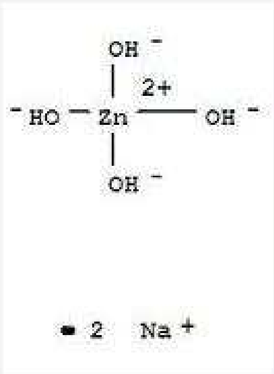




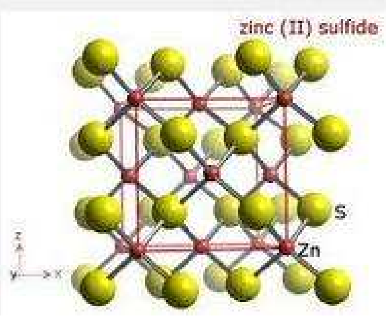


Zn(OH)2 + NaOH -> Na2ZnO2 (soluble) + H2O





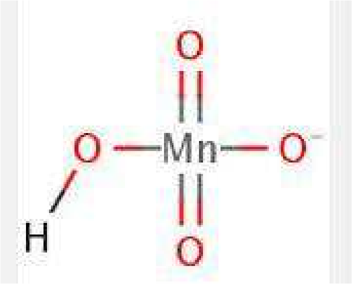
Na2ZnO2 + H2S -> ZnS (white) + NaOH

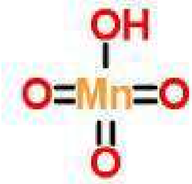


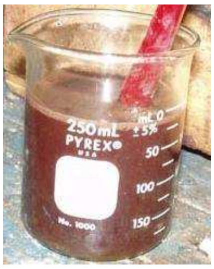
$$2KI + H_2O + O_3 \longrightarrow 2KOH + O_2 + I_2$$

Potassium iodide   Ozone   Potassium hydroxide   Iodine (Violet vapours)

Mn(NO3)2 + Pb3O4 + HNO3 -> HMnO4 (pink) + Pb(NO3)2 + H2O





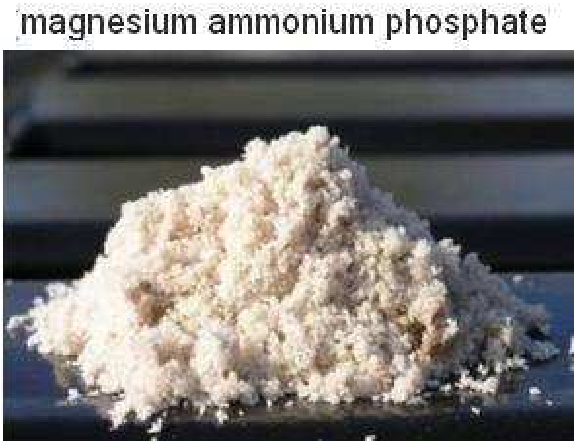


$Co(NO_2)_3 + KNO_2 \rightarrow K_3[Co(NO_2)_6]$ (yellow)

$Na_4Co(CO_3)_2 + H_2O + [O] \rightarrow Na_3Co(CO_3)_2$ (green coloration) + $NaOH$

$NiCO_3 + [O] \rightarrow Ni_2O_3$ (black) + CO_2

$MgCl_2 + Na_2HPO_4 + NH_4OH \rightarrow Mg(NH_4)PO_4$ (white) + NaCl + H_2O

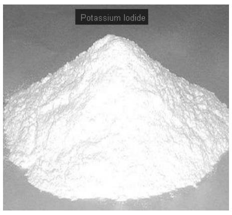


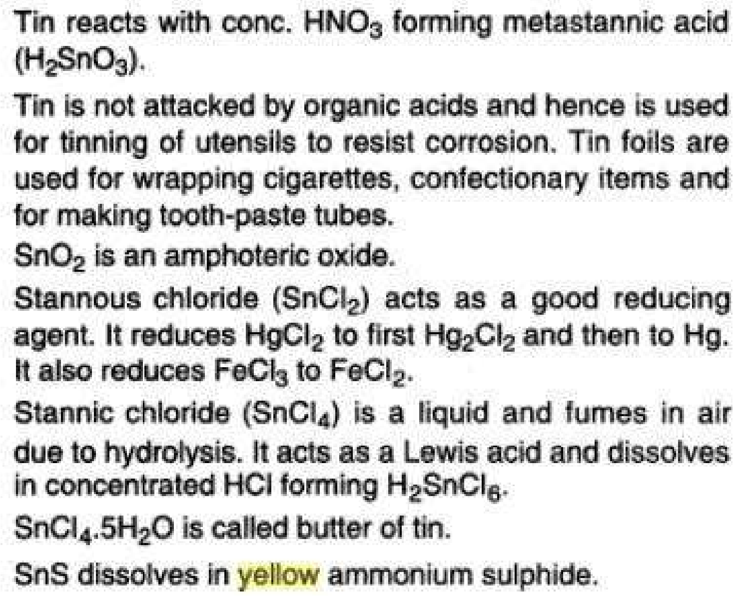
$K_2HgI_4 + NH_4Cl + KOH \rightarrow NH_2HgOHgI$ (Iodide of millon's base, brown ppt) + $KCl + H_2O$

Chromium oxide powder is green

Bismuth oxide is Yellow

KI Potassium Iodide is white





Tin reacts with conc. HNO_3 forming metastannic acid (H_2SnO_3).

Tin is not attacked by organic acids and hence is used for tinning of utensils to resist corrosion. Tin foils are used for wrapping cigarettes, confectionary items and for making tooth-paste tubes.

SnO_2 is an amphoteric oxide.

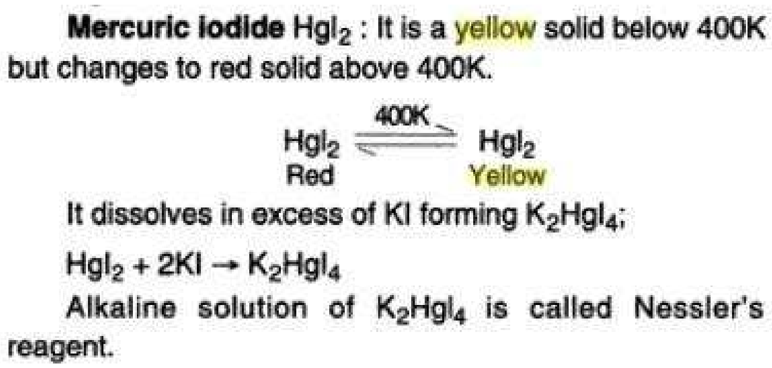
Stannous chloride ($SnCl_2$) acts as a good reducing agent. It reduces $HgCl_2$ to first Hg_2Cl_2 and then to Hg. It also reduces $FeCl_3$ to $FeCl_2$.

Stannic chloride ($SnCl_4$) is a liquid and fumes in air due to hydrolysis. It acts as a Lewis acid and dissolves in concentrated HCl forming H_2SnCl_6.

$SnCl_4.5H_2O$ is called butter of tin.

SnS dissolves in yellow ammonium sulphide.

Aluminum Oxide is white



Mercuric Iodide HgI_2 : It is a yellow solid below 400K but changes to red solid above 400K.

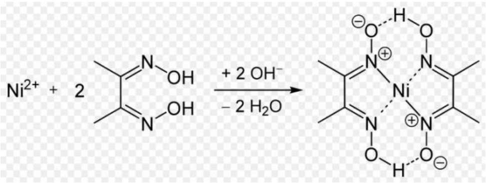
$$\underset{\text{Red}}{HgI_2} \xrightleftharpoons{400K} \underset{\text{Yellow}}{HgI_2}$$

It dissolves in excess of KI forming K_2HgI_4;

$$HgI_2 + 2KI \rightarrow K_2HgI_4$$

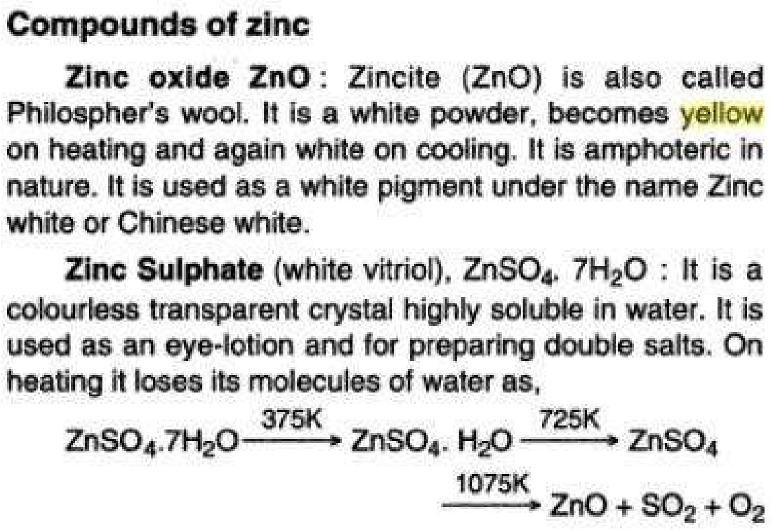
Alkaline solution of K_2HgI_4 is called Nessler's reagent.

Add NH4OH excess and dimethylglyoxime to NiCl2 then a rosy red ppt of nickel appears



A sample of Ni(dmgH)2

CuO + B2O3 (Glassy bead) -> Cu(BO2)2 (Copper metabolite - blue)



Compounds of zinc

Zinc oxide ZnO : Zincite (ZnO) is also called Philospher's wool. It is a white powder, becomes yellow on heating and again white on cooling. It is amphoteric in nature. It is used as a white pigment under the name Zinc white or Chinese white.

Zinc Sulphate (white vitriol), $ZnSO_4 . 7H_2O$: It is a colourless transparent crystal highly soluble in water. It is used as an eye-lotion and for preparing double salts. On heating it loses its molecules of water as,

$$ZnSO_4.7H_2O \xrightarrow{375K} ZnSO_4.\ H_2O \xrightarrow{725K} ZnSO_4 \xrightarrow{1075K} ZnO + SO_2 + O_2$$

Verdigris - basic copper acetate

This pigment was used to the Middle Ages in paints. Until the nineteenth century this was the most vibrant shade of green available and it was used by artists widely. The color comes from copper acetate and you should know that it is highly toxic to humans.

Compounds of Mercury

(1) Mercuric oxide, HgO : It is obtained as a red solid by heating mercury in air or oxygen for a long time at 673 K

$2Hg + O_2 \rightarrow 2\,HgO$ (red) or by heating mercuric nitrate alone or in the presence of Hg

Heat

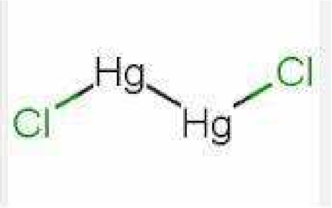
$2Hg(NO_3)_2 \rightarrow 2HgO + 4NO_2 + O_2$

red

When *NaOH* is added to a solution of $HgCl_2$, yellow precipitate of *HgO* are obtained.

$Hg_2Cl_2 + 2NaOH \rightarrow HgO \downarrow + H_2O + 2NaCl$

(yellow)



Red and yellow forms of *HgO* differ only in their particle size. On heating to 673 *K*, yellow form changes to red form.

$$\underset{Yellow}{HgO} \xrightarrow{673K} \underset{red}{HgO}$$

It is used in oil paints or as a mild antiseptic in ointments.

(2) Mercuric chloride, *$HgCl_2$* : It is obtained by treating *Hg* with Cl_2 or by heating a mixture of *NaCl*and $HgSO_4$ in presence of small amount of MnO_2 (which oxidizes any *Hg*(I) salts formed during the reaction).

Heat

$HgSO_4 + 2NaCl \rightarrow HgCl_2 + Na_2SO_4$

MnO_2

HgCl2 is a white crystalline solid and is commonly known as corrosive sublimate. It is a covalent compound since it dissolves in organic solvents like ethanol and ether.

It is extremely poisonous and causes death. Its best antidote is white of an egg.

When treated with stannous chloride, it is first reduced to white ppt. of mercurous chloride and then to mercury (black).

$2HgCl_2 + SnCl_2 \rightarrow Hg_2Cl_2 + SnCl_4$

white ppt.

$Hg_2Cl_2 + SnCl_2 \rightarrow 2Hg + SnCl_4$

grey

With ammonia it gives a white ppt. known as infusible white ppt.

$HgCl_2 + 2NH_3 \rightarrow Hg\,(NH_2)\,Cl + NH_4\,Cl$A dilute solution of $HgCl_2$ is used as an antiseptic.(3) Mercuric iodide, *HgI_2* : It is obtained when a required amount of *KI* solution is added to a solution of . $HgCl_2$$HgCl_2 + 2KI \rightarrow HgI_2 + 2KCl^{(red)}$Below 400 *K*, HgI_2 is red but above 400 *K*, it turns yellow

$$\underset{(\text{red})}{HgI_2} \underset{\text{below } 400\,K}{\overset{\text{above } 400\,K}{\rightleftharpoons}} \underset{(\text{yellow})}{HgI_2}$$

HgI_2 readily dissolves in excess of *KI* solution to form the $(HgI_4)^{2-}$ complex ion.$HgI_2 + 2KI \rightarrow$ K_2HgI_4 Red ppt. soluble colourless solution

An alkaline solution of $K_2[HgI_4]$ is called Nessler's reagent and is used to test NH^+_4 ions.

It gives a brown ppt. of $NH_2 - Hg - O\ Hg - I$ (Iodide of Millon's base) with NH^+_4 ions.

$2K_2[HgI_4] + NH_3 + 3KOH \rightarrow NH_2.\ HgO.\ HgI + 7KI + 2H_2O$

It is used in ointments for treating skin infections.

(4) Mercurous chloride, *Hg_2Cl_2* : It is obtained as under :

(a) $Hg_2(NO_3) + 2NaCl \rightarrow Hg_2Cl_2 + 2NaNo_3$

white ppt.

Heat in an iron retort

(b) $HgCl_2 + Hg \rightarrow Hg_2\,Cl_2$(condenses on cooling)

It is purified by sublimation.

Mercurous chloride is also called calomel. It is a white powder insoluble in H_2O. On heating, it decomposes to give $HgCl_2$ and *Hg*.

$$Hg_2\,Cl_2 \xrightarrow{Heat} Hgcl_2 + Hg$$

It dissolves in chlorine water forming mercuric chloride.

$$Hg_2Cl_2 + Cl_2 \rightarrow 2HgCl_2$$

With ammonia, it turns black due to the formation of a mixture of finely divided black *Hg* and mercuric amino chloride.

$$Hg_2Cl_2 + 2NH_3 \rightarrow Hg + NH_2\,HgCl + NH_4Cl$$

(black)

It is used to prepare standard calomel electrode and as a purgative in medicine.

(5) Mercuric sulphide, *HgS* : The solubility product of *HgS* is lower than that of *ZnS* and hence it gets precipitated as black solid when H_2S is passed through an acidic solution of any mercury (II) salt.

$$HgCl_2 + H_2S \rightarrow HgS + 2HCl$$

It is insoluble in water and *HCl* but dissolves in aqua regia (1 part conc.HNO_3 + 3 parts conc. *HCl*)

$$3HCl \quad + \quad HNO_3 \rightarrow NOCl \; + \; 2H_2O \quad + \quad 2\,[\,Cl\,]$$

Aqua regia   Nitrosyl chloride   Nacent chlorine

$$HgS + 2|\,Cl\,| \rightarrow HgCl_2 + S\downarrow$$

(Soluble)

On sublimation, its colour changes to red and hence it is used as a red pigment.

(6) Mercuric sulphate, *$HgSO_4$* : It is obtained when *HgS* is treated with conc.H_2SO_4.

$$Hg + 2H_2\,SO_4 \rightarrow HgSo_4 + SO_2 + 2H_2O$$

It is a white solid which decomposes on heating to give mercurous sulphate.

$$3HgSo_4 \xrightarrow{\Delta} Hg_2\,SO_4 + Hg + 2SO_2 + 2O_2$$

675K

It is used as a catalyst in the hydration of alkynes to give aldehydes or ketones. It is also used as a cosmetic under the name Vermillon and in ayurvedic medicine as makardhwaj.

(7) Amalgams : Mercury forms alloys commonly known as amalgams, with all metals except iron and platinum. Hence it is transported in iron containers.

(8) Alloy of transition metal : See in table discuss earlier in metallurgy.

xxxxxxxxxxxxxxxxxxxxx

The common Anions are divided into 3 groups for the purpose of identification:

Those which evolve gases with dilute hydrochloric acid :

Carbonate (CO_3^{2-})

bicarbonate(HCO_3^-)

Nitrite(NO_2^-)

Sulphide (S^{2-})

Sulphite(SO_3^{2-})

Thiosulphate($S_2O_3^{2-}$)

Those which do not react with dilute HCl, but which do evolve gases or volatile liquids with concentrated sulphuric acid:

Chloride (Cl^-)

Bromide (Br^-)

Iodide (I^-)

Nitrate (NO_3^-)

Those which do not react with either dilute hydrochloric acid or concentrated sulphuric acid:

Phosphate (PO_4^{3-})

Borate ($B_4O_7^{2-}$)

Sulphate (SO_4^{2-})

A Radical is an atom, or group of atoms that behaves like a single atom. Examples : H, K, Na, Cu, Pb, simple radicals ; NH 4 , SO 2 , NO 2 , compound radicals.

A Basic Radical is a metal, or any radical behaving like a metal. Examples : Ag, Cu, Ca, Na, and NH 4 the last

so classified because it behaves like such simple radicals as Na, K, and Li.

An Acid Radical is a nonmetal, or any radical behaving like a nonmetal. Examples : Cl, Br, I, SO 2 , NO2

An Acid is a substance containing H that may be replaced by a basic radical. When a nonmetal is combined with H only, as in HC1, HBr, H 2 S, the compound is said to be a binary (i.e. composed of two elements) acid, or an hydracid. When the nonmetal is combined with H and O, the compound is called a ternary (i.e. composed of three or more elements) acid, or an oxyacid ; and the characteristic name of the acid is formed by adding the termination -ic to the name of the nonmetal to denote more oxygen, or the termination -ous to denote less. Examples : H_2SO_4 sulphuric, and H_2SO_3 sulphurous acid.

These are classified in six different groups in accordance with certain properties which permit a group precipitation by a single reagent. A reagent used for precipitation of a whole group is called a general reagent.

A reagent used as a test for a single substance is known as a special reagent.

Experiments showing Method of Classification

Use the following solutions : $AgNO_3$, $Cu(NO3)_2$, $AsCl_3$, Fe_2Cl_6, $Co(N03)_2$, $Ba(N03)_2$, NaCl.

Take about 2 c.c. of each solution, and add HC1.*

Points to be Observed Whether a precipitate is formed.

Whether the precipitate is soluble in an excess of the reagent.

* Always use dilute acids unless otherwise directed.

GROUPING THE BASES

The color of each precipitate.

Whether there is change in color of the precipitate.

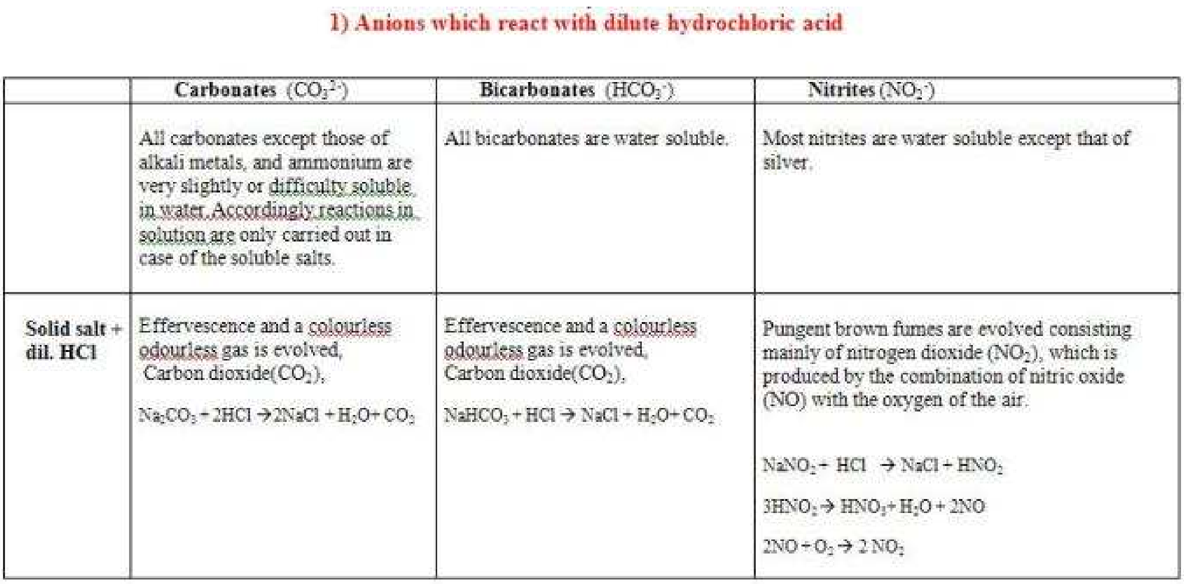
Change in color of solutions.

Chemical equations.

Tabulate all results carefully in your notebook, according to the following scheme :

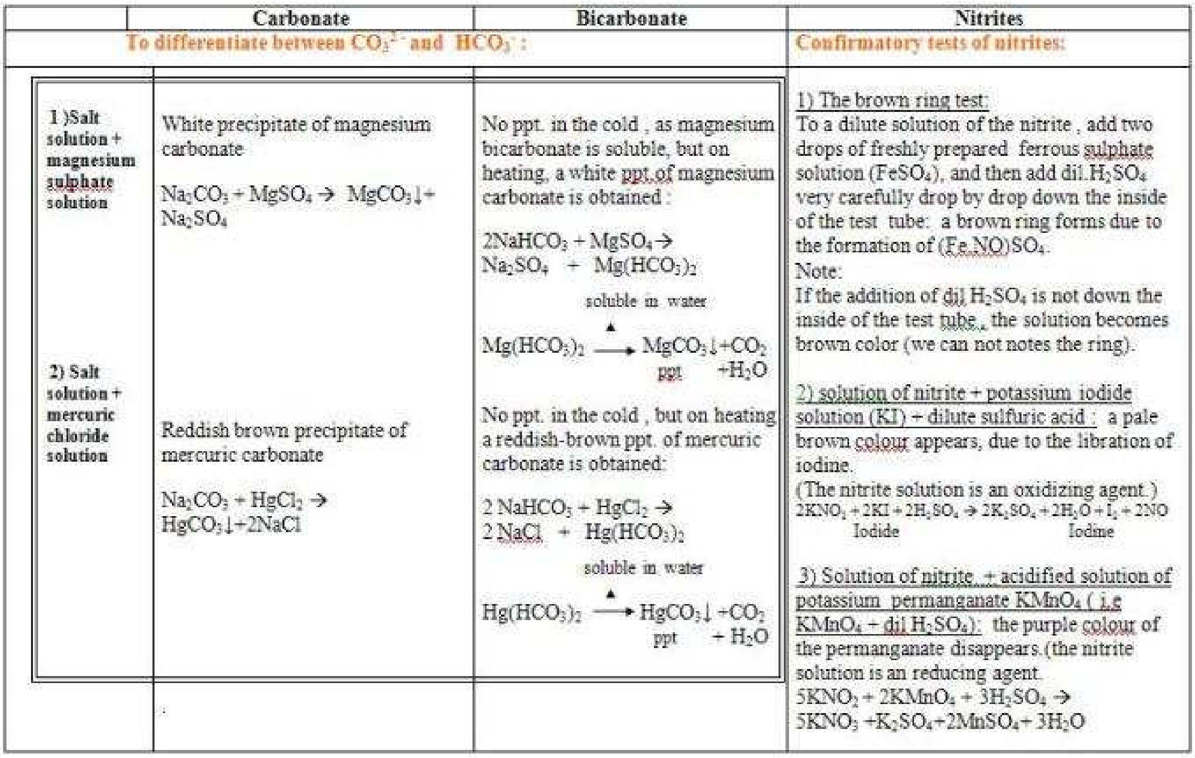
HC1 was added to the seven solutions with the following results :

Typical of Gr. I AgNO 3 + HC1 = AgCl w. pp. + I IN ( ) 8 .
Typical of Gr. II A $Cu(NO_3)_2$ + II Cl = No change.
Typical of Gr. II B AsCl 3 + HC1 = No change.
Typical of Gr. Ill Fe 2 Cl 6 + HC1 = No change.
Typical of Gr. IV Co(NO 3) 2 + IIC1 = No change.
Typical of Gr. V Ba(NO 3) 2 + HC1 = No change.
Typical of Gr. VI NaCl + HC1 = No change.



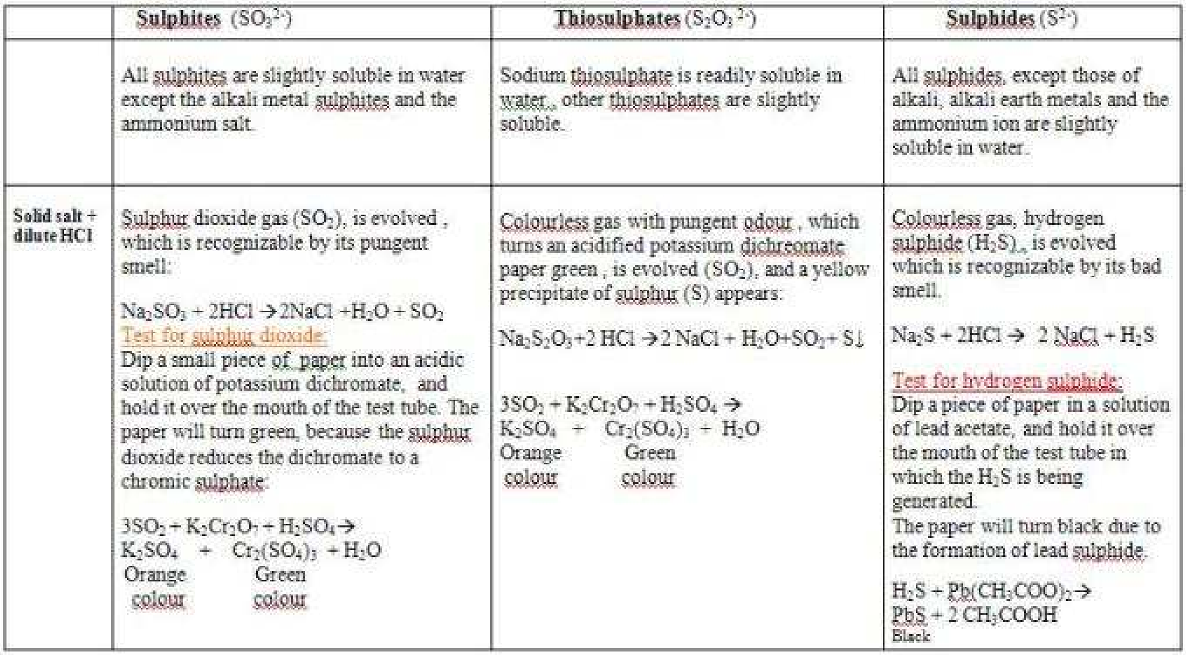
1) Anions which react with dilute hydrochloric acid

	Carbonates (CO_3^{2-})	**Bicarbonates** (HCO_3^-)	**Nitrites** (NO_2^-)
	All carbonates except those of alkali metals, and ammonium are very slightly or difficulty soluble in water. Accordingly reactions in solution are only carried out in case of the soluble salts.	All bicarbonates are water soluble.	Most nitrites are water soluble except that of silver.
Solid salt + dil. HCl	Effervescence and a colourless odourless gas is evolved, Carbon dioxide(CO_2). $Na_2CO_3 + 2HCl \rightarrow 2NaCl + H_2O + CO_2$	Effervescence and a colourless odourless gas is evolved, Carbon dioxide(CO_2). $NaHCO_3 + HCl \rightarrow NaCl + H_2O + CO_2$	Pungent brown fumes are evolved consisting mainly of nitrogen dioxide (NO_2), which is produced by the combination of nitric oxide (NO) with the oxygen of the air. $NaNO_2 + HCl \rightarrow NaCl + HNO_2$ $3HNO_2 \rightarrow HNO_3 + H_2O + 2NO$ $2NO + O_2 \rightarrow 2NO_2$



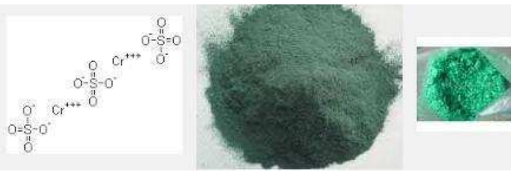
	Carbonate	**Bicarbonate**	**Nitrites**
To differentiate between CO_3^{2-} and HCO_3^- :			Confirmatory tests of nitrites:
1)Salt solution + magnesium sulphate solution	White precipitate of magnesium carbonate $Na_2CO_3 + MgSO_4 \rightarrow MgCO_3\downarrow + Na_2SO_4$	No ppt. in the cold , as magnesium bicarbonate is soluble, but on heating, a white ppt.of magnesium carbonate is obtained : $2NaHCO_3 + MgSO_4 \rightarrow Na_2SO_4 + \underset{\text{soluble in water}}{Mg(HCO_3)_2}$ $Mg(HCO_3)_2 \xrightarrow{\blacktriangle} \underset{ppt}{MgCO_3\downarrow} + CO_2 + H_2O$	1) The brown ring test: To a dilute solution of the nitrite , add two drops of freshly prepared ferrous sulphate solution ($FeSO_4$), and then add dil.H_2SO_4 very carefully drop by drop down the inside of the test tube: a brown ring forms due to the formation of $(Fe.NO)SO_4$. Note: If the addition of dil H_2SO_4 is not down the inside of the test tube., the solution becomes brown color (we can not notes the ring). 2) solution of nitrite + potassium iodide solution (KI) + dilute sulfuric acid : a pale brown colour appears, due to the libration of iodine. (The nitrite solution is an oxidizing agent.) $2KNO_2 + \underset{\text{Iodide}}{2KI} + 2H_2SO_4 \rightarrow 2K_2SO_4 + 2H_2O + \underset{\text{Iodine}}{I_2} + 2NO$ 3) Solution of nitrite. + acidified solution of potassium permanganate $KMnO_4$ (i.e $KMnO_4$ + dil H_2SO_4): the purple colour of the permanganate disappears.(the nitrite solution is an reducing agent. $5KNO_2 + 2KMnO_4 + 3H_2SO_4 \rightarrow 5KNO_3 + K_2SO_4 + 2MnSO_4 + 3H_2O$
2) Salt solution + mercuric chloride solution	Reddish brown precipitate of mercuric carbonate $Na_2CO_3 + HgCl_2 \rightarrow HgCO_3\downarrow + 2NaCl$	No ppt. in the cold , but on heating a reddish-brown ppt. of mercuric carbonate is obtained: $2NaHCO_3 + HgCl_2 \rightarrow 2NaCl + \underset{\text{soluble in water}}{Hg(HCO_3)_2}$ $Hg(HCO_3)_2 \xrightarrow{\blacktriangle} \underset{ppt}{HgCO_3\downarrow} + CO_2 + H_2O$	

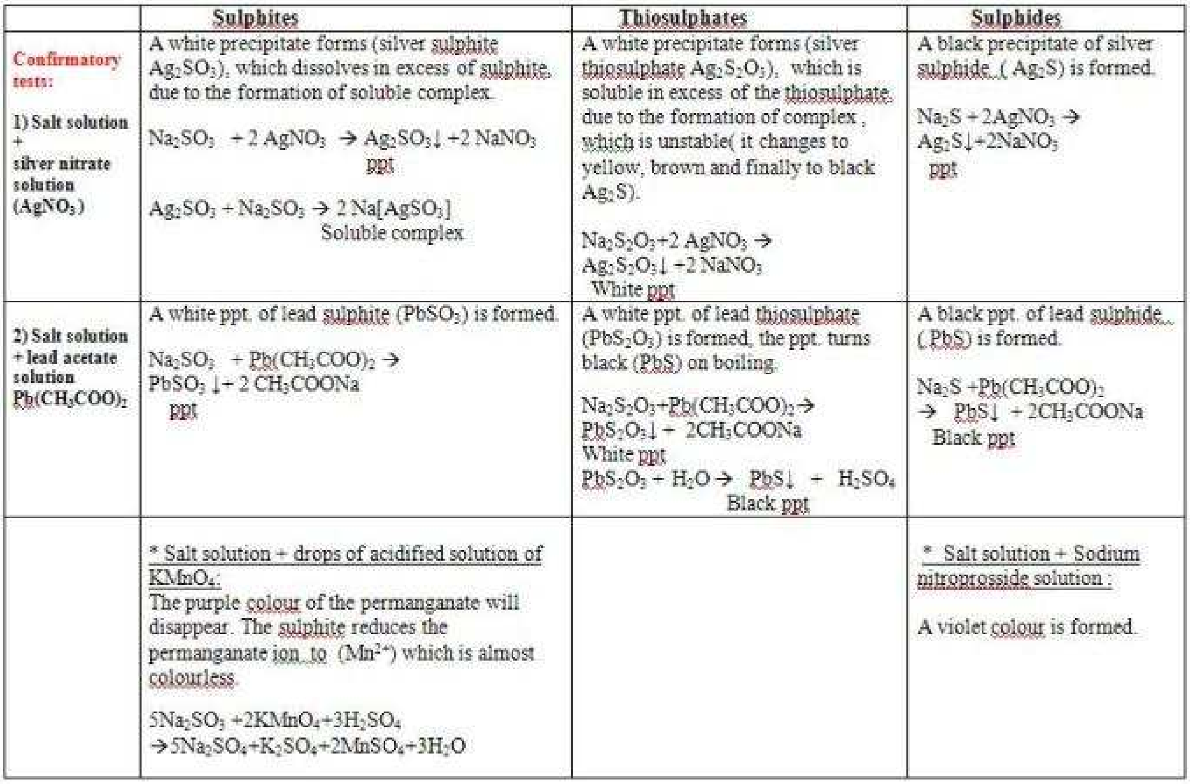
Is HgCO3 reddish Brown ?



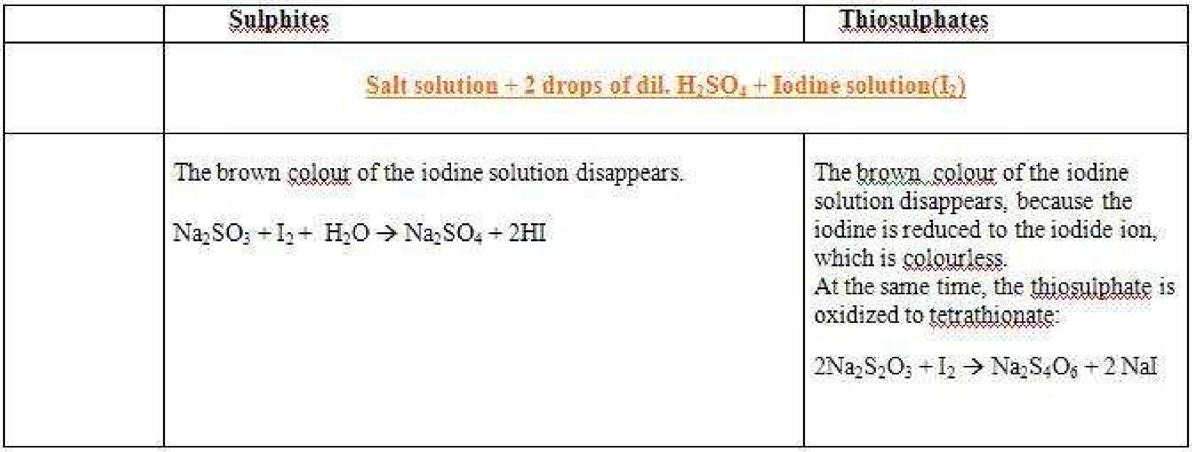
	Sulphites (SO_3^{2-})	**Thiosulphates** ($S_2O_3^{2-}$)	**Sulphides** (S^{2-})
	All sulphites are slightly soluble in water except the alkali metal sulphites and the ammonium salt.	Sodium thiosulphate is readily soluble in water., other thiosulphates are slightly soluble.	All sulphides, except those of alkali, alkali earth metals and the ammonium ion are slightly soluble in water.
Solid salt + dilute HCl	Sulphur dioxide gas (SO_2), is evolved , which is recognizable by its pungent smell: $Na_2SO_3 + 2HCl \rightarrow 2NaCl + H_2O + SO_2$ Test for sulphur dioxide: Dip a small piece of paper into an acidic solution of potassium dichromate, and hold it over the mouth of the test tube. The paper will turn green, because the sulphur dioxide reduces the dichromate to a chromic sulphate: $3SO_2 + K_2Cr_2O_7 + H_2SO_4 \rightarrow$ $K_2SO_4 + Cr_2(SO_4)_3 + H_2O$ Orange colour Green colour	Colourless gas with pungent odour , which turns an acidified potassium dichreomate paper green , is evolved (SO_2), and a yellow precipitate of sulphur (S) appears: $Na_2S_2O_3 + 2HCl \rightarrow 2NaCl + H_2O + SO_2 + S\downarrow$ $3SO_2 + K_2Cr_2O_7 + H_2SO_4 \rightarrow$ $K_2SO_4 + Cr_2(SO_4)_3 + H_2O$ Orange colour Green colour	Colourless gas, hydrogen sulphide (H_2S)., is evolved which is recognizable by its bad smell. $Na_2S + 2HCl \rightarrow 2NaCl + H_2S$ Test for hydrogen sulphide: Dip a piece of paper in a solution of lead acetate, and hold it over the mouth of the test tube in which the H_2S is being generated. The paper will turn black due to the formation of lead sulphide. $H_2S + Pb(CH_3COO)_2 \rightarrow$ $PbS + 2CH_3COOH$ Black

K2SO4 is white 

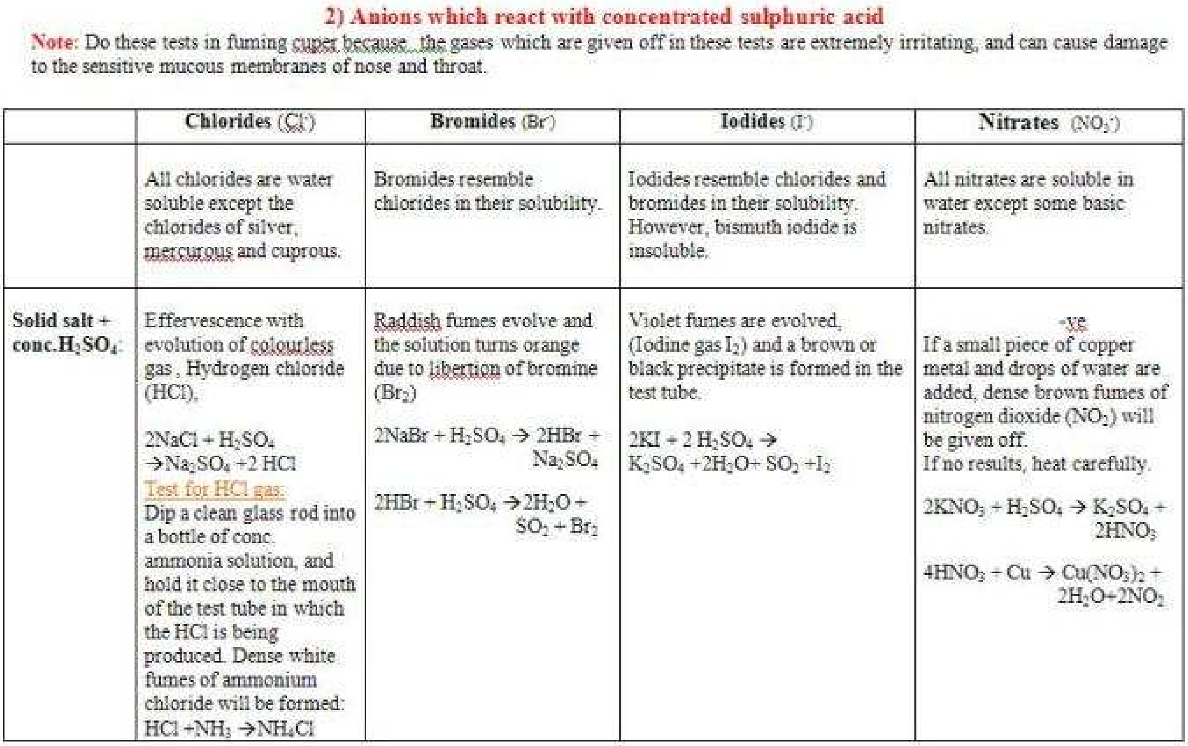
Cr2(SO4)3 is green 



	Sulphites	**Thiosulphates**	**Sulphides**
Confirmatory tests: **1) Salt solution + silver nitrate solution ($AgNO_3$)**	A white precipitate forms (silver sulphite Ag_2SO_3), which dissolves in excess of sulphite, due to the formation of soluble complex. $Na_2SO_3 + 2\,AgNO_3 \rightarrow Ag_2SO_3\downarrow + 2\,NaNO_3$ ppt $Ag_2SO_3 + Na_2SO_3 \rightarrow 2\,Na[AgSO_3]$ Soluble complex	A white precipitate forms (silver thiosulphate $Ag_2S_2O_3$), which is soluble in excess of the thiosulphate, due to the formation of complex , which is unstable(it changes to yellow, brown and finally to black Ag_2S). $Na_2S_2O_3 + 2\,AgNO_3 \rightarrow Ag_2S_2O_3\downarrow + 2\,NaNO_3$ White ppt	A black precipitate of silver sulphide (Ag_2S) is formed. $Na_2S + 2AgNO_3 \rightarrow Ag_2S\downarrow + 2NaNO_3$ ppt
2) Salt solution + lead acetate solution $Pb(CH_3COO)_2$	A white ppt. of lead sulphite ($PbSO_3$) is formed. $Na_2SO_3 + Pb(CH_3COO)_2 \rightarrow PbSO_3\downarrow + 2\,CH_3COONa$ ppt	A white ppt. of lead thiosulphate (PbS_2O_3) is formed, the ppt. turns black (PbS) on boiling. $Na_2S_2O_3 + Pb(CH_3COO)_2 \rightarrow PbS_2O_3\downarrow + 2CH_3COONa$ White ppt $PbS_2O_3 + H_2O \rightarrow PbS\downarrow + H_2SO_4$ Black ppt	A black ppt. of lead sulphide (PbS) is formed. $Na_2S + Pb(CH_3COO)_2 \rightarrow PbS\downarrow + 2CH_3COONa$ Black ppt
	* Salt solution + drops of acidified solution of $KMnO_4$: The purple colour of the permanganate will disappear. The sulphite reduces the permanganate ion to (Mn^{2+}) which is almost colourless. $5Na_2SO_3 + 2KMnO_4 + 3H_2SO_4 \rightarrow 5Na_2SO_4 + K_2SO_4 + 2MnSO_4 + 3H_2O$		* Salt solution + Sodium nitroprosside solution : A violet colour is formed.



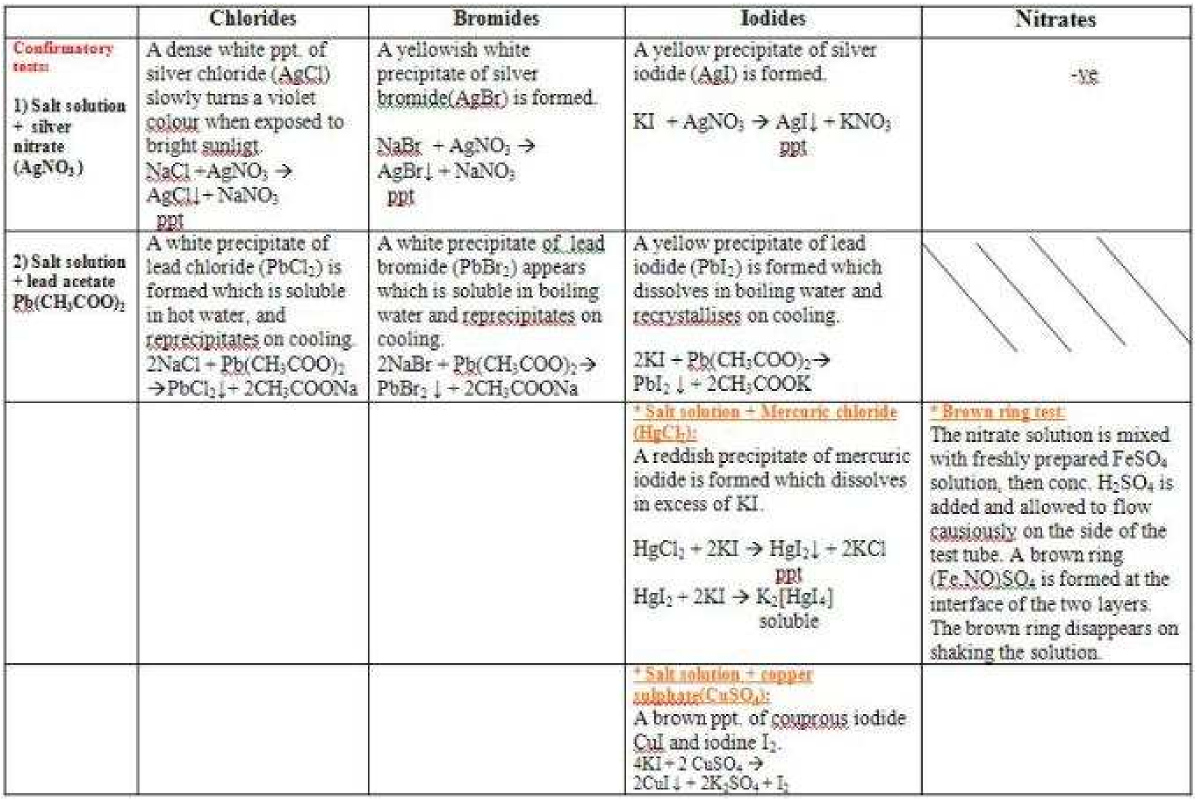
	Sulphites	**Thiosulphates**
	Salt solution + 2 drops of dil. H_2SO_4 + Iodine solution(I_2)	
	The brown colour of the iodine solution disappears. $Na_2SO_3 + I_2 + H_2O \rightarrow Na_2SO_4 + 2HI$	The brown colour of the iodine solution disappears, because the iodine is reduced to the iodide ion, which is colourless. At the same time, the thiosulphate is oxidized to tetrathionate: $2Na_2S_2O_3 + I_2 \rightarrow Na_2S_4O_6 + 2\,NaI$



2) Anions which react with concentrated sulphuric acid

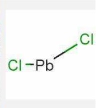
Note: Do these tests in fuming cuper because the gases which are given off in these tests are extremely irritating, and can cause damage to the sensitive mucous membranes of nose and throat.

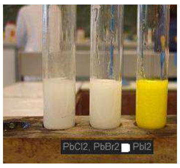
	Chlorides (Cl^-)	**Bromides** (Br^-)	**Iodides** (I^-)	**Nitrates** (NO_3^-)
	All chlorides are water soluble except the chlorides of silver, mercurous and cuprous.	Bromides resemble chlorides in their solubility.	Iodides resemble chlorides and bromides in their solubility. However, bismuth iodide is insoluble.	All nitrates are soluble in water except some basic nitrates.
Solid salt + conc.H_2SO_4:	Effervescence with evolution of colourless gas , Hydrogen chloride (HCl), $2NaCl + H_2SO_4 \rightarrow Na_2SO_4 + 2\ HCl$ Test for HCl gas: Dip a clean glass rod into a bottle of conc. ammonia solution, and hold it close to the mouth of the test tube in which the HCl is being produced. Dense white fumes of ammonium chloride will be formed: $HCl + NH_3 \rightarrow NH_4Cl$	Raddish fumes evolve and the solution turns orange due to libertion of bromine (Br_2) $2NaBr + H_2SO_4 \rightarrow 2HBr + Na_2SO_4$ $2HBr + H_2SO_4 \rightarrow 2H_2O + SO_2 + Br_2$	Violet fumes are evolved, (Iodine gas I_2) and a brown or black precipitate is formed in the test tube. $2KI + 2\ H_2SO_4 \rightarrow K_2SO_4 + 2H_2O + SO_2 + I_2$	-ve If a small piece of copper metal and drops of water are added, dense brown fumes of nitrogen dioxide (NO_2) will be given off. If no results, heat carefully. $2KNO_3 + H_2SO_4 \rightarrow K_2SO_4 + 2HNO_3$ $4HNO_3 + Cu \rightarrow Cu(NO_3)_2 + 2H_2O + 2NO_2$



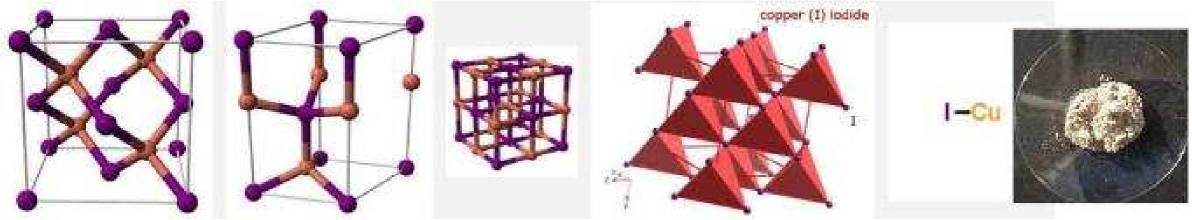
	Chlorides	**Bromides**	**Iodides**	**Nitrates**
Confirmatory tests: **1) Salt solution + silver nitrate ($AgNO_3$)**	A dense white ppt. of silver chloride (AgCl) slowly turns a violet colour when exposed to bright sunligt. $NaCl + AgNO_3 \rightarrow \underset{ppt}{AgCl\downarrow} + NaNO_3$	A yellowish white precipitate of silver bromide(AgBr) is formed. $NaBr + AgNO_3 \rightarrow \underset{ppt}{AgBr\downarrow} + NaNO_3$	A yellow precipitate of silver iodide (AgI) is formed. $KI + AgNO_3 \rightarrow \underset{ppt}{AgI\downarrow} + KNO_3$	-ve
2) Salt solution + lead acetate $Pb(CH_3COO)_2$	A white precipitate of lead chloride ($PbCl_2$) is formed which is soluble in hot water, and reprecipitates on cooling. $2NaCl + Pb(CH_3COO)_2 \rightarrow PbCl_2\downarrow + 2CH_3COONa$	A white precipitate of lead bromide ($PbBr_2$) appears which is soluble in boiling water and reprecipitates on cooling. $2NaBr + Pb(CH_3COO)_2 \rightarrow PbBr_2\downarrow + 2CH_3COONa$	A yellow precipitate of lead iodide (PbI_2) is formed which dissolves in boiling water and recrystallises on cooling. $2KI + Pb(CH_3COO)_2 \rightarrow PbI_2\downarrow + 2CH_3COOK$	
			* Salt solution + Mercuric chloride ($HgCl_2$): A reddish precipitate of mercuric iodide is formed which dissolves in excess of KI. $HgCi_2 + 2KI \rightarrow \underset{ppt}{HgI_2\downarrow} + 2KCl$ $HgI_2 + 2KI \rightarrow \underset{soluble}{K_2[HgI_4]}$	* Brown ring test: The nitrate solution is mixed with freshly prepared $FeSO_4$ solution, then conc. H_2SO_4 is added and allowed to flow causiously on the side of the test tube. A brown ring $(Fe.NO)SO_4$ is formed at the interface of the two layers. The brown ring disappears on shaking the solution.
			* Salt solution + copper sulphate($CuSO_4$): A brown ppt. of couprous iodide CuI and iodine I_2. $4KI + 2\ CuSO_4 \rightarrow 2CuI\downarrow + 2K_2SO_4 + I_2$	

AgCl is white

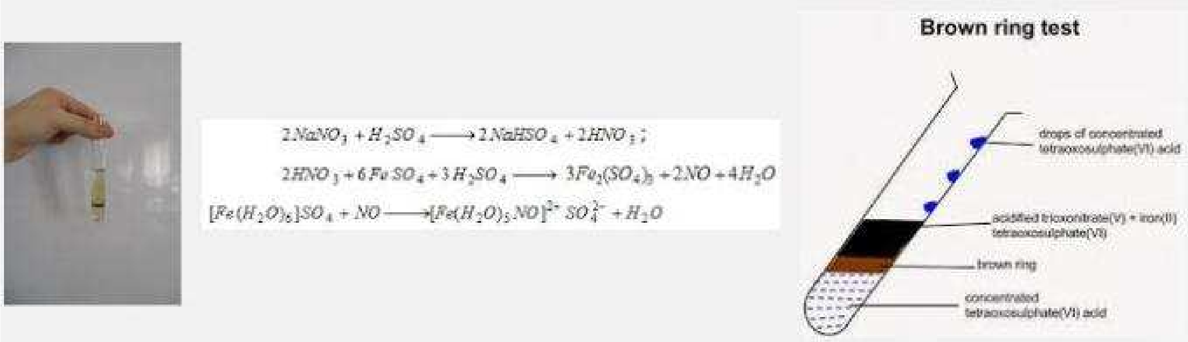




CuI Cuprous Iodide



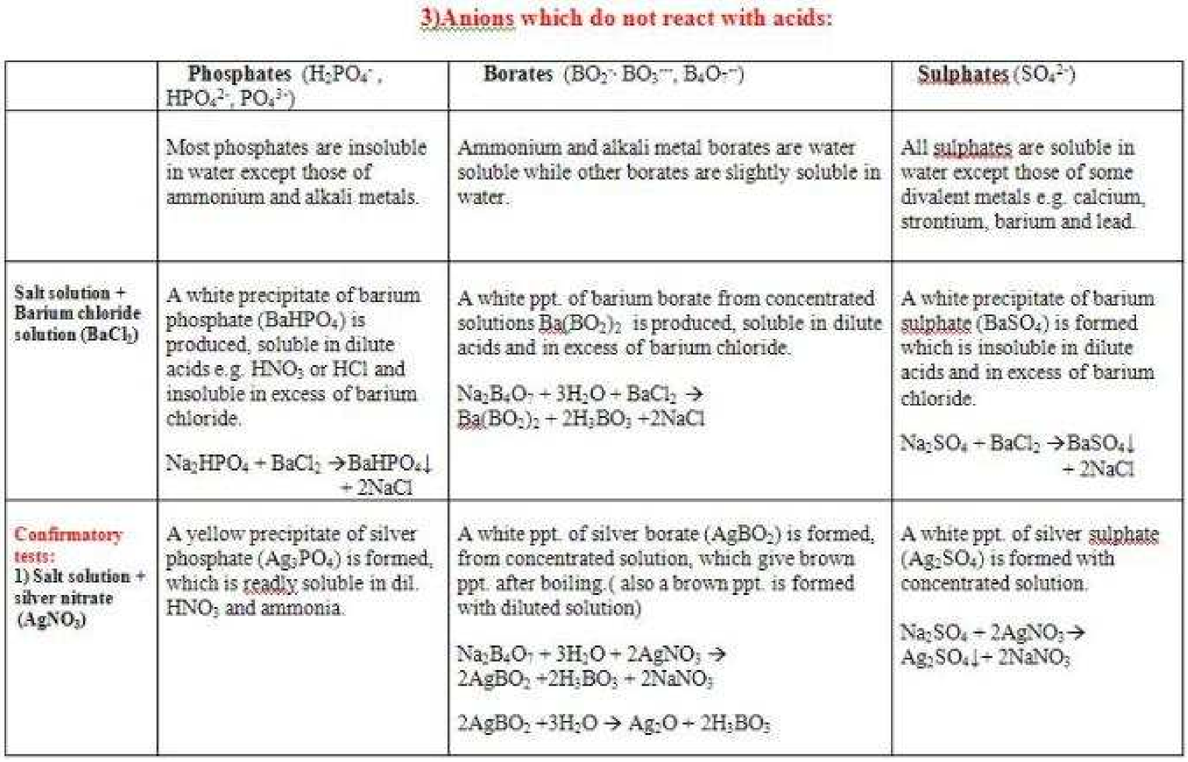
Brown Ring Test



$$2NaNO_3 + H_2SO_4 \longrightarrow 2NaHSO_4 + 2HNO_3;$$

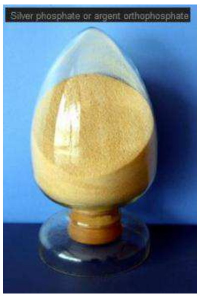
$$2HNO_3 + 6FeSO_4 + 3H_2SO_4 \longrightarrow 3Fe_2(SO_4)_3 + 2NO + 4H_2O$$

$$[Fe(H_2O)_6]SO_4 + NO \longrightarrow [Fe(H_2O)_5NO]^{2+} SO_4^{2-} + H_2O$$

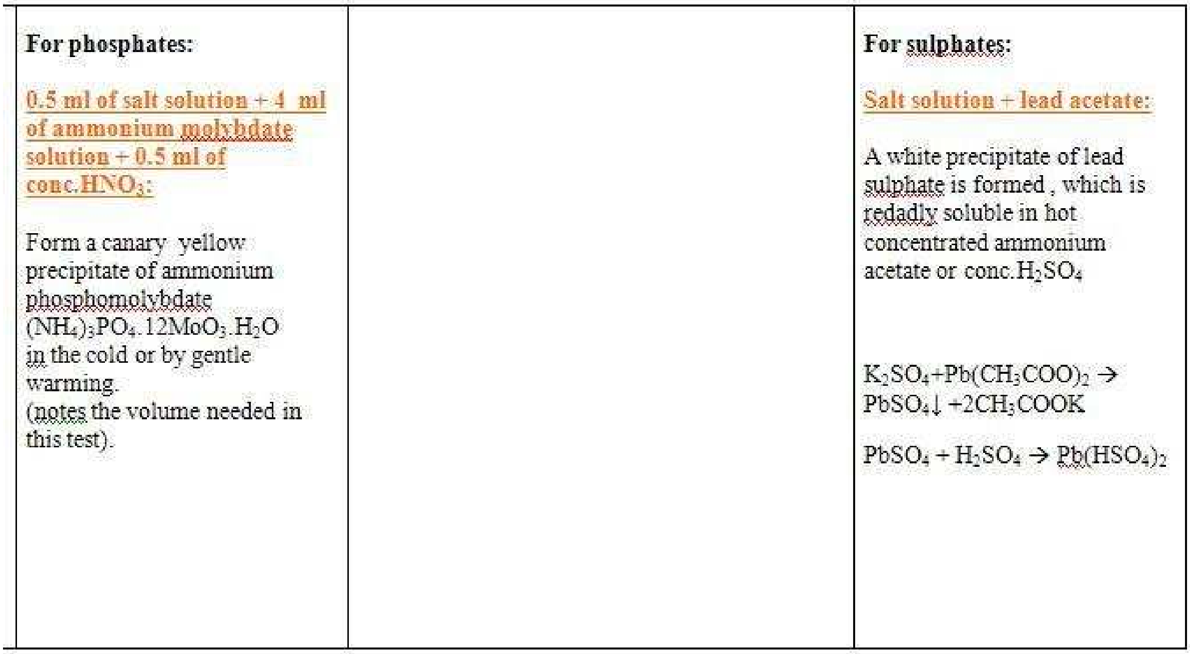


3)Anions which do not react with acids:

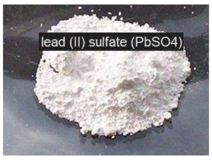
	Phosphates ($H_2PO_4^-$, HPO_4^{2-}, PO_4^{3-})	**Borates** (BO_2^-· BO_3^{---}, $B_4O_7^{--}$)	**Sulphates** (SO_4^{2-})
	Most phosphates are insoluble in water except those of ammonium and alkali metals.	Ammonium and alkali metal borates are water soluble while other borates are slightly soluble in water.	All sulphates are soluble in water except those of some divalent metals e.g. calcium, strontium, barium and lead.
Salt solution + Barium chloride solution ($BaCl_2$)	A white precipitate of barium phosphate ($BaHPO_4$) is produced, soluble in dilute acids e.g. HNO_3 or HCl and insoluble in excess of barium chloride. $Na_2HPO_4 + BaCl_2 \rightarrow BaHPO_4\downarrow + 2NaCl$	A white ppt. of barium borate from concentrated solutions $Ba(BO_2)_2$ is produced, soluble in dilute acids and in excess of barium chloride. $Na_2B_4O_7 + 3H_2O + BaCl_2 \rightarrow Ba(BO_2)_2 + 2H_3BO_3 + 2NaCl$	A white precipitate of barium sulphate ($BaSO_4$) is formed which is insoluble in dilute acids and in excess of barium chloride. $Na_2SO_4 + BaCl_2 \rightarrow BaSO_4\downarrow + 2NaCl$
Confirmatory tests: **1) Salt solution + silver nitrate ($AgNO_3$)**	A yellow precipitate of silver phosphate (Ag_3PO_4) is formed, which is readly soluble in dil. HNO_3 and ammonia.	A white ppt. of silver borate ($AgBO_2$) is formed, from concentrated solution, which give brown ppt. after boiling.(also a brown ppt. is formed with diluted solution) $Na_2B_4O_7 + 3H_2O + 2AgNO_3 \rightarrow 2AgBO_2 + 2H_3BO_3 + 2NaNO_3$ $2AgBO_2 + 3H_2O \rightarrow Ag_2O + 2H_3BO_3$	A white ppt. of silver sulphate (Ag_2SO_4) is formed with concentrated solution. $Na_2SO_4 + 2AgNO_3 \rightarrow Ag_2SO_4\downarrow + 2NaNO_3$

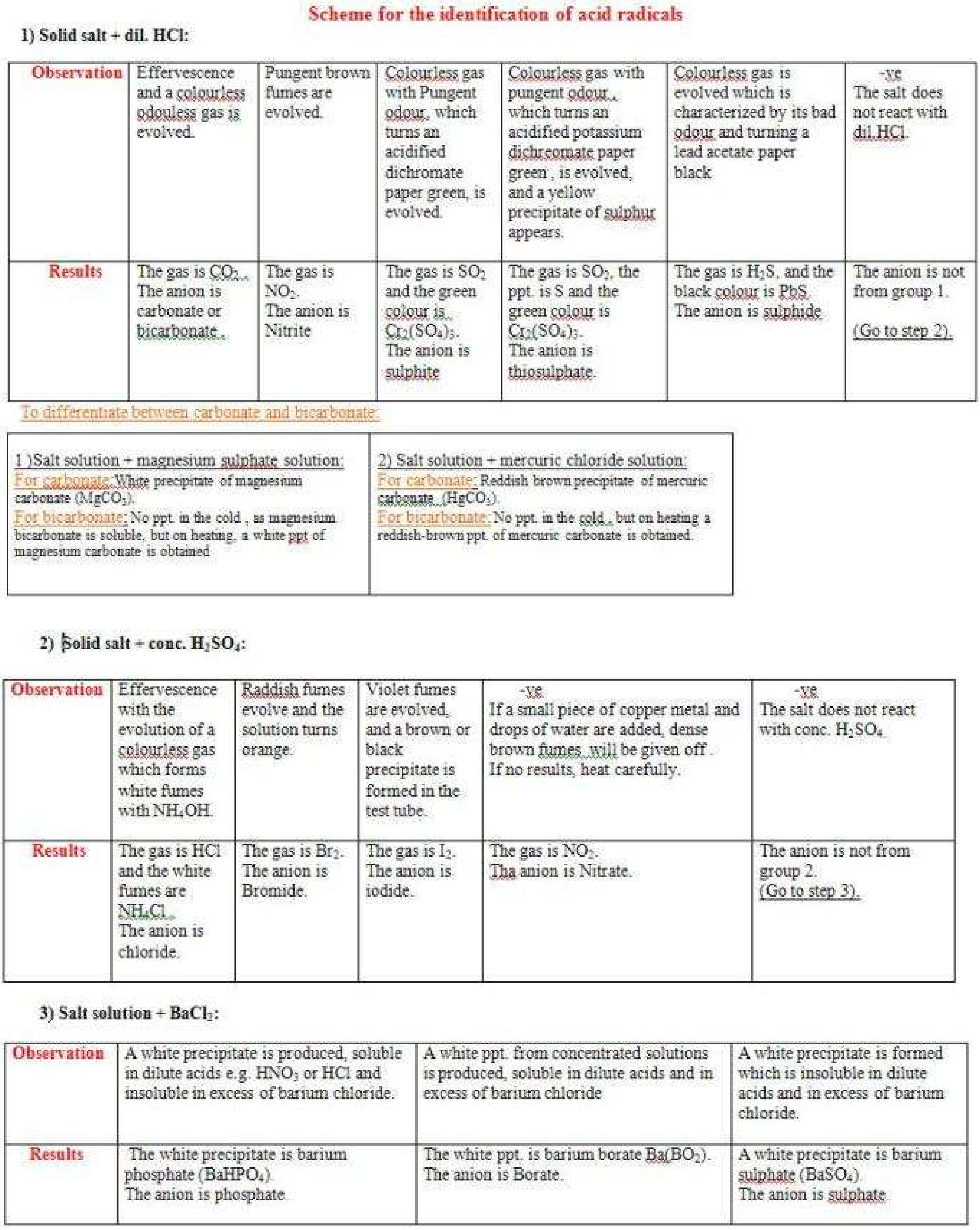


BaSO4 is white



For phosphates: **0.5 ml of salt solution + 4 ml of ammonium molybdate solution + 0.5 ml of conc.HNO_3:** Form a canary yellow precipitate of ammonium phosphomolybdate $(NH_4)_3PO_4.12MoO_3.H_2O$ in the cold or by gentle warming. (notes the volume needed in this test).		**For sulphates:** **Salt solution + lead acetate:** A white precipitate of lead sulphate is formed , which is redadly soluble in hot concentrated ammonium acetate or conc.H_2SO_4 $K_2SO_4 + Pb(CH_3COO)_2 \rightarrow PbSO_4\downarrow + 2CH_3COOK$ $PbSO_4 + H_2SO_4 \rightarrow Pb(HSO_4)_2$





Scheme for the identification of acid radicals

1) Solid salt + dil. HCl:

Observation	Effervescence and a colourless odouless gas is evolved.	Pungent brown fumes are evolved.	Colourless gas with Pungent odour, which turns an acidified dichromate paper green, is evolved.	Colourless gas with pungent odour, which turns an acidified potassium dichromate paper green , is evolved, and a yellow precipitate of sulphur appears.	Colourless gas is evolved which is characterized by its bad odour and turning a lead acetate paper black	-ve The salt does not react with dil.HCl.
Results	The gas is CO_2. The anion is carbonate or bicarbonate.	The gas is NO_2. The anion is Nitrite	The gas is SO_2 and the green colour is $Cr_2(SO_4)_3$. The anion is sulphite	The gas is SO_2, the ppt. is S and the green colour is $Cr_2(SO_4)_3$. The anion is thiosulphate.	The gas is H_2S, and the black colour is PbS. The anion is sulphide	The anion is not from group 1. (Go to step 2).

To differentiate between carbonate and bicarbonate:

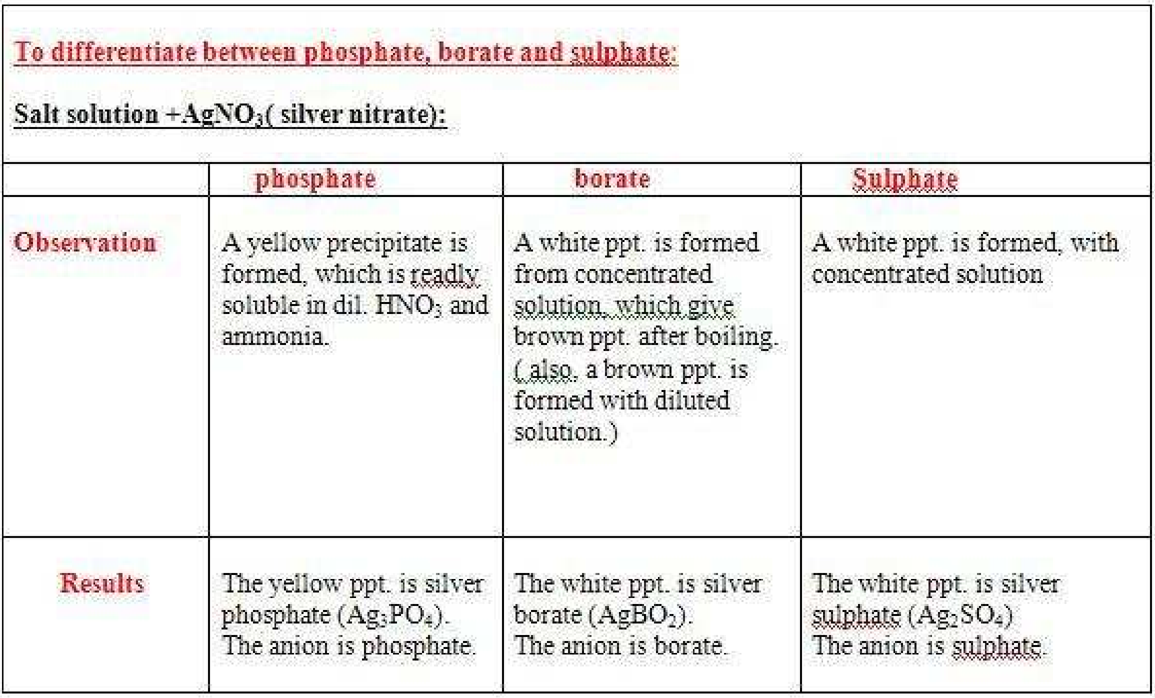
1)Salt solution + magnesium sulphate solution: For carbonate:White precipitate of magnesium carbonate ($MgCO_3$). For bicarbonate: No ppt. in the cold , as magnesium bicarbonate is soluble, but on heating, a white ppt of magnesium carbonate is obtained	2) Salt solution + mercuric chloride solution: For carbonate: Reddish brown precipitate of mercuric carbonate. ($HgCO_3$). For bicarbonate: No ppt. in the cold , but on heating a reddish-brown ppt. of mercuric carbonate is obtained.

2) Solid salt + conc. H_2SO_4:

Observation	Effervescence with the evolution of a colourless gas which forms white fumes with NH_4OH.	Raddish fumes evolve and the solution turns orange.	Violet fumes are evolved, and a brown or black precipitate is formed in the test tube.	-ve If a small piece of copper metal and drops of water are added, dense brown fumes will be given off . If no results, heat carefully.	-ve The salt does not react with conc. H_2SO_4
Results	The gas is HCl and the white fumes are NH_4Cl. The anion is chloride.	The gas is Br_2. The anion is Bromide.	The gas is I_2. The anion is iodide.	The gas is NO_2. Tha anion is Nitrate.	The anion is not from group 2. (Go to step 3).

3) Salt solution + $BaCl_2$:

Observation	A white precipitate is produced, soluble in dilute acids e.g. HNO_3 or HCl and insoluble in excess of barium chloride.	A white ppt. from concentrated solutions is produced, soluble in dilute acids and in excess of barium chloride	A white precipitate is formed which is insoluble in dilute acids and in excess of barium chloride.
Results	The white precipitate is barium phosphate ($BaHPO_4$). The anion is phosphate.	The white ppt. is barium borate $Ba(BO_2)$. The anion is Borate.	A white precipitate is barium sulphate ($BaSO_4$). The anion is sulphate.

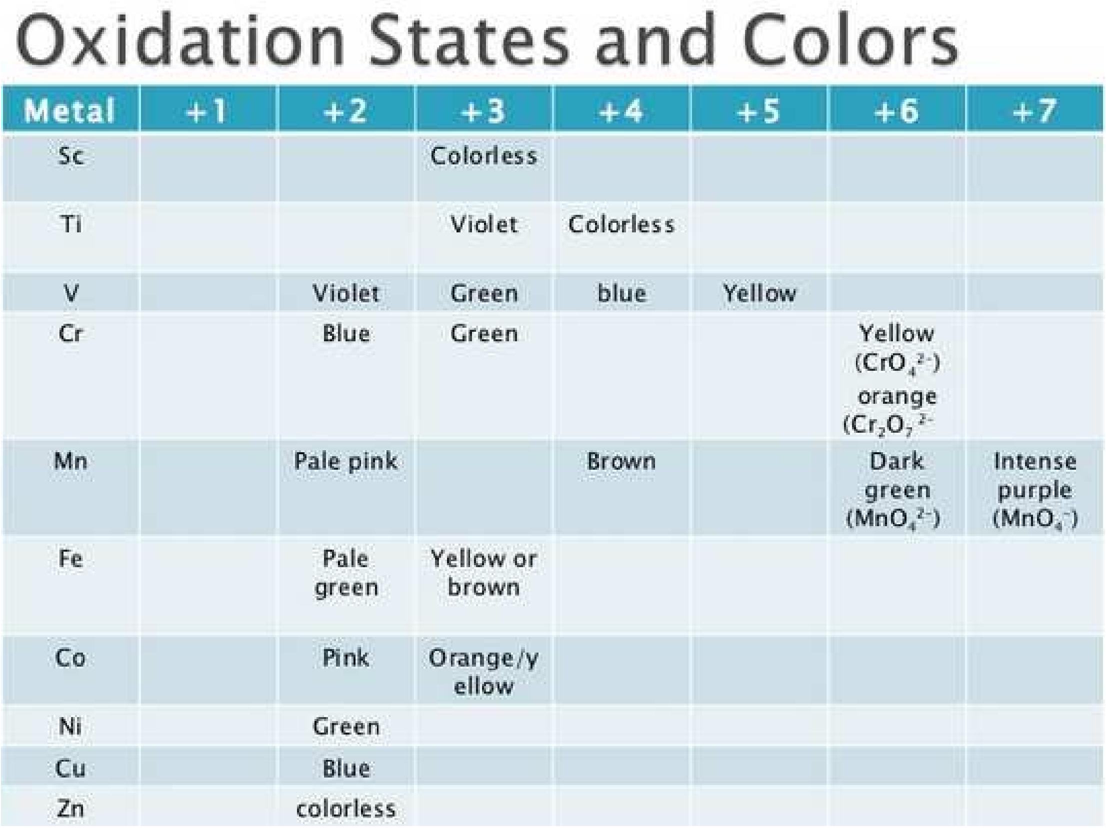


To differentiate between phosphate, borate and sulphate:

Salt solution +$AgNO_3$(silver nitrate):

	phosphate	borate	Sulphate
Observation	A yellow precipitate is formed, which is readly soluble in dil. HNO_3 and ammonia.	A white ppt. is formed from concentrated solution, which give brown ppt. after boiling. (also, a brown ppt. is formed with diluted solution.)	A white ppt. is formed, with concentrated solution
Results	The yellow ppt. is silver phosphate (Ag_3PO_4). The anion is phosphate.	The white ppt. is silver borate ($AgBO_2$). The anion is borate.	The white ppt. is silver sulphate (Ag_2SO_4) The anion is sulphate.

-



Oxidation States and Colors

Metal	+1	+2	+3	+4	+5	+6	+7
Sc			Colorless				
Ti			Violet	Colorless			
V		Violet	Green	blue	Yellow		
Cr		Blue	Green			Yellow (CrO_4^{2-}) orange $(Cr_2O_7^{2-}$	
Mn		Pale pink		Brown		Dark green (MnO_4^{2-})	Intense purple (MnO_4^-)
Fe		Pale green	Yellow or brown				
Co		Pink	Orange/yellow				
Ni		Green					
Cu		Blue					
Zn		colorless					

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Good Luck to you for your Preparations, References, and Exams

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